

DOCUMENT RESUME

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CE 035 300

TITLE Pharmacy Specialist, 10-8. Military Curriculum Materials for Vocational and Technical Education.

INSTITUTION Air Force Training Command, Sheppard AFB, Tex.; Ohio State Univ., Columbus. National Center for Research in Vocational Education.

SPONS AGENCY Office of Education (DHEW), Washington, D.C.

PUB DATE 18 Jul 75

NOTE 774p.; Some pages are marginally legible.

PUB TYPE Guides - Classroom Use - Guides (For Teachers) (052)

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DESCRIPTORS Behavioral Objectives; Course Descriptions; Curriculum Guides; Drug Abuse; Drug Therapy; *Drug Use; Learning Activities; Lesson Plans; *Pharmaceutical Education; Pharmacists; *Pharmacology; *Pharmacy; Postsecondary Education; Programed Instructional Materials; Textbooks; Workbooks

IDENTIFIERS Military Curriculum Project

ABSTRACT

These teacher and student materials for a postsecondary-level course in pharmacy comprise one of a number of military-developed curriculum packages selected for adaptation to vocational instruction and curriculum development in a civilian setting. The purpose stated for the 256-hour course is to train students in the basic technical phases of pharmacy and the minimum essential knowledge and skills necessary for the compounding and dispensing of drugs, the economical operation of a pharmacy, and the proper use of drugs, chemicals, and biological products. The course consists of three blocks of instruction. Block I contains four lessons: pharmaceutical calculations I and laboratory, inorganic chemistry, and organic chemistry. The five lessons in Block II cover anatomy and physiology, introduction to pharmacology, toxicology, drug abuse, and pharmaceutical and medicinal agents. Block III provides five lessons: pharmaceutical calculations I and II, techniques of pharmaceutical compounding, pharmaceutical dosage forms, and compounding laboratory. Instructor materials include a course chart, lesson plans, and a plan of instruction detailing instructional units, criterion objectives, lesson duration, and support materials needed. Student materials are eight study guides or workbooks with exercises and problems, three programed tests, and seven handouts. Suggested audiovisual aids are not provided. (YLB)

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MILITARY CURRICULUM MATERIALS

The military-developed curriculum materials in this course package were selected by the National Center for Research in Vocational Education Military Curriculum Project for dissemination to the six regional Curriculum Coordination Centers and other instructional materials agencies. The purpose of disseminating these courses was to make curriculum materials developed by the military more accessible to vocational educators in the civilian setting.

The course materials were acquired, evaluated by project staff and practitioners in the field, and prepared for dissemination. Materials which were specific to the military were deleted, copyrighted materials were either omitted or approval for their use was obtained. These course packages contain curriculum resource materials which can be adapted to support vocational instruction and curriculum development.

The National Center Mission Statement

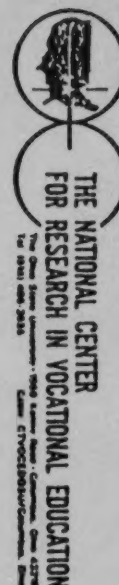
The National Center for Research in Vocational Education's mission is to increase the ability of diverse agencies, institutions, and organizations to solve educational problems relating to individual career planning, preparation, and progression. The National Center fulfills its mission by:

- Generating knowledge through research
- Developing educational programs and products
- Evaluating individual program needs and outcomes
- Installing educational programs and products
- Operating information systems and services
- Conducting leadership development and training programs

FOR FURTHER INFORMATION ABOUT Military Curriculum Materials

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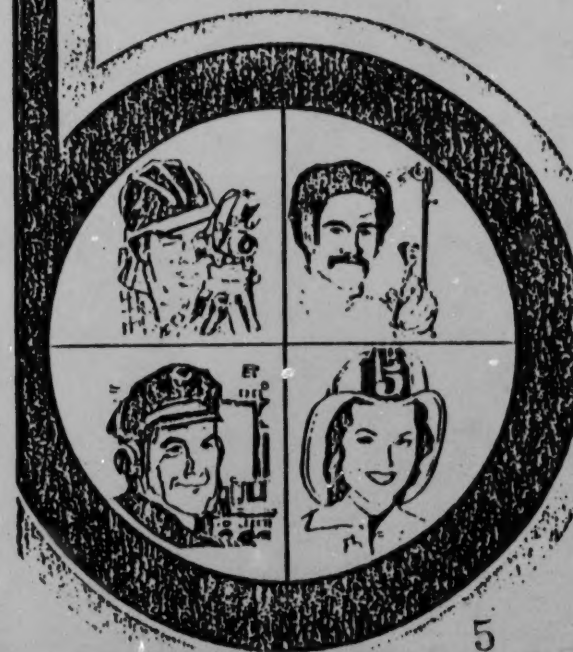
Program Information Office
The National Center for Research in Vocational
Education
The Ohio State University
1960 Kenny Road, Columbus, Ohio 43210
Telephone: 614/486-3655 or Toll Free 800/
848-4815 within the continental U.S.
(except Ohio)



Military Curriculum Materials for Vocational and Technical Education

Information and Field
Services Division

The National Center for Research
in Vocational Education



Military Curriculum Materials Dissemination Is . . .

an activity to increase the accessibility of military developed curriculum materials to vocational and technical educators.

This project, funded by the U.S. Office of Education, includes the identification and acquisition of curriculum materials in print form from the Coast Guard, Air Force, Army, Marine Corps and Navy.

Access to military curriculum materials is provided through a "Joint Memorandum of Understanding" between the U.S. Office of Education and the Department of Defense.

The acquired materials are reviewed by staff and subject matter specialists, and courses deemed applicable to vocational and technical education are selected for dissemination.

The National Center for Research in Vocational Education is the U.S. Office of Education's designated representative to acquire the materials and conduct the project activities.

Project Staff:

Wesley E. Budke, Ph.D., Director
National Center Clearinghouse

Shirley A. Chase, Ph.D.
Project Director

What Materials Are Available?

One hundred twenty courses on microfiche (thirteen in paper form) and descriptions of each have been provided to the vocational Curriculum Coordination Centers and other instructional materials agencies for dissemination.

Course materials include programmed instruction, curriculum outlines, instructor guides, student workbooks and technical manuals.

The 120 courses represent the following sixteen vocational subject areas:

Agriculture	Food Service
Aviation	Health
Building &	Heating & Air
Construction	Conditioning
Trades	Machine Shop
Clerical	Management &
Occupations	Supervision
Communications	Meteorology &
Drafting	Navigation
Electronics	Photography
Engine Mechanics	Public Service

The number of courses and the subject areas represented will expand as additional materials with application to vocational and technical education are identified and selected for dissemination.

How Can These Materials Be Obtained?

Contact the Curriculum Coordination Center in your region for information on obtaining materials (e.g., availability and cost). They will respond to your request directly or refer you to an instructional materials agency closer to you.

CURRICULUM COORDINATION CENTERS

EAST CENTRAL

Rebecca S. Douglass
Director
100 North First Street
Springfield, IL 62777
217/782-0759

MIDWEST

Robert Patton
Director
1515 West Sixth Ave.
Stillwater, OK 74704
405/377-2000

NORTHEAST

Joseph F. Kelly, Ph.D.
Director
225 West State Street
Trenton, NJ 08625
609/292-6562

NORTHWEST

William Daniels
Director
Building 17
Airdustrial Park
Olympia, WA 98504
206/753-0879

SOUTHEAST

James F. Shill, Ph.D.
Director
Mississippi State University
Drawer DX
Mississippi State, MS 39762
601/325-2510

WESTERN

Lawrence F. H. Zane, Ph.D.
Director
1776 University Ave.
Honolulu, HI 96822
808/948-7834

PHARMACY SPECIALIST

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PHARMACY SPECIALIST

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Availability:
Military Curriculum Project, The Center
for Vocational Education, 1960 Kenny
Rd., Columbus, OH 43210

Expires July 1, 1978

Course Description

This course trains students in the basic technical phases of pharmacy and the minimum essential knowledge and skills necessary for the compounding and dispensing of drugs, the economical operation of a pharmacy, and the proper use of drugs, chemicals, and biological products. The course consists of three blocks covering 256 hours of instruction.

Block I — *Fundamentals of Pharmacy* contains four lessons involving 64 hours of instruction. One orientation lesson and two lessons on pharmacy administration were deleted because they discuss specific military operations and clinical procedures. The included lesson topics and respective hours follow:

- Pharmaceutical Calculations I (18 hours)
- Pharmaceutical Calculations I Laboratory (6 hours)
- Pharmaceutical Inorganic Chemistry (18 hours)
- Pharmaceutical Organic Chemistry (22 hours)

Block II — *Pharmacology* contains five lessons covering 86 hours of instruction. One lesson on the dispensing laboratory was deleted.

- Anatomy and Physiology (18 hours)
- Introduction to Pharmacology (2 hours)
- Toxicology (2 hours)
- Drug Abuse (4 hours)
- Pharmaceutical and Medicinal Agents (60 hours)

Block III — *Pharmaceutical Preparations and Manufacture* contains five lessons covering 102 hours of instruction.

- Pharmaceutical Calculations II (16 hours)
- Pharmaceutical Calculations II Laboratory (3 hours)
- Techniques of Pharmaceutical Compounding (8 hours)
- Pharmaceutical Dosage Forms (34 hours)
- Compounding Laboratory (42 hours)

This course contains both teacher and student materials. Printed instructor materials include a course chart; lesson plans for each block of instruction; and a plan of instruction detailing instructional units, criterion objectives, the duration of the lessons, and support materials needed. Student materials provided include four study guide/workbooks, one programmed text and two handouts for Block I; one study guide/workbook, two programmed texts and three handouts for Block II; and three study guide/workbooks and two handouts for Block III.

All text materials are provided for this course. Audiovisuals suggested for use include four slide sets, eight transparency sets, and nine films. The audiovisuals are not provided.

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COURSE CHART		
NUMBER 3ABR90530	PDS CODE ARC	DATE 3 December 1975
COURSE TITLE Pharmacy Specialist		
ATC OPR AND APPROVAL DATE SGHE, 3 March 1975	CENTER OPR Sheppard/SHCS/MSOXC	SUPERSLOES COURSE CHART 3ABR90530, 18 July 1975
DEPARTMENT OPR Department of Biomedical Sciences		APPLICABLE TRAINING STANDARD STS 905X0, 28 February 1975
LOCATION OF TRAINING Sheppard AFB, Texas 76311		COURSE SECURITY CLASSIFICATION UNCLASSIFIED
INSTRUCTIONAL DESIGN Group/Lock Step		TARGET READING GRADE LEVEL FOR PREPARATION OF TRAINING LITERATURE 11.3
LENGTH OF TRAINING (12 Weeks, 0 Days)		Hours
Technical Training		460
Classroom/Laboratory (C/L)		360
Complementary Technical Training (CTT)		100
Related training		20
Standard Traffic Safety (AFR 50-24)		12
Local Conditions Course (Course II) (AFR 50-24)		2
Commander's Calls/Briefings		4
End of Course Appointments; Predeparture Safety Briefing (ATCR 127-1)		2
TOTAL		480
REMARKS		
Applicable safety is integrated throughout the course.		
Effective date: 5 January 1976 with class 760102. All previously enrolled classes will continue to be governed by course chart dated 18 July 1975.		
TABLE I - MAJOR ITEMS OF EQUIPMENT		
Class A Prescription Balances Laboratory Magnetic Stirrer - Hot Plate Filter Tank Unit with Mixer Water Distilling Apparatus Suppository Molds Prescription Bottle Filling Machine Laminar Flow Station Typewriters Prescription Numbering Machines Prescription Label Imprinter Tablet-Capsule Counting Machine		

COURSE CHART - TABLE II - TRAINING CONTENT

3ABR90530

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NOTE: Include time spent on technical training (TT) (classroom/laboratory (C/L) and complementary technical training (CTT) and related training (RT). Exclude time spent on individual assistance (remedial instruction). A single entry of time shown for a unit is C/L time. When a double entry is shown, the second entry is CTT time.

Wk OF TNG	Hrs Per Day	1	2	3	4	5	6	7	8
1 2 3 4(2/5)		Course Material - UNCLASSIFIED 122 Hours TT					14 Hours RT		
		BLOCK I - Fundamentals of Pharmacy							
		Welcome and Orientation (2 hrs); Pharmaceutical Calculations I (18/6 hrs); Pharmaceutical Calculations I Laboratory (6 hrs); Pharmaceutical Inorganic Chemistry (18/3 hrs); Pharmaceutical Organic Chemistry (22/5 hrs); Pharmacy Administration (22/4 hrs); Pharmacy Administration Laboratory (6 hrs); Measurement Test and Test Critique (10 hrs)					18 Hours CTT		
		104 Hours C/L							
4(3/5) 5 6 7 8 9(.2/5)		Course Material - UNCLASSIFIED 182 Hours TT					8 Hours CTT		
		BLOCK II - Pharmacology					2 Hours RT		
		Anatomy and Physiology (18/6 hrs); Introduction to Pharmacology (2/2 hrs); Toxicology (2 hrs); Drug Abuse (4 hrs); Pharmaceutical and Medicinal Agents (60/22 hrs); Dispensing Laboratory (42/14 hrs); Measurement Test and Test Critique (10 hrs)					36 Hours CTT		
		138 Hours C/L					2 Hours RT		
9(4.8/5) 10 11 12		Course Material - UNCLASSIFIED 156 Hours TT							
		BLOCK III - Pharmaceutical Preparations and Their Manufacture							
		Pharmaceutical Calculations II (16/6 hrs); Pharmaceutical Calculations II Laboratory (3/2 hrs); Techniques of Pharmaceutical Compounding (8/2 hrs); Pharmaceutical Dosage Forms (34/12 hrs); Compounding Laboratory (42/16 hrs); Measurement Test and Test Critique (11 hrs); Equipment Turn-In (1 hr); Course Critique (2 hrs); Graduation (1 hr)					38 Hours CTT		
		118 Hours C/L					2 Hours RT		

PLAN OF INSTRUCTION
(Technical Training)

10-8

PHARMACY SPECIALIST



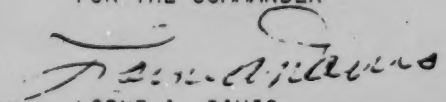
SHEPPARD TECHNICAL TRAINING CENTER

18 July 1975- Effective 31 July 1975 with Class 750731

FOREWORD

1. PURPOSE. This plan of instruction prescribes the qualitative requirements for Course Number 3ABR90530, Pharmacy Specialist, in terms of criterion objectives presented by units/modules of instruction, and shows duration, correlation with the training standard, support materials, and instructional guidance. It was developed under the provisions of ATCR 50-5, Instructional Systems Development, and ATCR 52-7, Plans of Instruction.
 2. COURSE DESCRIPTION. This 12 week technical training course trains airmen to perform duties prescribed in AFM 39-1 for Pharmacy Specialist, AFSC 90530. It includes training in the basic technical phases of pharmacy and the minimum essential knowledge and skills necessary for compounding and dispensing of drugs, the economical operation of an Air Force Pharmacy, the proper use of drugs, chemicals, and biological products of the Federal Catalog of Medical Materiel. In addition, related training consists of driver education, supplemental military training, commander's calls/briefing, end of course appointments and a predeparture safety briefing.
 3. EQUIPMENT ALLOWANCE AND AUTHORIZATION. Training equipment required to conduct this course, and for which accountability must be maintained, is found in the Report of Medical and Non-Medical In-Use Equipment and is listed under custody account number 285588.
- NOTE: Group size is shown in parentheses after equipment listed in column 3 of numbered pages of this POI.
4. MULTIPLE INSTRUCTOR REQUIREMENTS. Units of instruction which require more than one instructor per instructional group are identified in the multiple instructor annex to this POI.
 5. REFERENCES. This plan of instruction is based on SPECIALTY TRAINING STANDARD 905X0, 28 February 1975, Change 1, 24 July 1975 and COURSE CHART 18 July 1975.
 6. POI OVERLAP DURING PHASE-IN. All classes enrolled prior to 31 July 1975 will continue to be governed by POI 3ABR90530, 26 June 1974.

FOR THE COMMANDER


LORNE A. DAVIS
Chief, Training Operations Division

Supersedes Plan of Instruction 3ABR90530, 26 June 1974
OPR: Department of Biomedical Sciences
DISTRIBUTION: Listed on Page A.

MODIFICATIONS

Lesson 1 of this publication has (have) been deleted in adapting this material for inclusion in the "Trial Implementation of a Model System to Provide Military Curriculum Materials for Use in Vocational and Technical Education." Deleted material involves extensive use of military forms, procedures, systems, etc. and was not considered appropriate for use in vocational and technical education.

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PLAN OF INSTRUCTION (Continued)			
1. UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2. DURATION (HOURS)	3. SUPPORT MATERIALS AND GUIDANCE	
2. <u>Pharmaceutical Calculations I</u> a. Solve problems pertaining to basic mathematical operations, metric system, apothecary system, avoirdupois system, and ratio and proportion. b. Solve problems pertaining to conversion of weights and measures, and calculation of doses.	24 (18/6)	<u>Column 1 Reference</u> 2a 2b	<u>STS Reference</u> 14a, 14b, 14c, 14d 14a, 14b, 14c, 14d
	(10)	<u>Instructional Materials</u> SW 3ABR90530-I-1, Pharmaceutical Calculations I HO 3ABR90530-I-1 thru 12, Pharmaceutical Calculations-I	
	(8)	<u>Audio Visual Aids</u> Flip Chart Set, Pharmaceutical Calculations I Transparency Set, Pharmaceutical Calculations I	
		<u>Training Methods</u> Lecture/Discussion (15 hrs) Demonstration (3 hrs) Outside Assignments (6 hrs) <u>Instructional Environment/Design</u> Classroom (18 hrs) Home Study (6 hrs) Group/Work Step <u>Instructional Guidance</u> Discuss and demonstrate Basic Mathematical Operations, Metric System, Apothecary System, Ratio and Proportion. Conduct a two hour Pharmaceutical Calculations Laboratory (as indicated in Block I, Unit 3) followed by Measurement Test I-1 and Critique. Discuss and demonstrate conversion of weights and measures and calculation of doses. Conduct a four hour Calculations laboratory (as indicated in Block I, Unit 3) and administer Measurement Test I-2 and Critique.	
PLAN OF INSTRUCTION NO. 3ABR90530		DATE 18 JUL 1975	PAGE NO. 2

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
		<u>Training Methods</u> Lecture/Discussion (16 hrs) Performance (2 hrs) Outside Assignments (3 hrs) <u>Instructional Environment/Design</u> Classroom (16 hrs) Laboratory (2 hrs) Home Study (3 hrs) Group/Work Step <u>Instructional Guidance</u> Discuss the Basic Concepts, Composition and Properties of Matter; Classification of the Elements, Formulas and Naming Inorganic Compounds Writing and Balancing Equations; Molar and Normal Solutions; Oxygen, Hydrogen, Water and Peroxides, Alkali and Alkaline Earth Metals; Halogens; Sulfur, Nitrogen and Boron; and Miscellaneous Inorganic Elements.	
6. <u>Measurement Test and Test Critique</u>	2		
7. <u>Pharmaceutical Organic Chemistry</u>	27 (22/5)	<u>Column 1 Reference</u> 7a 7b	<u>STS Reference</u> 15j, 15k 15j, 15k
a. Identify the basic concepts, principles and definitions of pharmaceutical organic chemistry.	(11)	<u>Instructional Materials</u> WB 3ABR90530-1-2, Fundamentals of Pharmacy	
b. Select the properties of pharmaceutical organic chemical compounds.	(11)	<u>Audio Visual Aids</u> Transparency Set, Pharmaceutical Organic Chemistry Flip Chart Set, Pharmaceutical Organic Chemistry <u>Training Methods</u> Lecture/Discussion (22 hrs) Outside Assignments (5 hrs)	
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. I	PAGE NO. 4

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
8. <u>Measurement Test and Test Critique</u>	2	<u>Instructional Environment/Design</u> Classroom (22 hrs) Home Study (5 hrs) Group/Lock Step <u>Instructional Guidance</u> Discuss Organic Chemistry; Aliphatic Hydrocarbons, Alcohols, Aldehydes, Ketones and Esters; Aliphatic Acids, Esters and Salts; Surfactants, Aromatic Acids and Derivatives, Aliphatic and Aromatic Halogenated Compounds; Amines and Amides; Amino Acids and Proteins, Carbohydrates; Glycosides; Alkaloids; Steroids and Miscellaneous Organic Compounds.	
MODIFICATIONS			
<u>Lessons 9 + 10</u> of this publication has (have) been deleted in adapting this material for inclusion in the "Trial Implementation of a Model System to Provide Military Curriculum Materials for Use in Vocational and Technical Education." Deleted material involves extensive use of military forms, procedures, systems, etc. and was not considered appropriate for use in vocational and technical education.			
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. 1	PAGE NO. 5

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
10. <u>Pharmacy Administration Laboratory</u> a. Given instructor assistance and placed in the dispensing pharmacy, fill prescriptions and complete ward, bulk compounding and supply forms in accordance with AFM 168-4 for legend and controlled drugs.	6	<u>Instructional Environment/Design</u> Classroom (22 hrs) Home Study (4 hrs) Group/Lock Step <u>Instructional Guidance</u> Discuss Pharmacy References, Pharmacy Law, Pharmacy Supply, Prescription Reading and Terminology. Conduct a Pharmacy Administration Laboratory as explained in Block I, Unit 10. <u>Column 1 Reference</u> 10a. <u>STS Reference</u> 5a, 5b, 5c, 7b, 7d(2), 7d(3), 7d(4), 7d(5), 7d(7), 7d(8), 7d(9), 7d(10), 9a(3), 9a, 9g, 10b, 11b, 11d, 11f, 11g <u>Instructional Materials</u> Extract AFM 168-4, Administration of Medical Activities AF Form 579, Ward Alcohol and Narcotics Register AF Form 582, Pharmacy Stock Record AF Form 781, Multiple Item Prescription AF Form 115a, Register of Control Numbers DD 1348-6, NON-FSN Requisition (Manual) Pharmacy Administration Reference File HQ 3ABR90530-1-14, Pharmacy Administration <u>Training Equipment</u> Dispensing Pharmacy (7) Typewriter (3) Numbering Machines (7) Prescription Files (7) Telephones (7) <u>Training Methods</u> Performance (6 hrs) <u>Instructional Environment/Design</u> Laboratory (6 hrs) Group/Lock Step	
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 1 JUL 1975	BLOCK NO. I	PAGE NO. 6

PLAN OF INSTRUCTION		COURSE TITLE	
BLOCK TITLE Pharmacology		Pharmacy Specialist	
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION HOURS	3 SUPPORT MATERIALS AND GUIDANCE	
1. ✓ <u>Anatomy and Physiology</u>	24 (18/6)	<u>Column 1 Reference</u> 1a 1b 1c 1d, 1e, 1f, 1g 1h 1i 1j 1k 1l <u>STS Reference</u> 16a 16c 16b, 16d 16d, 16e 16d, 16e, 16f 16d, 16e 16d, 16e, 16g 16d, 16e 16d, 16e <u>Instructional Materials</u> PT 3ABR90530-II-1, Prefixes, Roots and Suffixes of Medical Terminology HO 3ABR90530-II-1, Anatomical Drawings SW 3ABR90530-II-1b, Anatomy and Physiology <u>Audio Visual Aids</u> Transparency Set #1, Anatomy and Physiology <u>Training Equipment</u> American Froese Anatomical Chart (20) FA Manual (1) <u>Training Methods</u> Lecture/Discussion (18 hrs) Outside Assignments (6 hrs) <u>Instructional Environment/Design</u> Classroom (18 hrs) Home Study (6 hrs) Group/Lock Step	
a. Identify the etymology of selected medical terms and choose the meaning of terms, word roots, combining forms, suffixes and prefixes.	(0/6)		
b. Identify selected cells, tissues and glands pertaining to the human body.	(2)		
c. Identify selected basic facts and terms about the skeletal system.	(1)		
d. Identify selected basic facts and terms about the muscular system.	(2)		
e. Identify selected basic facts and terms about the nervous system.	(3)		
f. Identify selected basic facts and terms about the circulatory system.	(2)		
g. Identify selected basic facts and terms about the respiratory system.	(1)		
h. Identify selected basic facts and terms about the digestive system.	(2)		
i. Identify selected basic facts and terms about the endocrine system.	(1)		
j. Identify selected basic facts and terms about the urinary system.	(1)		
PLAN OF INSTRUCTION NO. 3ABR90530	18 JUL 1975	BLOCK NO. II	PAGE 8

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
→k. Identify selected basic facts and terms about the reproductive system. →l. Identify selected basic facts and terms about the eye and ear. 2. ✓ <u>Introduction to Pharmacology</u> →a. Identify the basic principles of pharmacology. 3. ✓ <u>Toxicology</u> →a. Classify the symptoms of drug toxicity.	(1) (2) 4 (2/2) 2	<u>Instructional Guidance</u> Identify the major organs and describe the functions of each system. Also relate the importance of each system to the body. Have students label parts of the human body in their workbooks. <u>Column 1 Reference</u> 2a <u>STS Reference</u> 17a, 17b, 17c <u>Instructional Materials</u> WB 3ABR90530-II-1, Pharmacology <u>Audio Visual Aids</u> Transparency Set #1, Pharmacology <u>Training Methods</u> Lecture/Discussion (2 hrs) Outside Assignments (2 hrs) <u>Instructional Environment/Design</u> Classroom (2 hrs) Home Study (2 hrs) Group/Work Step <u>Column 1 Reference</u> 3a <u>STS Reference</u> 4d(4), 18a, 18b, 18c, 18d <u>Instructional Materials</u> WB 3ABR90530-II-1, Pharmacology <u>Audio Visual Aids</u> Transparency Set, Toxicology <u>Training Methods</u> Lecture/Discussion (2 hrs)	
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. II	PAGE NO. 9

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PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
4. <u>Drug Abuse</u> ✓a. Describe the drugs subject to abuse and the symptoms of drug abuse.	4	<u>Instructional Environment/Design</u> Classroom (2 hrs) Group/Lock Step <u>Column 1 Reference</u> <u>STS Reference</u> 4a 11 <u>Instructional Materials</u> WB 3ABR90530-II-1, Pharmacology <u>Audio Visual Aids</u> Unnumbered Films, Weed; Acid 11:59 - Last Minute to Choose; Speedscene - The Problem of Amphetamine Abuse; The Perfect Drug; Hooks; Drug Abuse <u>Training Methods</u> Lecture/Discussion (4 hrs) <u>Instructional Environment/Design</u> Classroom (4 hrs) Group/Lock Step <u>Instructional Guidance</u> Discuss films with students. Show one of the above which has not been previously shown in basic training.	
	82 (60/22) (18) (10)	<u>Column 1 Reference</u> <u>STS Reference</u> 5a, 5b, 5c 17a, 17b, 17c 5d 17a, 17b, 17c <u>Instructional Materials</u> WB 3ABR90530-II-1, Pharmacology <u>Audio Visual Aids</u> Transparency Sets #2, 3, 4, 5; Pharmacology Commercial Films, The Digestive System; Ascariasis; Fundamentals of the Nervous System; Halothane; Work of the Heart; The Blood; Common Heart Disorders; Congestive Heart Failure; Endocrine System; Menstrual Cycle; Vitamins and Some Deficiency Diseases; Immunization; It's A Plot; Abnormal Behavior	
5. <u>Pharmaceutical and Medicinal Agents</u> ✓a. Classify and describe the properties of locally acting drugs, gastrointestinal drugs, local anesthetics and anti-infective drugs. ✓b. Classify and describe the properties of drugs acting on the central nervous system.		PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975 BLOCK NO. 11 PAGE NO. 10

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
c. Classify and describe the properties of drugs acting on the autonomic nervous system and circulatory system. d. Classify and describe the properties of drugs acting on the endocrine system and miscellaneous therapeutic drugs.	(16) (16)	<u>Training Methods</u> Lecture/Discussion (60 hrs) Outside Assignments (22 hrs) <u>Instructional Environment/Design</u> Classroom (60 hrs) Home Study (22 hrs) Group/Lock Step <u>Instructional Guidance</u> Discuss anatomy and physiology of the human body, Introduction to Pharmacology and Toxicology. Administer measurement test II-1. Discuss locally acting drugs, gastrointestinal drugs, local anesthetic drugs and anti-infective drugs. Conduct a 12 hour dispensing laboratory utilizing these drugs as explained in Block II, Unit 6. Administer measurement test II-2 and critique. Discuss drugs acting on the central nervous system. Conduct a 6 hour dispensing laboratory utilizing these drugs as explained in Block II, Unit 6. Administer measurement test II-3. Discuss drugs acting on the autonomic nervous system and circulatory system. Conduct a 12 hour dispensing laboratory utilizing these drugs as explained in Block II, Unit 6. Administer measurement test II-4 and critique. Discuss drugs acting on the endocrine system and miscellaneous therapeutic drugs. Conduct a 12 hour dispensing laboratory utilizing these drugs as explained in Block II, Unit 6. Administer measurement test II-5 and critique.	
6. <u>Dispensing Laboratory</u> a. Given instructor assistance and placed in the dispensing pharmacy (model pharmacy and pharmacology research area), correctly interpret, fill and label prescriptions in accordance with AEM 168-4 and complete handouts for locally acting drugs, gastrointestinal drugs, local anesthetics and anti-infective drugs.	56 (42/14) (12)	<u>Column 1 Reference</u> 6a, 6b, 6c 6d <u>Instructional Materials</u> HO 3ABR90530-II-1 thru 7, Pharmacology <u>STS Reference</u> 5a, 5b, 5c, 7d(6), 7d(8), 9a, 9g, 10b, 11b, 11d, 11f, 11g, 13a, 13b, 13c, 13h, 13i, 18a, 18c, 18d 6b, 9a, 10b, 11b, 11d, 11f, 11g, 13a, 13b, 13c, 13d, 13i, 18a, 18c, 18d	
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. II	PAGE NO. 11

PLAN OF INSTRUCTION		COURSE TITLE	
		Pharmacy Specialist	
BLOCK TITLE			
Pharmaceutical Preparations and Their Manufacture			
1	2	3	
UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	DURATION (HOURS)	SUPPORT MATERIALS AND GUIDANCE	
1. ✓ <u>Pharmaceutical Calculations II</u> a. Solve problems in reducing and enlarging formulas, specific gravity, percentage preparations, concentration and dilution, alligation, and temperature conversion.	20 (16/4)	<u>Column 1 Reference</u> 1a <u>Instructional Materials</u> SW 3ABR90530-III-1, Pharmaceutical Preparations HO 3ABR90530-III-1 thru 2, Pharmaceutical Preparations <u>Audio Visual Aids</u> Transparency Set #1, Pharmaceutical Calculations II <u>Training Methods</u> Lecture/Discussion (16 hrs) Outside Assignments (4 hrs) <u>Instructional Environment/Design</u> Classroom (16 hrs) Home Study (4 hrs) Group/Lock Step <u>Instructional Guidance</u> Discuss reducing and enlarging of formulas, specific gravity, percentage preparation, concentration and dilution of stock solutions and stock triturations, alligation, temperature conversion. Conduct a three hour pharmaceutical calculations laboratory (as indicated in Block I, Unit 2) followed by Measurement Test III-1 and Critique.	
PLAN OF INSTRUCTION NO.	DATE	BLOCK NO.	PAGE NO.
3ABR90530	18 JUL 1975	III	13

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
2. <u>Pharmaceutical Calculations II Laboratory</u> <u>a.</u> Given information pertaining to reducing and enlarging formulas, specific gravity, percentage preparations, concentration and dilution, alligation, and temperature conversion, solve problems in each area in SW 3ABR90530-III-1 with 60 percent accuracy.	3	<u>Column 1 Reference</u> <u>STS Reference</u> 2a 14e, 14f, 14g, 14h, 14i, 14j, 14k <u>Instructional Materials</u> SW 3ABR90530-III-1, Pharmaceutical Preparations <u>Audio Visual Aids</u> Transparency Set #2, Pharmaceutical Calculations II <u>Training Methods</u> Performance (3 hrs) <u>Instructional Environment/Design</u> Laboratory (3 hrs) Group/Lock Step <u>Instructional Guidance</u> This laboratory will be conducted over subject areas indicated in Block III Unit 1. All work in SW 3ABR90530-III-1 which is related to POI objective 2a must be completed in class under the supervision of the instructor.	
3. <u>Measurement Test and Test Critique</u>	3		
4. <u>Techniques of Pharmaceutical Compounding</u> <u>a.</u> Identify laboratory equipment, equipment user maintenance procedures, metrology procedures, incompatibilities, and methods of comminution.	8	<u>Column 1 Reference</u> <u>STS Reference</u> 4a T2a, T2b, T2g, 19a, 19b, 19c, 19d <u>Instructional Materials</u> WB 3ABR90530-III-2, Pharmaceutical Preparations <u>Audio Visual Aids</u> Transparency Set #1, Pharmaceutical Preparations <u>Training Methods</u> Lecture/Demonstration (3 hrs)	
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. III	PAGE NO. 14

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
5. <u>Pharmaceutical Dosage Forms</u> ✓a. Identify the properties, preparation techniques and incompatibilities of waters, spirits, solutions, and syrups. ✓b. Identify the properties, preparation techniques and incompatibilities of eye, ear and nose preparations, elixirs, tinctures, mixtures, magmas, suspension, gels, lotions, and liniments. ✓c. Identify the properties, preparation techniques and incompatibilities of powders capsules, emulsions, ointments, pastes, creams, and suppositories. ✓d. Identify the properties, preparation techniques and incompatibilities of parenterals, bulk compound, prepackaged items and intravenous admixtures. ✓e. Identify information as classified, unclassified, of possible intelligence value, Top Secret, Secret, Confidential, or For Official Use Only.		<u>Instructional Environment/Design</u> Classroom (8 hrs) Group/Lock Step <u>Instructional Guidance</u> Demonstrate and discuss laboratory equipment, metrology, incompatibilities and comminution. Review pertinent forms and emphasize safety procedures.	
	40 (34/6)	<u>Column 1 Reference</u>	<u>STS Reference</u>
	(4)	5a	12c(1), 12c(2), 12c(3), 12c(4), 12f, 13h, 19a, 19b, 19c, 19d
		5b	12c(5), 12c(6), 12c(7), 12c(8), 12c(9), 12c(10), 12c(11), 12c(12), 12c(13), 12f, 13g, 13h, 19a, 19b, 19c, 19d
		5c	12c(14), 12c(15), 12c(16), 12c(17), 12c(18), 12c(19), 12c(20), 12f, 13h, 19a, 19b, 19c, 19d
	(8)	5d 5e 5f	12f, 13e, 13h, 19a, 19b, 19c, 19d 3a, 3b 3c, 3d
	(10)	<u>Instructional Materials</u> WB 3ABR90530-III-2, Pharmaceutical Preparations SW 3ABR9XXXX, Communication Security	
	(12)	<u>Audio Visual Aids</u> Slides, Ohio State University Admixture Slides and Cassette Tape	
		<u>Training Methods</u> Lecture/Discussion (34 hrs) Outside Assignments (6 hrs)	
	(0/.5)	<u>Instructional Environment/Design</u> Classroom (34 hrs) Home Study (6 hrs) Group/Lock Step	
PLAN OF INSTRUCTION	3ABR90530	DATE	18 JUL 1975
		BLOCK NO.	III
		PAGE NO.	15

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
f. Select and recommend mode of transmission dictated by security and expediency required, and observe security precautions involved in communications.	(0/.5)	Instructional Guidance Discuss and demonstrate the preparation of waters, spirits, solutions, and syrups. Conduct a compounding laboratory as explained in Block III, Unit 6. Administer Measurement Test III-2 and Critique. Discuss and demonstrate the preparation of eye, ear, and nose preparation, elixirs and tinctures, mixtures, magmas, suspensions, gels, lotions and liniments. Conduct a compounding laboratory as explained in Block III Unit 6. Administer Measurement Test III-3 and Critique. Discuss and demonstrate the preparation of powders, capsules, emulsions, ointments, pastes, creams and suppositories. Conduct a compounding laboratory as explained in Block III, Unit 6. Administer Measurement Test III-4 and Critique. Discuss and demonstrate parenterals, bulk compounding, prepackaging and intravenous admixtures. Conduct a laboratory as explained in Block III, Unit 6. Administer Measurement Test III-5 and Critique. These hours may vary in scheduling due to a lack of availability of equipment.	
6. <u>✓Compounding Laboratory</u>	50 (42/8)	Column 1 Reference 6a 6b 6c 6d 6e	
a. Given instructor assistance, necessary references and selected formulas; compound waters, spirits, solutions, and syrups in accordance with AF Form 2380 and AF Form 2381. Then package and label in accordance with AFM 168-4.	(4)	STS Reference 4a, 4b, 4c, 4d(1), 4d(2), 4d(3), 10a, 10b, 10c, 12a, 12b, 12c(1), 12c(2), 12c(3), 12c(4), 12d, 12e, 12f, 12g, 12h, 13a, 13b, 13c, 13d, 13e, 13f, 13i, 19a, 19b, 19c, 19d 4b, 4d(1), 4d(2), 4d(3), 10a, 10b, 10c, 12a, 12b, 12c(5), 12c(6), 12c(7), 12c(10), 12c(12), 12d, 12e, 12f, 12g, 12h, 13a, 13b, 13c, 13d, 13f, 13i, 19a, 19b, 19c, 19d 4d(1), 4d(2), 4d(3), 10a, 10b, 10c, 12a, 12b, 12c(15), 12c(16), 12c(17), 12c(19), 12d, 12e, 12f, 12g, 13a, 13b, 13c, 13d, 13f, 13g, 13i, 19a, 19b, 19c, 19d 4d(1), 4d(2), 4d(3), 10a, 10b, 10c, 12a, 12b, 12f, 12g, 12h, 13a, 13b, 13c, 13e, 13f, 13h, 13i, 19a, 19b, 19c, 19d 4b, 4d(1), 4d(2), 4d(3), 5a, 5b, 5c, 7d(6), 7d(8), 9g, 10a, 10b, 10c, 11b, 11d, 11f, 11g, 12a, 12b, 12c(8), 12c(11), 12c(13), 12d, 12e, 12f, 12g, 12h, 13a, 13b, 13c, 13d, 13f, 13h, 13i, 18c, 18d, 19a, 19b, 19c, 19d	
b. Given instructor assistance, necessary references and selected formulas; compound ear and nose preparations, elixirs, tinctures, magmas, suspensions, lotions and liniments in accordance with AF Form 2380 and AF Form 2381. Then package and label preparation in accordance with AFM 168-4.	(10)		
c. Given instructor assistance, necessary references and selected formulas; compound powders, capsules, emulsions, ointments, creams and suppositories in accordance with AF Form 2380 and AF Form 2381. Then package and label			
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. III	PAGE NO. 16

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
preparations in accordance with AFM 168-4. d. Given instructor assistance, necessary references and selected prescriptions; compound intravenous admixtures, correcting any incompatibilities, using accepted methods and techniques as outlined in checklist 3ABR90530-III-6d. e. Given instructor assistance, rotate through the outpatient, inpatient, supply and administrative work areas of the USAF Regional Hospital Sheppard Pharmacy in accordance with local directives and policies.	(10)	<u>Instructional Materials</u> WB 3ABR90530-III-3, Pharmaceutical Preparations Book, The United States Pharmacopeia, Committee of Revision, United States Pharmacopeia Convention Book, National Formulary, National Formulary Board, American Pharmaceutical Association Book, Remington's Pharmaceutical Sciences, Mack Publishing Company	
	(12)	<u>Training Equipment</u> Compounding laboratory (20) Laboratory equipment (1) Laminar flow hood (7) IV admixture materials (1)	
	(6)	Chemicals (1) Class A balances (1) Typewriters (4) Prescriptions (1) Alsop filter-tank unit (20) Tablet counting machine (20) Bottle filling machine (20) Label imprinter (20) <u>Training Methods</u> Performance (42 hrs) Outside Assignments (8 hrs) <u>Instructional Environment/Design</u> Laboratory (42 hrs) Home Study (8 hrs) Group/Lock Step <u>Instructional Guidance</u> Students compound the preparations demonstrated in Block III, Unit 5. Students are placed at work benches and using properly prepared forms in WB 3ABR90530-III-3 with instructor assistance, properly prepare selected formulas. Preparations are carefully checked for quality and correctness. During 3 days of the 12th week, divide the class into three groups. One group will work at the USAF Regional Hospital Sheppard. Two instructors	
PLAN OF INSTRUCTION NO. 3ABR90530	DATE 18 JUL 1975	BLOCK NO. III	PAGE NO. 17

PLAN OF INSTRUCTION (Continued)			
1 UNITS OF INSTRUCTION AND CRITERION OBJECTIVES	2 DURATION (HOURS)	3 SUPPORT MATERIALS AND GUIDANCE	
		<u>Instructional Guidance (Cont'd)</u> are required with the group at the Sheppard AFB Hospital since the groups will be working at two different locations within the hospital pharmacy. The second group will participate in the intravenous admixture program. The third group will participate in the incompatibilities laboratory. Rotate each group through each of the areas. Instructors will check and initial AF Form 2380 and AF Form 2381 in WB 3ABR90530-III-3 prior to students starting preparations.	
7. <u>Related Training</u> (identified in course chart)	22		
8. <u>Measurement Test and Test Critique</u>	8		
9. <u>Equipment Turn-In</u>	1		
10. <u>Course Critique</u>	2		
11. <u>Graduation</u>	1		
PLAN OF INSTRUCTION NO. 3ABR90530		DATE 18 JUL 1975	BLOCK NO. III
		PAGE NO. 18	

MODIFICATIONS

Annex 1-5 of this publication has (have) been deleted in adapting this material for inclusion in the "Trial Implementation of a Model System to Provide Military Curriculum Materials for Use in Vocational and Technical Education." Deleted material involves extensive use of military forms, procedures, systems, etc. and was not considered appropriate for use in vocational and technical education.

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Pharmaceutical Calculations I, Block I: Pharmaceutical Calculations
Performance - 2 instructors

The following training method will apply to all performance hours in pharmaceutical calculations I, block I.

In order for students to achieve the proficiency level required, they must be able to interpret the problem, select the correct formula and follow the correct procedures for solving pharmaceutical problems, with instructor assistance. They must then be able to solve selected problems independently.

Each instructor will work with a group of 10 students. He will lend individual assistance, make on-the-spot corrections, and determine when the student is capable of solving the problems independently. Finally, he will administer an evaluation quiz to determine if the objective has been met.

Week 1, Day 3, Hours 1-2 - Performance - 2 instructors

Week 1, Day 5, Hours 1-4 - Performance - 2 instructors

Pharmaceutical Inorganic Chemistry, Block I: Pharmaceutical Inorganic Chemistry
Performance - 2 instructors

The rationale for multiple instructors and training methods are identical to that for Pharmaceutical Calculations I, Block I.

Week 2, Day 7, Hours 5-6 - Performance - 2 instructors

Pharmacy Administration, Block I: Pharmacy Administration Laboratory -
Performance - 4 instructors

To achieve this function, the class is divided into three distinct functional areas in three separate physical locations. Students rotate through each of these 3 areas.

The functional areas and instructor requirements are as follows:

- a. Model Pharmacy - 2 instructors
 1. One instructor inside model pharmacy
 2. One instructor outside model pharmacy
- b. Inpatient Dispensing - 1 instructor
- c. Supply procedures - 1 instructor

Model Pharmacy: Inside instructor assists students in receiving, interpreting, preparing, labeling, dispensing and filing prescriptions.

Outside instructor functions in the role of physician, patient and evaluator of final preparations.

Inpatient Dispensing: Instructor assists students in receiving, preparing and dispensing ward and clinic orders. Involved in maintaining records on Schedule II controlled drugs, inpatient medication labels, and inspecting ward pharmaceuticals.

Supply Procedures: Instructor assists students in proper methods of inventorying drugs and equipment, procedures for suspending unsuitable items, determining stock levels, procedures in purchasing nonstock listed medications, receiving and storing bulk pharmaceuticals, biologicals, narcotics and other controlled drugs.

Week 4, Day 17, Hours 1-6 - Performance - 4 instructors

Pharmacology, Block II: Pharmacology Performance - 4 instructors

The following training method will apply to all performance hours in pharmacology, block II.

After listening to classroom lectures concerning a class of drugs, students will be divided into 3 separate groups as in the pharmacy administration laboratory and will perform in the model pharmacy, inpatient dispensing and supply procedures areas. They will apply all knowledge, procedures and techniques gained in the course thus far to filling prescriptions and dispensing actual drugs in the following categories:

- a. Locally-acting drugs
- b. Anti-infective drugs
- c. Drugs acting on the Central Nervous System
- d. Drugs acting on the Autonomic Nervous System
- e. Drugs acting on the Circulatory System
- f. Drugs acting on the Endocrine System
- g. Miscellaneous Therapeutic Drugs

Instructor function in the three performance areas in the same manner as they do in the performance phase of the Pharmacy Administration laboratory.

Week 6, Day 26, Hours 1-6 Performance - 4 instructors

Week 6, Day 27, Hours 1-6 - Performance - 4 instructors

Week 6, Day 30, Hours 1-6 - Performance - 4 instructors

Week 7, Day 34, Hours 1-6 - Performance - 4 instructors

Week 7, Day 35, Hours 1-6 - Performance - 4 instructors

Week 8, Day 39, Hours 1-6 - Performance - 4 instructors

Week 8, Day 40, Hours 1-6 - Performance - 4 instructors

Pharmaceutical Calculations II, Block III: Pharmaceutical Calculations Performance - 2 instructors.

The rationale for multiple instructors in this block is identical to that for Pharmaceutical Calculations, Block I.

Week 9, Day 44, Hours 1-3 - Performance - 2 instructors.

Techniques of Pharmaceutical Compounding, Block III: Techniques of Pharmaceutical compounding performance - 6 instructors

After listening to classroom lectures and demonstration students will be assigned projects to complete at the workbenches involving the principles and techniques of metrology. Close supervision of performance is vital to insuring student proficiency in this very important area. An instructor at each of the 5 work benches is required to enable the continual monitoring of student technique, accuracy and safety practices. The sixth laboratory instructor will serve as primary lecturer, evaluator and coordinator of bench instructors.

Block 10, Day 47, Hours 1-4 - Performance - 6 instructors

Compounding Laboratory, Block III: Compounding Laboratory Performance - 6 instructors.

The following training method will apply to all performance hours in Compounding Laboratory, Block III.

After listening to classroom lectures concerning pharmaceutical dosage forms the students will be placed in the laboratory and compound representative dosage forms. The specialty training standard requires that the students be able to:

- a. Weigh and measure drugs and chemicals
- b. Combine ingredients and prepare dosage forms
- c. Combine and compound stock and extemporaneous preparations
- d. Provide quality control data on manufactured and prepackaged preparations
- e. Identify and correct physical and chemical incompatibilities
- f. Identify toxic dose of ingredients

Students must develop correct compounding techniques as the health and welfare of the patient are at stake. Close supervision of performance is vital to insuring student proficiency in this very important area. An instructor at each of the 5 workbenches is required to enable the continual monitoring of student compound technique, accuracy and safety practices. Specifically, his duties include monitoring:

- a. Compounding technique and accuracy
 1. Selection of proper compounding material
 2. Selection of correct ingredients
 3. Accurate measurement of drugs
 4. Correct order of combining ingredients
 5. Correct packaging and labeling
- b. Safety practices

1. Flammable materials
2. Caustic materials
3. Explosive combinations
4. Equipment such as Fisher burners, pipetting techniques and steam baths

The sixth laboratory instructor will serve as primary lecturer, preparation evaluator and coordinator of bench instructors. He will also monitor the use of the controlled substances used in the laboratory.

Week 10, Day 49, Hours 3-6 - Performance - 6 instructors

Week 10, Day 50, Hours 1-6 - Performance - 6 instructors

Week 11, Day 53, Hours 1-6 - Performance - 6 instructors

Week 11, Day 54, Hours 1-4 - Performance - 6 instructors

Field Trip, Intravenous Admixture Performance and Incompatibility Research: During the last week of the course, to achieve the criterion objectives, the students are divided into three distinct functional areas in three separate physical locations. Each day a student group rotates through these areas.

The functional areas and instructor requirements are as follows:

- Hospital field trip - 2 instructors
- Incompatibility research - 1 instructor
- Intravenous Admixture performance - 3 instructors

Hospital Field Trip: The students will be assigned and rotated through outpatient dispensing, inpatient dispensing, bulk compounding, prepackaging, intravenous admixture, supply and records management areas of the hospital pharmacy by the instructors. The instructors will monitor the students performance in each of these areas. When actively dispensing medication to a patient, the student must be under the direct supervision of an instructor. One instructor will always be present at the dispensing window, the other to monitor the students in the areas mentioned above.

Incompatibility research: Instructor assists students in selecting and using correct pharmaceutical reference compendia to determine the correct dosage forms, doses, indications, contraindications, side effects, incompatibilities and drug interactions using a set of selected prescriptions.

Intravenous Admixture Performance: Extremely close supervision of performance is vital to insure student proficiency in this very important area. The group will be subdivided into 3 subgroups. Each subgroup will research the intravenous admixture prescriptions for incompatibilities, properly prepare the I.V. tray, prepare the I.V. admixture, check the I.V. admixture under the light/dark field, check the I.V. tray and label the I.V. admixture. Each subgroup will perform under simulated conditions and actual use of the laminar flow hood. The maximum number of students that can simultaneously use the flow hood is two.

Week 12, Day 57, Hours 1-6 - Performance - 5 instructors

Week 12, Day 58, Hours 1-6 - Performance - 5 instructors

Week 12, Day 59, Hours 1-6 - Performance - 5 instructors

APPROVAL OFFICE AND DATE MSDB <i>Wilson 8 Nov 74</i>		INSTRUCTOR	
COURSE NUMBER 3AB(90530)		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER I		BLOCK TITLE Fundamentals of Pharmacy	
LESSON TITLE Pharmacy Orientation			
LESSON DURATION			
CLASSROOM/Laboratory 2 / 0	~~LABORATORY~~ Complementary 0	TOTAL 2 hrs	
POI REFERENCE			
PAGE NUMBER 1	PAGE DATE 18 July 75	PARAGRAPH 1	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	8 Nov 74		
<i>[Signature]</i>	25 April 75		
<i>[Signature]</i>	10 May 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	SW3ABR90530-I-1, Course Orientation PT 3ABR90530-II-IA
CRITERION OBJECTIVES AND TEACHING STEPS			
1. <u>Course Orientation</u> a. Welcome of class and introduction of staff b. Overview of course c. Course supervisor's briefing (Teaching steps listed in Part II)			

APPROVAL OFFICE AND DATE MSDB <i>Wilson 3 Sept 74</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 1		BLOCK TITLE Fundamentals of Pharmacy	
LESSON TITLE Pharmaceutical Calculations I			
LESSON DURATION			
CLASSROOM/Laboratory 18 hrs/0 hrs	~~xxxxxxx~~ Complementary 68 hrs	TOTAL 2428 hrs	
POI REFERENCE			
PAGE NUMBER 2	PAGE DATE 18 July 75	PARAGRAPH 2a, b	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>R. E. Clegg</i>	3 SEPT 7	<i>R. E. Clegg</i>	6 OCT 1975
<i>R. E. Clegg</i>	1 MAR 5		
<i>R. E. Clegg</i>	27 MAR 5		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	Pharmaceutical Calculation - 1 Flip Chart Set Pharmaceutical Calculations Transparency Set SW3ABR90530-I-1, Fundamentals of Pharmacy WS1-1 Reduce all Fractions
CRITERION OBJECTIVES AND TEACHING STEPS			
2a. Solve problems pertaining to basic mathematical operations, metric system, apothecary system, avoirdupois system, and ratio and proportion. 2b. Solve problems pertaining to conversion of weights and measures, and calculation of doses. (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)

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APPROVAL OFFICE AND DATE MSDB <i>Wilson 3 Sept 74</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER I		BLOCK TITLE Fundamentals of Pharmacy	
LESSON TITLE Pharmaceutical Calculations I Laboratory			
LESSON DURATION			
CLASSROOM/Laboratory 0 hrs/6 hrs	~~COMPLEMENTARY~~ Complementary 0 hrs	TOTAL 6 hrs	
POI REFERENCE			
PAGE NUMBER 3	PAGE DATE 18 July 75	PARAGRAPH 3a, b.	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>R. E. Leggett</i>	3 Sept 74	<i>R. E. Leggett</i>	6 OCT 1975
<i>R. E. Leggett</i>	3 Mar 75		
<i>R. E. Leggett</i>	21 Aug 75		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	SW3ABR90530-I-1
CRITERION OBJECTIVES AND TEACHING STEPS			
3a. Given instructor assistance, solve problems in each each in SW3ABR90530-I-1, with a 60% accuracy in Basic Mathematical Operations, Metric System, Apothecary System, Avoirdupois System, and Ration and Proportion. 3b. Given instructor assistance, solve problems in each area in SW3ABR90530-I-1, with a 60% accuracy in conversion of weights and measurements and calculation of doses. (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)

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APPROVAL OFFICE AND DATE MSDR Wilson 3 Sept 74		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER I		BLOCK TITLE Fundamentals of Pharmacy	
LESSON TITLE Pharmaceutical Inorganic Chemistry			
LESSON DURATION			
CLASSROOM/Laboratory 16 hrs/2 hrs	Complementary 6 hrs	TOTAL 24	
POI REFERENCE			
PAGE NUMBER 3	PAGE D 18 July 75	PARAGRAPH 5	
STS/CTS REFERENCE			
NUMBER STS 905X0			
SUPERVISOR APPROVAL 28 Feb 75			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	3 SEPT 7	<i>[Signature]</i>	6 OCT 1975
<i>[Signature]</i>	1 MAR 3		
<i>[Signature]</i>	27 Aug 5		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	SW3ABR90530-1-1- Fundamentals of Pharmacy Transparency Set Pharmaceutical Inor- ganic Chemistry Unnumbered Film Explaining Matter, Atoms and Molecules
CRITERION OBJECTIVES AND TEACHING STEPS			
5a. Identify the basic concepts, principles, and definitions of pharmaceutical inorganic chemistry. b. Select the properties of pharmaceutical inorganic chemical elements and compounds. c. Given the names of specific inorganic elements, correctly write and balance simple chemical equations. Instructor assistance is permitted. d. Given the necessary data, correctly calculate the milliequivalent concentration of electrolyte solutions. Instructor assistance is permitted. (Teaching steps II)			

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE MSTNB Wilson 26 Sept 74		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 1		BLOCK TITLE Fundamentals of Pharmacy	
LESSON TITLE Pharmaceutical Organic Chemistry			
LESSON DURATION			
CLASSROOM /Laboratory 22 hrs/0 hrs	~~2200000000~~ Complementary 5 hrs	TOTAL 27 hrs	
POI REFERENCE			
PAGE NUMBER 4	PAGE DATE 18 July 75	PARAGRAPH 7	
STS/CTS REFERENCE			
NUMBER STS 905X0			
SUPERVISOR APPROVAL			
SIGNATURE		DATE	SIGNATURE
<i>[Signature]</i>		27 Sept 74	<i>[Signature]</i>
<i>[Signature]</i>		27 Mar 75	
<i>[Signature]</i>		12 Sept 75	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	NB 3ABR90530-I-2 Fundamentals of Pharmacy Transparency Set Pharmaceutical Organic Chemistry Flip Chart Set, Pharm- aceutical Organic Chemistry
CRITERION OBJECTIVES AND TEACHING STEPS			
7a. Identify the basic concepts, principles, and definitions of pharmaceutical organic chemistry. b. Select the properties of pharmaceutical organic chemical compounds. (Teaching steps listed in Part II)			

WORKSHEETS AND HANDOUTS - BLOCK I

COURSE 10-8

SABR90530 Pharm Cal-1

Worksheet SABR90530-1-1

1. Add: $5/6 + 1/2 + 1/6 + 1/3$ Answer _____
2. Add: $3/8 + 5/7 + 1\ 1/2 + 2\ 3/4$ Answer _____
3. Subtract: $3/7$ from $5/6$ Answer _____
4. Subtract: $1\ 5/16$ from $5\ 6/24$ Answer _____
5. Multiply: $5/8 \times 6/30$ Answer _____
6. Multiply: $3/8 \times 4/9 \times 8/15$ Answer _____
7. Divide: $5/16$ by $7/8$ Answer _____
8. Divide: $1\ 6/7$ by $1\ 3/5$ Answer _____
9. Convert $5/9$ to a decimal fraction Answer _____
10. Convert $1\ 3/16$ to a decimal fraction Answer _____
11. Convert .135 to a simple fraction Answer _____
12. Convert .625 to a simple fraction Answer _____
13. Add: $1.35 + .697 + .573 + 3.2153$ Answer _____
14. Add $6.3 + 9.721 + .611 + .0035$ Answer _____
15. Subtract .037 from 1.67 Answer _____
16. Subtract 5.335 from 10.224 Answer _____
17. Multiply 103.65×15.11 Answer _____
18. Multiply 19.66×5.25 Answer _____
19. Divide 103.6 by 7.5 Answer _____
20. Divide 6.3 by .773 Answer _____
21. Given $X = y/z$ solve for y answer _____
22. Given $12 = 15/A$ solve for A answer _____
23. Write the following in Arabic Numbers:

MCDLXI	=	_____
XXIV	=	_____
MCCX	=	_____
LXXIV	=	_____
CCXVII	=	_____
24. Write the following in Roman Numerals:

635	=	_____
37	=	_____
1974	=	_____
338	=	_____
99	=	_____

1. Add: $6/4 + 13/15 + 1/9 + 4/8$ Answer _____
2. Add: $2/3 + 14/16 + 1\ 5/6 + 5\ 5/16$ Answer 5 _____
3. Subtract: $1/3$ from $14/15$ Answer _____
4. Subtract: $1\ 3/8$ from $5\ 7/14$ Answer _____
5. Multiply: $6/10 \times 5/14$ Answer _____
6. Multiply: $1\ 3/5 \times 15/16 \times 6/9$ Answer _____
7. Divide: $5/6 \div 1/12$ Answer _____
8. Divide: $1\ 3/5 \div 4/7$ Answer _____
9. Convert $5/8$ to a decimal fraction Answer _____
10. Convert $3\ 9/16$ to a decimal fraction Answer _____
11. Convert .035 to a simple fraction Answer _____
12. Convert .655 to a simple fraction Answer _____
13. Add: $135.606 + .039 + 1.776 + 66.66$ Answer _____
14. Add: $27.005 + 1.375 + 10.6 + 1.396$ Answer _____
15. Subtract: 1.567 from 3.01 Answer _____
16. Subtract: .036 from 1.0066 Answer _____
17. Multiply: $9.66 \times .75$ Answer _____
18. Multiply: $.035 \times 6.69$ Answer _____
19. Divide: $10.5 \div 5.35$ Answer _____
20. Divide: $.776 \div 1.359$ Answer _____
21. Given $C = A/B$ Solve for B Answer _____
22. Given $9 = 12/x$ Solve for x Answer _____
23. Write the following in Arabic Numbers:

MMCDXLVI	+=	_____
XXXVIII	=	_____
MCCLIX	=	_____
DCCXXXII	=	_____
MCXXL	=	_____
24. Write the following in Roman Numerals:

765	=	_____	999	=	_____
3655	=	_____	36	=	_____
75	=	_____			

1. Add: $5/14 + 6/7 + 1/8 + 3/28$ Answer _____
2. Add: $3/8 + 5/16 + 3\ 4/7 + 1/6$ Answer _____
3. Subtract: $1/3$ from $6/13$ Answer _____
4. Subtract: $5/8$ from $1\ 6/7$ Answer _____
5. Multiply: $3/7 \times 1\ 3/5$ Answer _____
6. Multiply: $1\ 7/8 \times 3/5 \times 9/12$ Answer _____
7. Divide: $6/7 \div 9/21$ Answer _____
8. Divide: $1/8 \div 3\ 5/16$ Answer _____
9. Convert: $7/8$ to a decimal fraction Answer _____
10. Convert: $12/13$ to a decimal fraction Answer _____
11. Convert: $.325$ to a simple fraction Answer _____
12. Convert: $.777$ to a simple fraction Answer _____
13. Add: $5.037 + 1.798 + 555 + 10.003$ Answer _____
14. Add: $20.1 + 15.09 + 9.667 + 1.0037$ Answer _____
15. Subtract: 10.77 from 11.035 Answer _____
16. Subtract: $.097$ from 1.01 Answer _____
17. Multiply: 9.73×10.11 Answer _____
18. Multiply: 1.07×6.735 Answer _____
19. Divide: $1.395 \div 16.711$ Answer _____
20. Divide: $19.01 \div 16.335$ Answer _____
21. Given $Z/P = W$ Solve for Z Answer _____
22. Given $15/T = 21$ Solve for T Answer _____
23. Write the following in Arabic Numbers:

MCMXXVII = _____	CDXIV = _____
XLVI = _____	LV = _____
MMCXI = _____	
24. Write the following in Roman Numerals:

125 = _____	1965 = _____
37 = _____	91 = _____
1333 = _____	

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HANDOUT 3ABR90530-1-2-1
April 1976

PHARMACEUTICAL CALCULATIONS - I

1. Change the following to milligrams:
 - a. 39.1 Gm
 - b. 125 Hg
 - c. .035 Mcg
 - d. .01 cg
 - e. .075 Gm
2. Change the following to liters
 - a. 35 Kl
 - b. 1.07 Dl
 - c. .03 dl
 - d. 19.77 mcl
 - e. .03 cl
3. Add: 1.25 dl + 12L = 13 Kl + 125 mcl + 25 cl - express in ml.
4. Add: 25 M + 1.07 mm + 120 cm + .005 Hm - express in dm.
5. Subtract: 1.25 dg from .01 Hg - express in Gm.
6. Subtract: .035 dl from 250 L - express in L.
7. Multiply: 25 ml x 50
8. Multiply: 115 x .01 cg
9. Restate to a lower denomination in the apothecary system.
 - a. Reduce 10, 3 fl~~3~~, 6 fl~~3~~ to mx
 - b. Reduce 17~~5~~, 10~~3~~ to ~~3~~
 - c. Reduce 4c, 2 qt, 10 to fl~~3~~
10. Restate to a higher denomination in the apothecary system.
 - a. 3985 gr
 - b. 10,125 gr
 - c. 61,955 mx

Answer _____

Answer _____

This supersedes WS 3ABR90530-1-4

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DO NOT USE ON THE JOB

11. Add: 1c, 3 qt, 2 fl~~3~~, 40 mx + 2c, 3 qt, 6 fl~~3~~, 20 mx
12. Add: 3c, 1 qt, 10, 12 fl~~3~~, 1 c, 3 qt, 10, 4 fl~~3~~, 6 fl~~3~~
13. Subtract: 3c, 1 qt, 10, 14 fl~~3~~, 6 fl~~3~~, - 3 qt, 2 fl~~3~~, 7 fl~~3~~, 20 mx
14. Subtract: 1 c, 2 qt, 10, 7 fl~~3~~, 6 fl~~3~~ - 3 qt, 10, 8 fl~~3~~, 2 fl~~3~~
15. Restate to a lower denomination in the Avoirdupois System.
- 2 lb, 5 oz, 400 gr to gr
 - 1 lb, 12 oz, 125 gr to gr
 - 6 lb, 218.75 gr to oz
16. Restate to a higher denomination in the Avoirdupois System.
- 6735 gr
 - 8927 gr
 - 1225 gr
17. 6 lb of tomatoes cost \$1.25. How many pounds can you buy for \$6.00?
18. 3 fl~~3~~ of a preparation contains 6 Gm of active ingredient. How much would be needed to prepare 14 fl~~3~~ of the preparation?
19. A pharmacist has 3 ~~lb~~ of medication. How many 3 gr tablets can be prepared from the total?
20. You have \$1.50. How many Gm's can you buy if 6 Gm cost \$10.00?

1. Change the following to milliliters:

45.6 L
 625 Hl
 9.01 mcl
 .005 cl
 1,025.6 L

2. Change the following to meters:

27 Km
 16.3 Dm
 .012 dm
 21.035 mcm
 5,033.635 cm

3. Add: 15.03 dl + 1.03 L + 1.077 Kl + 1.0011 mcl + 303 Hl
 (express in ml).

4. Add: 6.6 M + 103.6 mm + .967 cm + .005 Hm (express in dm).

5. Subtract: 9.999 dg from 10 Hg (express in grams).

6. Subtract: 103,596 ml from .9 Kl (express in L).

7. Multiply: 37.5 x 66 ml Answer L

8. Multiply: 113.6 x 1.01 cg Answer Cm

9. Restate to a lower denomination in the Apothecary System:

- a. Reduce 1c, 10, 5 fl $\frac{3}{4}$, 6 fl $\frac{3}{4}$ to mx
- b. Reduce 2 $\frac{1}{16}$ 9 $\frac{3}{4}$, to $\frac{3}{4}$
- c. Reduce 3 qt, 10, 6 fl $\frac{3}{4}$ to fl $\frac{3}{4}$

10. Restate to a higher denomination in the Apothecary System.

- a. 2375 gr
- b. 16,125 gr
- c. 60,655 mx

11. Add: 2c, 2 qt, 6 fl $\frac{3}{4}$, 25 mx + 1 c, 1 qt, 8 fl $\frac{3}{4}$, 20 mx

12. Add: 1 $\frac{1}{16}$, 10 $\frac{3}{4}$, 5 $\frac{3}{4}$, 1 $\frac{1}{2}$, 10 gr + 3 $\frac{1}{16}$, 1 $\frac{3}{4}$, 5 $\frac{3}{4}$, 2 $\frac{1}{2}$ 10 gr.

13. Subtract: 4 c, 2 qt, 14 fl $\frac{3}{4}$, 5 fl $\frac{3}{4}$ - 3 qt, 1 fl $\frac{3}{4}$, 5 fl $\frac{3}{4}$, 20 mx.

14. Subtract: 2 $\frac{1}{16}$, 11 $\frac{3}{4}$, 6 $\frac{3}{4}$, 2 $\frac{1}{2}$, 10 gr - 1 $\frac{1}{16}$, 7 $\frac{3}{4}$, 1 $\frac{1}{2}$, 15 gr.

15. Restate to a lower denomination in the Avoirdupois System.
- a. 3 lb, 10 oz, 250 gr to gr
 - b. 1 lb, 13 oz, 425 gr to gr
 - c. 4 lb, 218.75 gr to oz
16. Restate to a higher denomination in the Avoirdupois System.
- a. 5355 gr
 - b. 8525 gr
 - c. 14,437.5 gr
17. 1 lb of oranges cost \$2.50. How many pounds can you buy for \$1.75?
18. 5 fl ~~3~~ of a preparation contains 3 gm of active ingredient. How many gm of active ingredient would be needed to prepare 12 fl of the preparation?
19. A pharmacist has 4 ~~lb~~ of medication. How many 2 gr tablets can be prepared from the total?
20. You have \$2.25. How many Gm can you buy if 6 Gm cost \$25.00?

1. Change the following to millimeters:

65.5 M
427 Hm
10.13 mcm
.015 cm
4,326 .1 M

2. Change the following to grams:

13 Kg
21.05 Dg
.137 dg
36.755 mcg
7,036.111 cg

3. Add: 11.6 dm + .137 M + 12.66 Km + 125.1 mcm + 325 M (express in mm).

4. Add: 7.7 L + 125.6 ml + .037 cl + .0777 Kl (Express in dl).

5. Subtract: 625 ml from 1 L (express in liters).

6. Subtract: 253.6 L from 33.36 Hl (express in ml).

7. Multiply: 33.1 cl x 426 (express in L).

8. Multiply: 6.113 x 25 mg (express in Gm).

9. Restate to a lower denomination in the Apothecary System.

a. Reduce 1 c, 2 qt, 1 O, 12 fl , 6 fl to mx.
b. Reduce 1 ^{lb}/₁₆ , 6 ^z/₃ , 6 ^z/₃ , 1 ^z/₃ , 11 gr to gr
c. Reduce 1 qt, 10 ^z/₃ , 11 fl ^z/₃ , to fl ^z/₃

10. Restate to a higher denomination in the Apothecary System.

a. 5476 gr
b. 8995 gr
c. 10125 mx

11. Add: 3c, 3 qt, 10 fl ^z/₃ , 3 fl ^z/₃ , 25 mx + 1 qt, 6 fl ^z/₃ , 5 fl ^z/₃ , 35 mx.

12. Add: 2 ^{lb}/₁₆ , 11 ^z/₃ , 5 ^z/₃ , 2 ^z/₃ , 10 gr + 2 ^{lb}/₁₆ , 1 ^z/₃ , 3 ^z/₃ , 1 ^z/₃ , 10 gr.

13. Subtract: 2c, 10, 5 fl ^z/₃ , 1 fl ^z/₃ , 3 mx - 1 c, 2 qt, 7 fl ^z/₃ , 2 fl ^z/₃ .

14. Subtract: 4 ^{lb}/₁₆ , 5 ^z/₃ , 2 ^z/₃ , 1 ^z/₃ , 11 gr - 1 ^{lb}/₁₆ , 11 ^z/₃ , 7 ^z/₃ ,
2 ^z/₃ , 10 gr.

15. Restate to a lower denomination in the Avoirdupois System.

a. 2 lb, 12 oz, 235 gr to gr
b. 1 lb, 10 oz, 10 gr to gr
c. 1 lb, 218.75 gr to oz

16. Restate to a higher denomination in the Avoirdupois System:
- a. 6125 gr
 - b. 10666 gr
 - c. 13,125 gr
17. 5 lb of peaches cost \$10.25. How many pounds can you buy for \$3.25?
18. 7 fl ³/₄ of a prescription contains 3 Gm of active ingredient. How many Gm of active ingredient would be needed to prepare 16 fl ³/₄ of the preparation?
19. A pharmacist has 5 lb of medication. How many 4 gr tablets can be prepared from the total?
20. You have \$2.25. How many meters can you buy if 3 meters cost \$15.35?

PHARMACEUTICAL CALCULATIONS I

1. Convert 6 fl $\bar{3}$ to ml.
2. Convert 15 Gm to gr.
3. Convert 250 ml to fl $\bar{3}$.
4. Convert 1 fl $\bar{3}$ 20 mx to ml.
5. If a mixture weighing 30 Gm is divided into 100 dosage forms, how many grains will each dose weigh?
6. Convert each of the following to the Metric System: ml or mg
 - a. 1/60 gr
 - b. 2 fl
 - c. 3/8 gr
 - d. 30 mx
 - e. 1/200 gr
7. Convert each of the following to the Apothecaries unit: $\bar{3}$ or fl $\bar{3}$
 - a. 150 ml
 - b. 0.3 ml
 - c. .001 Gm
 - d. 065. mg
8. A certain drug is available in 15, 25 and 30 mg tablets. Express these amounts in Apothecaries system. (gr)
9. Convert 50 micrograms to grains.
10. If 2 fl $\bar{3}$ of a solution contain 7 1/2 gr of a chemical, how many grams would be contained in 125 ml of solution?
11. If a chemical costs \$3.50 a pound (AV) what is the cost of 15 Gm?
12. How many 6.5 mg tablets can be obtained from $\bar{3}$ ss of a chemical?

This supersedes WS 3ABR90530-I-7

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13. A prescription calls for $\frac{4}{5}$ gr of Atropine Sulfate to be divided into 80 doses. How many milligrams will each dose weigh?
14. In the compounding of a prescription a pharmacist used $\frac{1}{4}$ gr of Atropine Sulfate. How many 0.000325 Gm doses were prescribed on the prescription?
15. A certain elixir contains 0.325 Gm of Potassium Thiocyanate per fl ~~3~~. At \$1.75 per pound, what is the cost of the Potassium Thiocyanate required to make 1 gallon of the elixir?
16. A formula for a cough syrup calls for $\frac{1}{8}$ gr of Codeine Phosphate per fl ~~3~~. How many Gm of Codeine Phosphate should be used in preparing one pint of the cough syrup?
17. A prescription calls for 2 grains of Ephedrine Bitartrate. If 1 Gm of Ephedrine Bitartrate cost \$2.00, what is the cost of the amount needed in the prescription?
18. Convert 2 c, 3 qt, 1 θ to ml.
19. Convert 1 lb, 3 oz to mg.
20. Convert 15 Gm to oz.

1. Convert 8 fl $\overline{3}$ to ml.
2. Convert 21 Gm to gr.
3. Convert 350 ml to fl $\overline{3}$.
4. Convert 5 fl $\overline{3}$ 15 mx to ml.
5. If a mixture weighing 15 Gm is divided into 50 dosage forms, how many grains will each weigh?
6. Convert each of the following to the Metric System: ml or mg
 - a. 1/15 gr
 - b. 5 fl $\overline{3}$
 - c. 6/7 gr
 - d. 15 mx
 - e. 125 $\overline{3}$
7. Convert each of the following to the Apothecaries units: $\overline{3}$ or f $\overline{3}$
 - a. 125 ml
 - b. 015. ml
 - c. 01.5 Gm
 - d. 135. mg
8. A certain drug is available in 20, 30, and 40 mg tablets. Express these amounts in Apothecaries system. (gr)
9. Convert 25 micrograms to grains.
10. If 3 fl $\overline{3}$ of a solution contains 10 gr of a chemical, how many grams would be contained in 125 ml of solution?
11. If a chemical cost \$2.75 a pound (AV), what is the cost of 7 Gm?
12. How many 6.5 mg tablets can be obtained from $\overline{3}$ ss of a chemical?
13. A prescription calls for 1 1/4 gr of atrophine sulfate to be divided into 80 doses. How many milligrams will each dose weigh?
14. In the compounding of a prescription a pharmacist used 1/2 gr of atropine sulfate. How many 0.000650 Gm doses were prescribed on the prescription?
15. A certain elixir contains 0.125 Gm of potassium thiocyanate per fl $\overline{3}$. At \$1.75 per pound, what is the cost of the potassium thiocyanate required to make 1 gallon of the elixir?
16. A formula for a cough syrup calls for 1/4 gr of codeine phosphate per fl $\overline{3}$. How many Gm of codeine phosphate should be used in preparing 1 qt of the cough syrup?

17. A prescription calls for 4 grains of ephedrine bitartrate. If 2 Gm of ephedrine bitartrate cost \$4.00, what is the cost of the amount needed in the prescription?
18. Convert 3 c, 1 qt, 1 0 to ml.
19. Convert 5 lb, 11 oz to mg.
20. Convert 350 Gm to oz.

1. Convert 11 fl $\frac{3}{4}$ to ml.
2. Convert 35 Gm to gr.
3. Convert 275 ml to fl $\frac{3}{4}$.
4. Convert 5 fl $\frac{3}{4}$ 20 mx to ml.
5. If a mixture weighing 10 Gm is divided into 25 dosage forms, how many grains will each dose weigh?
6. Convert each of the following to the Metric System. ml or mg
 - a. $\frac{1}{20}$ gr
 - b. 7 fl
 - c. $\frac{5}{7}$ gr
 - d. 26 mx
 - e. 35
7. Convert each of the following to the Apothecaries Units. $\frac{3}{4}$ or $\frac{1}{2}$
 - a. 75 ml
 - b. 0.25 ml
 - c. .025 Gm
 - d. 205. mg
8. A certain drug is available in 5, 10 & 15 mg tablets. Express these amounts in apothecaries system. (gr)
9. Convert 10 micrograms to grains.
10. If 5 fl $\frac{3}{4}$ of a solution contains 9 gr of a chemical, how many grams would be contained in 100 ml of solution?
11. If a chemical cost \$1.50 a pound (AV), what is the cost of 3.5 Gm?
12. How many 13 mg tablets can be obtained from $\frac{3}{4}$ li of a chemical?
13. A prescription calls for $\frac{3}{5}$ gr of atropine sulfate to be divided into 60 doses. How many milligrams will each dose weigh?
14. In the compounding of a prescription a pharmacist used $\frac{1}{8}$ gr of atropine sulfate. How many 0.000325 Gm doses were prescribed on the prescription?
15. A certain elixir contains 0.275 Gm of potassium thiocyanate per fl $\frac{3}{4}$ i. At \$1.75 per pound, what is the cost of the potassium thiocyanate required to make 1 gallon of the elixir?
16. A formula for a cough syrup calls for $\frac{1}{2}$ gr of codeine phosphate per fl $\frac{3}{4}$ i. How many Gm of codeine phosphate should be used in preparing 1 gallon of the cough syrup?

17. A prescription calls for 1 grain of ephedrine bitartrate. If $\frac{1}{2}$ Gm of ephedrine bitartrate costs \$1.50 what is the cost of the ephedrine bitartrate needed in the prescription?
18. Convert 1 c, 3 qt, 1 O , 5 fl z to ml.
19. Convert 6 lb, 6 oz to mg.
20. Convert 225 Gm to oz.

1. A certain elixir contains 0.325 Gm of Potassium Thiocyanate per 1 fl $\frac{3}{4}$. At \$1.75 per pound (Avoir), what is the cost of the potassium thio-cyanate required to make 1 gallon of the elixir?
2. A formula for a cough syrup calls for 1/8 gr of codeine phosphate per 1 fl $\frac{3}{4}$. How many grams of codeine phosphate should be used in preparing one pint of the cough syrup?
3. If the cost of 10 Gm of merbromin is \$1.25, what is the cost of 4 1/2 gr?
4. Convert 3 gal, 1 pt, 10 fl $\frac{3}{4}$ to milliliters.
5. If a preparation contains 5 Gm of a drug in 500 ml, how many Gm are contained in each tablespoonful dose?
6. How many grams of a chemical are required to make 120 ml of a solution, each teaspoonful of which will contain 3 mg of the chemical?
7. How many 15 minimum doses are contained in 60 ml of a tincture?
8. If the dose of a drug is 1/16 gr, how many doses are contained in 1 fl $\frac{3}{4}$?
9. If a medicine is to be taken three times daily, and if 180 ml are to be taken in four days, how many tablespoonfuls should be prescribed for each dose?
10. How many teaspoonfuls per dose must be taken if 8 fl $\frac{3}{4}$ of a medicine are to be taken three times a day for eight days?
11. What is the dosage in teaspoonfuls if 240 ml of a medicine contain 48 doses?
12. If a prescription contains 0.24 Gm of codeine phosphate in 120 ml, how much is contained in each teaspoonful dose?
13. How many grains of a chemical are contained in each capsule if a mixture containing 1 1/4 gr of the chemical is divided into 30 capsules?
14. If 180 ml of a cough mixture contain 3/4 gr of Dilaudid, how much is contained in 1 teaspoonful of the mixture?
15. How many grams of a chemical are required to make 120 ml of a mixture each teaspoonful of which is to contain 2.5 mg?
16. Rx Codeine Phosphate 0.24 Gm
Sodium Citrate 4.0 Gm
Chloroform 0.5 ml
Tolu Syrup Qsad 120.0 ml

Sig: Fl $\frac{3}{4}$ i

How many mg of codeine phosphate and how much chloroform are contained in each dose?

17. Rx Phenobarbital 0.6 mg
Belladonna Tinc 12.0 ml
Peppermint Water
Qs ad 120.0 ml

Sig: 1 tsp T.I.D.

How much phenobarbital & how much Belladonna Tincture will be contained in each dose?

18. A powder is divided into 36 capsules. If each capsule contains 0.5 mg of one ingredient, 15 mg of a second, and enough of a third to make 0.300 Gm, how much of each was there in the original powder?
19. A solution contains 30 mg of a chemical per 120 ml and has a dose of 10 drops. If the dispensing dropper calibrates 25 drops per ml, how many mcg of the chemical are contained in each dose?
20. The dose of a drug is 5 mg per kilogram of body weight. How many grams should be given to a child weighing 55 lb?
21. The usual rectal dose of tribromoethanol is 0.06 ml for each kilogram of body weight. How many milliliters should be given to a person weighing 150 lb?
22. If the usual adult dose of a drug is 324 mg, what is the dose for a child 6 years old?
23. If the usual adult dose of a drug is 5 ml, what is the dose in fl *z* of a child 4 years old?
24. If the usual adult dose is 6 fl *z*, what is the dose, in milliliters, for a child weighing 75 lbs?
25. The usual adult dose of Benadryl Elixir is 2 tablespoonfuls, what is the dose in milliliters for a child weighing 120 lbs?

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HANDOUT 3ABR90530-I-2-2
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PHARMACEUTICAL CALCULATIONS I

1. A certain elixir contains 0.225 Gm of Potassium Thiocyanate per fl³i. At \$1.75 per pound (AV), what is the cost of the Potassium Thiocyanate required to make 1 gallon of the elixir?
2. A formula for cough syrup calls for 1/4 gr of Codeine Phosphate per fl³ii. How many Gm of Codeine Phosphate should be used in preparing one pint of the cough syrup?
3. If the cost of 15 Gm of Merbromin is \$1.50, what is the cost of 5 gr?
4. Convert 1 gal, 1 qt, 5 fl³ to ml.
5. If a preparation contained 3 Gm of a drug in 250 ml, how many Gm are contained in each tablespoonful dose?
6. How many Gm of a chemical are required to make 150 ml of a solution, each table-
spoonful of which will contain 5 mg of the chemical?
7. How many 10 minimum doses are contained in 30 ml of a tincture?
8. If the dose of a drug is 1/20 gr, how many doses are contained in fl³i?
9. If a medicine is to be taken three times daily and if 250 ml are to be taken in four
days, how many tablespoonfuls should be prescribed for each dose?
10. How many teaspoonfuls per dose must be taken if fl³v of a medicine are to be taken
three times a day for 3 days?
11. What is the dosage in tablespoonfuls, if 300 ml of a medicine contain 40 doses?
12. If a prescription contains 0.45 Gm of Codeine Phosphate in 150 ml, how much is con-
tained in each teaspoonful dose?
13. How many grains of a chemical are contained in each capsule if a mixture containing
3/4 gr of the chemical is divided into 30 capsules?
14. If 150 ml of a cough mixture contain 3/4 gr of Dilaudid, how much is contained in
1 teaspoonful of the mixture?
15. How many grams of a chemical are required to make 150 ml of a mixture, each teaspoonful
of which is to contain 5 mg?
16. Rx Codeine Phosphate 0.36 Gm
 Sodium Citrate 5.0 Gm
 Chloroform 0.75 ml
 Tolu Syrup QSAD 125.0 ml

Sig: fl³ii

How many mg of Codeine Phosphate and how much Chloroform are contained in each dose?

This supersedes WS 3ABR90530-I-11

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17. Rx Phenobarbital 0.5 mg
 Belladonna Tinc 15.0 ml
 Peppermint Water
 QSAD 125.0 ml

Sig: Fl̄ 11 T.I.D.

How much Phenobarbital and how much Belladonna Tincture will be contained in each dose?

18. A powder is divided into 25 capsules. If each capsule contain 0.6 mg of one ingredient, 17 mg of a second, and enough of a third to make 0.50 Gm, how much of each was there in the original powder?

19. A solution contains 25 mg of a chemical per 150 ml and has a dose of 5 drops. If the dispensing dropper calibrates 30 drops per ml, how many mcg of the chemical are contained in each dose?

20. The dose of a drug is 7 mg per kilogram of body weight. How many grams should be given to a child weighing 75 lbs?

21. The usual rectal dose of Tribromoethanol is 0.15 ml for each kilogram of body weight. How many milliliters should be given to a person weighing 125 lbs?

22. If the usual adult dose of a drug is 7.5 ml, what is the dose in fl̄ of a child 6 years old?

23. If the usual adult dose of a drug is 250 mg, what is the dose for a child 8 years old?

24. If the usual adult dose is 8 fl̄, what is the dose in milliliters, for a child weighing 80 lbs?

25. The usual dose of Benedryl elixir is 2 teaspoonfuls, what is the dose in milliliters for a child weighing 100 lbs?

17. Rx Phenobarbital 0.75 mg
 Belladonna Tinc 12.5 ml
 Peppermint Water
 QSAD 150.0 ml

Sig: Fl $\frac{3}{4}$ ss T.I. D.
 How much phenobarbital and how much belladonna tincture will be contained in each dose?

18. A powder is divided into 30 capsules. If each capsule contains 0.75 mg of one ingredient, 18 mg of a second, and enough of a third to make 0.70 Gm, how much of each was there in the original powder?
19. A solution contains 35 mg of a chemical per 175 ml and has a dose of 7 drops. If the dispensing dropper calibrates 25 drops per ml, how many mcg of the chemical are contained in each dose?
20. The dose of a drug is 10 mg per kilogram of body weight. How many grams should be given to a child weighing 80 lb?
21. The usual rectal dose of tribromoethanol is 0.05 ml for each kilogram of body weight. How many milliliters should be given to a person weighing 115 lb?
22. If the usual adult dose of a drug is 300 mg, what is the dose for a child 10 years old?
23. If the usual adult dose of a drug is 10 ml, what is the dose in fl $\frac{3}{4}$ of a child 8 years old?
24. If the usual adult dose is 10 fl $\frac{3}{4}$, what is the dose in milliliters for a child weighing 50 lb?
25. The usual adult dose of Benedryl elixir is 1 tablespoonful, what is the dose in milliliters for a child weighing 95 lb?

1. A certain elixir contains 0.125 Gm of Potassium Thiocyanate per fl ~~3~~ ss. At \$1.50 per pound (AV), what is the cost of the potassium thiocyanate required to make 2 gallons of the elixir?
2. A formula for a cough syrup calls for 1/2 gr of codeine phosphate per fl ~~3~~ i. How many Gm of codeine phosphate should be used in preparing one gallon of the cough syrup?
3. If the cost of 12.5 Gm of merbromin is \$.75, what is the cost of 100 gr?
4. Convert 3 gal, 3 qt, 10 fl ~~3~~ to ml.
5. If a preparation contains 2.5 Gm of a drug in 100 ml, how many Gm are contained in each tablespoonful dose?
6. How many gr of a chemical are required to make 250 ml of a solution, each teaspoonful of which will contain 4 mg of the chemical?
7. How many 5 minimum doses are contained in 15 ml of a tincture?
8. If the dose of a drug is 1/10 gr, how many doses are contained in ~~3~~ 1?
9. If a medicine is to be taken twice daily, and if 100 ml are to be taken in three days, how many teaspoonfuls should be prescribed for each dose?
10. How many teaspoonfuls per dose must be taken if f ~~3~~ vii of medicine are to be taken three times a day for 6 days?
11. What is the dosage in tablespoonfuls, if 150 ml of a medicine contains 25 doses?
12. If a prescription contains 0.30 Gm of codeine phosphate in 100 ml, how much is contained in each teaspoonful dose?
13. How many grains of a chemical are contained in each capsule if a mixture containing 2 gr of the chemical is divided into 20 capsules?
14. If 250 ml of a cough mixture contain 1 gr of Dilaudid, how much is contained in 2 teaspoonfuls of the mixture?
15. How many grams of a chemical are required to make 200 ml of a mixture, each tablespoonful of which is to contain 12 mg?
16. Rx Codeine Phosphate 0.30 Gm
 Sodium Citrate 7.50 Gm
 Chloroform 0.95 ml
 Tolu Syrup QSAD 150.0 ml

Sig: Fl ~~3~~ ss

How many mg of codeine phosphate and how much chloroform are contained in each dose?

DEPARTMENT OF BIOMEDICAL SCIENCES

PHARMACY SPECIALIST

BLOCK I

10-8

FUNDAMENTALS OF PHARMACY

March 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
SHEPPARD AIR FORCE BASE, TEXAS

Designed For ATC Course Use

7

DO NOT USE ON THE JOB

FUNDAMENTALS OF PHARMACY

OBJECTIVE

Solve problems pertaining to basic mathematical operations.

INTRODUCTION

Pharmacy, to many of you might mean the corner drug store where you were able to buy just about anything from cosmetics, to magazines, to pop-up toasters. It normally had a high counter, somewhere in the rear of the store, which concealed everything except the head and white coated shoulders of the pharmacist. There were always two or three clerks and the pharmacist who seemed to be concocting some mysterious formula. The Air Force Pharmacy is unlike any pharmacy most of you might be acquainted with. There are no clerks, cosmetics, school supplies; just trained pharmacy personnel, medications and chemicals. You have been selected to serve as a member in one of the most important fields within the medical service. Each year, millions of prescriptions are filled by Air Force pharmacy specialists and technicians around the world. Without the pharmacy service personnel, the overall care of the patients would not be fulfilled. In turn, the mission of the Air Force would be affected. The purpose of this course is to prepare you to effectively meet the demands of modern pharmacy in your role as a Pharmacy Specialist. Today's pharmacy is changing and the emphasis on compounding pills and certain dosage forms, and mysterious formulations is being replaced by subjects such as drug distribution systems, unit doses, strip packaging, drug information centers, drug interactions, I.V. admixtures and drug stability. Today we are increasingly concerned with the patient receiving quality medications. It is not sufficient to accurately weigh some ingredients on a prescription balance and incorporate them into a weighed quantity of ointment base or dissolve them in a flavored vehicle. Hospital pharmacies, large and small, that engage in bulk compounding and prepackaging cannot insure quality products unless they have adequately trained personnel.

INSTRUCTIONS

The instructor supervisor will welcome the class, introduce them to the pharmacy career field and discuss the purpose and policies of the course, including class schedule, graduation date and administrative requirements of the School of Health Care Sciences. He will brief the students on policies of both school and squadron. He will inform them of their responsibilities as students and airmen concerning such items as uniforms, appearance, conduct and behavior. He will issue text books, course materials and select a class leader. The NCOIC will explain course content, examination, critiques, remedial training, counseling, elimination, assignments and course completion.

This supersedes SW 3ABR90530-I-1, dated December 1973

The course supervisor will discuss honor graduates, training objectives, physical facilities and functions and introduce the instructor staff.

INFORMATION

During the many years the pharmacy course has been in operation, certain administrative and scholastic policies (which affect your academic progress, need, activities, and interests) have been established. These policies are briefly outlined below to provide you with a ready reference of your individual inquiries. Also outlined in this section is a brief description of the subject areas covered in this course. The Organization Structure, Duties and Responsibilities and Pharmacy Career Ladder will also be discussed.

1F-75-1104

ADMINISTRATIVE REQUIREMENTS

Records provide essential information for administering and directing student affairs in the section. Consequently, requested information should be completed and supplied as quickly as possible.

Leave will not normally be granted; however, emergency leave will be processed in the usual manner through your squadron.

Absences from class must be coordinated with the flight or class leader and instructor and approved by the course supervisor.

Sick call will be on an appointment basis through the base dispensary. Appointments will be made by the student squadron orderly room before class in the morning or by the pharmacy training section during school hours.

Telephone calls will be limited to those of emergency nature when originating in the pharmacy training section. Incoming emergency calls will be handled as expeditiously as possible for the benefit of the person concerned. Pay telephones are provided on the first floor of the building for routine calls.

Flight or class leader is selected for each class. The class leader acts as spokesman for the class. He also promotes group spirit and effort. Performance of certain routine duties by the students are essential for efficient classroom administration and housekeeping. The class leader will supervise the details for accomplishing these tasks. A complete briefing on details and housekeeping responsibilities will be given by the staff.

TRAINING MATERIALS

Materials issued for your class use are study guides and workbooks (SW), and pertinent Air Force Publications as well as the following technical references:

United States Pharmacopoeia
 National Formulary
 Remington's Pharmaceutical Sciences
 Physicians' Desk Reference
 Rogers' Inorganic Pharmaceutical Chemistry
 Cutting's Handbook of Pharmacology

Study Guides and Workbooks provide, or refer to, general material which supports a unit of instruction. They also contain material concerned with specialized procedures, work to be completed, problems to be solved, and questions to be answered. You must study the applicable guide and references to be prepared for the instruction and discussion which takes place during a specific instructional period.

Air Force Publications and Technical References issued to you are those which support the various units of instruction. The references found in each SW identify the reference applicable to the unit of instruction.

SCHOLASTIC CONDITIONS AFFECTING STUDENT PROGRESS

Grades. You will be graded in both raw and percentage grades. Each block has a minimum passing score. Anytime you receive a score below this fixed grade, you have done unsatisfactory work and will be in danger of being withdrawn from the course. The course supervisor will determine appropriate action in each case of unsatisfactory work accomplishment.

Honor Graduates. The top 10 percent of the class maintaining an academic average above 90 percent throughout the course.

Probationary Continuation is the means by which a student deficient in some aspect of his training may continue with his class in the expectation that the deficiency will be corrected.

Elimination may result from academic deficiency, that is, an average grade in a block of less than the minimum grade set for that block. Additionally, administrative elimination may be due to extended absences (such as leave or hospitalization) punitive actions, factors over which the student had control, (i.e., display of improper conduct, failure to accomplish work assignments, uncooperative attitude), and other controllable actions.

Faculty Board is a fact finding agency which considers students with training deficiencies that seem to warrant termination of training. The board exercises no punitive or administrative action; nevertheless, the board's proceedings may become a basis for such actions.

Remedial Training is conducted for all students with grades of less than 10 percent above the minimum passing grade for the block. Remedial training will be conducted per instructions of the course supervisor.

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Counseling is an integral part of instruction and is employed by the course supervisor or his representative to assist and guide the student. If you encounter problems, academic or otherwise, the course supervisor is available to lend whatever assistance and resources he possesses toward the solution of these problems.

ORGANIZATIONAL STRUCTURE

The Medical Service of the Air Force is made up of the Surgeon General, and the personnel in his office, all other qualified officers and airmen serving in medical career fields, civilian personnel employed to support the Medical Service Mission, and civilian consultants.

The primary mission of the Medical Service is to provide the medical support necessary to maintain the highest possible degree of combat readiness and effectiveness of the Air Force. This mission includes medical support for field operations and combat, medical and dental care, veterinary service, flight medicine, military public health, and occupational medicine for all personnel.

The Director of Base Medical Services has the responsibility of the medical mission at his base. He is the senior Air Force physician of the base medical unit. The Air Force hospital system is represented by installations ranging from small dispensaries manned by one or two doctors to large hospitals with professional staffs numbering the scores. All personnel required to carry out the mission of the base medical services will be under his professional supervision. (Page 10 will show the Hospital/Dispensary Organization Structure. The relationship of professional and support personnel to the Hospital Commander is shown on this chart.)

You are interested in the organization of the Pharmacy Service and its contribution to the medical mission. By referring to page 8 (Organization - Hospital Services) you will see that the Commander of the Base Medical Unit or Director of Base Medical Services has a physician who acts as a coordinator of the professional services in the hospital. This physician is known as the Chief of Hospital Services. The professional departments, such as surgery, nursing, radiology, laboratory, and pharmacy are his responsibilities. Problems associated with the management and administration of the pharmacy should be coordinated with the Chief of Hospital Services. We will examine the organization of the pharmacy services in more detail.

The hospital commander appoints a commissioned officer who is also a graduate, licensed pharmacist as Pharmacy Officer. If a commissioned graduate, licensed pharmacist of the Biomedical Sciences Corps is not available, A Civil Service Employee who is a graduate licensed pharmacist is appointed as Chief of the Pharmacy Service. If no Civil Service Pharmacist is available, the commander will designate a physician as "Pharmacy Officer."

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The Pharmacy Officer will be responsible for the overall mission and operation of the pharmacy and is designated as the consultant in pharmacy for the staff of the medical treatment facility. The Pharmacy Officer will have the assistance of a noncommissioned officer known as the NCOIC of Pharmacy. The NCOIC is responsible to the pharmacy officer for the management and administration functions including the supervision of assigned pharmacy specialists and technicians. Page 9 shows the relationship of the pharmacy service in the hospital.

The pharmacy in an Air Force Hospital has many functions. It provides the hospital staff with information on all drugs stocked in the pharmacy. It stores, manufactures, and dispenses pharmaceuticals to patients. The pharmacy maintains prescribed records of stored and dispensed pharmaceuticals. The pharmacy must comply with the federal regulations governing the storage and issue of specific drugs, narcotics, and poisons. One of the additional functions is to conduct on-the-job training of assigned duty personnel. To provide these services the pharmacy must have well trained personnel.

DUTIES AND RESPONSIBILITIES

PHARMACY SPECIALIST

Specialty Summary

Requisitions, stocks compounds, and dispenses medicinal preparations.

Duties And Responsibilities

COMPOUNDS AND DISPENSES MEDICINAL PREPARATIONS. Interprets prescriptions and formulas to determine content and therapeutic, chemical, and physical compatibility of ingredients. Confers with writer of prescription on any questions that arise to prevent possibility of error in desired therapeutic action. Calculates amounts of ingredients required for prescriptions or formulations. Weighs, measures, and combines drugs and chemicals according to accepted pharmaceutical methods. Prepares, packages or bottles, and labels prescriptions as ordered by physicians, dentists, or veterinarians. Manufactures, labels, and stores preparations according to official United States compendia and other reference literature. Issues medicaments to patients, wards, and clinics.

REQUISITIONS AND STORES PHARMACY SUPPLIES. Inventories drugs and equipment periodically. Ascertains supply requirements and prepares supply requisitions. Receives incoming pharmaceuticals in bulk, separates, and stores. Safeguards items such as chemicals, drugs, biological products, narcotics, and alcohol. Rotates stocks to insure freshness and potency.

PERFORMS GENERAL PHARMACY TASKS. Posts and maintains pharmacy records, including special files required in dispensing of narcotics and alcohol. Cleans and arranges pharmacy, equipment, and supplies.

SUPERVISES PHARMACY PERSONNEL. Assigns work and evaluates performance for attaining desired standards. Conducts on-the-job training in compounding, requisitioning, storing, and dispensing medicinal preparations.

Specialty Qualifications

KNOWLEDGE. Knowledge of principles of chemistry; pharmaceutical arithmetic; pharmacology; and medical ethics is mandatory. Possession of mandatory knowledge will be determined in accordance with AFM 35-1.

EDUCATION. High school courses in biology and chemistry are desirable.

EXPERIENCE. Experiences in functions such as compounding, storing, or dispensing preparations is mandatory.

TRAINING. Completion of a basic pharmacy course is desirable.

OTHER.

1. A minimum aptitude level of General 60 is mandatory.
2. This AFSC may be awarded to WAF airmen.

Speciality Data

1. Grade Spread:
Sergeant and staff sergeant - - - - - 90550
Airman first class - - - - - 90530
2. Related D.O.T. Jobs:
Pharmacy Clerk - - - - - 074.387
Pharmaceutical Detail Man - - - - - 266.158
3. Related DOD Occupational Subgroup - - 312

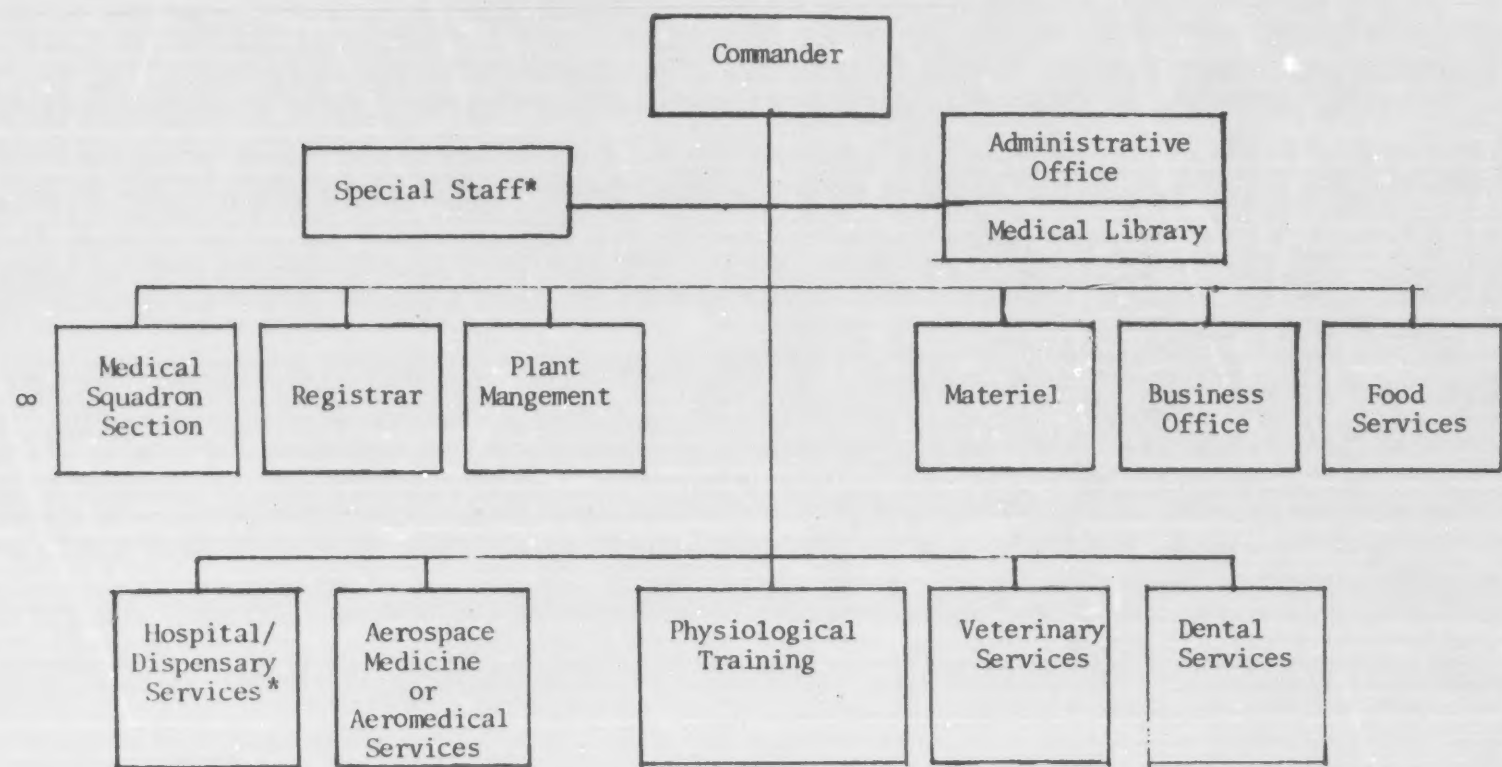
PHARMACY CAREER LADDER

Promotion is a prime area of concern for all of us. The pharmacy field, like most others, requires its personnel to advance to a specific level of proficiency and training in order to qualify for upgrading. Advancement in skill levels are accomplished either through resident schools or dual channel concept of OJT. Specific time requirements between levels must also be met in order for successful upgrading. The specific skill levels, titles, and corresponding ranks are listed below:

90010	Pharmacy Helper	- - - - -	Airman
90530	Pharmacy Apprentice	- - - - -	A1C
90550	Pharmacy Specialist	- - - - -	Sgt & SSgt
90570	Pharmacy Technician	- - - - -	TSgt & MSgt
90590	Pharmacy Superintendent	- - -	SMSgt & CMSgt

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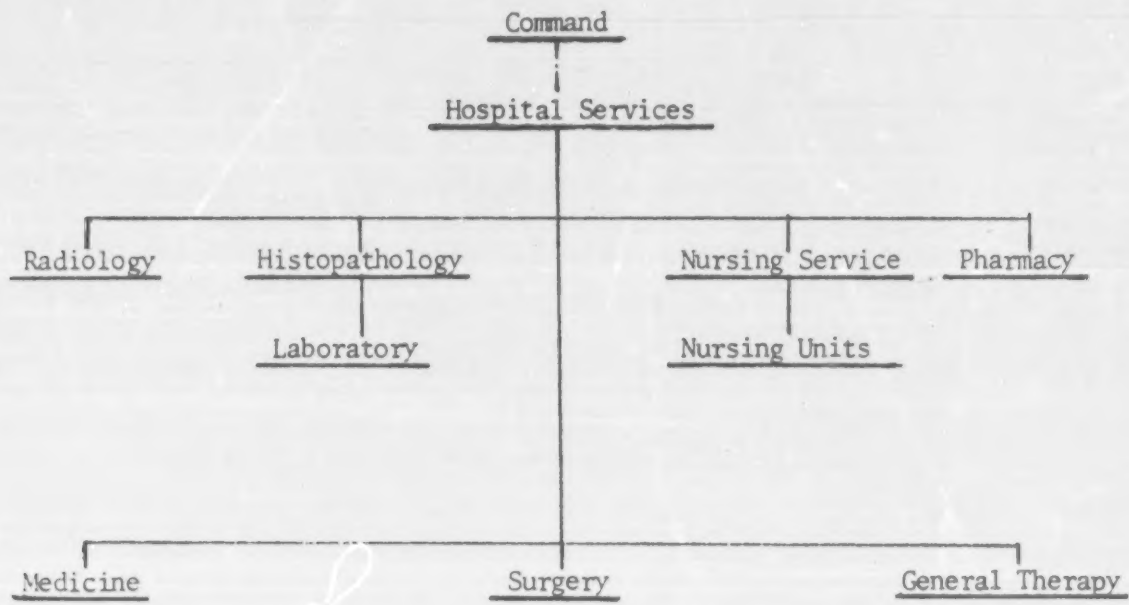
CHART 1
HOSPITAL/DISPENSARY ORGANIZATION STRUCTURE



* When authorized by HQ USAF for specialty centers.

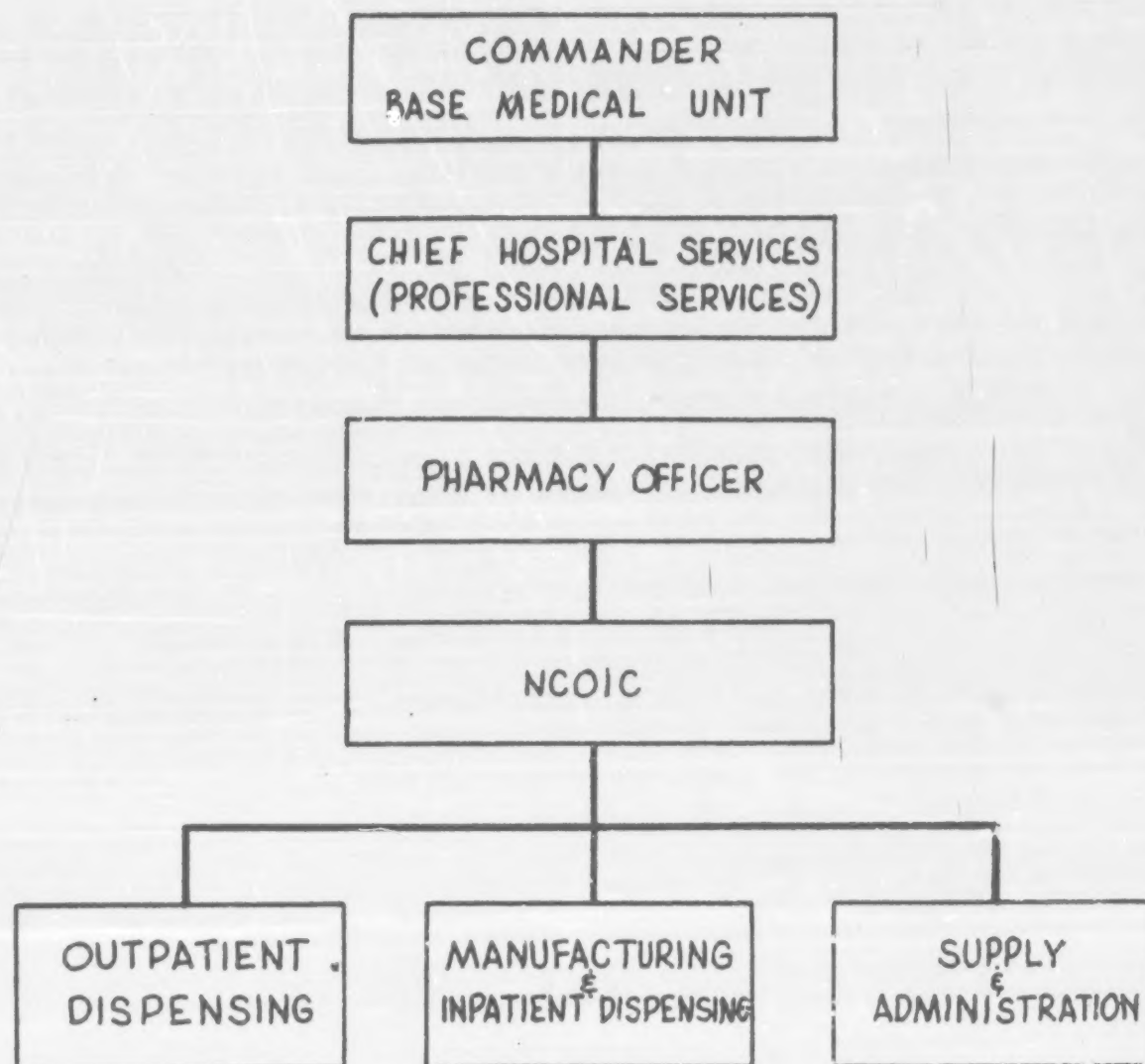
65

CHART 2
ORGANIZATION - HOSPITAL SERVICES



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2F-75-1104



BASIC MATHEMATICAL OPERATIONS

Although, we tend to think of fractions and decimal fractions as a very simple subject, it is the starting point of all math problems. Therefore a review will be advantageous to you as a foundation from which other, more complex, pharmaceutical problems may be worked.

Each type problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

Solve Problems Involving Simple Fractions

1. This is not a test. This is a learning situation. In this PT on fractions, you will be learning at your own speed.

2. Two types of programming are used in this PT.

a. LINEAR. In this portion, you will go from "frame" to "frame", using a piece of paper to cover upcoming frames and answers. In each frame, you are given information and then a question to answer or a problem to solve. Your answer can be checked to the left of the next frame. "Peeking" is not an advantage. If you make an error, strike out your incorrect answer, reread the frame, and write the correct answer.

b. SCRAMBLED. In this portion, you will be given problems to solve and asked to select the answer from a list of answers. Circle the answer you choose and go to the page as your answer directs. Follow directions closely. If you select an incorrect answer, do not erase, but put an "X" through the circle. Rework the problem again and circle another answer.

3. READ ALL INFORMATION CAREFULLY. BE SURE YOU UNDERSTAND WHAT IS SAID BEFORE YOU TRY TO ANSWER THE QUESTION. If you wish, you may turn back in the PT for review at any time.

	1. A fraction is a part of a whole. $\frac{3}{4}$ is a fraction and therefore is a part of a _____.
whole	2. Part of a whole is the definition of a _____.
fraction	3. The definition of a fraction is stated as: _____ of a _____.
part whole	4. Define a fraction. _____.
part of a whole	5. Fractions have two parts - - a numerator (above the line) and a denominator (below the line). Example: $\frac{3}{8}$ - - numerator - - denominator In the fraction $\frac{2}{3}$, the number 3 below the line is the _____ and the number 2 above the line is the _____.
denominator numerator	6. All fractions have denominators and numerators. In the fractions $\frac{2}{3}$ and $\frac{11}{12}$, the 3 and 12 are _____ and the 2 and 11 are _____.

A

Wrong! $12 \times 3 = 36$, but you must now do step 2. Add this product (36) to the numerator; retain the denominator to get the improper fraction. Go back to page 22, Frame 29, and select another answer.

B

Nope! You will still have to go to lower terms. You reduced by dividing two into the numerator and denominator but you must now find a number to further reduce $\frac{21}{27}$ and then you'll have it. Return to page 20A, select the other answer, and continue.

C

$\frac{2}{5}$ is the correct answer.

Now try another problem. $\frac{3}{8} \div \frac{2}{3} = \underline{\hspace{2cm}}$

If your answer is:

$$\frac{1}{4}$$

$$4 \text{ or } \frac{4}{1}$$

$$\frac{9}{16}$$

Go to page:

29C

37B

32B

denominators numerators	7. The denominator tells how many equal parts the whole has been divided into. In the fraction $\frac{9}{10}$, the denominator indicates the whole has been divided into _____ equal parts.
10	8. Under the figures below, write the number that would be used as the denominator of a fraction.     a. _____ b. _____ c. _____ d. _____
a. 4 b. 2 c. 3 d. 4	9. In the fraction below, circle the denominator and explain what it indicates. $\frac{15}{16}$ _____ _____
16 denominator Tells (indicates) how many parts the whole has been divided into.	10. The numerator (number above the line) of a fraction shows "how many parts of the whole are being considered." In the fraction $\frac{2}{3}$, the numerator indicates that _____ parts of the whole are being considered and the denominator indicates that the whole has divided into _____ equal parts.
2 3	11. In the fraction $\frac{13}{14}$, the number of parts being considered is _____ and the part of the fraction that tells us this is called the _____.

A

Wrong! Multiplication and addition are correct but you must place this sum over the denominator of the fraction. Return to page 21, Frame 29, and select another answer.

B

Right! Now try this. Reduce $\frac{14}{56}$ to its lowest term.

If your answer is:

Go to page:

$$\frac{1}{4}$$

178

$$\frac{7}{28}$$

198

C

No! You forgot to obtain the reciprocal of the divisor (invert the divisor), before you multiplied. Go back to page 38, frame 57; review the procedure again, then rework the problem from frame 59 again and select the correct answer.

13 numerator	12. The number of parts being considered is indicated by the _____ of a fraction.
numerator	13. Under the figures below, write the fractions. The number of parts being considered are shaded.     a. _____ b. _____ c. _____ d. _____
a. $\frac{1}{3}$ b. $\frac{3}{4}$ c. $\frac{2}{3}$ d. $\frac{1}{4}$	14. In the fraction below, write what each number is called and what it indicates: $\frac{6}{7}$ 6 -- _____ _____ 7 -- _____ _____
6 numerator. Indicates how many parts of the whole are being considered 7 denominator. Indicates how many equal parts the whole has been divided into.	15. There are three types of common fractions -- proper, improper, and mixed numbers. The three types of common fractions are mixed numbers, _____ and _____ fractions.
proper and improper	16. The difference between proper and improper fractions is the size of the numerator. The numerator of an improper fraction is <u>always the same as or larger</u> than the denominator; therefore, in a proper fraction, the numerator is _____ that the denominator.

A

Correct. Now change $15 \frac{1}{5}$ to an improper fraction.

If your answer is:

Go to page:

$$\frac{76}{5}$$

19A

$$\frac{75}{5}$$

21A

B

Good! You might have started with dividing by two (2) and doing several steps, but 14 divides into 14 and 56 evenly. To reduce an improper fraction such as $\frac{8}{4}$ or $\frac{9}{5}$, you simply divide the denominator into the numerator. Reduce $\frac{9}{5}$ to its lowest terms.

If your answer is:

Go to page:

$$\frac{9}{5}$$

21A

$$1 \frac{4}{5}$$

23B

smaller (less than)	17. $\frac{7}{8}$ is a proper fraction because the _____ is _____ than the denominator.
numerator smaller(less)	18. $\frac{8}{7}$ and $\frac{8}{8}$ are improper fractions because the _____ are _____ than the denominators.
numerators are same as or greater (are the same as or larger)	19. In the list below, place a "P" by the proper fractions and "I" by the improper fractions. a. $\frac{12}{17}$ c. $\frac{4}{5}$ b. $\frac{9}{7}$ d. $\frac{12}{12}$
a. P b. I c. P d. I	20. A mixed number is a whole number combined with a proper fraction. $3\frac{5}{6}$ is a whole number (3) and a proper fraction $\frac{(5)}{(6)}$; therefore, $3\frac{5}{6}$ is a _____.
mixed number	21. To review definitions, match the following types of fractions with the correct statement or statements by writing the letter of the statement by the number of the fraction. All letters are to be used. 1. Proper fraction A. Numerator greater than the denominator 2. Mixed number B. Numerator less than the denominator 3. Improper fraction C. Whole number and a proper fraction D. Numerator equal to denominator

A

Right! $\frac{76}{5}$ is correct. You can check your answers by changing the improper fraction back to the mixed number. Change $7\frac{1}{4}$ to an improper fraction and check your answer.

$$7\frac{1}{4} = \frac{\quad}{\text{(improper fraction)}} = \frac{\quad}{\text{(mixed number)}}$$

Go to page 24, Frame 30, to check answer and continue from there.

B

You reduced - - but not to the lowest terms. Return to page 15B and find the number that will reduce the $\frac{7}{28}$ and then you'll have the correct answer that will allow you to continue.

B. 1. C. 2. D. A 3.	22. In the list below, place a 'P' by the proper fractions, an 'I' by the improper fractions, and a 'M' by the mixed numbers. a. $3\frac{1}{2}$ b. $\frac{9}{5}$ c. $12\frac{2}{3}$ d. $\frac{3}{4}$ e. $\frac{22}{29}$ f. $\frac{7}{7}$ g. $\frac{79}{75}$
a. M e. P b. I f. I c. M g. I d. P	23. An improper fraction can be changed to a mixed number by dividing the denominator into the numerator. The fraction $\frac{21}{10}$ can be changed to a mixed number by dividing the numerator (number) by (number) the denominator.
21 10	24. To change the improper fraction $\frac{21}{10}$ to a mixed number, follow two steps: (1) Divide the numerator by the denominator to get the whole number: $\begin{array}{r} 2 \\ 10 \overline{) 21} \\ \underline{20} \\ 1 \end{array}$ the whole number (2) Place the remainder over the denominator to get the proper fraction: $\frac{1}{10}$ the proper fraction. Then $\frac{21}{10} =$ (mixed number)
$2\frac{1}{10}$	25. Now change the improper fraction $\frac{26}{5}$ to a mixed number. Show your work. / _____ (mixed number)

A

Wrong! You forgot to add the numerator to the product of the whole number times the denominator. If you now see your error, go back to page 17A and select the other answer and follow directions. If you need the rule again, return to page 21, Frame 28, and start again from there.

B

No ... To reduce an improper fraction, you simply change it to a whole number or to a whole number and a fraction (mixed number) by dividing the numerator by the denominator. Now go back to page 17B and reduce properly.

C

Negative. You have simply added numerators retained highest denominator, and reduced. You must change to equivalent fractions. Re-read rule on page 28, Frame 40, and rework problem from page 30, Frame 43, again.

$\begin{array}{r} 5 \\ 5 \overline{) 26} \\ \underline{25} \\ 1 \end{array}$ $5 \frac{1}{5}$ (mixed number)	26. Try another. Change $\frac{54}{11}$ to a mixed number. $\frac{\quad}{\quad} = \frac{\quad}{\quad}$
$4 \frac{10}{11}$ - - if you missed this one, re-read and rework Frames 22 thru 26, then continue.	27. An improper fraction can be changed to a mixed number. So can a <u>mixed number</u> be changed to an <u>improper fraction</u> . Therefore, an improper fraction is interchangeable with a _____ number.
mixed	28. Changing mixed numbers to improper fractions requires three steps: Example: Change $4 \frac{3}{5}$ to an improper fraction. <u>Step</u> (1) Multiply the whole number by the denominator of the fraction. $4 \times 5 = 20$ (2) Add the product to the numerator. $20 + 3 = 23$. (3) Place the sum over the denominator of the fraction. Then $4 \frac{3}{5} = \frac{\quad}{\quad}$ (improper fraction)
$\frac{23}{5}$	29. Change $12 \frac{2}{3}$ to an improper fraction. If your answer is: Go to page: $\frac{36}{3}$ 13A $\frac{38}{2}$ 15A $\frac{38}{3}$ 17A

A

No! $\frac{2}{4}$ can be reduced to $\frac{1}{2}$ by dividing two (2) into both the numerator and denominator. Remember the rule, a fraction is in its lowest terms only when the number one (1) is the only number that divides evenly into both the numerator and denominator. Return to page 24, Frame 31, and select the correct answer.

B

$1\frac{4}{5}$ is correct. If we ask you to reduce the fraction $\frac{8}{4}$, would you answer 2? You would have been correct there, too. Now turn to top of page 26, Frame 32, and continue the program.

C

No. You've added numerators but have not changed fractions to equivalent fractions. Read rule again on page 9 Frame 40, then rework problem on page 30, Frame 43. Select another answer.

If you came to this page directly from the previous page, you have not followed the directions given in the previous frame. From this point (unless otherwise directed) in the lesson, you will proceed by the scrambled method. Do Not read the frames in sequence, but after selecting an answer, refer to the proper page or frame as directed. Return to page 22, Frame 29, check your answer, and refer to the page as directed.

$7 \frac{1}{4} = \frac{29}{4}$ $\frac{29}{4} = 7 \frac{1}{4}$	30. Change each of the following improper fractions to mixed numbers and the mixed numbers to improper fractions: a. $1 \frac{4}{9}$ c. $10 \frac{11}{12}$ b. $\frac{21}{8}$ d. $\frac{49}{3}$
a. $\frac{13}{9}$ b. $2 \frac{5}{8}$ c. $\frac{131}{12}$ d. $16 \frac{1}{3}$	31. A fraction is in its <u>lowest terms</u> when the number one (1) is the <u>only</u> number that divides evenly into both the numerator and denominator. (NOTE: Dividing both the numerator and denominator by the same number does not change the value of the fraction) Select the fraction below that is in its lowest terms. If your answer is: Go to page: $\frac{2}{4}$ 23A $\frac{6}{9}$ 25A $\frac{3}{7}$ 27A

A

Wrong! $\frac{6}{9}$ can be further reduced. Three (3) is the largest number that divides evenly into both the numerator (6) and the denominator (9). $\frac{6}{9}$, then, reduced to lowest possible terms, is $\frac{2}{3}$. Now return to page 24, Frame 31, and select the correct answer.

B

Right! $1\frac{1}{4}$ is the correct answer. Try another, reduce to lowest terms.

Add $\frac{1}{2} + \frac{1}{2} + \frac{4}{5} + \frac{3}{20} =$ _____

If your answer is:

$$1\frac{9}{10}$$

$$1\frac{19}{20}$$

Go to page:

29B

31B

You came from page 238	32. Reduce each of the following fractions to lowest terms: a. $\frac{12}{4} =$ c. $\frac{64}{72} =$ b. $\frac{21}{49} =$ d. $\frac{17}{51} =$
a. 3 or b. $\frac{3}{7}$ c. $\frac{8}{9}$ d. $\frac{1}{3}$	33. To <u>add</u> or <u>subtract</u> fractions, they must be <u>like</u> fractions. <u>Like</u> fractions have the same number for a denominator. $\frac{7}{12} + \frac{5}{12}$ or $\frac{7}{12} - \frac{5}{12}$ are like fractions because they have the same number for a _____
denominator	34. Fractions must have like (common) denominators before you can _____ or _____ them.
add subtract (any order)	35. When fractions have common denominators, you add or subtract numerators and retain the common denominator. Example: $\frac{7}{12} + \frac{5}{12} = \frac{12}{12}$ reduced = 1 Then $\frac{7}{12} - \frac{5}{12} =$ _____ reduced = _____.
$\frac{2}{12}$	36. Before fractions with <u>unlike</u> denominators can be added or subtracted, they must be changed to their <u>lowest common denominator</u> (LCD). <u>LCD</u> is the lowest number that is divisible by each denominator. Example: $\frac{2}{5} + \frac{1}{20}$ or $\frac{2}{5} - \frac{1}{20}$ The lowest number divisible by each denominator is 20; therefore, 20 is the _____.

4E-75-1104

A

$\frac{3}{7}$ is correct. One (1) is the only number that divides evenly into both 3 and 7.

Let's try a larger fraction. Reduce $\frac{42}{54}$ to its lowest terms.

If your answer is:

Go to page:

$$\frac{21}{27}$$

138

$$\frac{7}{9}$$

158

B

No. Not quite. Your addition is correct but you must have overlooked the "reduce answers to lowest terms." Go back to page 30, Frame 43, reduce, and pick the correct answer.

C

$\frac{6}{15}$ is wrong. You borrowed one (1) from 16, which gave you the fraction $\frac{15}{15}$, but now you must add $\frac{15}{15} + \frac{8}{15}$, then do your subtraction. Return to page 31A, rework the problem, and select another answer.

84

least or lowest common denominator or (LCD)	37. Again, the lowest number divisible by each denominator of fractions to be added or subtracted is called the _____.
LCD	38. Determine the lowest common denominator (LCD) for these fractions: $\frac{1}{2}$ $\frac{1}{4}$, the LCD is _____. $\frac{2}{7}$ $\frac{1}{42}$, the LCD is _____.
4 42	39. Find the LCD for the fractions below: a. $\frac{5}{8} + \frac{1}{16} + \frac{1}{4}$, the LCD is _____. b. $\frac{4}{7} - \frac{1}{49}$, the LCD is _____.
a. 16 b. 49	40. After the LCD has been determined, change all fractions to equivalent fractions of the same denominator; then add or subtract. Example: $\frac{2}{7} + \frac{1}{42}$, the LCD is 42. To change $\frac{2}{7}$ to LCD 42: Divide 7 into 42; the quotient is 6. Multiply 6 by the numerator 2 and place the product (12) over the LCD. $\frac{2}{7} = \frac{12}{42}$. Now we can add. $\frac{12}{42} + \frac{1}{42} = \frac{13}{42}$ reduced is $\frac{13}{42}$. Change the fractions below so they have the same LCD. a. $\frac{1}{3} + \frac{5}{6} + \frac{1}{12} = \underline{\hspace{1cm}} + \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$ b. $\frac{4}{5} - \frac{1}{3} = \underline{\hspace{1cm}} - \underline{\hspace{1cm}}$

A

You have the correct fraction but made a mistake in the addition of whole numbers. Now return to page 33B and work the problem again. Do not just pick the other answer without first re-working the problem to find your error.

B

Incorrect. You've made a mistake someplace in changing fractions to equivalent fractions of the same denominator. Return to page 28, Frame 40, re-read the rule, then go back to page 18B and choose the other answer.

C

$\frac{1}{4}$ is wrong. You did not obtain the reciprocal of the divisor.

$\frac{2}{3}$ inverted is $\frac{3}{2}$ and the reciprocal of $\frac{2}{3}$ is also $\frac{3}{2}$.

Go back to page 13C, rework the problem, and select the correct answer.

a. $\frac{4}{12} + \frac{10}{12} + \frac{1}{12}$ b. $\frac{12}{15} - \frac{5}{15}$	41. Find the LCD and change the fractions below to equivalent fractions. a. $\frac{1}{9} + \frac{1}{81} + \frac{2}{3} = \underline{\hspace{2cm}} + \underline{\hspace{2cm}} = \underline{\hspace{2cm}}$ b. $\frac{4}{5} - \frac{5}{8} = \underline{\hspace{2cm}} - \underline{\hspace{2cm}}$								
a. $\frac{9}{81} + \frac{1}{81} + \frac{54}{81}$ b. $\frac{32}{40} - \frac{25}{40}$	42. The rule again for adding and subtracting fractions. (1) Change fractions to common denominators. (2) Add or subtract numerators. (3) Keep common denominator. (4) Reduce answers to lowest terms. At your left and below are the LCD problems from the last frame. Complete the problems. a. $\frac{9}{81} + \frac{1}{81} + \frac{54}{81} = \underline{\hspace{2cm}}$ reduced $\underline{\hspace{2cm}}$ b. $\frac{32}{40} - \frac{25}{40} = \underline{\hspace{2cm}}$ reduced $\underline{\hspace{2cm}}$								
$\frac{64}{81}$ reduced is $\frac{64}{81}$ $\frac{7}{40}$ reduced is $\frac{7}{40}$	43. Does it all come back to you now? Solve this problem and reduce answer to lowest terms. $\frac{1}{28} + \frac{6}{7} + \frac{5}{14} = \underline{\hspace{2cm}}$ If your answer is: <table border="0"> <tr> <td>$1 \frac{1}{4}$</td> <td>Go to page: 25B</td> </tr> <tr> <td>$1 \frac{7}{28}$</td> <td>27C</td> </tr> <tr> <td>$\frac{12}{28}$</td> <td>23C</td> </tr> <tr> <td>$\frac{3}{7}$</td> <td>21C</td> </tr> </table>	$1 \frac{1}{4}$	Go to page: 25B	$1 \frac{7}{28}$	27C	$\frac{12}{28}$	23C	$\frac{3}{7}$	21C
$1 \frac{1}{4}$	Go to page: 25B								
$1 \frac{7}{28}$	27C								
$\frac{12}{28}$	23C								
$\frac{3}{7}$	21C								
You came from page 28A a. $1 \frac{2}{7}$ d. $1 \frac{3}{4}$ b. $1 \frac{1}{8}$ c. $13 \frac{63}{104}$	44. When multiplying two or more fractions, multiply numerators of the fractions to obtain numerator of the product. To obtain the numerator of the product in the problem $\frac{2}{3} \times \frac{2}{3}$, multiply <u>(number)</u> times <u>(number)</u>.								

47

A

Very good. Work the following problem by subtracting mixed numbers.

Reduce answer to lowest term. $16 \frac{8}{15} - 15 \frac{3}{5} = \underline{\hspace{2cm}}$

If your answer is:

Go to page:

$$1 \frac{14}{15}$$

33A

$$\frac{14}{15}$$

35A

$$\frac{6}{15}$$

27C

Can't be solved

37C

B

Good. $1 \frac{19}{20}$ is correct. Now try one on subtraction and reduce answer

to lowest terms. $\frac{4}{13} - \frac{3}{39} = \underline{\hspace{2cm}}$.

If your answer is:

Go to page:

$$\frac{3}{13}$$

33B

$$\frac{3}{39}$$

42A

C

No. $6 \frac{29}{36}$ is incorrect. Again you forgot to invert the divisor. The

divisor $\frac{1}{6}$ is changed to $\frac{7}{6}$ and inverted is $\frac{6}{7}$. Now go back to page

39B and select another answer.

2 2	45. Like the numerator, the denominator of the product is obtained by multiplying the denominators of the fractions. In the problem $\frac{2}{3} \times \frac{4}{5}$, the numerator of the product is obtained by multiplying _____ times _____ and the denominator is obtained by multiplying _____ times _____.
2 X 4 3 X 5	46. The rule, then, for multiplying fractions is: "Multiply numerators of the fractions to obtain the _____ of the product and multiply the denominators to obtain the _____ of the product." Solve this problem: $\frac{2}{3} \times \frac{2}{5} = \underline{\hspace{2cm}}$
numerator denominator $\frac{4}{15}$	47. The word "of" is sometimes used in place of the multiplication sign "X". $\frac{2}{3}$ of 15 = 10 can be written as $\frac{2}{3} \times \frac{15}{1} = \frac{30}{3} = 10$. Solve this problem and reduce: $\frac{5}{8}$ of 40 = _____ reduced _____.
$\frac{200}{8}$ reduced = 25	48. If the problem contains more than two fractions, multiply all the numerators and multiply all the denominators. Example: $\frac{2}{5} \times \frac{1}{3} \times \frac{2}{3} \times \frac{1}{4} = \frac{4}{180} \text{ reduced } \frac{1}{45}$ Solve this problem: $\frac{3}{5} \times \frac{4}{7} \times \frac{1}{2} = \underline{\hspace{2cm}} \text{ reduced } = \underline{\hspace{2cm}}$

A

You've forgotten the rule on borrowing. $16 \frac{8}{15} = 16 \frac{8}{15} = 15 \frac{23}{15}$

$$- 15 \frac{3}{5} = 15 \frac{9}{15} = 15 \frac{9}{15}$$

You cannot subtract $\frac{9}{15}$ from $\frac{8}{15}$, so you have to borrow a whole number

(1). $1 = \frac{15}{15}$, which you now add to the $\frac{8}{15}$. Don't forget now that you borrowed a whole number from 16. Go back to page 31A. Rework the problem and select the correct answer.

B

Good. Now for the rule for adding and subtracting mixed numbers:

1. Change fractions to like fractions (LCD). 2. Add/subtract the fractions. 3. Add/subtract the whole numbers. 4. Reduce answers to lowest terms. Example: $1 \frac{1}{3} + 3 \frac{11}{12}$ and $7 \frac{1}{2} - 4 \frac{1}{5}$.

$$1 \frac{1}{3} = 1 \frac{4}{12}$$

$$7 \frac{1}{2} = 7 \frac{5}{10}$$

$$+ 3 \frac{11}{12} \quad 3 \frac{11}{12}$$

$$- 4 \frac{1}{5} = 4 \frac{2}{10}$$

$$4 \frac{15}{12} = 4 + 1 \frac{3}{12} = 5 \frac{1}{4}$$

$$3 \frac{3}{10}$$

Now add these fractions: $7 \frac{1}{9} + 6 \frac{5}{18} + \frac{1}{6} =$ _____.

If your answer is:

Go to page:

$$14 \frac{5}{9}$$

29A

$$13 \frac{5}{9}$$

31A

$\frac{12}{70}$ reduced = $\frac{6}{35}$	49. "Cancellation" is a "short cut" used in multiplying fractions. The short cut in multiplying fractions is called _____.
cancellation	50. Cancellation is much the same as reducing. The first step is to select a numerator and denominator that can be divided evenly by the same number. The problem $\frac{5}{10} \times \frac{2}{5} \times \frac{4}{10}$ can be reduced to: $\frac{1}{\cancel{10}} \times \frac{\cancel{2}}{5} \times \frac{2}{\cancel{10}}$ $\frac{1}{5} \times \frac{2}{1} \times \frac{2}{5}$ The next step is to multiply the numerators and the denominators $\frac{1}{5} \times \frac{2}{1} \times \frac{2}{5} = \frac{2}{25}$ reduced is $\frac{2}{25}$. Solve the problem below by cancellation. Show work. $\frac{5}{8} \times \frac{4}{7} \times \frac{1}{5} = \underline{\hspace{2cm}}$
$\frac{1}{\cancel{8}} \times \frac{\cancel{4}}{7} \times \frac{1}{\cancel{8}} = \frac{1}{14}$ $\frac{1}{2} \times \frac{1}{1}$	51. When you use the cancellation method, the basic principle is: Dividing both the numerator and the denominator by the same number does not change the value of a fraction. The value of a fraction is not changed when the _____ and the _____ are _____ by the same number.
numerator denominator (either order) divided	52. In the problem $\frac{2}{15} \times \frac{3}{8}$, the 2 and 8 can be cancelled by dividing each by _____ and the 3 and 15 cancelled by dividing each by _____. The answer to the problem, then, is _____.

A

Very good. The idea here was to see if you remember how to borrow.

Solve the addition and subtraction problems below. Answers must be in lowest terms.

a. $\frac{1}{21} + \frac{4}{7} + \frac{2}{3} =$

b. $3\frac{3}{8} - 2\frac{1}{4} =$

c. $11\frac{1}{8} + 1\frac{3}{13} + \frac{1}{2} + \frac{3}{4} =$

d. $14\frac{1}{6} - 12\frac{5}{12} =$

Go to page 30, Frame 44, to check answers and continue from there.

B

5 is the correct answer. Try one more.

$$5\frac{4}{7} \div 3 =$$

If your answer is:

$$1\frac{6}{7}$$

$$16\frac{5}{7}$$

$$\frac{13}{7}$$

Go to page:

40B

41A

39A

$\frac{2}{3} \times \frac{1}{20}$	53. In the problem $\frac{10}{13} \times \frac{26}{50} \times \frac{7}{21}$, the 10 and 50 are cancelled by dividing each by _____; the 13 and 26 are cancelled by dividing each by _____; and $\frac{7}{21}$ can be reduced to _____. Now solve the problem, showing your cancellation. $\frac{10}{13} \times \frac{26}{50} \times \frac{7}{21} = \underline{\hspace{2cm}}$
10; 13; $\frac{1}{3}$ $\frac{10}{13} \times \frac{26}{50} \times \frac{7}{21} =$ $\frac{1}{1} \times \frac{2}{5} \times \frac{1}{3} =$ ans. $\frac{2}{15}$	54. Solve the following problems, using cancellation where applicable. Reduce answers to lowest terms. a. $\frac{2}{5} \times \frac{3}{10} \times \frac{7}{9} =$ b. $\frac{12}{16} \times \frac{8}{24} \times \frac{8}{10} =$
a. $\frac{7}{75}$ b. $\frac{1}{5}$	55. In order to multiply fractions and mixed numbers, the mixed numbers must be changed to improper fractions. Example: $2\frac{1}{2} \times \frac{3}{5} \times 1\frac{1}{3}$ will be changed to $5 \times \frac{1}{2} \times \frac{4}{3} = 5 \text{ reduced is } 1\frac{1}{4}$ Solve the following problems, using cancellation where applicable, and reduce answers to lowest terms: a. $3\frac{1}{3} \times 5\frac{1}{2} \times \frac{9}{10} =$ b. $4\frac{1}{2} \times 3\frac{1}{3} \times 2\frac{2}{8} =$ c. $\frac{3}{4}$ of 80 =

A

Not quite. $\frac{35}{7}$ is an improper fraction and for the answer to be completely correct (lowest terms), you must now change your answer to a mixed number. Return to page 39B recheck your work, and reduce answer to lowest terms.

B

4 or $\frac{4}{1}$ is incorrect. You obtained the reciprocal of the dividend. You're to obtain the reciprocal of the divisor and then proceed as in multiplication. Now go to page 13C, rework the problem, and select the correct answer.

C

You've forgotten the rule on borrowing. True, you can't subtract $15\frac{9}{15}$ from $16\frac{8}{15}$ unless you borrow. Why not take one (1) from 16 and add the fraction $\frac{15}{15}$ to $\frac{8}{15}$? Now you can subtract, but don't forget the (1) you borrowed. Go back to page 31A, rework the problem, and select another answer.

a. $16 \frac{1}{2}$ b. $33 \frac{3}{4}$ c. 60	56. Solve the problems below, cancelling where applicable, and reduce answers to lowest terms. a. $\frac{3}{5}$ of $2 \frac{5}{8} =$ b. $3 \frac{1}{2} \times 2 \frac{1}{4} \times \frac{2}{3} =$ c. $\frac{1}{6}$ of 24 = d. $2 \frac{1}{8} \div 3 \frac{3}{4} \times 1 \frac{1}{3} =$
a. $1 \frac{23}{40}$ b. $5 \frac{1}{4}$ c. 4 d. $10 \frac{5}{8}$	57. Dividing common fractions requires two steps: Example: $\frac{2}{7} \div \frac{1}{3} =$ Dividend Divisor (1) Obtain reciprocal of divisor - - $\frac{3}{1}$ (invert divisor) (2) Multiply the dividend by the reciprocal of the divisor - - $\frac{2}{7} \times \frac{3}{1} = \frac{6}{7}$ Then $\frac{2}{7} \div \frac{1}{3} =$ _____.
$\frac{6}{7}$	58. Fill in the steps to find $\frac{5}{9} \div \frac{3}{4}$. (1) Obtain reciprocal of divisor (invert the divisor). _____ (2) Multiply the dividend by the reciprocal of the divisor. _____ Then $\frac{5}{9} \div \frac{3}{4} =$ _____.
$\frac{4}{3}$ $\frac{5}{9} \times \frac{4}{3}$ $\frac{20}{27}$	59. Solve this problem: $\frac{3}{10} \div \frac{3}{4}$ If your answer is: Go to page: $\frac{2}{5}$ 13C $\frac{9}{40}$ 15C $\frac{12}{30}$ 40A

A

$\frac{13}{7}$ is unacceptable, because answers will always be reduced to their lowest terms. Return to page 35B and select the correct answer that is in its lowest terms.

B

$\frac{9}{16}$ is correct.

Dividing with mixed numbers requires three steps: (1) Change the mixed number or mixed numbers to improper fractions. (2) Obtain the reciprocal of the divisor (invert divisor). (3) Multiply the dividend by the reciprocal of the divisor.

Try this problem: $5 \frac{5}{6} \div 1 \frac{1}{6} =$

If your answer is:

Go to page:

5

35B

$6 \frac{29}{36}$

31C

$\frac{35}{7}$

31A

A

Not quite right. You must not have cancelled the 3's after obtaining the reciprocal of the divisor and you haven't reduced to the lowest terms. Go back to page 38 frame 59, and correct your mistake. Then select the correct answer.

B

$1\frac{6}{7}$ is correct. Divide the following fractions and reduce answers to lowest terms:

a. $\frac{5}{8} \div \frac{3}{4} =$

b. $22 \div 6\frac{7}{8} =$

c. $2\frac{1}{6} \div 4\frac{1}{2} =$

d. $\frac{8}{21} \div 3\frac{3}{7} =$

GO TO PAGE 41B TO CHECK YOUR ANSWERS.

63-75-1104

A

No! Does it sound reasonable that 3 is contained in $5\frac{4}{7}$ - - - 16 and $\frac{5}{7}$ times? You forgot to obtain the reciprocal of the divisor before you multiplied. Go back to page 358, invert the divisor, multiply, and then select the correct answer.

B

Answers from page 408; a. $\frac{5}{6}$ b. $3\frac{1}{5}$ c. $\frac{13}{27}$ d. $\frac{1}{9}$

If you had any answers other than those above, you must rework the problem(s) on page 408. When you've gotten all correct, solve these problems:

a. $5\frac{2}{3} \div 9\frac{5}{9} =$

b. $5\frac{2}{5} \times 2\frac{1}{2} \times 4\frac{2}{3} =$

c. $21\frac{1}{16} + 9\frac{3}{8} + 8\frac{1}{2} + \frac{3}{4} =$

d. $3\frac{3}{16} - 1\frac{3}{4} =$

GO TO PAGE 43A TO CHECK YOUR ANSWERS.

A

Never! The only way you could have arrived at this answer was to have reduced the numerator and not the denominator. Return to page 31B, work the problem again, and select the correct answer.

6F-75-1104

A

Answers from page 34B:

- a. $\frac{51}{86}$ b. 63 c. $39\frac{11}{16}$ d. $1\frac{7}{16}$

If you missed any problem, you must rework and recheck. After all problems are correct, read the rules again that are on the pages listed below and then go to page 43B.

Problem:	Go to page:
a. (division)	39B
b. (multiplication)	36, Frame 55
c. (addition)	33B
d. (subtraction and borrowing)	33A

After you've read the rules again, go to page 43B.

B

You have completed the Programmed Lesson on fractions. For some, the program was just a review; for others, it has been a process of learning.

A SELF-TEST ON FRACTIONS COMMENCES ON PAGE 44.

SELF-TEST ON FRACTIONS

1. Write the definition of a fraction.

2. Identify the two parts of the fraction $\frac{7}{8}$ and explain what each part shows.

7 - -

8 - -

3. Identify the proper fractions, the improper fractions, and the mixed numbers in the following list by placing a "P" by the proper fractions, an "I" by the improper fractions, and an "M" by the mixed number.

a. $\frac{15}{16}$

f. $\frac{300}{299}$

b. $\frac{19}{17}$

g. $\frac{10}{11}$

c. $2\frac{4}{5}$

h. $\frac{7}{12}$

d. $\frac{9}{7}$

i. $6\frac{3}{7}$

e. $77\frac{2}{3}$

j. $\frac{5}{6}$

4. Change the mixed numbers to improper fractions and the improper fractions to mixed numbers.

a. $3\frac{2}{3}$

d. $\frac{19}{15}$

b. $\frac{11}{10}$

e. $7\frac{7}{8}$

c. $12\frac{4}{5}$

5. Reduce the following fractions to their lowest terms:

a. $\frac{18}{81}$

d. $\frac{3}{7}$

b. $\frac{9}{12}$

e. $\frac{14}{21}$

c. $\frac{21}{63}$

f. $\frac{16}{64}$

6. Solve the following ADDITION and SUBTRACTION problems. Reduce answers to lowest terms.

a. $\frac{1}{2} + \frac{1}{2} =$

d. $2\frac{3}{8} - 1\frac{5}{8} =$

b. $\frac{5}{7} - \frac{2}{3} =$

e. $6\frac{7}{10} - 4\frac{4}{5} =$

c. $\frac{3}{8} + \frac{3}{4} =$

f. $11\frac{3}{4} + 19\frac{5}{8} + 9\frac{1}{2} + \frac{3}{16} =$

7. Multiply the following fractions, cancelling where applicable. Reduce answers to lowest terms.

a. $\frac{1}{2} \times \frac{3}{4} \times \frac{2}{3} =$

b. $4\frac{2}{3} \times 5\frac{1}{4} \times 2\frac{2}{3} =$

c. $\frac{3}{4} \times 5\frac{1}{2} =$

d. $\frac{1}{8}$ of 16 =

8. Divide the following fractions, cancelling where applicable. Reduce answers to lowest terms.

a. $\frac{7}{8} \div \frac{7}{16} =$

d. $\frac{4}{5} \div 2\frac{7}{15} =$

b. $15 \div 4\frac{1}{5} =$

c. $4\frac{2}{3} \div 12\frac{4}{9} =$

GO TO NEXT PAGE FOR ANSWERS.

ANSWERS TO SELF-TEST

1. A fraction is part of a whole.
2. 7 - - Numerator. Indicates how many parts of the whole are being considered.
8 - - Denominator. Indicates how many equal parts the whole has been divided into.
3. a. P; b. I; c. M; d. I; e. M; f. I; g. P; h. P; i. M;
j. P
4. a. $\frac{11}{3}$; b. $1\frac{1}{10}$; c. $\frac{64}{5}$; d. $1\frac{4}{15}$; e. $\frac{63}{8}$
5. a. $\frac{2}{9}$; b. $\frac{3}{4}$; c. $\frac{1}{3}$; d. $\frac{3}{7}$; e. $\frac{2}{3}$; f. $\frac{1}{4}$
6. a. 1; b. $\frac{1}{21}$; c. $1\frac{1}{8}$; d. $\frac{3}{4}$; e. $1\frac{9}{10}$; f. $41\frac{1}{16}$
7. a. $\frac{1}{4}$; b. $65\frac{1}{3}$; c. $4\frac{1}{8}$; d. 2
8. a. 2; b. $3\frac{4}{7}$; c. $\frac{3}{8}$; d. $\frac{12}{37}$

Solve Problems Involving Decimal Fractions

Directions: Read each frame carefully, then, write in the answer; be sure that you are satisfied with your answer before you write it in.

This section was designed to provide you with a review of the multiplication and division of decimals. When you have completed this section, you should be able to convert within the system without any difficulty. Remember:

- 0.1 one place to right of decimal is 1/10
- 0.01 two places to right of decimal is 1/100
- 0.001 three places to right of decimal is 1/1000
- 0.010 NOTE: This is the same as 1/100. You may drop the last zero.

1. When multiplying decimal numbers, you must remember to count off the TOTAL number of decimal places in the answer.

EXAMPLE: 1.50 (Contains two places)
 x .5 (Contains one place)
 .750 After adding the places of both parts of the problem you can see that the answer must contain three places.

2. Work the following problems and place your answers in the blanks provided.

a. 7.50
 x0.4 ()

b. 8.471
 x10.12 ()

3. Work this problem: 3.760
 x .40 Your answer ()

4. When multiplying decimal numbers, the problem (multiplier and multiplicand) and the answer will contain () amount of number(s) placed to the right of the decimal point.

5. Study this problem

$$\begin{array}{r} .364 \\ \times .02 \\ \hline 728 \end{array}$$

You know that there must be five places to the right of the decimal point; therefore, you must add two zeros. The answer, then, would be ().

6. Solve this problem:

$$\begin{array}{r} .322 \\ \times .02 \\ \hline \end{array}$$

Your answer: ()

7. To divide a whole number by a decimal; convert the decimal (divisor) to a whole number by moving the decimal all the way to the right. Move the decimal in the whole number (dividend) the same number of places to the right. Divide as usual, placing the decimal directly above the decimal point in the dividend. NOTE THE ARROWS BELOW.

Example:

$$\begin{array}{r} 356 \\ .712 \wedge \end{array}$$

$$.712 \wedge \overline{356.000} \wedge$$

8. Convert this decimal number to a whole number and place the decimal in the answer.

$$\begin{array}{r} 4 \\ 0.44 \end{array}$$

$$0.44 / \overline{4.}$$

9. To check the answer after dividing, multiply the product by the divisor. The answer after multiplying, should be the same as the dividend.

EXAMPLE: (divisor) $.04 / \overline{6.50.0}$ (product) (dividend)

$$\begin{array}{r} 4 \\ 25 \\ 24 \\ \hline 10 \\ 8 \\ \hline 20 \\ 20 \\ \hline 0 \end{array}$$

Checking your answer:

$$\begin{array}{r} 162.5 \text{ (product)} \\ .04 \text{ (divisor)} \\ \hline 6.500 \text{ (dividend)} \end{array}$$

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10. Solve these problems and check the answers.

a.
$$\begin{array}{r} 5 \\ 0.44 \end{array}$$

b.
$$\begin{array}{r} 8 \\ .55 \end{array}$$

c.
$$\begin{array}{r} 91.2 \\ \times .21 \\ \hline \end{array}$$

d.
$$\begin{array}{r} .463 \\ \times .02 \\ \hline \end{array}$$

11. To solve this problem, $\frac{3.50}{.02} \times .02 = 3.50$

Divide
$$\begin{array}{r} 175. \\ .02 \overline{) 3.50} \\ \underline{2} \\ 15 \\ \underline{14} \\ 10 \\ \underline{10} \\ 0 \end{array}$$

and multiply
$$\begin{array}{r} 175 \\ .02 \\ \hline 3.50 \end{array}$$

Solve this problem: $\frac{4.40}{.02} \times .3 = (\quad)$

12. To solve the above problem, first ($\quad$) and then ($\quad$).

13. Solve these problems:

a. $\frac{6.50}{.04} \times .4 = (\quad)$

b. $\frac{3.50}{.03} \times .3 = (\quad)$

14. Solve these problems and check the answers. Record your answers below each problem. Remember, you must divide to check your answers.

a.
$$\begin{array}{r} 0.69 \\ \times .04 \\ \hline \end{array}$$

b.
$$\begin{array}{r} .256 \\ \times .57 \\ \hline \end{array}$$

c.
$$\begin{array}{r} .756 \\ \times .04 \\ \hline \end{array}$$

15. To round off a decimal number, increase the last place number by one when the next figure is five or greater; leave the last place number the same when the next figure is less than five.

Example:

- a. Round off 1.876643 to two places after the decimal point.
1.876643. The third place number is five or more, so seven is increased by one, and 1.876643 becomes 1.88.
- b. Round off 1.432329 to four places after the decimal point.
1.432329. The fifth place number is four or less, so three remains the same and 1.432329 becomes 1.4323.
- c. Round off 1.875429 to four places after the decimal point.
1.875429. The fifth place number is (). Since it is less than () 1.875429 becomes ().
Round off 1.875449 to four places after the decimal point.
Answer here. ().

16. Round off 3.4357810 to three places after the decimal point.

Answer here. ()

17. Round off 12.1314 to two places after the decimal point.

Answer here. ()

18. Round off 11.25 to one place after the decimal point.

Answer here. ()

19. Solve these problems and round off your answers to one place after the decimal point.

a. $\frac{32.43}{.02} \times 2.44 =$

b. $\frac{2.652}{.03} \times 4.345$

77-75-1104

WORK SHEET

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ANSWERS TO DECIMAL FRACTION PROGRAMMED TEXT

1. Non response
2. a. 3
b. 85.72652
3. 1.50400
4. The same
5. .00728
6. .00644
7. No response
8. 0.44/4.00
9. No response
10. a. 11.363 and 5.00
b. 14.545 and 8.00
c. 19.152 and 91.2
d. .00926 and .463
11. 220 and 66.0
12. Divide and multiply
13. a. 162.5 and 65
b. 116.66 and 34.998
14. a. .0276 and .69
b. .14592 and .256
c. .03024 and .756
15. 2; 5; 1.8754; 1.8754
16. 3.436
17. 12.13
18. 11.3
19. a. 1621.5 and 3956.5
b. 88.4 and 384.1

Convert Simple Fractions to Decimal Fractions

EXAMPLE: Convert $\frac{3}{4}$ to a decimal fraction.

Step 1. Write down the simple fraction: $\frac{3}{4}$

Step 2. Divide the numerator by the denominator: $4 \overline{)3.000}$

Step 3. Your answer: .75

PRACTICE PROBLEMS

1. Convert $\frac{1}{20}$ to a decimal fraction.
2. Convert $\frac{1}{80}$ to a decimal fraction.
3. Convert $\frac{25}{1000}$ to a decimal fraction.
4. Convert $\frac{2}{3}$ to a decimal fraction.

Convert Decimal Fractions To Simple Fractions

EXAMPLE: Convert 0.06 to a simple fraction.

Step 1. Write down the decimal fraction: 0.06

Step 2. This is read as six-hundredths. So place the 6 over 100: $\frac{6}{100}$

Step 3. Reduce this fraction to lowest terms: $\frac{3}{50}$

Step 4. Your answer: $\frac{3}{50}$

PRACTICE PROBLEMS

1. Convert 0.125 to a simple fraction.

2. Convert 0.005 to a simple fraction.

3. Convert 0.250 to a simple fraction.

4. Convert 1.25 to a simple fraction

Roman Numerals

M = 1000

D = 500

C = 100

L = 50

X = 10

V = 5

i = 1

ss = 1/2

PRACTICE PROBLEMS

Convert the following to Arabic numerals:

1. DC = _____

2. XVi = _____

3. CX = _____

4. XL = _____

Convert the following to Roman numerals:

1. 25 = _____

2. 38 = _____

3. 14 = _____

4. 151 = _____

Rearrange Equations

DEFINITION: An equation is a statement of equality.

NOTE: To solve for any term in an equation; that term must stand alone.

EXAMPLE: Using the equation, $A = \frac{B}{C}$; solve for "B".

Step 1. Always write the complete equation before starting:

$$A = \frac{B}{C}$$

Step 2. Multiply both sides by "C":

$$AC = \frac{BC}{C}$$

Step 3. Cancel the "C's" on the right:

$$AC = \frac{B\cancel{C}}{\cancel{C}}$$

Step 4. Your new equation:

$$AC = B$$

PRACTICE PROBLEMS

1. Solve for X. $A = \frac{X}{Y}$

2. Solve for N $2 = \frac{N}{6}$

3. Solve for Z $4 = \frac{7}{A}$

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Rearrange Equations, Con't

EXAMPLE: Using the equation, $X = \frac{Y}{Z}$; Solve for Z.

Step 1. Write the complete equation: $X = \frac{Y}{Z}$

Step 2. Multiply both sides by "Z": $XZ = \frac{Y}{Z} Z$

Step 3. Cancel the "Z's" on the right side: $XZ = \frac{Y\cancel{Z}}{\cancel{Z}}$

Step 4. Now your equation looks like: $XZ = Y$

Step 5. Divide both sides by "X": $\frac{XZ}{X} = \frac{Y}{X}$

Step 6. Cancel the "X's" on the left side: $\frac{\cancel{X}Z}{\cancel{X}} = \frac{Y}{X}$

Step 7. Your new equation: $Z = \frac{Y}{X}$

PRACTICE PROBLEMS

1. Solve for B. $A = \frac{C}{B}$

2. Solve for N. $3 = \frac{6}{N}$

DEPARTMENT OF BIOMEDICAL SCIENCES

PHARMACY SPECIALIST

10-8

BLOCK I
FUNDAMENTALS OF PHARMACY

March 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
SHEPPARD AIR FORCE BASE, TEXAS

Designed For ATC Course Use

DO NOT USE ON THE JOB

Department of Biomedical Science
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas

WB 3ABR90530-I-1
March 1975

FUNDAMENTALS OF PHARMACY

OBJECTIVES

1. Solve problems pertaining to basic mathematical operations, metric system, apothecary system, avoirdupois system, and Ratio and Proportion.
2. Solve problems pertaining to conversion of weights and measures, and calculation of doses.

PROCEDURES

Systems of measurements in the the past were based on traditional standards, such as the length of the King's foot or the weight of a grain of wheat. For their time and technology they were adequate. The need for a more exacting and universal system of measurement brought about the creation and standardization of the Metric system.

Instructions

Each type problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

DEFINITION: METER is the basic unit of length (39.37) inches)

Liter is the basic unit of volume (The volume of the cube of 1/10 of a meter.

Gram is the basic unit of weight (equal to the weight of one cubic centimeter of water at 4 degrees centigrade).

DEFINITION: Latin prefixes:

deci is equal to 1/10 of the basic unit.

centi is equal to 1/100 of the basic unit.

milli is equal to 1/1000 of the basic unit.

micro is equal to 1/1,000,000 of the basic unit.

This supersedes SW 3ABR90530-I, December 1973

DEFINITION: GREEK PREFIXES

deka is equal to ten times the basic unit

hecto is equal to one hundred times the basic unit.

kilo is equal to one thousand times the basic unit.

ABBREVIATIONS:

Meter = M

Liter = L

Gram = Gm

deci = d

centi = c

milli = m

micro = mc

Deka = D

Hecto = H

Kilo = K

Solve Problems Within The Metric System

This is a programmed text to help you to learn the Metric System.
Follow the directions carefully and do not "skip around".

Directions: Read each frame carefully, then, write in the answer;
be sure that you are satisfied with your answer before
you write it in.

1. In the metric system, weight is expressed in grams, linear measurement is expressed in meters, and liquid volume is expressed in liters.

The system which uses grams, meters and liters is called the
() system.

2. The primary units of measurements in the metric system are (),
(), and ().

3. Which of the following units of measurement belong to the metric system: (Circle your answers below)

- | | |
|-----------|----------|
| a. pound | d. liter |
| b. gram | e. yard |
| c. gallon | f. meter |

4. The gram, which is a much smaller unit than our commonly used pound, is the basic metric unit used to measure (Circle your answer below)

- | | | |
|-----------|-----------|-----------|
| a. volume | b. length | c. weight |
|-----------|-----------|-----------|

5. Length, in the household system, is measured in inches, feet, yards, etc. In the metric system, however, the primary unit for the measurement of length is the meter.

With the metric system, length is measured in ().

6. When using the metric system to measure the length of an item, you would record its length as so many ().

7. In the metric system, the primary unit of weight is the ();
the primary unit of length is the ().

8. The primary metric unit of measurement used to measure volume is the liter.

Which of the following is used to measure volume in the metric system?
(Circle your answer below)

- | | |
|------------|-----------|
| a. pounds | d. grams |
| b. gallons | e. liters |
| c. meters | f. inches |

9. In the common household system, pints, quarts and gallons are used to measure volume. In the metric system, however, the primary unit used to measure volume is the ().
10. When items are weighed by the metric system, their weight is expressed in ().
11. The length of an item measured by the metric system is expressed in ().
12. The volume of liquids measured by the metric system is expressed in ().
13. When metric measurements are written, the amount is written as a numeral followed by the unit. Study these examples:

Four meters is written as 4 meters.
Four liters is written as 4 liters.
Twelve grams is written as 12 grams.

Now write the following measurements:

- a. Four grams ()
- b. Eight liters ()
- c. Nine meters ()

14. If the measurement contains a fraction, the fraction is written as a decimal. Study these examples:

4 $\frac{1}{4}$ meters is written as 4.25 meters.
4 $\frac{3}{4}$ liters is written as 4.75 liters.
4 $\frac{1}{8}$ grams is written as 4.125 grams.

Now, write the following measurements.

- a. Five and one-half grams ()
- b. Three and one-fourth meters ()
- c. Four and three-quarter liters ()

15. Write the primary metric unit used to measure weight, length and volume.

- a. weight ()
- b. length ()
- c. volume ()

16. You should also know the abbreviations for the three basic metric units of measurement. Abbreviations of the basic units are always capitalized.

17. The abbreviation for gram(s) is Gm.

Using the abbreviation, write 12 grams. ()

18. The abbreviation for meter(s) is M.

Using the abbreviation, write 2 meters. ()

19. The abbreviation for liter(s) is L.

Using the abbreviation, write 1 liter. ()

20. Using abbreviations, write:

200 liters ()

17 meters ()

16 grams ()

21. In addition to the three basic units you have just studied, the metric system has other units which are subdivisions of the basic units. Let us now study some of those subdivisions which are frequently used.

22. The common subdivision of the gram is the milligram (.001 of a gram) The abbreviation for the milligram is mg.

Using the abbreviation, write 12 milligrams. ()

23. When the prefix milli (m) is used with a basic unit (Gm., L., etc.) and the figure is less than 1000, the amount expressed is less than the basic unit.

Example: 500 mg. = .5 of a gram
 250 ml. = .25 of a liter
 700 mm. = .7 of a meter

24. When the prefix milli (m) is used with a basic unit and the figure is greater than 1000, the amount expressed is more than the basic unit.

EXAMPLE: 1,500 mg. = 1.5 grams
 2,500 ml. = 2.5 liters
 1,700 mm. = 1.7 meters

Complete the following:

- a. 350 milligrams = () grams
 - b. 2,300 milliliters = () liters
 - c. 1,800 milligrams = () grams
 - d. 300 millimeters = () meters
 - e. 450 milliliters = () liters
25. A meter may be divided into 100 parts; each part, then is one centimeter (.01 of a meter). The abbreviation for centimeter is cm. The abbreviation for cubic centimeter is cc.
- Using the abbreviation, write 1 centimeter. ()
- Using the abbreviation, write 4 cubic centimeters. ()
26. Using the abbreviation, write 500 cubic centimeters. ()
- Using the abbreviation, write 400 cubic centimeters. ()
27. The common subdivision of the liter is the milliliter, or .001 of a liter. The abbreviation for milliliter is ml.
- Using the abbreviation, write 200 milliliters ()
28. Using the abbreviation, write 4 milliliters. ()
29. Write the abbreviations for meter (), gram (), liter (), cubic centimeter (), milliliter (), milligram () and centimeter ().

36. A kiloliter is equal to () liters. A milliliter is equal to what part of a liter? ()
37. To express 1,000 grams, 1,000 liters and 1,000 meters, you may use the same prefix which is ()
38. To express .001 of a gram, .001 of a liter and .001 of a meter, you may use the prefix ().
39. As you recall, 1 milliliter is used to express .001 of a liter. Another way to express that same amount is 1 cubic centimeter, abbreviated 1 cc. This is true because 1 cc. occupies the same space and has the same volume as 1 milliliter.

One cc. is () one milliliter.

40. Do not get the two prefixes confused. Remember that the prefix milli means .001; the prefix centi means .01.

In the spaces below, write five cubic centimeters and eight centimeters using abbreviations.

() ()

41. To convert grams to milligrams, multiply the number of grams by 1000 or move the decimal three places to the right.

Example: 0.15 Gm. = 150 mg.

$$\begin{array}{r} 0.15 \\ \times 1,000 \\ \hline 150.00 \end{array}$$

42. Convert 2.5 grams to milligrams. ()

43. To convert milligrams to grams, divide the number of grams by 1000 or move the decimal three places to the left.

EXAMPLE: 150 mg. = 0.15 Gm.

$$\begin{array}{r} 0.15 \\ 1000 \overline{) 150.00} \\ \underline{100.0} \\ 5000 \\ \underline{5000} \end{array}$$

850 mg. = () Gm.

44. Now that you know the prefixes, work the following problems for practice. Check your responses.

- a. 500 milligrams is the same as () grams.
- b. 2 grams is the same as () milligrams.
- c. 500 centigrams is the same as () milligrams.
- d. 350 milligrams is the same as () centigrams.
- e. 250 milliliters is the same as () liters.
- f. 180 liters is the same as () milliliters.
- g. 420 millimeters is the same as () meters.
- h. 3.5 meters is the same as () millimeters.
- i. 500 kilograms is the same as () grams.
- j. 4,500 grams is the same as () kilograms.
- k. 1 kilogram is the same as () centigrams.
- l. 2,500 centiliters is the same as () kiloliters.
- m. 3.5 kiloliters is the same as () liters.
- n. 1.6 meters is the same as () kilometers.

Answers for Metric System

1. metric
2. grams, meters liters
3. gram; liter; meter
4. weight
5. meters
6. meters
7. gram, meter
8. liters
9. liter
10. grams
11. meters
12. liters
13. 4 grams; 8 liters; 9 meters
14. 5.5 grams; 3.25 meters; 4.75 liters
15. grams; meters; liters
16. No response
17. 12 Gm.
18. 2M
19. 1 L.
20. 200 L; 17M; 16 Gm
21. No response
22. 12 mg
23. No response
24.
 - a. .35
 - b. 2.3
 - c. 1.8
 - d. .3
 - e. .45
25. 1 cm; 4 cc
26. 500 cc; 400 cc
27. 200 ml
28. 4 ml
29. M; Gm; L; cc; ml; mg; cm
30.
 - a. 15 cc
 - b. 10 Gm
 - c. 9 mg
 - d. 5 L
 - e. 1 cc
 - f. 17 ml
 - g. 14 cm
31. kilo
32. are; are not
33. kilometers
34. 1,000
35. 1,000
36. 1,000; .001
37. kilo
38. milli
39. equal to (or same as)
40. 5 cc; 8 cm
41. No response
42.
$$\begin{array}{r} 2.5 \text{ Gm.} \\ \times 1000 \\ \hline 2500.0 = 2500 \text{ mg.} \end{array}$$
43. .850
44.
 - a. .5 grams
 - b. 2,000
 - c. 5,000
 - d. 35
 - e. .25
 - f. 180,000
 - g. .42
 - h. 3,500
 - i. 500,000
 - j. 4.5
 - k. 100,000
 - l. .025
 - m. 3,500
 - n. .0016

PROCEDURES

One of the oldest system of weights and measures is the Apothecary system and although antiquated and no longer official it is still used extensively in medicine. Therefore your complete comprehension is necessary.

Instructions

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

APOTHECARY TABLE OF WEIGHTS

20 grains	1 scruple
3 scruples	1 drachm
8 drachms	1 ounce
12 ounce	1 pound

APOTHECARY TABLE OF FLUID MEASURE (VOLUME)

60 minims	1 fluidrachm
8 fluidrachms	1 fluidounce
16 fluidounces	1 pint
2 pints	1 quart
4 quarts	1 gallon

DEFINITION OF APOTHECARY SYMBOLS

MINIM	℥
FLUIDRACHM	fl ℥
FLUIDOUNCE	fl ℥
PINT	.pt., ②
QUART	qt
GALLON	C or Cong
GRAIN	.gr
SCRUPLE	℥
DRACHM. .DRAM	℥
OUNCE	℥
POUND	℥

Restate To A Lower Denomination In The Apothecary System

EXAMPLE: Reduce 3 fl $\frac{3}{4}$ 2 fl $\frac{3}{4}$ to mx

Step 1. Copy the value from the problem carefully,

3 fl $\frac{3}{4}$ 2 fl $\frac{3}{4}$

NOTE: Each value will be reduced separately

Step 2. First reduce the 2 fl $\frac{3}{4}$
How many mx are in each
fl $\frac{3}{4}$? _____ How many
fl $\frac{3}{4}$ are you reducing? _____
Multiply 60 times 2 to
find the number of mx in
2 fl $\frac{3}{4}$.

$$\begin{array}{r} 60 \\ \times 2 \\ \hline 120 \end{array}$$

Step 3. Rewrite the problem using
120 mx for the 2 fl $\frac{3}{4}$

3 fl $\frac{3}{4}$ 120 mx

Step 4. Now reduce the 3 fl $\frac{3}{4}$
How many mx are in each
fl $\frac{3}{4}$? _____. How
many fl $\frac{3}{4}$ are you chang-
ing? _____. Multiply 480
times 3 to find the number
of mx in 3 fl $\frac{3}{4}$.

$$\begin{array}{r} 480 \\ \times 3 \\ \hline 1440 \end{array}$$

Step 5. Rewrite the problem using
1440 mx for the 3 fl $\frac{3}{4}$

Step 6. Add up the mx and your
answer is:

1560 mx

PRACTICE PROBLEMS

1. Convert the following to a lower denomination in the Apothecary System.

a. Cong ii, pt ii, fl $\overline{\text{z}}\text{ii}$ to fl z

b. $\overline{\text{z}}\text{xvi}$, $\overline{\text{z}}\text{xxxii}$ to gr

c. Oiii , mx 480 to fl z

d. ss , $\overline{\text{z}}\text{xxiv}$ to $\overline{\text{z}}$

Restate To A Higher Denomination In The Apothecary System

EXAMPLE: Change 5840 gr to weighable units.

Step 1. Copy the values from the problem carefully.

5840 gr

Step 2. Study this number. What is the largest unit that this could be changed to? _____
How many grains does this unit contain? _____

Step 3. Now to find the number of pounds, divide 5760 into 5840. The number of pounds in 5840 gr is _____ and the number of grains left over is 80
Now the amount is rewritten using the pound and grains.

$$\begin{array}{r} 1 \\ 5760 \overline{) 5840} \\ \underline{5760} \\ 80 \text{ gr left} \end{array}$$

1 lb. 80 gr.

Step 4. Study the 80 gr. What is the largest unit that this amount can be changed to? _____
How many grains does 1 dram contain? _____

Step 5. Now to find the number of drams, divide 60 into 80. The number of drams in 80 gr. is _____ and the number of grains left over is, _____.

$$\begin{array}{r} 1 \\ 60 \overline{) 80} \\ \underline{60} \\ 20 \text{ gr left} \end{array}$$

Step 6. Now rewrite the problem again using the pound, dram and grains

1 lb 1 3 20 gr

Step 7. Study the 20 gr. What is the largest unit that this can be changed to? _____
How many grains does this unit have in it? _____

Step 8. Now to find the number of scruples in 20 gr, divide 20 into 20.
The number of scruples is _____.

Step 9. Rewrite the problem using the number of scruples. Your answer.

1 lb 1 3 1 3

PRACTICE PROBLEMS

1. Convert the following to weighable Apothecary units.

a. 3440 gr.

b. 1650 gr.

c. 950 gr.

d. 695 gr.

AVOIRDUPOLIS SYSTEM

PROCEDURES

The Avoirdupois System is the official system of commerce and you are indirectly related to commerce in ordering bulk drugs through medical supply channels. You must have a complete understanding of this system to facilitate transactions with supply.

Instructions

Each type problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete problems assigned. SHOW ALL WORK!

AVOIRDUPOLIS TABLE OF WEIGHT

437.5 grains 1 ounce
16 ounces 1 pound

Definition of Avoirdupois Symbols

grain gr
ounce oz
pound lb

To restate Avoirdupois units to a higher or lower denomination follow the same procedures you used in restating within the Apothecary System.

PRACTICE PROBLEMS

1. Reduce the following to weighable Avoirdupois denominations.
 - a. 7540 gr _____
 - b. 1560 gr _____
 - c. 856 gr _____
 - d. 466 gr _____
2. How many 10 grain capsules can be made from $\frac{1}{2}$ lb of iron crystals?
3. How many 5 grain capsules of Aspirin can be made from 4 oz of Aspirin powder?
4. How many $\frac{1}{2}$ gr tablets of Codeine can be made from $\frac{1}{8}$ oz of codeine powder?
5. How many grains of chemical are left in a 1 oz bottle after enough of it has been used to make 2000 tablets each containing $\frac{1}{200}$ grain of the chemical.

RATIO AND PROPORTION

PROCEDURES

If it were possible to choose the most useful method of solving mathematical problems, ratio and proportion would probably be selected. Nearly 80 percent of the problems you will encounter in Pharmacy can be solved using this method.

Instructions

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

DEFINITION: A ratio is the numerical comparison of two similar quantities.

DEFINITION: A proportion is a statement of the equality of two ratios.

Solving Problems Using Ratio and Proportion

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EXAMPLE: How many feet per second is a car traveling at 90 mph, if at 60 mph it is traveling 88 feet per second?

Step 1. Read the question. Determine what is asked (the number of feet per second at 90mph) and call this the "problem". Now determine what information is given (88 feet per second at 60 mph).

Step 2. Write the "problem" on one line, using "X" for the unknown.

90 mph X ft/sec

Step 3. Write the given information on the line under the problem. Be sure to place 60 mph under the 90 mph (the first ratio) and the 88 ft/sec under the "X" ft/sec (the second ratio)

90 mph X ft/sec
60 mph 88 ft/sec

Step 4. Now draw a line between the 90 mph and the 60 mph and another line between the "X" ft/sec and the 88 ft/sec. Then place an equal sign in the center.

$\frac{90 \text{ mph}}{60 \text{ mph}} = \frac{X \text{ ft/sec}}{88 \text{ ft/sec}}$

Step 5. Cross multiply (90 times 88 and 60 times "X"). Giving the products. (Note: the ft/sec and the mph are not used here)

$$60 \cdot X = 90 \cdot 88$$

$$60 X = 7920$$

Step 6. Divide by the number next to the "X"

$$X = \frac{7920}{60}$$

Step 7. Your answer (Note: the ft/sec is placed next to the answer because "X" is the number of ft/sec)

$$X = 132 \text{ ft/sec}$$

PRACTICE PROBLEMS

MAKE VALID RATIOS BETWEEN THESE QUANTITIES

1. 1 yard and 2 feet _____
2. 4 hours and 120 minutes _____
3. 2 feet and 6 inches _____
4. 100 Grams and 10 Kilograms _____
5. Butter sells 3 lb. for 98¢. How much will 2 lb. cost?
6. A drug cost \$6.98 for 12 ounces. How much will three and $\frac{3}{4}$ ounces cost the pharmacist?
7. 20 gallons of gasoline will run your car 235 miles. How far should you go on six and $\frac{1}{2}$ gallons?

8. The airliner travels 600 mph and you will fly 1230 miles. How long will your trip be?

9. The item sells for \$4.25 a dozen and you only have \$2.00. How many can you buy?

CONVERSION OF WEIGHTS AND MEASURES

PROCEDURES

Even though AFM 168-4 states that all prescriptions should be written in the Metric System, some physicians will continue to write in one of the other systems. The responsibility will rest on you to convert these prescriptions to the Metric System.

Instructions

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

CONVERSION EQUIVALENTS:

64.8 mg	=	1 gr.
1 Gm.	=	15.432 gr.
31.1 Gm.	=	1 ounce (Apoth)
28.35 Gm.	=	1 oz. (Av.)
454 Gm.	=	1 lb. (Av)
1 Kg.	=	2.2 lb. (Av)
1 ml.	=	16.23 minim
29.57 ml.	=	1 fl ounce (Apoth)
473 ml.	=	1 pint

COMMON EQUIVALENTS:

1 Teaspoonful	=	1 dram = 5 ml
1 Tablespoonful	=	1/2 fl ounce = 15 ml

NOTE: The common equivalents are used only when interpreting prescriptions.

Convert From The Common Systems to The Metric System

EXAMPLE: Convert 4 fl ~~z~~ to ml.

Step 1. Identify the problem; in this case it is to convert 4 fl ~~z~~ to ml. Write the problem down, use an "X" for the unknown.

$$4 \text{ fl } \cancel{z} = x \text{ ml.}$$

Step 2. Be sure that the common system quantity is in one denomination, make any changes now.

$$4 \text{ fl } \cancel{z} = x \text{ ml.}$$

Step 3. Choose a conversion equivalent that possesses both the denominations present in the problem. In this case use, 1 fl ~~z~~ = 29.57 ml.

Step 4. Write the conversion equivalent under the problem. Be sure to place the 1 fl ~~z~~ under the 4 fl ~~z~~ and the 29.57 ml. under the x ml.

$$\begin{array}{l} 4 \text{ fl } \cancel{z} = x \text{ ml.} \\ 1 \text{ fl } \cancel{z} = 29.57 \text{ ml} \end{array}$$

Step 5. Draw a line between the fl ~~z~~'s and another between the ml's.

$$\begin{array}{l} 4 \text{ fl } \cancel{z} = x \text{ ml.} \\ \hline 1 \text{ fl } \cancel{z} = 29.57 \text{ ml.} \end{array}$$

Step 6. Cross multiply

$$1x = 118.28$$

Divide; by the number next to the "x"

$$\frac{1x}{1} = \frac{118.28}{1}$$

Step 7. Your answer, be sure to attach the proper "label" to it.

$$x = 118.28 \text{ ml.}$$

PRACTICE PROBLEMS

1. How many grams are in 246 grains?

2. How many ml are contained in ~~fl~~ 11?

3. Convert one dram and 20 minims to ml.

4. Convert 3 gallons, 1 pint, 10 fl~~z~~ to ml.

5. A formula for a cough syrup calls for $\frac{1}{8}$ gr of Codeine phosphate per fluid dram. How many grams would be used in preparing a pint of this cough syrup?

6. Convert $1/1000$ gr to mcg.
7. If fl ~~3~~ i of a cough syrup contains 10 gr of Sodium Citrate how many grams will it contain?
8. A prescription calls for $3/4$ gr of a medication, how many mg will be dispensed?
9. Convert 3 ~~10~~, 15 ~~3~~, 6 ~~3~~ to grams.
10. How many 500 mg doses could be obtained from $3/4$ lb of a drug?

Convert From The Metric System To The Common SystemsEXAMPLE: Convert 324 mg. to gr.

Step 1. Identify the problem;
in this case it is to
convert 324mg. to gr.
Write it down, use an
"x" for the unknown:

$$324 \text{ mg.} = x \text{ gr.}$$

Step 2. Choose a conversion
equivalent that possesses
both the denominations
present in the problem.
In this case use;
64.8 mg. = 1 gr:

Step 3. Write the conversion
equivalent under the
problem. Be sure to
place 64.8 mg under
the "x" gr:

$$\begin{array}{l} 324 \text{ mg.} = x \text{ gr.} \\ 64.8 \text{ mg} = 1 \text{ gr.} \end{array}$$

Step 4. Draw a line between the
two mg's and another between
the two gr's:

$$\frac{324 \text{ mg.}}{64.8 \text{ mg.}} = \frac{x \text{ gr.}}{1 \text{ gr.}}$$

Step 5. Solve by the ratio and
proportion method.

$$\frac{324 \text{ mg.}}{64.8 \text{ mg.}} = \frac{x \text{ gr.}}{1 \text{ gr.}}$$

Cross multiply:

$$64.8x = 324$$

Divide by the number
next to the "x":

$$x = \frac{324}{64.8}$$

Step 6. Your answer, be sure
to attach the proper
"label" to it:

$$x = 5 \text{ gr.}$$

PRACTICE PROBLEMS

1. Convert 250 ml to fluid ounces.
2. Convert 4.5 liters to fluid ounces.
3. How many mg are there in 6 1/2 ~~g~~ ?
4. Convert 6.6 pounds to kilograms.

5. How many 6.5 mg tablets can be obtained from 1/2 ounce (Apoth) of a chemical?
6. If a mixture weighing 30 grams is divided into 100 doses, how many grains will each dose weigh?
7. How many 1/8 gr tablets can be made from 3 grams of drug?
8. A certain drug is available in 16.2 mg tablets. Express this as a fraction of a grain.
9. How many teaspoonfuls are there in 0.5 Kiloliters and 500 milliliters?

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CALCULATION OF DOSES

PROCEDURES

Everytime you fill a prescription you must determine many things within a few minutes. Has the doctor prescribed enough medication or the right strength medication or could this prescription be for a child, how much would he get? In many instances the Physician will leave the variables for you to calculate.

Instructions

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

DEFINITION: A DOSE is the amount of preparation a patient takes at one time.

The formulas used in calculating

1. The number of dose

$$\text{Number of dose} = \frac{\text{Total preparation}}{\text{Size of each dose}}$$

2. The size of each dose

$$\text{Size of each dose} = \frac{\text{Total preparation}}{\text{Number of doses}}$$

3. The total preparation

$$\text{Total preparation} = \text{Number of doses} \times \text{each dose}$$

The formulas used in calculating children's dosages

1. Young's rule

$$\frac{\text{Age in years}}{\text{Age in years} + 12} \times \text{Adult dose} = \text{Child's dose}$$

2. Clark's rule

$$\text{Child's dose} = \frac{\text{Weight in pounds} \times \text{adult dose}}{150}$$

Calculation Of The Number Of Doses In A Preparation

EXAMPLE: Find the number of doses in 120 ml. if each dose is one teaspoonful (5 ml.).

Step 1. Write the complete formula:

$$\# = \frac{\text{Total}}{\text{Size}}$$

Step 2. Assign values to the appropriate terms:

Total = 120 ml.
Size = 5 ml.
= X

Step 3. Rewrite the formula, substituting the assigned values for the terms:

$$X = \frac{120}{5}$$

Step 4. Solve by the process indicated:

$$X = \frac{120}{5}$$

Step 5. Your answer:

$$X = 24 \text{ doses}$$

PRACTICE PROBLEMS

PRACTICE PROBLEMS

1. How many 15 minim doses are contained in 60 ml of a tincture?

2. If 180 ml of medicine is to be taken and each dose contains 2 tablespoonfuls, how many doses will this 180 ml contain?

3. How many 250 mgm doses can be obtained from one-half ounce (apoth) of a chemical?

4. The physician prescribes 8 fluid ounces (apoth) of Penicillin to be take in 10 ml doses. How many doses will the patient receive?

5. How many $\frac{3}{4}$ ss doses could you get from one pound (apoth) of a drug?

Calculate The Size Of Each Dose

EXAMPLE: What is the size of each dose if a patient is given 300 ml. and instructed to take the medicine once daily for 20 days?

Step 1. Write the complete formula:

$$\text{Size} = \frac{\text{Total}}{\#}$$

Step 2. Assign values to the appropriate terms:

$$\begin{aligned} \text{Total} &= 300 \\ \# &= 20 \end{aligned}$$

Step 3. Rewrite the formula, Substituting the assigned values for the terms.

$$\text{Size} = \frac{300}{20}$$

Step 4. Solve by the process indicated.

$$\text{Size} = \frac{300}{20}$$

Step 5. Your Answer:

$$\text{Size} = 15 \text{ ml. or } 1 \text{ tablespoonful}$$

PRACTICE PROBLEMS

1. What is the dose a patient will take if he receives 3 grams and is told to take it four times a day?

2. 20 doses are to be obtained from $\frac{3}{4}$ lbs of a chemical. How many mg is each dose?

3. 40 grams of a drug are to be divided into 500 doses. What is the strength of each dose?

4. One pound (Apoth) of chemical will make 60 doses. How many mg will each dose contain?

5. 6 fl $\frac{3}{4}$ are to be divided into 20 doses. How many ml will each dose be?

CALCULATE THE TOTAL AMOUNT OF A PREPARATION

EXAMPLE: How many ml. should be dispensed if the patient is to take 2 teaspoonfuls three times a day for one day?

Step 1. Write the complete formula:

$$\# = \frac{\text{Total}}{\text{Size}} \quad \text{Total} = \text{Size} \times \#$$

Step 2. Assign values to the appropriate terms:

$$\begin{aligned} \text{Size} &= 2 \text{ teasp.} = 10 \text{ ml} \\ \# &= 3 \text{ (doses)} \end{aligned}$$

Step 3. Rewrite the formula, substituting the assigned values for the terms:

$$\text{Total} = 10 \times 3$$

Step 4. Solve by the process indicated:

$$\text{Total} = 10 \times 3$$

Step 5. Your Answer:

$$\text{Total} = 30 \text{ ml.}$$

PRACTICE PROBLEMS

1. The prescription calls for the patient to take one teaspoonful four times a day for ten days. How many ml will you dispense?
2. The dose is one tablespoonful every six hours for one week. How many ml will you dispense?
3. The patient uses $\frac{3}{4}$ ii of a powder three times a day for soaks. He is to use this for 12 days. How many grams will be dispensed?
4. The patient will take 350 mg in each dose six doses a day for 14 days. How many total grams will be received?
5. 0.3 mg is the dose to be taken daily for 30 days. How many grams will you dispense?

Calculate the Dose Of A Drug When Given The Patients Weight and The Amount Of Drug Required Per Kilogram Of Body Weight.

EXAMPLE: The dose of a drug is 10 mg./1 Kg. How much should a patient weighing 154 lb. take?

Step 1. Convert the patient's weight in Kg: $154 \text{ lb} = 70 \text{ Kg.}$

Step 2. Write down the given dose: $10 \text{ mg} / 1 \text{ Kg.}$

Step 3. Write the patient's weight in Kg. under the 1 Kg. Then write "x" mg under the 10 mg: $\begin{array}{l} 10 \text{ mg.} / 1 \text{ Kg.} \\ x \text{ mg.} / 70 \text{ Kg.} \end{array}$

Step 4. Draw a line between the mg's and another between the Kg's: $\begin{array}{l} 10 \text{ mg.} / 1 \text{ Kg.} \\ \hline x \text{ mg.} / 70 \text{ Kg.} \end{array}$

Step 5. Solve by the ratio and proportion method: $\frac{10 \text{ mg.}}{x \text{ mg.}} = \frac{1 \text{ Kg.}}{70 \text{ Kg.}}$

Cross-multiply $1x = 700$

Divide by the number next to the "x": $x = \frac{700}{1}$

Step 6. Your Answer' $x = 700 \text{ mg}$

PRACTICE PROBLEMS

1. The patient weighs 190 pounds and the dose of the drug is 0.5 mg/Kilogram of body weight. How many mg will the patient take?

2. The average dose is 6.3 mg/Kilogram of body weight and the patient weighs 97 pounds. How many mg will she take?

3. The dose is $\frac{1}{4}$ gr./kilogram of body weight. The patient weighs 127 pounds. How many mg will he take?

4. The average dose is $\frac{1}{8}$ gr/Kilogram of body weight to be taken every six hours for 10 days. The patient weighs 81 Kilograms. How many total grams will the patient take?

Calculation Of Children's Doses Using Young's Rule

YOUNG'S RULE

$$\frac{\text{AGE IN YEARS}}{\text{AGE IN YEARS} + 12} \times \text{ADULT DOSE} = \text{CHILD'S DOSE}$$

EXAMPLE: How many mg. of a medication should a 4 year old child take if the adult dose is 250 mg?

Step 1. Write the complete formula.

$$\text{Child's Dose} = \frac{\text{Age}}{\text{Age} + 12} \times \text{Adult Dose}$$

Step 2. Assign values to the appropriate terms.

Age = 4
Adult Dose = 250
Child's Dose =

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$CD = \frac{4}{4 + 12} \times 250$$

Step 4. Solve by the processes indicated.

$$CD = \frac{4}{16} \times 250$$

Reduce fraction

$$CD = \frac{1}{4} \times 250$$

Multiply

$$CD = \frac{1}{4} \times \frac{250}{1}$$

Divide

$$CD = \frac{250}{4}$$

Step 5. Your answer

$$CD = 62.5 \text{ mg.}$$

PRACTICE PROBLEMS

1. If the usual adult dose of a drug is 0.25 Gm., what is the dose for a 9 year old child?
2. If the usual adult dose of a liquid medication is 5ml., how many ml. should a child 8 years old be given?
3. The adult dose is .6 Gm. How many mg. should a 2 year old child take?
4. A child of 10 years would take how many mg of a medication having the adult dose of 250 mg.
5. An adult would take a tablespoonful of this medication; how many ml. should a 5 year old take?
6. If the adult dose of a medication is 7 gr., how many milligrams should a 9 year old child take?

Calculation Of Children's Doses Using Clark's Rule

$$\text{CLARK'S RULE} \quad \text{CHILD'S DOSE} = \frac{\text{WEIGHT IN POUNDS} \times \text{ADULT DOSE}}{150}$$

EXAMPLE: An infant weighing 30 lbs will receive how many mg of a medication having an adult dose of 500 mg?

Step 1. Write the complete formula.

$$\text{Child's Dose} = \frac{\text{Weight} \times \text{Adult Dose}}{150}$$

Step 2. Assign values to the appropriate terms.

$$\text{Weight} = 30$$

$$\text{Adult Dose} = 500$$

$$\text{Child's Dose} = \text{ChD.}$$

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$\text{ChD} = \frac{30 \times 500}{150}$$

Step 4. Solve by the processes indicated.

$$\text{ChD} = \frac{30 \times 500}{150}$$

Multiply

$$\text{ChD} = \frac{30 \times 500}{150}$$

Divide

$$\text{ChD} = \frac{15000}{150}$$

Step 5. Your answer.

$$\text{ChD} = 100 \text{ mg}$$

PRACTICE PROBLEMS

1. The adult dose of a medication is 324 mg. How many mg will a 60 lb child take?
2. A child weighs 25 lbs. and is 18 months old. The adult dose is two tablespoonfuls. What is the child's dose?
3. A child weighing 83 lbs. would take how many mg if the adult dose is 5 gr?
4. An infant weighs 15 lbs. and the adult dose is 100 mg. What is the child's dose?
5. How many ml does a 55 lb child take if the adult dose is 2 teaspoonsful?
6. One Gram is the adult dose. How many mg does a 46 lb child take?

ADDITIONAL PRACTICE PROBLEMS FOR BASIC MATHEMATICAL OPERATIONS

1. Add: $\frac{5}{6} + \frac{1}{2} + \frac{1}{6} + \frac{1}{3} =$

Answer _____

2. Subtract: $\frac{7}{8}$ from 16 =

Answer _____

3. Divide $\frac{3}{10}$ by $\frac{1}{5} =$

Answer _____

4. Convert $\frac{2}{5}$ to a decimal fraction.

Answer _____

5. Add: $.15 + 3.14 + 13.25 + 0.034 =$

Answer _____

6. Multiply: 6.42×3.8

Answer _____

7. Convert .75 to a common fraction.

Answer _____

8. Write the following in arabic numbers :

Answer _____

xii = _____

ix = _____

xxvi = _____

MCMLX = _____

XLIX = _____

MXL = _____

9. Write the following as Roman numerals:

Answer _____

19 = _____

54 = _____

400 = _____

34 = _____

75 = _____

1970 = _____

10. Rearrange this formula to solve for C:

Answer _____

$$A = B \times C$$

11. Rearrange this formula to solve for B

Answer _____

$$A = \frac{B}{C}$$

12. Multiply: 2156 times 1.0023

Answer _____

13. Divide 1.01 by .98

Answer _____

ADDITIONAL PROBLEMS FOR THE METRIC SYSTEM

14. Convert the following to milligrams:

5 Grams = _____

50 decigrams = _____

10 micrograms = _____

3 centigrams = _____

.5 Grams = _____

15. Add the following and express your answer in Grams:

50 mg + 300 cg + 20 dg + 10 Gm =

Answer _____

16. Add the following and express your answer in milligrams:

.6 Gm + 0.25 cg + 0.125 Gm + 0.5 dg

Answer _____

ADDITIONAL PRACTICE PROBLEMS FOR THE METRIC SYSTEM

1. Convert the following to milligrams:

5 grams = _____

3 centigrams = _____

50 decigrams = _____

10 Dekagrams = _____

2. Add the following and express the answer in grams:

50 milligrams + 300 centigrams + 25 decigrams + 30 grams

Answer _____

3. Add the following and express the following in milligrams:

0.6 grams + 0.25 centigrams + 0.125 grams + 0.5 decigrams

Answer _____

4. Perform the following indicated problems:

Subtract 32 mg from 1.2 grams

Answer _____

Multiply 10 Kilograms X 8 and express the answer as grams.

Answer _____

Divide 45 Grams by 3.4 and express the answer in milligrams.

Answer _____

5. Restate the following:

125 mcg to milligrams = _____

85 deciliters to milliliters = _____

125 hectograms to centigrams = _____

6. Without reference write the prefixes of the metric system and what part or parts of the basic unit each represents.

ADDITIONAL PRACTICE PROBLEMS FOR THE APOTHECARY SYSTEM

1. Reduce the following to grains
 - a. $\mathfrak{z}$ ii, $\mathfrak{z}$ iss
 - b. $\mathfrak{z}$ iv $\mathfrak{z}$ iv gr iv

2. Restate the following in weighable Apothecary denominations.
 - a. 158 gr _____
 - b. 175 gr _____
 - c. 75 gr _____

3. Reduce the following to minims.
 - a. $\odot$ ii, fl $\mathfrak{z}$ v
 - b. qt i, $\odot$ ss, fl $\mathfrak{z}$ vii

4. Convert the following to fl $\mathfrak{z}$
 - a. mx 120, fl $\mathfrak{z}$ 16, $\odot$ iv
 - b. qt iii, $\mathfrak{z}$ viii, fl $\mathfrak{z}$ ii ss

5. Add: 1 $\overline{103}$ 2 $\overline{3}$ 3 $\overline{91}$ to 2 $\overline{103}$ 15 $\overline{3}$ 7 $\overline{31}$

Answer _____

6. Subtract: 2 Gal - 3 qt, 2 pt, 10 fl $\overline{3}$, 6 f $\overline{3}$

Answer _____

7. How many bottles, each containing fl $\overline{3}$ iv, can be obtained from $\overline{f3}$ ii of Iodine Tincture?

Answer _____

8. How many gr $\frac{1}{4}$ tablets can be made from $\overline{3}$ $\frac{1}{8}$ of Morphine Sulfate?

Answer _____

9. A cough syrup contains $\mathfrak{z}$ ss of ammonia chloride in fl $\mathfrak{z}$ iv. How many grains should be used in preparing one gallon of the syrup?

Answer _____

10. What is the volume in fluid ounces of a mixture containing $\frac{1}{2}$ gallon of one liquid, one pint of another and fl $\mathfrak{z}$ 96 of a third?

Answer _____

11. A pharmacist had $\frac{1}{2}$ gallon of alcohol. At different times he dispensed $\mathfrak{f}\mathfrak{z}$ iss, $\odot$ i, $\mathfrak{f}\mathfrak{z}$ iv. What volume was left?

Answer _____

ADDITIONAL PRACTICE PROBLEMS FOR THE AVOIRDUPOIS SYSTEM

1. How many $1/120$ grain tablets can be made from $1/8$ oz of a powder?

Answer _____

2. How much chemical is left in a $1\ 1/2$ oz bottle after enough has been taken out to make 1000 tablets of $1/100$ grain each?

Answer _____

3. How many $1/4$ gr capsules can you make from $1\ 1/4$ oz of a chemical?

Answer _____

4. How many 2 grain tablets could be made from 2 oz of Aspirin powder?

Answer _____

5. How many grains are left in a $1/4$ lb bottle after enough of it has been used to make 150 tablets, each containing $1/300$ gr.?

Answer _____

ADDITIONAL PRACTICE PROBLEMS FOR RATIO AND PROPORTION

1. If cold capsules were 12 for \$1.98, (b) 25 for \$3.25 and (c) 100 for \$10.95, which would be the best buy?

Answer _____

2. A formula for 1250 capsules calls for 3.25 Gms of a chemical. How much of the chemical would be used to make 350 capsules?

Answer _____

3. If 125 gallons of a mouth rinse contains 20 Grams of a coloring agent how many Grams will 160 gallons contain?

Answer _____

4. If 3 doses of a liquid preparation contain 7.5 grains of a substance, how many grains will 32 doses contain?

Answer _____

5. If 50 tablets contain 0.625 grams of an active ingredient, how many tablets can be prepared from 31.25 grams of the ingredient?

Answer _____

6. How many grains of a substance are needed for 350 tablets if 75 tablets contain 3 grains of the substance?

Answer _____

ADDITIONAL PRACTICE PROBLEMS FOR CONVERSION OF WEIGHT AND MEASURES

1. Convert 53 lb (AV) to Kg.

Answer _____

2. How many grains are in a .5 Gm tablet.

Answer _____

3. How many Kg do you weigh?

Answer _____

4. How many ml are there in 3 fl $\bar{z}$?

Answer _____

5. 1/200 gr is equivalent to how many mcg?

Answer _____

6. Convert 5000 ml to Apothecary units.

Answer _____

7. Compare an Apothecary grain to an Avoirdupois grain.

Answer _____

8. What is the difference, in grams, between an Apothecary pound and an Avoirdupois pound?

Answer _____

9. Convert 1 lb 2 oz (AV) to Apothecary units.

Answer _____

10. How many grains are there in 25 mcg?

Answer _____

11. A doctor orders a patient to take three $\frac{1}{8}$ gr. tablets per day.
How many mg. will this equal per day?

Answer _____

12. How many ml will the patient take daily?
Sig: Take fl $\frac{3}{4}$ daily

Answer _____

13. What directions will you give the patient for the prescription?
Sig: 2.5 ml daily

Answer _____

14. A doctor orders 12 fl $\frac{3}{4}$ be given to a patient. How many ml. will
you dispense?

Answer _____

15. How many ml are there in 20 gals?

Answer _____

16. A 4 fl $\frac{3}{4}$ prescription bottle will hold how many ml?

Answer _____

17. A 2 $\frac{3}{4}$ powder jar will hold how many grams?

Answer _____

18. Convert $\frac{1}{4}$ gr to mg.

Answer _____

19. Convert 1 qt to liters.

Answer _____

20. How many grains are in a .250 Gm tablet?

Answer _____

ADDITIONAL PRACTICE PROBLEMS FOR CALCULATION OF DOSES

1. How many doses will this prescription contain?

ETH 120 ml

Sig: $\overline{3}$ qid

Answer _____

2. How many doses will this prescription contain?

Tetracycline Tab #40 250 mg

Sig 500 mg qid

Answer _____

3. How many doses will this prescription contain?

Atarax Syrup 16 fl $\overline{3}$

Sig: $\overline{3}$ q 4h

Answer _____

4. What is the size of each dose in this prescription?

Kaopectate 1 pt

Sig: Divide equally into 32 doses

Answer _____

5. How many grams should you dispense for this prescription?

PenVK 125 mg

Sig: tab $\overline{\text{ii}}$ qid x 10 d

Answer _____

6. How many fl $\overline{\text{z}}$ should you dispense for this prescription?

Tetracycline Syrup

Sig: $\overline{\text{zi}}$ Tid for 2 weeks

Answer _____

7. How many Grams should you dispense for this prescription?

Valium 5 mg

Sig: 2.5 mg at bedtime for 5 days

Answer _____

8. The dose of a drug is $1/10$ gr. per Kg. of body weight. How many milligrams should be given to a person weighing 70 Kg.

Answer _____

9. The adult dose is 500 mg, how much should be given to a 50 pound child?

Answer _____

10. The adult dose is 30 ml. How much should be given to a 6 year old?

Answer _____

11. If the usual adult dose of a drug is 0.25 Gm what is the dose for a child 9 years old?

Answer _____

12. If the usual adult dose of paregoric is 5 ml what is the dose for a child 8 years old?

Answer _____

13. If the usual dose for an adult is .6 Gm what is the dose for a 2 year old child?

Answer _____

14. The usual dose of a certain solution is 0.5 ml. (a) what is the dose for a child 4 years old? Answer _____ (b) if the solution is to be dispensed in a dropper bottle, the dropper of which calibrates 24 drops per ml, how many drops should be given to obtain the correct dose for the child?

Answer _____

15. The usual dose of a drug is $\frac{1}{60}$ grain for an adult, (a) calculate the dose for a 25-lb child (b) the dose for an infant of 1 year and (c) the dose for a child weighing 50 lb.

Answer _____

Answer _____

Answer _____

16. The usual adult dose of a drug is 0.6 Gm. What is the dose for a child weighing 20 lbs? A child weighing 10 lbs?

Answer _____

Answer _____

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Technical Training

10-8

Pharmacy Specialist

PHARMACEUTICAL CALCULATIONS I

March 1976



SCHOOL OF HEALTH CARE SCIENCES, USAF
Department of Biomedical Sciences
Sheppard Air Force Base, Texas 76311

Designed for ATC Course Use

DO NOT USE ON THE JOB

PHARMACEUTICAL CALCULATIONS I

OBJECTIVES

- Solve problems pertaining to basic mathematical operations, metric system, apothecary system, avoirdupois system, and ratio and proportion.
- Solve problems pertaining to conversion of weights and measures, and calculations of doses.

INTRODUCTION

Systems of measurements in the past were based on traditional standards, such as the length of the King's foot or the weight of a grain of wheat. For their time and technology they were adequate. The need for a more exacting and universal system of measurement brought about the creation and standardization of the Metric System.

INFORMATION

MEASUREMENT

Definitions

Meter is the basic unit of length (39.37 inches)

Liter is the basic unit of volume. (The volume of the cube is 1/10 of a meter.)

Gram is the basic unit of weight (equal to the weight of one cubic centimeter of water at 4 degrees centigrade).

LATIN PREFIXES

deci is equal to 1/10 of the basic unit.

centi is equal to 1/100 of the basic unit.

milli is equal to 1/1000 of the basic unit.

micro is equal to 1/1,000,000 of the basic unit.

GREEK PREFIXES

deka is equal to ten times the basic unit.

hecto is equal to one hundred times the basic unit.

kilo is equal to one thousand times the basic unit.

Abbreviations

Meter = M

Liter = L

This supersedes WB 3ABR90530-I-1, March 1975

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Gram = Gm

deci = d

centi = c

milli = m

micro = mc

Deka = D

Hecto = H

Kilo = K

THE METRIC SYSTEM

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INSTRUCTIONS

Each type problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

This Study Guide Workbook is to help you learn the metric system. Follow the directions carefully and do not "skip around."

Directions: Read each problem carefully, then write in the answer: Be sure that you are satisfied with your answer before you write it in.

1. In the metric system, weight is expressed in grams, linear measurement is expressed in meters, and liquid volume is expressed in liters.
The system which uses grams, meters and liters is called the () system.
2. The primary units of measurements in the metric system are (), (), and ().
3. Which of the following units of measurement belong to the metric system? (Circle your answers below.)
a. pound d. liter
b. gram e. yard
c. gallon f. meter
4. The gram, which is a much smaller unit than our commonly used pound, is the basic metric unit used to measure (Circle your answer below.)
a. volume b. length c. weight
5. Length, in the household system, is measured in inches, feet, yards, etc. In the metric system, however, the primary unit for the measurement of length is the meter.
With the metric system, length is measure in ().
6. When using the metric system to measure the length of an item, you would record its length as so many ().
7. In the metric system, the primary unit of weight is the (); the primary unit of length is the ().
8. The primary metric unit of measurement used to measure volume is the liter.
Which of the following is used to measure volume in the metric system? (Circle your answer below.)
a. pounds d. grams
b. gallons e. liters
c. meters f. inches
9. In the common household system, pints, quarts and gallons are used to measure volume. In the metric system, however, the primary unit used to measure volume is the ().
10. When items are weighed by the metric system, their weight is expressed in ().

11. The length of an item measured by the metric system is expressed in ().
12. The volume of liquids measured by the metric system is expressed in ().
13. When metric measurements are written, the amount is written as a numeral followed by the unit. Study these examples:

Four meters is written as 4 meters.
 Four liters is written as 4 liters.
 Twelve grams is written as 12 grams.

Now write the following measurements:

- a. Four grams ()
- b. Eight liters ()
- c. Nine meters ()
14. If the measurement contains a fraction, the fraction is written as a decimal. Study these examples:

4 $\frac{1}{4}$ meters is written as 4.25 meters.
 4 $\frac{3}{4}$ liters is written as 4.75 liters.
 4 $\frac{1}{8}$ grams is written as 4.125 grams.

Now, write the following measurements.

- a. Five and one-half grams ()
- b. Three and one-fourth meters ()
- c. Four and three-quarter liters ()
15. Write the primary metric unit used to measure weight, length and volume.
 - a. weight ()
 - b. length ()
 - c. volume ()
16. You should also know the abbreviations for the three basic metric units of measurement. Abbreviations of the basic units are always capitalized.
17. The abbreviation for gram(s) is Gm.
 Using the abbreviation, write 12 grams. ()
18. The abbreviation for meter(s) is M.
 Using the abbreviation, write 2 meters. ()
19. The abbreviation for liter(s) is L.
 Using the abbreviation, write 1 liter. ()
20. Using abbreviations, write:
 200 liters () 17 meters () 16 grams ()

21. In addition to the three basic units you have just studied, the metric system has other units which are subdivisions of the basic units. Let us now study some of those subdivisions which are frequently used.

22. The common subdivision of the gram is the milligram (.001 of a gram). The abbreviation for the milligram is mg.

Using the abbreviation, write 12 milligrams. ()

23. When the prefix milli (m) is used with a basic unit (Gm., L., etc.) and the figure is less than 1000, the amount expressed is less than the basic unit.

Example: 500 mg = .5 of a gram
 250 ml = .25 of a liter
 700 mm = .7 of a meter

24. When the prefix milli (m) is used with a basic unit and the figure is greater than 1000, the amount expressed is more than the basic unit.

Example: 1,500 mg = 1.5 grams
 2,500 ml = 2.5 liters
 1,700 mm = 1.7 meters

Complete the following:

- a. 350 milligrams = () grams
- b. 2,300 milliliters = () liters
- c. 1,800 milligrams = () grams
- d. 300 millimeters = () meters
- e. 450 milliliters = () liters

25. A meter may be divided into 100 parts; each part then is one centimeter (.01 of a meter). The abbreviation for centimeter is cm. The abbreviation for cubic centimeter is cc.

Using the abbreviation, write 1 centimeter. ()

Using the abbreviation, write 4 cubic centimeters. ()

26. Using the abbreviation, write 500 cubic centimeters. ()

Using the abbreviation, write 400 cubic centimeters. ()

27. The common subdivision of the liter is the milliliter, or .001 of a liter. The abbreviation for milliliter is ml.

Using the abbreviation, write 200 milliliters. ()

28. Using the abbreviation, write 4 milliliters. ()

29. Write the abbreviations for meter (), gram (),
 liter (), cubic centimeter (), milliliter
 (), milligram (), and centimeter ().
30. Using the correct abbreviations, rewrite each of the following:
- 15 cubic centimeters ()
 - 10 grams ()
 - 9 milligrams ()
 - 5 liters ()
 - 1 cubic centimeter ()
 - 17 milliliters ()
 - 14 centimeters ()
31. Just as it has subdivisions to express measurements less than the primary units, the metric system also has units to express measurements larger than the primary units. Those larger units are expressed by the prefix kilo which means 1,000. For example, 1 kilometer = 1,000 meters, 1 kilogram = 1,000 grams, and 1 kiloliter = 1,000 liters. The prefix that means 1,000 is ().
32. The abbreviation of kilogram is Kg, kilometer is Km, and kiloliter is Kl. Abbreviations of prefixes whose values are larger than the basic units: (Circle your answer below.)
- are capitalized
 - are not capitalized
33. A length of 5,000 meters expressed in kilometers would be written as 5 ().
34. An object that weighs 1 kilogram weighs how many grams? ()
35. As you have already learned, an item which is shorter, or which weighs less than the primary unit may be expressed by the prefix milli. A milligram is .001 of a gram. How many milligrams are required to make up one gram? (Circle your answer below.)
- 10
 - 100
 - 1,000
 - 10,000
36. A kiloliter is equal to () liters. A milliliter is equal to what part of a liter? ()
37. To express 1,000 grams, 1,000 liters and 1,000 meters, you may use the same prefix, which is ()
38. To express .001 of a gram, .001 of a liter and .001 of a meter, you may use the prefix ().
39. As you recall, 1 milliliter is used to express .001 of a liter. Another way to express that same amount is 1 cubic centimeter, abbreviated 1 cc. This is true because 1 cc occupies the same space and has the same volume as 1 milliliter.
- One cc is () one milliliter.

40. Do not get the two prefixes confused. Remember that the prefix milli means .001; the prefix centi means .01.

In the spaces below, write five cubic centimeters and eight centimeters using abbreviations.

() ()

41. To convert grams to milligrams, multiply the number of grams by 1,000 or move the decimal three places to the right.

Example: 0.15 Gm = 150 mg

$$\begin{array}{r} 0.15 \\ \times 1,000 \\ \hline 150.00 \end{array}$$

42. Convert 2.5 grams to milligrams. ()

43. To convert milligrams to grams, divide the number of grams by 1,000 or move the decimal three places to the left.

Example: 150 mg = 0.15 Gm

$$\begin{array}{r} 0.15 \\ 1000 \overline{) 150.00} \\ \underline{100.0} \\ 50.00 \\ \underline{50.00} \\ 0 \end{array}$$

850 mg = () Gm

44. Now that you know the prefixes, work the following problems for practice. Check your responses.

- a. 500 milligrams is the same as () grams.
- b. 2 grams is the same as () milligrams.
- c. 500 centigrams is the same as () milligrams.
- d. 350 milligrams is the same as () centigrams.
- e. 250 milliliters is the same as () liters.
- f. 180 liters is the same as () milliliters.
- g. 420 millimeters is the same as () meters.
- h. 3.5 meters is the same as () millimeters.
- i. 500 kilograms is the same as () grams.
- j. 4,500 grams is the same as () kilograms.
- k. 1 kilogram is the same as () centigrams.
- l. 2,500 centiliters is the same as () kiloliters.
- m. 3.5 kiloliters is the same as () liters.
- n. 1.6 meters is the same as () kilometers.

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Answers for Metric System

1. metric
2. grams, meters, liters
3. gram; liter; meter
4. weight
5. meters
6. meters
7. gram, meter
8. liters
9. liter
10. grams
11. meters
12. liters
13. 4 grams; 8 liters; 9 meters
14. 5.5 grams; 3.25 meters; 4.75 liters
15. grams; meters; liters
16. No response
17. 12 Gm.
18. 2M
19. 1 L.
20. 200 L; 17M; 16 Gm
21. No response
22. 12 mg
23. No response
24.
 - a. .35
 - b. 2.3
 - c. 1.8
 - d. .3
 - e. .45
25. 1 cm; 4 cc
26. 500 cc; 400 cc
27. 200 ml
28. 4 ml
29. M; Gm; L; cc; ml; mg; cm
30.
 - a. 15 cc
 - b. 10 Gm
 - c. 9 mg
 - d. 5 L
 - e. 1 cc
 - f. 17 ml
 - g. 14 cm
31. kilo
32. are; are not
33. kilometers
34. 1,000
35. 1,000
36. 1,000; .001
37. kilo
38. milli
39. equal to (or same as)
40. 5 cc, 8 cm
41. No response
42. 2.5 Gm.
$$\frac{1000}{2500.0} = 2500 \text{ mg.}$$
43. .850
44.
 - a. .5 grams
 - b. 2,000
 - c. 5,000
 - d. 35
 - e. .25
 - f. 180,000
 - g. .42
 - h. 3,500
 - i. 500,000
 - j. 4.5
 - k. 100,000
 - l. .025
 - m. 3.500
 - n. .0016

APOTHECARY SYSTEM

One of the oldest system of weights and measures is the Apothecary system and although antiquated and no longer official it is still used extensively in medicine. Therefore your complete comprehension is necessary.

APOTHECARY TABLE OF WEIGHTS

20 grains	1 scruple
3 scruples	1 drachm
8 drachms	1 ounce
12 ounces	1 pound

APOTHECARY TABLE OF FLUID MEASURE (VOLUME)

60 minims	1 fluidrachm
8 fluidrachms.	1 fluidounce
16 fluidounces.	1 pint
2 pints.	1 quart
4 quarts	1 gallon

Definition of Apothecary Symbols

Minim	mx
Fluidrachm.	fl ℥
Fluidounce.	fl ℥
Pint.	pt. ①
Quart	qt
Gallon	C or Cong
Grain	gr
Scruple	℥
Drachm, Dram.	℥
Ounce	℥
Pound	℔

INSTRUCTIONS

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

Restate To A Lower Denomination In The Apothecary System

E Example: Reduce 3 fl $\bar{z}$ 2 fl $\bar{z}$ to mx

Step 1. Copy the value from the problem carefully, 3 fl $\bar{z}$ 2 fl $\bar{z}$

NOTE: Each value will be reduced separately.

Step 2. First reduce the 2 fl $\bar{z}$.

How many mx are in each
fl $\bar{z}$? _____ How many
fl $\bar{z}$ are you reducing? _____
Multiply 60 times 2 to
find the number of mx in
2 fl $\bar{z}$.

$$\begin{array}{r} 60 \\ \times 2 \\ \hline 120 \end{array}$$

Step 3. Rewrite the problem using
120 mx for the 2 fl $\bar{z}$.

3 fl $\bar{z}$ 120 mx

Step 4. Now reduce the 3 fl $\bar{z}$.

How many mx are in each
fl $\bar{z}$? _____ How
many fl $\bar{z}$ are you chang-
ing? _____ Multiply 480
times 3 to find the number
of mx in 3 fl $\bar{z}$.

$$\begin{array}{r} 480 \\ \times 3 \\ \hline 1440 \end{array}$$

Step 5. Rewrite the problem using
1440 mx for the 3 fl $\bar{z}$.

Step 6. Add up the mx and your answer is: 1560 mx

Practice Problems

1. Convert the following to a lower denomination in the Apothecary System.

a. Cong ii, pt ii, fl $\bar{z}$ ii to fl $\bar{z}$

b. $\bar{z}$ xvi, $\bar{z}$ xxxii to gr

c. $\odot$ iii, mx 480 to fl $\bar{3}$

c. $\bar{10}$ ss, $\bar{3}$ xxiv to $\bar{3}$

Restate To A Lower Denomination In The Apothecary System

Example: Change 5840 gr to weighable units.

Step 1. Copy the values from the problem carefully.

5840 gr

Step 2. Study this number. What is the largest unit that this could be changed to?
How many grains does this unit contain?

Step 3. Now to find the number of pounds, divide 5760 into 5840. The number of pounds in 5840 gr is _____ and the number of grains left over is 80. Now the amount is rewritten using the pound and grains.

$$\begin{array}{r} 1 \\ 5760 \overline{) 5840} \\ \underline{5760} \\ 80 \text{ gr left} \end{array}$$

1 lb. 80 gr.

Step 4. Study the 80 gr. What is the largest unit that this amount can be changed to? _____ How many grains does 1 dram contain? _____

Step 5. Now to find the number of drams, divide 60 into 80. The number of drams in 80 gr. is _____ and the number of grains left over is _____.

$$\begin{array}{r} 1 \\ 60 \overline{) 80} \\ \underline{60} \\ 20 \text{ gr left} \end{array}$$

Step 6. Now rewrite the problem again using the pound, dram and grains.

1 lb 1 $\bar{3}$ 20 gr

Step 7. Study the 20 gr. What is the largest unit that this can be changed to?
How many grains
 does this unit have in it? _____

Step 8. Now to find the number of scruples in 20 gr, divide 20 into 20. The number of scruples is _____.

Step 9. Rewrite the problem using the number of scruples. Your answer:

1 lb 1 $\bar{3}$ 13

Practice Problems

1. Convert the following to weighable Apothecary units.

a. 3440 gr.

b. 1650 gr.

c. 950 gr.

d. 695 gr.

AVOIRDUPOIS SYSTEM

The Avoirdupois System is the official system of commerce and you are indirectly related to commerce in ordering bulk drugs through medical supply channels. You must have a complete understanding of this system to facilitate transactions with supply.

AVOIRDUPOIS TABLE OF WEIGHT

437.5 grains 1 ounce
16 ounces 1 pound

Definition Of Avoirdupois Symbols

grain gr
ounce oz
pound lb

To restate Avoirdupois units to a higher or lower denomination, follow the procedure as you used in restating within the Apothecary System.

INSTRUCTIONS

Each type problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete problems assigned. SHOW ALL WORK!

Practice Problems

1. Reduce the following to weighable Avoirdupois denominations.

- a. 7540 gr _____
- b. 1560 gr _____
- c. 856 gr _____
- d. 466 gr _____

2. How many 10 grain capsules can be made from 1/2 lb of iron crystals?

3. How many 5 grain capsules of aspirin can be made from 4 oz of aspirin powder?

4. How many $\frac{1}{2}$ gr tablets of codeine can be made from $\frac{1}{8}$ oz of codeine powder?

5. How many grains of chemical are left in a 1 oz bottle after enough of it has been used to make 2000 tablets each containing $\frac{1}{200}$ grain of the chemical?

1B- 76-1248

RATIO AND PROPORTION

If it were possible to choose the most useful method of solving mathematical problems, ratio and proportion would probably be selected. Nearly 80 percent of the problems you will encounter in Pharmacy can be solved using this method.

Definitions

A ratio is the numerical comparison of two similar quantities.

A proportion is a statement of the equality of two ratios.

INSTRUCTIONS

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK.

Solving Problems Using Ratio and Proportion

Example: How many feet per second is a car traveling at 90 mph, if at 60 mph it is traveling 88 feet per second?

Step 1. Read the question. Determine what is asked (the number of feet per second at 90 mph) and call this the "problem." Now determine what information is given (88 feet per second at 60 mph).

Step 2. Write the "problem" on one line, using "X" for the unknown.

90 mph X ft/sec

Step 3. Write the given information on the line under the problem. Be sure to place 60 mph under the 90 mph (the first ratio) and the 88 ft/sec under the "X" ft/sec (the second ratio).

90 mph X ft/sec
60 mph 88 ft/sec

Step 4. Now draw a line between the 90 mph and the 60 mph and another line between the "X" ft/sec and the 88 ft/sec. Then place an equal sign in the center.

$\frac{90 \text{ mph}}{60 \text{ mph}} = \frac{X \text{ ft/sec}}{88 \text{ ft/sec}}$

Step 5. Cross multiply (90 times 88 and 60 times "X"), giving the products. (Note: the ft/sec and the mph are not used here.)

$60 \cdot X = 90 \cdot 88$
 $60 X = 7920$

Step 6. Divide by the number next to the "X."

$X = \frac{7920}{60}$

Step 7. Your answer
(Note: the ft/sec is placed next to
the answer because "X" is the number
of ft/sec.)

$X = 132 \text{ ft/sec}$

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Practice Problems

Make valid ratios between these quantities.

1. 1 yard and 2 feet _____
2. 4 hours and 120 minutes _____
3. 2 feet and 6 inches _____
4. 100 Grams and 10 Kilograms _____
5. Butter sells 3 lb. for 98¢. How much will 2 lb. cost?
6. A drug costs \$6.98 for 12 ounces. How much will three and $\frac{3}{4}$ ounces cost the pharmacist?

7. Twenty gallons of gasoline will run your car 235 miles. How far should you go on six and $\frac{1}{2}$ gallons?

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8. The airliner travels 600 mph and you will fly 1230 miles. How long will your trip be?

9. The item sells for \$4.25 a dozen and you only have \$2.00. How many can you buy?

CONVERSION OF WEIGHTS AND MEASURES

Even though AFM 168-4 states that all prescriptions should be written in the Metric System, some physicians will continue to write in one of the other systems. The responsibility will rest on you to convert these prescriptions to the Metric System.

CONVERSION EQUIVALENTS:

64.8 mg = 1 gr.
 1 Gm = 15.432 gr.
 31.1 Gm = 1 ounce (Apoth)
 28.35 Gm = 1 oz (Av)
 454 Gm = 1 lb (Av)
 1 Kg = 2.2 lb (Av)
 1 ml = 16.23 minim
 29.57 ml = 1 fl ounce (Apoth)
 473 ml = 1 pint

COMMON EQUIVALENTS:

1 Teaspoonful = 1 dram = 5 ml
 1 Tablespoonful = 1/2 fl ounce = 15 ml

NOTE: The common equivalents are used only when interpreting prescriptions.

INSTRUCTIONS

Each type of problems you may encounter will be explained by the instructor. Fill in each blank in the example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

Convert From The Common Systems To The Metric System

Example: Convert 4 fl $\bar{z}$ to ml.

Step 1. Identify the problem; in this case it is to convert 4 fl $\bar{z}$ to ml. Write the problem down, use an "X" for the unknown.

$$4 \text{ fl } \bar{z} = x \text{ ml.}$$

Step 2. Be sure that the common system quantity is in one denomination. Make any changes now.

$$4 \text{ fl } \bar{z} = x \text{ ml.}$$

Step 3. Choose a conversion equivalent that possesses both the denominations present in the problem. In this case use 1 fl $\bar{z}$ = 29.57 ml.

Step 4. Write the conversion equivalent under the problem. Be sure to place the 1 fl $\bar{z}$ under the 4 fl $\bar{z}$ and the 29.57 ml. under the x ml.

$$\begin{array}{l} 4 \text{ fl } \bar{z} = x \text{ ml.} \\ 1 \text{ fl } \bar{z} = 29.57 \text{ ml.} \end{array}$$

Step 5. Draw a line between the fl $\frac{3}{4}$'s and another between the ml's.

$$\frac{4 \text{ fl } \frac{3}{4}}{1 \text{ fl } \frac{3}{4}} = \frac{x \text{ ml.}}{29.57 \text{ ml.}}$$

Step 6. Cross multiply.

$$1x = 118.28$$

Divide by the number next to the "x"

$$\frac{1x}{1} = \frac{118.28}{1}$$

Step 7. Your answer. Be sure to attach the proper "label" to it.

$$x = 118.28 \text{ ml.}$$

Practice Problems

1. How many grams are in 246 grains?

2. How many ml are contained in fl $\frac{3}{4}$ li?

3. Convert one dram and 20 minims to ml.

4. Convert 3 gallons, 1 pint, 10 fl $\frac{3}{4}$ to ml.

5. A formula for a cough syrup calls for $\frac{1}{8}$ gr of codeine phosphate per fluid dram. How many grams would be used in preparing a pint of this cough syrup?

6. Convert $\frac{1}{1000}$ gr to mcg.

7. If fl $\frac{3}{4}$ i of a cough syrup contains 10 gr of sodium citrate, how many grams will it contain?

8. A prescription calls for $\frac{3}{4}$ gr of a medication. How many mg will be dispensed?

9. Convert 3 lb, 15 $\frac{3}{4}$, 6 $\frac{3}{4}$ to grams.

10. How many 500 mg doses could be obtained from $\frac{3}{4}$ lb of a drug?

Convert From The Metric System To The Common Systems

Example: Convert 324 mg. to gr.

Step 1. Identify the problem. In this case it is to convert 324 mg. to gr. Write it down. Use an "x" for the unknown.

$$324 \text{ mg.} = x \text{ gr.}$$

Step 2. Choose a conversion equivalent that possesses both the denomination present in the problem. In this case use:
64.8 mg. = 1 gr.

Step 3. Write the conversion equivalent under the problem. Be sure to place 64.8 mg under the "x" gr.

$$\begin{array}{l} 324 \text{ mg.} = x \text{ gr.} \\ 64.8 \text{ mg} = 1 \text{ gr.} \end{array}$$

Step 4. Draw a line between the two mg's and another between the two gr's.

$$\begin{array}{l} 324 \text{ mg.} = x \text{ gr.} \\ \hline 64.8 \text{ mg.} = 1 \text{ gr.} \end{array}$$

Step 5. Solve by the ratio and proportion method.

$$\begin{array}{l} 324 \text{ mg.} = x \text{ gr.} \\ \hline 64.8 \text{ mg.} = 1 \text{ gr.} \end{array}$$

Cross multiply.

$$64.8x = 324$$

Divide by the number next to the "x".

$$x = \frac{324}{64.8}$$

Step 6. Your answer. Be sure to attach the proper "label" to it.

$$x = 5 \text{ gr.}$$

Practice Problems

1. Convert 250 ml to fluid ounces.

2. Convert 4.5 liters to fluid ounces.

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3. How many mg are there in 6 1/2 ³?

4. Convert 6.6 pounds to kilograms.

5. How many 6.5 mg tablets can be obtained from 1/2 ounce (Apoth) of a chemical?

6. If a mixture weighing 30 grams is divided into 100 doses, how many grains will each dose weigh?

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7. How many $\frac{1}{8}$ gr tablets can be made from 3 grams of drug?

8. A certain drug is available in 16.2 mg tablets. Express this as a fraction of a grain.

9. How many teaspoonfuls are there in 0.5 kiloliters and 500 milliliters.

CALCULATION OF DOSES

Everytime you fill a prescription you must determine many things within a few minutes. Has the doctor prescribed enough medication or the right strength medication or could this prescription be for a child? How much would he get? In many instances the physician will leave the variables for you to calculate.

INSTRUCTIONS

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in example section as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

Definition

A dose is the amount of preparation a patient takes at one time.

Formulas Used In Calculating

1. The number of dose

$$\text{Number of dose} = \frac{\text{Total preparation}}{\text{Size of each dose}}$$

2. The size of each dose

$$\text{Size of each dose} = \frac{\text{Total preparation}}{\text{Number of doses}}$$

3. The total preparation

$$\text{Total preparation} = \text{Number of doses} \times \text{each dose}$$

Formulas Used In Calculating Children's Dosages

1. Young's rule

$$\frac{\text{Age in years}}{\text{Age in years} + 12} \times \text{Adult dose} = \text{Child's dose}$$

2. Clark's rule

$$\text{Child's dose} = \frac{\text{Weight in pounds} \times \text{adult dose}}{150}$$

Calculation Of The Number Of Doses In A Preparation

Example: Find the number of doses in 120 ml. if each dose is one teaspoonful (5 ml).

- Step 1. Write the complete formula.

$$\# = \frac{\text{Total}}{\text{Size}}$$

- Step 2. Assign values to the appropriate terms.

$$\begin{aligned} \text{Total} &= 120 \text{ ml.} \\ \text{Size} &= 5 \text{ ml.} \\ \# &= \end{aligned}$$

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$X = \frac{120}{5}$$

Step 4. Solve by the process indicated.

$$X = \frac{120}{5}$$

Step 5. Your answer:

$$X = 24 \text{ doses}$$

Practice Problems

1. How many 15 minim doses are contained in 60 ml of a tincture?

2. If 180 ml of medicine is to be taken and each dose contains 2 tablespoonfuls, how many doses will this 180 ml contain?

3. How many 250 mgm doses can be obtained from one-half ounce (Apoth) of a chemical?

4. The physician prescribes 8 fluid ounces (Apoth) of penicillin to be taken in 10 ml doses. How many doses will the patient receive?

5. How many $\frac{3}{4}$ ss doses could you get from one pound (Apoth) of a drug?

Calculate The Size Of Each Dose

Example: What is the size of each dose if a patient is given 300 ml. and instructed to take the medicine once daily for 20 days?

Step 1. Write the complete formula.

$$\text{Size} = \frac{\text{Total}}{\#}$$

Step 2. Assign values to the appropriate terms:

$$\begin{aligned} \text{Total} &= 300 \\ \# &= 20 \end{aligned}$$

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$\text{Size} = \frac{300}{20}$$

Step 4. Solve by the process indicated.

$$\text{Size} = \frac{300}{20}$$

Step 5. Your answer:

Size = 15 ml. or 1 tablespoonful

Practice Problems

1. What is the dose a patient will take if he receives 3 grams and is told to take it four times a day?

2. Twenty doses are to be obtained from $\frac{3}{4}$ iss of a chemical. How many mg is each dose?

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Calculate The Total Amount Of A Preparation

Example: How many ml. should be dispensed if the patient is to take 2 teaspoonfuls three times a day for one day?

Step 1. Write the complete formula.

$$\# = \frac{\text{Total}}{\text{Size}} \quad \text{Total} = \text{Size} \times \#$$

Step 2. Assign values to the appropriate terms.

$$\begin{aligned} \text{Size} &= 2 \text{ teaspoonful} = 10 \text{ ml} \\ \# &= 3 \text{ (doses)} \end{aligned}$$

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$\text{Total} = 10 \times 3$$

Step 4. Solve by the process indicated.

$$\text{Total} = 10 \times 3$$

Step 5. Your answer:

$$\text{Total} = 30 \text{ ml.}$$

Practice Problems

1. The prescription calls for the patient to take one teaspoonful four times a day for ten days. How many ml will you dispense?

2. The dose is one tablespoonful every six hours for one week. How many ml will you dispense?

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3. The patient uses $\frac{3}{4}$ of a powder three times a day for soaks. He is to use this for 12 days. How many grams will be dispensed?

4. The patient will take 350 mg in each dose six doses a day for 14 days. How many total grams will be received?

5. 0.3 mg is the dose to be taken daily for 30 days. How many grams will you dispense?

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3. The patient uses $\frac{3}{4}$ ii of a powder three times a day for soaks. He is to use this for 12 days. How many grams will be dispensed?

4. The patient will take 350 mg in each dose six doses a day for 14 days. How many total grams will be received?

5. 0.3 mg is the dose to be taken daily for 30 days. How many grams will you dispense?

Calculate The Dose Of A Drug When Given The Patient's Weight And The Amount Of Drug Required For Kilogram Of Body Weight

Example: The dose of a drug is 10 mg/1 Kg. How much should a patient weighing 154 lb. take?

- Step 1. Convert the patient's weight in Kg. $154 \text{ lb} = 70 \text{ Kg.}$
- Step 2. Write down the given dose. 10 mg/1 Kg
- Step 3. Write the patient's weight in Kg. under the 1 Kg. Then write "x" mg under the 10 mg. $\begin{array}{l} 10 \text{ mg./1 Kg.} \\ x \text{ mg./70 Kg.} \end{array}$
- Step 4. Draw a line between the mg's and another between the Kg's. $\frac{10 \text{ mg.}}{x \text{ mg.}} / \frac{1 \text{ Kg.}}{70 \text{ Kg.}}$
- Step 5. Solve by the ratio and proportion method. $\frac{10 \text{ mg.}}{x \text{ mg.}} = \frac{1 \text{ Kg.}}{70 \text{ Kg.}}$
- Cross multiply. $1x = 700$
- Divide by the number next to the "x." $x = \frac{700}{1}$
- Step 6. Your answer: $x = 700 \text{ mg}$

Practice Problems

1. The patient weighs 190 pounds and the dose of the drug is 0.5 mg/Kilogram of body weight. How many mg will the patient take?
2. The average dose is 6.3 mg/Kilogram of body weight and the patient weighs 97 pounds. How many mg will the patient take?

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3. The dose is $1/4$ gr/Kilogram of body weight. The patient weighs 127 pounds. How many mg will the patient take?

4. The average dose is $1/8$ gr/Kilogram of body weight to be taken every six hours for 10 days. The patient weighs 81 Kilograms. How many total grams will the patient take?

Calculation Of Children's Doses Using Young's Rule

Young's rule: $\frac{\text{Age In Years}}{\text{Age In Years} + 12} \times \text{Adult Dose} = \text{Child's Dose}$

Example: How many mg of a medication should a 4 year old child take if the adult dose is 250 mg?

Step 1. Write the complete formula.

$$\text{Child's Dose} = \frac{\text{Age}}{\text{Age} + 12} \times \text{Adult Dose}$$

Step 2. Assign values to the appropriate terms.

Age = 4
Adult Dose = 250
Child's Dose =

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$CD = \frac{4}{4 + 12} \times 250$$

Step 4. Solve by the processes indicated.

$$CD = \frac{4}{16} \times 250$$

Reduce fraction

$$CD = \frac{1}{4} \times 250$$

Multiply

$$CD = \frac{1}{4} \times \frac{250}{1}$$

Divide

$$CD = \frac{250}{4}$$

Step 5. Your answer:

$$CD = 62.5 \text{ mg.}$$

Practice Problems

1. If the usual adult dose of a drug is 0.25 Gm., what is the dose for a 9 year old child?

2. If the usual adult dose of a liquid medication is 5 ml., how many ml. should a child 8 years old be given?

3. The adult dose is .6 Gm. How many mg should a 2 year old child take?

4. A child of 10 years would take how many mg of a medication having the adult of 250 mg.

5. An adult would take a tablespoonful of this medication. How many ml should a 5 year old take?

6. If the adult dose of a medication is 7 gr., how many milligrams should a 9 year old child take?

Calculation Of Children's Doses Using Clark's Rule

Clark's rule: $\text{Child's Dose} = \frac{\text{Weight In Pounds} \times \text{Adult Dose}}{150}$

Example: An infant weighing 30 lbs will receive how many mg of a medication having an adult dose of 500 mg?

Step 1. Write the complete formula. $\text{Child's Dose} = \frac{\text{Weight} \times \text{Adult Dose}}{150}$

Step 2. Assign values to the appropriate terms.

Weight = 30
Adult Dose = 500
Child's Dose = ChD

Step 3. Rewrite the formula, substituting the assigned values for the terms.

$$\text{ChD} = \frac{30 \times 500}{150}$$

Step 4. Solve by the processes indicated.

$$\text{ChD} = \frac{30 \times 500}{150}$$

· Multiply

$$\text{ChD} = \frac{30 \times 500}{150}$$

Divide

$$\text{ChD} = \frac{15000}{150}$$

Step 5. Your answer:

$$\text{ChD} = 100 \text{ mg}$$

Practice Problems

1. The adult dose of a medication is 324 mg. How many mg will a 60 lb child take?

2. A child weighs 25 lbs. and is 18 months old. The adult dose is two tablespoonfuls. What is the child's dose?

3. A child weighing 83 lbs. would take how many mg if the adult dose is 5 gr?

4. An infant weighs 15 lbs. and the adult dose is 100 mg. What is the child's dose?

5. How many ml does a 55 lb child take if the adult dose is 2 teaspoonsful?

6. One Gram is the adult dose. How many mg does a 46 lb child take?

Additional Practice Problems For Basic Mathematical Operations

1. Add: $5/6 + 1/2 + 1/6 + 1/3 =$

Answer _____

2. Subtract: $7/8$ from 16 =

Answer _____

3. Divide: $3/10$ by $1/5 =$

Answer _____

4. Convert $2/5$ to a decimal fraction.

Answer _____

5. Add: $.15 + 3.14 + 13.25 + 0.034 =$

Answer _____

6. Multiply: $6.42 \times 3.8 =$

Answer _____

7. Convert .75 to a common fraction.

Answer _____

8. Write the following in arabic numbers:

xii = _____

ix = _____

xxvi = _____

MCMLX = _____

XLix = _____

MXL = _____

9. Write the following as Roman numerals:

19 = _____

54 = _____

400 = _____

34 = _____

75 = _____

1970 = _____

10. Rearrange this formula to solve for C:

$$A = B \times C$$

Answer _____

11. Rearrange this formula to solve for B:

$$A = \frac{B}{C}$$

Answer _____

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12. Multiply: 2156 times 1.0023

Answer _____

13. Divide 1.01 by .98

Answer _____

Additional Problems For The Metric System

14. Convert the following to milligrams:

5 Grams = _____

50 decigrams = _____

.0 micrograms = _____

3 centigrams = _____

.5 Grams = _____

15. Add the following and express your answer in Grams: Answer _____

 $50 \text{ mg} + 300 \text{ cg} + 20 \text{ dg} + 10 \text{ Gm} =$

16. Add the following and express your answer in milligrams: Answer _____

 $.6 \text{ Gm} + 0.25 \text{ cg} + 0.125 \text{ Gm} = 0.5 \text{ dg}$

Additional Practice Problems For The Metric System

1. Convert the following to milligrams:

5 grams = _____

3 centigrams = _____

50 decigrams = _____

10 Dekagrams = _____

2. Add the following and express the answer in grams: Answer _____

50 milligrams + 300 centigrams + 25 decigrams + 30 grams

3. Add the following and express the following in milligrams: Answer _____

0.6 grams + 0.25 centigrams + 0.125 grams + 0.5 decigrams

4. Perform the following indicated problems: Answer _____

Subtract 32 mg from 1.2 grams

Multiply 10 Kilograms X 8 and express the answer in grams. Answer _____

Divide 45 Grams by 3.4 and express the answer in milligrams. Answer _____

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5. Restate the following:

125 mcg to milligrams = _____

85 deciliters to milliliters = _____

125 hectograms to centigrams = _____

6. Without reference, write the prefixes of the metric system and what part or parts of the basic unit each represents.

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Additional Practice Problems For The Apothecary System

1. Reduce the following to grains:

- a. $\bar{z}$ ii, $\bar{z}$ iss
 b. $\bar{z}$ iv $\bar{z}$ iv gr iv

2. Restate the following in weighable Apothecary denominations:

- a. 158 gr _____
 b. 175 gr _____
 c. 75 gr _____

3. Reduce the following to minims:

- a. $\odot$ ii, fl $\bar{z}$ v
 b. qt i, $\odot$ ss, fl $\bar{z}$ vii

4. Convert the following to fl $\bar{z}$:

- a. mx 120, fl $\bar{z}$ 16, $\odot$ iv
 b. qt iii, fl $\bar{z}$ viii, fl $\bar{z}$ ii ss

5. Add: 1 $\bar{z}$ 2 $\bar{z}$ 3 gr 1 to 2 $\bar{z}$ 15 $\bar{z}$ 7 gr 3

6. Subtract:
2 Gal - 3 qt, 2 pt, 10 fl $\bar{3}$, 6 fl $\bar{3}$

Answer _____

7. How many bottles, each containing fl $\bar{3}$ iv, can be obtained from fl $\bar{3}$ ii of iodine tincture?

Answer _____

8. How many gr $\frac{1}{4}$ tablets can be made from $\bar{3}$ $\frac{1}{8}$ of morphine sulfate?

Answer _____

9. A cough syrup contains $\bar{3}$ ss of ammonia chloride in fl $\bar{3}$ iv. How many grains should be used in preparing one gallon of the syrup?

Answer _____

10. What is the volume in fluid ounces of a mixture containing $\frac{1}{2}$ gallon of one liquid, one pint of another and fl $\frac{3}{4}$ 96 of a third? Answer _____

11. A pharmacist had $\frac{1}{2}$ gallon of alcohol. At different times he dispensed f $\frac{3}{4}$ iss, $\odot$ i, f $\frac{3}{4}$ iv. What volume was left? Answer _____

Additional Practice Problems For The Avoirdupois System

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1. How many $\frac{1}{120}$ grain tablets can be made from $\frac{1}{8}$ oz of a powder? Answer _____

2. How much chemical is left in a $1 \frac{1}{2}$ oz bottle after enough has been taken out to make 1000 tablets of $\frac{1}{100}$ grain each? Answer _____

3. How many $\frac{1}{4}$ gr capsules can you make from $1 \frac{1}{4}$ oz of a chemical? Answer _____

4. How many 2 grain tablets could be made from 2 oz of aspirin powder? Answer _____

5. How many grains are left in a $\frac{1}{4}$ lb bottle after enough of it has been used to make 150 tablets, each containing $\frac{1}{300}$ gr? Answer _____

Additional Practice Problems For Ratio And Proportion

1. If cold capsules were (a) 12 for \$1.98, (b) 25 for \$3.25, and (c) 100 for \$10.95, which would be the best buy? Answer _____

2. A formula for 1250 capsules calls for 3.25 Gms of a chemical. How much of the chemical would be used to make 350 capsules? Answer _____

3. If 125 gallons of a mouth rinse contains 20 Grams of a coloring agent, how many Grams will 160 gallons contain? Answer _____

4. If 3 doses of a liquid preparation contain 7.5 grains of a substance, how many grains will 32 doses contain? Answer _____

5. If 50 tablets contain 0.625 grams of an active ingredient, how many tablets can be prepared from 31.25 grams of the ingredient? Answer _____

6. How many grains of a substance are needed for 350 tablets if 75 tablets contain 3 grains of the substance? Answer _____

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Additional Practice Problems For Conversion Of Weight And Measures

1. Convert 50 lb (AV) to Kg.

Answer _____

2. How many grains are in a .5 Gm tablet?

Answer _____

3. How many Kg do you weigh?

Answer _____

4. How many ml are there in 3 fl *3*?

Answer _____

5. 1/200 gr is equivalent to how many mcg?

Answer _____

6. Convert 5000 ml to Apothecary units.

Answer _____

7. Compare an Apothecary grain to an Avoirdupois grain.

Answer _____

8. What is the difference, in grams, between an Apothecary pound and an Avoirdupois pound?

Answer _____

9. Convert 1 lb 2 oz (AV) to Apothecary units.

Answer _____

10. How many grains are there in 25 mcg?

Answer _____

11. A doctor orders a patient to take three $\frac{1}{8}$ gr tablets per day. How many mg. will this equal per day?

Answer _____

12. How many ml will the patient take daily?
Sig: Take fl 3 daily

Answer _____

13. What directions will you give the patient for this prescription?
Sig: 2.5 ml daily

Answer _____

14. A doctor orders 12 fl 3 be given to a patient. How many ml. will you dispense?

Answer _____

15. How many ml are there in 20 gals?

Answer _____

16. A 4 fl $\frac{3}{4}$ prescription bottle will hold how many ml? Answer _____

17. A 2 $\frac{3}{4}$ powder jar will hold how many grams? Answer _____

18. Convert $\frac{1}{4}$ gr to mg. Answer _____

19. Convert 1 qt to liters. Answer _____

20. How many grains are in a .250 Gm tablet? Answer _____

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Additional Practice Problems For Calculation Of Doses

1. How many doses will this prescription contain?

Answer _____

ETH 120 ml

Sig: *3i* qid

2. How many doses will this prescription contain?

Answer _____

Tetracycline Tab 250 mg

#40

Sig: 500 mg qid

3. How many doses will this prescription contain?

Answer _____

Atarax Syrup

16 fl *3*

Sig: *3ii* q 4h

4. What is the size of each dose in this prescription? Answer _____

Kaopectate 1 pt

Sig: Divide equally into 32 doses

5. How many grams should you dispense for this prescription? Answer _____

PenVK 125 mg

Sig: tab $\overline{\text{ii}}$ qid x 10 d

6. How many fl $\overline{3}$ should you dispense for this prescription? Answer _____

Tetracycline Syrup

Sig: $\overline{3\text{t}}$ Tid for 2 weeks

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7. How many Grams should you dispense for this prescription?

Answer _____

Valium 5 mg

Sig: 2.5 mg at bedtime for 5 days

8. The dose of a drug is 1/10 gr per Kg. of body weight. How many milligrams should be given to a person weighing 70 Kg?

Answer _____

9. The adult dose is 500 mg. How much should be given to a 50 pound child?

Answer _____

10. The adult dose is 30 ml. How much should be given to a 6 year old?

Answer _____

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11. If the usual adult dose of a drug is 0.25 Gm, what is the dose for a child 9 years old?

Answer _____

12. If the usual adult dose of paregoric is 5 ml, what is the dose for a child 8 years old?

Answer _____

13. If the usual dose for an adult is .6 Gm, what is the dose for a 2 year old child?

Answer _____

14. The usual dose of a certain solution is 0.5 ml.
(a) What is the dose for a child 4 years old?

Answer _____

- (b) If the solution is to be dispensed in a dropper bottle, the dropper of which calibrates 24 drops per ml, how many drops should be given to obtain the correct dose for the child?

Answer _____

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15. The usual dose of a drug is $1/60$ grain for an adult.

(a) Calculate the dose for a 25 lb child.

Answer _____

(b) The dose for an infant of 1 year.

Answer _____

(c) The dose for a child weighing 50 lb.

Answer _____

16. The usual adult dose of a drug is 0.6 Gm.
What is the dose for a child weighing
20 lbs? A child weighing 10 lbs?

Answer _____

Answer _____

DEPARTMENT OF BIOMEDICAL SCIENCES

10-8

PHARMACY SPECIALIST

FUNDAMENTALS OF PHARMACY

May 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
SHEPPARD AIR FORCE BASE, TEXAS

Designed For ATC Course Use

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1B-75-1348

FUNDAMENTALS OF PHARMACY

OBJECTIVE

Given information pertaining to pharmaceutical organic chemistry complete questions in SW 3ABR90530-I-3 to classify organic compounds shown in the instructional guidance and describe their properties. Each day you will review the material presented in class, then answer the appropriate questions pertaining to that day's lecture.

EQUIPMENT

- Selected flipcharts
- Selected transparencies
- Overhead projector

PROCEDURE

Defining and Identifying Organic Chemistry and Compounds

The object of this lesson is to acquaint you with some basic fundamentals of pharmaceutical organic chemistry. Specifically, you will:

1. Identify
 - a. The definition of organic chemistry.
 - b. Three sources of organic compounds.
 - c. Selected properties of organic compounds.
 - d. Principles and types of covalent bonding.
 - e. Types of molecular formulas used in organic chemistry.
2. Identify and define the types of carbon atoms.

This supersedes WB 3ABR90530-I-3, April 1974.

QUESTIONS

1. _____ - The study of Carbon.
2. _____ - The study of compounds containing Carbon and Hydrogen and their derivatives.
3. _____ - Organic compounds obtained from plant and animal sources.
4. _____ - Compounds made entirely from raw elements or by chemical action of naturally occurring compounds to form different compounds.
5. _____ - A combination of natural substances and synthetic compounds.
6. _____ compounds are the most important source for pharmaceuticals.
7. The three sources of organic compounds are:
 - a.
 - b.
 - c.
8. General properties of organic compounds as compared to inorganic compounds.
 - a.
 - b.
 - c.
 - d.
 - e.
 - f.
 - g.
9. _____ - A type of chemical bonding in which each atom donates one or more valence electrons to be shared by the two.
10. _____ is the principle type of bonding found in organic compounds.

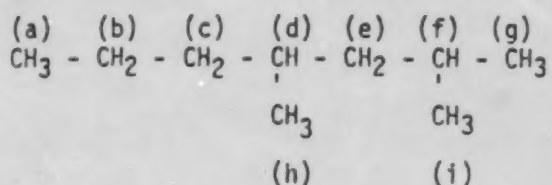
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11. _____ bonding between two carbon atoms signifies the sharing of two electrons.
12. C - C is an example of a _____ bond.
13. _____ bonding between two carbon atoms signifies the sharing of four electrons.
14. C = C is an example of a _____ bond.
15. _____ bonding between two carbon atoms signifies the sharing of six electrons.
16. C $\equiv$ C is an example of a _____ bond.
17. The three types of covalent bonding are:
- -
 -
18. _____ molecular formulas show the complete atomic relationship.
19. _____ molecular formulas show partial atomic relationship.
20. _____ molecular formulas show NO atomic relationship.
21. Identify the following types of molecular formulas.
- CH₃ - CH₂ - CH₃ _____
 - $$\begin{array}{ccccccc}
 & \text{H} & & \text{H} & & \text{H} & \\
 & | & & | & & | & \\
 \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{H}
 \end{array}$$

$$\begin{array}{ccccccc}
 & | & & | & & | & \\
 & \text{H} & & \text{H} & & \text{H} &
 \end{array}$$

 - C₃H₈ _____
22. The best way to show the complete atomic relationship of a molecule is by the use of the _____ formula.
23. _____ - Two or more compounds with the same empirical formula but different graphic structure and physical properties.

24. _____ Carbon Atoms have one of its valence electrons satisfied by another carbon atom.
25. _____ Carbon Atoms have two of its valence electrons satisfied by two other carbon atoms.
26. _____ Carbon Atoms have three of its valence electrons satisfied by three other carbon atoms.
27. The three types of carbon atoms are:
- -
 -
28. Name each type of carbon atom in the following illustration.



-
-
-
-
-
-
-
-
-

ALIPHATIC HYDROCARBONS

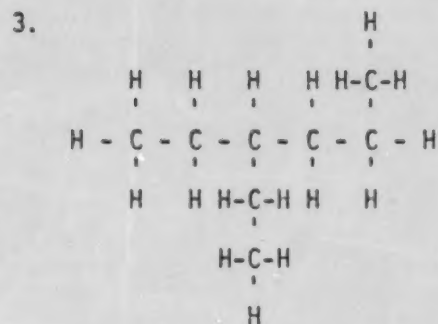
The purpose of this lesson is to acquaint you with the classification and properties of Aliphatic Hydrocarbons and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically you will:

1. Define
 - a. Hydrocarbon
 - b. Aliphatic Hydrocarbon
 - c. Radical
2. Identify
 - a. Classifications of Aliphatic Hydrocarbons
 - b. General Formulas
 - c. Selected properties
3. Identify selected pharmaceuticals belonging to these classes.
4. Using selected rules of the IUC System to name organic compounds.

QUESTIONS

1. _____ - Compounds which contain ONLY carbon and hydrogen.

2. _____ - Compounds which contain ONLY carbon and hydrogen and are formed in straight or branched open chains.



The above example is an _____.

4. Are Aliphatic Hydrocarbons cyclic in structure? YES or NO

5. _____ is the principle source of the Aliphatic Hydrocarbons.

6. _____ - A group that preserves its identity throughout a reaction.

7. List the General formulas for each of the following series, the

ALKANES

General Formula _____

General Properties _____

- a. _____
 b. _____
 c. _____
 d. _____
 e. _____
 f. _____

ALKANE (Methane) SERIES

 CH_4 _____ CH_3-CH_3 _____ $\text{CH}_3-\text{CH}_2-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_2-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_3-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_4-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_5-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_6-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_7-\text{CH}_3$ _____ $\text{CH}_3-(\text{CH}_2)_8-\text{CH}_3$ _____

ALKENES

General Formula _____

General Properties _____

- a. _____
 b. _____

ALKENE (Olefin) SERIES

 $\text{CH}_2=\text{CH}_2$ _____ $\text{CH}_2=\text{CH}-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-\text{CH}_2-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-(\text{CH}_2)_2-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-(\text{CH}_2)_3-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-(\text{CH}_2)_4-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-(\text{CH}_2)_5-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-(\text{CH}_2)_6-\text{CH}_3$ _____ $\text{CH}_2=\text{CH}-(\text{CH}_2)_7-\text{CH}_3$ _____

general properties of each and name the member of each series.

ALKYNES

General Formula _____

General Properties _____

a. _____

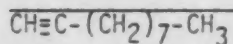
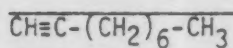
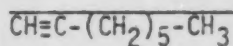
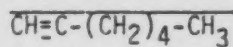
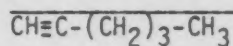
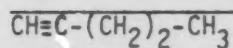
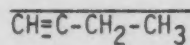
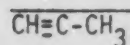
b. _____

RADICAL

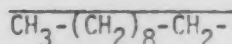
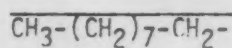
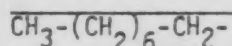
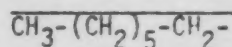
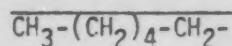
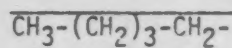
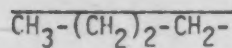
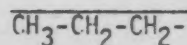
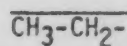
General Formula _____

General Properties _____

ALKYNE (Acetylene) SERIES



ALKYLS



8. _____ Series are compounds built of the same elements with similar chemical and physical properties varying regularly from member to member by a common amount (CH_2).

9. A _____ is a member within the series.

10. The _____ are the least reactive of the Aliphatic Hydrocarbons because of hydrogen saturation and single bonds.

11. The ANE ending identifies the _____ Series.

12. The _____ are the most reactive Aliphatic Hydrocarbons because of a triple bond.

13. The _____ are derived from the Alkanes and have double bonds.

14. In naming the radicals, the base name is derived from the Alkanes and the ending is changed to _____.

15. Three important pharmaceuticals belonging to the Aliphatic Hydrocarbon class.

a. Light Mineral Oil N.F. (Light Petrolatum)

Use: _____

b. Mineral Oil U.S.P. (Heavy Petrolatum)

Use: _____

c. White Petrolatum U.S.P. (Petrolatum or Vaseline)

Use: _____

16. The pharmaceutical that is never taken internally and is used in cosmetics is _____.

17. The Aliphatic Hydrocarbon which is used as an ointment base and is a pharmaceutical necessity is _____.

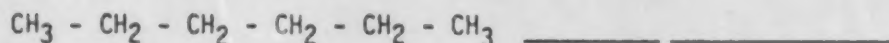
18. The pharmaceutical Aliphatic Hydrocarbon which is used as a laxative is _____.

The following questions are to be answered by using the IUC System in naming organic compounds.

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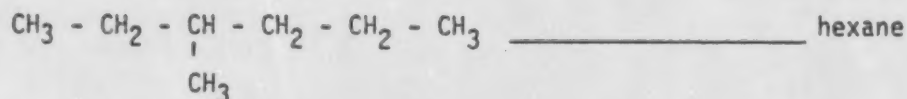
19. Identify the _____ continuous chain of carbon atoms in the formula. This chain will then be called the _____ and named by its respective _____ name.

20. What is the base name for the following compound?



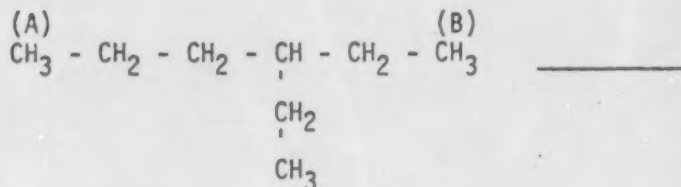
21. The names of the branched radicals are named as prefixes to the _____ name.

22. What is the name of the radical attached to the base of the following compound?



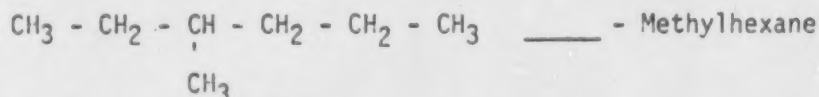
23. Number the _____ atoms from the end which will give the _____ their lowest number.

24. From which end of the following compound would you number?



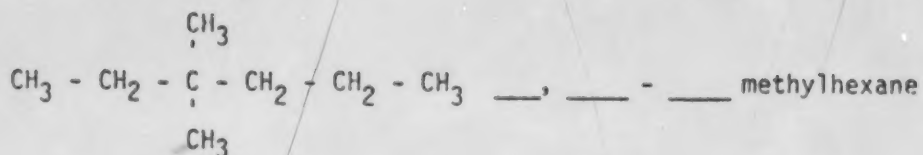
25. The position of the radical is indicated by the _____ of the carbon atom to which it is attached.

26. What would be the number that is attached to the following named compound?



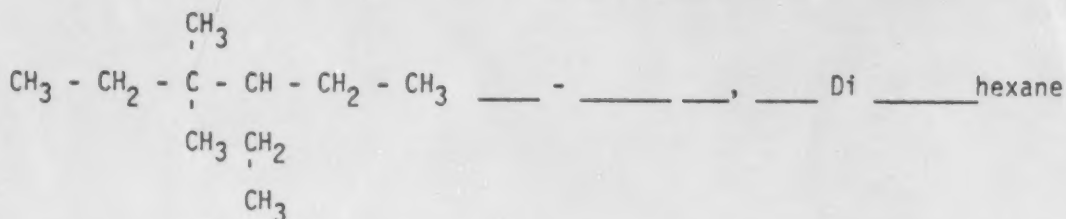
27. If _____ radicals are attached to the same carbon atom, the number of the carbon atom to which they are attached is repeated and the numerical prefix is added to the radical name.

28. Place the proper information in the blanks for the following compound.



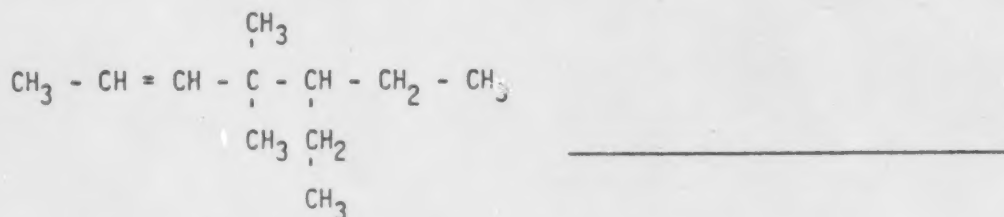
29. If _____ radicals are attached to either the same or different carbons, the position of each radical is indicated by the number of the carbon atom to which it is attached and named in alphabetical order.

30. Complete the blanks in naming the following compound.



31. If the compound contains a _____ bond, the position of the bond is indicated by the number of the carbon atom which contains the multiple bond closest to a terminal carbon atom, included in the name and the base will be given its respective Alkene or Alkyne name.

32. Name the following compound.



INFORMATION

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ALCOHOLS, ALDEHYDES, KETONES AND ETHERS

The purpose of this lesson is to acquaint you with the properties and uses of the alcohols, aldehydes, ketones, and ethers and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and

1. Define and give general formulas for the following:
 - a. Alcohols
 - b. Aldehydes
 - c. Ketones
 - d. Ethers
2. Select properties of these classes.
3. Select pharmaceuticals belonging to these classes and their uses.

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QUESTIONS

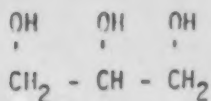
1. _____ - Compounds consisting of two distinct parts, an aliphatic radical and a hydroxyl radical.
2. _____ is the formula for the hydroxyl radical.
3. _____ is the general formula for alcohols.
4. _____ are alcohols which contain one hydroxyl (OH) radical per molecule.
5. List three types of monohydroxy alcohols and define each.
 - a. _____ The hydroxyl (OH) radical is attached to a primary carbon atom.
 - b. _____ The hydroxyl (OH) radical is attached to a secondary carbon atom.
 - c. _____ The hydroxyl (OH) radical is attached to tertiary carbon atom.
6. Monohydroxy alcohols are classified by the manner of which the hydroxyl group is _____.
7. _____ are alcohols which contain two or more hydroxyl (OH) radicals per molecule.
8. Identify the following monohydroxy alcohols.

- a.
$$\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3 - \text{C} - \text{CH}_2 - \text{CH}_3 \\ | \\ \text{OH} \end{array}$$

- b.
$$\text{CH}_3 - \text{CH}_2 - \underset{\text{CH}_3}{\underset{|}{\text{CH}}} - \text{CH}_2 - \text{OH}$$

- c.
$$\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3 - \text{CH} - \text{CH} - \text{CH}_3 \\ | \quad | \\ \text{CH}_3 \quad \text{OH} \end{array}$$

9. Identify the following alcohol.



10. Increasing the length of the hydrocarbon chain, _____ its solubility in water, and solubility in organic solvents increases.
11. Increasing the number of hydroxyl (OH) radicals per molecule _____ its solubility in water.
12. Alcohols become _____ with the increase of the hydroxy (OH) radicals.
13. Alcohols are _____ up to 11 carbon atoms, and become _____ after 12 carbon atoms.
14. The following is a list of five important compounds classified as alcohols and give their use.

a. Methyl Alcohol (Methanol or Wood Alcohol)

(1) Never used in compounding because it is extremely poisonous, both internally and externally.

(2) Use: _____

b. Alcohol U.S.P. (Ethyl Alcohol, Ethanol or Grain Alcohol)

Use: _____

c. Isopropyl Alcohol N.F. (Isopropanol)

Use: _____

d. Glycerin U.S.P. (Glycerol)

Use: _____

e. Propylene Glycol U.S.P.

(1) Substitute for Glycerin

(2) Use: _____

15. The alcohol that is used as a vehicle for internal and external use and a solvent is _____.
16. The alcohol that is NEVER used in compounding is _____.
17. An alcohol that is used as a sweetening vehicle and solvent is _____.
18. _____ are the oxidation products of primary alcohols.
19. _____ is the general formula for Aldehydes.
20. _____ is a Carbonyl Radical and is always present in Aldehydes.
21. Aldehydes are soluble in _____ and only slightly soluble in _____.
22. Aldehydes with lower molecular weight are colorless _____ having a _____ odor.
23. Under two carbon atoms Aldehydes are a _____ with a choking, painful odor.
24. Two selected pharmaceuticals belonging to this class are:
- a. Formaldehyde Solution U.S.P. (Formalin)
 - (1) Subject to polymerization
 - (2) Use: _____
 - b. Chloral Hydrate U.S.P. ("Mickey Finn")
 - Use: _____
25. _____ are the oxidation products of secondary alcohols.
26. _____ is the general formula for Ketones.
27. Ketones are _____ reactive than Aldehydes.

28. Low molecular weight Ketones are _____ and have pungent odors, but high molecular weight Ketones may be either _____ or _____.

29. The radical that is always present in Aldehydes and Ketones is _____.

30. Two selected pharmaceuticals belonging to the Ketone class are:

a. Acetone N.F. (Dimethyl Ketone)

Use: _____

b. Camphor U.S.P.

(1) Forms eutectic mixtures

(2) Use: _____

31. _____ are the dehydration products of two alcohols.

32. _____ is the general formula for Ethers.

33. Ethers are _____ reactive than Ketones and Aldehydes but are good organic _____ but are not very soluble in water.

34. As the molecular weight increases, Ethers become:

a. _____

b. _____

c. _____

35. Two selected pharmaceuticals belonging to the Ether class are:

a. Ether U.S.P. (Ethyl Ether)

Use: _____

b. Ethyl Oxide N.F.

(1) The same as Ether U.S.P., except it contains impurities (96 - 97% Ether).

(2) Use: _____

ALIPHATIC ACIDS, ESTERS, FATS AND FIXED OILS, AND SALTS

The purpose of this lesson is to acquaint you with the properties and uses of organic acids and their derivatives, and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify:

1. Define and give general formulas for the following:
 - a. Acids
 - b. Esters
 - c. Salts
2. Selected properties of these classes.
3. Selected pharmaceuticals belonging to these classes and their uses.

QUESTIONS

1. _____ - Compounds which contain an aliphatic radical and one or more carboxyl radicals per molecule.

2. _____ is the general formula for Aliphatic Acids.



3. - C - OH is the general formula for a _____ radical.

4. - OH is the general formula for a _____ radical.

5. Types of aliphatic acids

a. _____ are assigned to the number of carboxylic acid to get its name.

b. The prefixes are as follows:

(1) _____ - The acid contains one carboxyl radical per molecule.

(2) _____ - The acid contains two carboxyl radicals per molecule.

(3) _____ - The acid contains three carboxyl radicals per molecule.

(4) _____ - The acid contains many carboxyl radicals per molecule.

6. The _____ tell how many carboxyl radicals per molecule.

7. In general, organic acids are _____ acids.

8. The solubility in water of organic acids _____ as the molecular weight increases.

9. Acids react with metals and bases or other alkalies to produce _____.

10. Organic acids are strong acids. (TRUE) (FALSE)

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11. Three selected pharmaceuticals belonging to the Aliphatic Acid class are:

a. Acetic Acid U.S.P. 36 - 37% (Vinegar)

Use: _____

b. Trichloroacetic Acid U.S.P.

(1) Strongest of the organic carboxylic acids

(2) Use: _____

c. Undecylenic Acid N.F. 1 - 10% (One of the active ingredients in Desenex (2%) Foot Powder and Ointment)

Use: _____

12. _____ - Products formed from the reaction between an alcohol and an acid, an acid chloride or an acid anhydride.

13. _____ is the general formula for Esters.

14. Esters can be either _____ or _____.

15. Esters are essentially _____ in water, but may hydrolyze when exposed to moisture for a period of time.

16. Esters may be either fats or fixed oils.

a. Both are _____ esters of fatty acids.

b. They are distinguished by their _____ range (20°C).

17. _____ are solid glyceryl esters.

18. _____ are liquid glyceryl esters.

19. Three selected pharmaceuticals belonging to the Ester, Fats and Fixed oils class are:

a. Glyceryl Trinitrate U.S.P. (Nitroglycerin)

Use: _____

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b. Castor Oil U.S.P.

(1) Fixed oil

(2) Use: _____

c. Theobroma Oil U.S.P. (Cocoa Butter)

(1) Fat

(2) Use: _____

20. _____ - Products formed from the reaction between organic acids with metals and bases or other alkalies, or salts of weaker acids.

21. _____ is the general formula for Salts.

22. Salts can be identified by _____ bonding.

23. Salts are _____ crystal solids.

24. Salts have _____ melting points (300°C - 400°C).

25. Because salts possess ionic bonding they are soluble in water and _____ in organic solvents.

26. Two selected pharmaceuticals belonging to the Salt class are:

a. Zinc Undecylenate N.F. (Another ingredient in Desenex 20%).

Use: _____

b. Magnesium Sulfate U.S.P. (Epsom Salts)

Use: _____

SURFACTANTS

The purpose of this lesson is to acquaint you with the classification and properties of surface active agents used in pharmacy and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will:

1. Define surface tension.
2. Define surfactants.
3. Identify selected properties and classification of surfactants:
 - a. Anionic
 - b. Cationic
 - c. Nonionic
4. Identify selected pharmaceuticals belonging to these classes and their uses.

QUESTIONS

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1. _____ - The attraction of molecules in a liquid (cohesion).
2. _____ - Or surface acting agents are intended to modify the surface tension of a liquid in contact with other liquids or solids.
3. _____ surfactants owe their action to the negative charged portion of the molecule.
4. Anionic surfactants affect the _____ portion of a dipolar molecule.
5. The _____ of anionic surfactants in water is greatly influenced by the length of the _____, and they are _____ with cationic surfactants.
6. Three selected pharmaceuticals belonging to the Anionic Surfactants class are:
 - a. Official Soaps (All are cleansing agents)
 - (1) Hard Soap N.F. (Castile soap)
 - (2) Green Soap N.F. (Medicinal Soft Soap)
 - (3) Detergents
 - b. Dioctyl Sodium Sulfosuccinate N.F. (Colace)
Use: _____
 - c. Dioctyl Calcium Sulfosuccinate N.F. (Surfak)
Use: _____
7. _____ surfactants owe their action to the positive charged portion of the molecule.
8. Cationic surfactants affect the _____ portion of a dipolar molecule.

9. Cationic surfactants are _____, _____, and _____.
10. Cationic surfactants are incompatible with _____ surfactants.
11. Two selected pharmaceuticals belonging to the Cationic Surfactants class are:
- a. Benzalkonium Chloride Solution U.S.P. (Zephiran)
Use: _____
 - b. Cetyl Pyridinium Chloride N.F. (Cepacol)
Use: _____
12. _____ surfactants are those which carry NO ionic effect.
13. If non-ionic surfactants carry no charge, they DO NOT _____.
14. Non-ionic surfactants may be either _____ (water loving) or _____ (oil loving).
15. Two selected pharmaceuticals belonging to the Non-ionic Surfactants class are:
- a. Polysorbate 80 U.S.P. (Tween 80)
Use: _____
 - b. Sorbitan Monooleate (Span)
Use: _____
16. Anionic surfactants are _____ with Cationic surfactants.

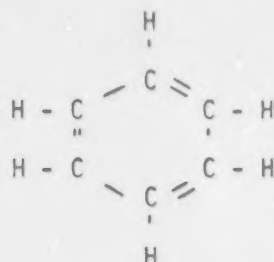
AROMATIC HYDROCARBONS

The purpose of this lesson is to acquaint you with the properties and uses of aromatic hydrocarbons and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and:

1. Define and give a general formula for aromatic hydrocarbons.
2. Select general properties of aromatic hydrocarbons.
3. Select pharmaceuticals belonging to this class and their uses.

QUESTIONS

1. _____ - Compounds which have six carbon atoms, three double bonds, and three single bonds.
2. The example below is the _____ formula for an aromatic hydrocarbon.



This structure is also known as the _____ ring.

3. C_6H_{2n-6} is the general formula for _____.
4. _____ is the principle source of aromatic hydrocarbons.
5. Liquid aromatic hydrocarbons are _____ than water, and have _____ odors.
6. Solid aromatic hydrocarbons have _____ odors.
7. The three types of hydrogen substitution on the Benzene ring are as follows:
- _____ - A single hydrogen is replaced by an element or radical.
 - _____ - Two hydrogens are replaced by an element or radical. This can bring about three possible isomers.
 - _____ "straight line"
 - _____ "beyond"
 - _____ "opposite"
 - _____ - Three hydrogens are replaced by an element or radical. This also brings about three possible isomers.

8. Identify the type of substitution and name the compound.

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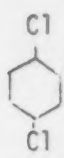
Example:



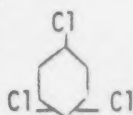
Monosubstitution

Chlorobenzene

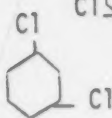
a.



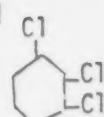
b.



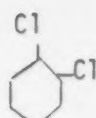
c.



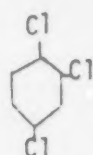
d.



e.



f.



9. In naming compounds of Disubstitution the prefix that is added to the element or radical is

a. Meta (m) meaning _____.

b. Ortho (o) meaning _____.

c. Para (p) meaning _____.

10. When using _____ substitution of hydrogen on the Benzene ring, the prefix tri- is given to the base name and the appropriate three numbers are given to designate the carbon atom to which each substitution the element or radical is attached.

11. Six selected pharmaceuticals belonging to the Aromatic Hydrocarbon class are:

a. Benzene (Benzol)

Use: _____

b. Toluene

Use: _____

c. Xylene (Xylol)

Use: _____

d. Naphthalene (Sublimes)

Use: _____

e. Anthracene

Structure:

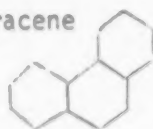


Use: _____

f. Phenanthrene

(1) Isomer of Anthracene

(2) Structure:



(3) Use: _____

AROMATIC ACIDS AND DERIVATIVES

The purpose of this lesson is to acquaint you with the properties and uses of the more common aromatic acids, esters, and salts and will further acquaint you with basic fundamentals of organic pharmaceutical chemistry. Specifically, you will identify:

1. Define and give general formula representing aromatic acids.
2. Selected general properties of aromatic acids.
3. Selected pharmaceuticals belonging to this class and their uses.

QUESTIONS

1. _____ - Compounds which contain an aromatic radical and a carboxyl radical.
2. The aromatic radical (R) represents the _____ ring.
3. $\begin{array}{c} \text{O} \\ \parallel \\ -\text{C}-\text{OH} \end{array}$ is the _____ radical.
4. _____ is the general formula for Aromatic Acids.
5. Most Aromatic Acids are NOT soluble in water, but they react with bases to produce water _____ salts.
6. Four selected pharmaceuticals belonging to the Aromatic Acid class are:
 - a. Benzoic Acid U.S.P.
Use: _____
 - b. Salicylic Acid U.S.P.
Use: _____
 - c. Methyparaben and Propylparaben
Use: _____
 - d. Aspirin U.S.P. (Acetylsalicylic Acid)
 - (1) Aspirin is unstable if moist, it slowly hydrolyzes into
 - (a) _____
 - (b) _____
 - (2) Use: _____
7. _____ are items which lessen pain.
8. _____ are items which reduce temperature.
9. _____ are items which remove the outer horny layer of skin.
10. If Aspirin is stored improperly, it will hydrolyze into _____ and _____.
11. If you open a new bottle of Aspirin and it smells of Vinegar would you use it or dispense it to a patient? YES NO

ALIPHATIC AND AROMATIC HALOGENATED COMPOUNDS

The purpose of this lesson is to acquaint you with the properties and uses of aliphatic and aromatic halogenated compounds and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify:

1. Define and give general formulas representing aliphatic and aromatic halogenated compounds.
2. Selected general properties of aliphatic and aromatic halogenated compounds.
3. Selected pharmaceuticals belonging to these classes and their uses.

QUESTIONS

1. List the four Halogens learned in inorganic chemistry.
 - a. _____
 - b. _____
 - c. _____
 - d. _____
2. _____ are compounds which a Halogen has replaced a hydrogen.
3. _____ halides are the combination of an aliphatic radical and a halogen.
4. _____ is the general formula for Alkyl halides.
5. Alkyl halides are _____ compounds.
6. Alkyl halides have physical states of either a _____ or _____.
7. Alkyl halides possess a _____ sweet odor and taste and they are _____ with water.
8. The alkyl halides are less _____ than their corresponding hydrocarbons and become even lesser as the degree of halogenation increases.
9. Two selected pharmaceuticals belonging to the Aliphatic Halogenated (Alkyl Halides) class are:
 - a. Chloroform N.F.
 - (1) Air, sunlight or open flame causes chloroform to oxidize into _____ and hydrogen chloride.
 - (2) Use: _____
 - b. Halothane U.S.P. (Fluothane)
Use: _____

10. Another name for an aliphatic halogenated compound is _____

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11. _____ halides are the combination of an aromatic radical and a halogen.

12. _____ is the general formula for Aryl halides.

13. Aryl halides have a _____ but not unpleasant odor and are the most _____ of the halogenated compounds.

14. Three selected pharmaceuticals belonging to the Aryl Halides class are:

a. Gamma Benzene Hexachloride U.S.P. (Kwell)

Use: _____

b. Chlorophenothane U.S.P. (DDT)

Use: _____

c. Iodochlorhydroxyquin U.S.P. (Vioform)

Use: _____

15. Alkyl Halides represent an _____ halogenated compound.

16. Aryl Halides represent an _____ halogenated compound.

AMINES AND AMIDES

The purpose of this lesson is to acquaint you with the properties and uses of amines and amides and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify:

1. Define and give general formulas representing:
 - a. Ammonia
 - b. Ammonium Radical
 - c. Amino Radical
 - d. Amines
 - e. Amides
2. Selected general properties of the Amines and Amides.
3. Classify and select pharmaceuticals belonging to the amines and amides and their uses:
 - a. Analgesics
 - b. Local Anesthetics
 - c. Barbiturates
 - d. Antihistamines
 - e. Sulfonamides
 - f. Autonomic Nervous System Drugs

QUESTIONS

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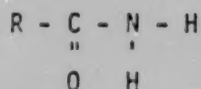
1. _____ is Nitrogen (valence 3) with all its available electrons bonded by hydrogen.
2. _____ is the general formula for Ammonia.
3. _____ radical is Nitrogen (valence 5) bonded by four hydrogens and donating an electron for ionic bonding.

Example: - NH_4^+

4. _____ radical in Nitrogen (valence 3) bonded with two hydrogens with an available electron for bonding with an element or radical.
5. _____ is the general formula for an amino radical.
6. _____ are derivatives of ammonia by replacement of a hydrogen by an aliphatic (alkyl) radical.
7. _____ is the general formula for Amines.
8. _____ are derivatives of ammonia by replacement of a hydrogen by an ACYL Radical.

O
||

9. -C-NH₂ is the general formula for an _____ Radical.
10. The general formula for the _____ is RCONH_2 . This is represented graphically:



_____ Radical

11. Amines are _____ and _____ in aqueous solutions.
12. Amines react with acids to produce _____.
13. Amides are _____ and _____ in aqueous solutions as a result of hydrolysis.

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14. _____ are drugs which lessen pain.
15. Two selected pharmaceuticals belonging to the Analgesic class are:
- a. Phenacetin U.S.P. (Acetophenetidin)
 - b. Acetaminophen N.F. (Tylenol, Tempra)
16. _____ are compounds which render nerve fibers temporarily incapable of conducting impulses.
17. Two selected pharmaceuticals belonging to the Local Anesthetics class are:
- a. Procaine Hydrochloride U.S.P. (Novocaine)
 - (1) Least _____ and most widely used.
 - (2) Use: _____
 - b. Lidocaine Hydrochloride U.S.P. (Xylocaine)
 - (1) Twice as potent and no more _____ than Procaine.
 - (2) Use: _____
18. _____ are related to the Amines and Amides and are used as sedatives and hypnotics.
19. _____ the act or process of calming.
20. _____ an item that induces sleep.
21. Five selected pharmaceuticals belonging to the Barbiturate class are:
- a. Phenobarbital U.S.P. (Luminal)
 - (1) Long-acting
 - (2) Use: _____
 - b. Sodium Amobarbital U.S.P. (Amytal Sodium)
 - (1) Intermediate-acting
 - (2) Use: _____

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- c. Sodium Pentobarbital U.S.P. (Nembutal)
(1) Short-acting
(2) Use: _____
- d. Secobarbital U.S.P. (Seconal)
(1) Short-acting
(2) Use: _____
- e. Thiopental Sodium U.S.P. (Pentothal Sodium)
(1) Ultra Short-acting
(2) Use: _____
22. _____ are synthetic derivatives of ethanolamine which prevents the effects of histamine.
23. Two selected pharmaceuticals belonging to the Antihistamine class are:
- a. Diphenhydramine Hydrochloride U.S.P. (Benadryl)
Use: _____
- b. Chlorpheniramine Maleate U.S.P. (Chlor-Trimeton Maleate)
Use: _____
24. _____ are synthetic derivatives of p-aminobenzene-sulfonamide which are used for their antimicrobial properties.
25. Two selected pharmaceuticals belonging to the Sulfonamide class are:
- a. Sulfisoxazole U.S.P. (Gantrisin)
(1) Oral tablets
(2) Use: _____
- b. Acetyl Sulfisoxazole N.F. (Gantrisin)
(1) A tasteless, pediatric suspension (liquid)
(2) Use: _____

26. _____ (ANS) drugs stimulate the sympathetic and parasympathetic nervous system.

27. _____ hormones stimulate the sympathetic nervous system and are similar in structure to the natural occurring Epinephrine.

28. Three selected pharmaceuticals exhibiting Sympathetic action are:

a. Epinephrine U.S.P. (Adrenalin)

Use: _____

b. Ephedrine Sulfate U.S.P.

Use: _____

c. Phenylephrine Hydrochloride U.S.P. (Neo-synephrine)

Use: _____

29. _____ hormones stimulate the parasympathetic nervous system and are similar in structure to the natural occurring Acetylcholine.

30. Two selected pharmaceuticals belonging to the Parasympathetic Hormone class are:

a. Bethanechol Chloride U.S.P. (Urecholine Chloride)

(1) Comes in oral tablets and injection.

(2) Use: _____

b. Methacholine Chloride N.F. (Mechoyl Chloride)

(1) Injection Only

(2) Use: _____

31. The six classifications of Amines and Amides mentioned have been:

a.

b.

c.

d.

e.

f.

32. Drugs which lessen pain are _____.

33. Drugs that calm are called _____.

34. Drugs that induce sleep are called _____.

INFORMATION

AMINO ACIDS AND PROTEINS

The purpose of this lesson is to acquaint you with the properties and uses of amino acids and proteins and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify:

1. Define and give general formula representing amino acids.
2. Define:
 - a. Peptides
 - b. Proteins
3. Selected properties of proteins.
4. Selected pharmaceuticals belonging to the protein class and their use.

QUESTIONS

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1. _____ are difunctional compounds containing an amino radical and an acid radical.
2. $R - NH_2$ is the _____ radical.
3. $R - \overset{\overset{O}{\parallel}}{C} - OH$ is the _____ radical.
4. $R - \overset{\overset{H}{\mid}}{\underset{\underset{NH_2}{\mid}}{C}} - \overset{\overset{O}{\parallel}}{C} - OH$ is the general formula for an _____.
5. _____ are the combination of two or more amino acids with the removal of a water molecule. This loss of water molecule and combination (continuous) is known as the Peptide Linkage.
6. _____ are polypeptides forming high molecular weight polymers of amino acids by the peptide linkage.
7. All proteins contain the following elements:
 - a. _____
 - b. _____
 - c. _____
 - d. _____
8. Some proteins contain:
 - a. _____
 - b. _____
9. Most proteins are _____ in water but not in organic solvents.
10. Proteins are subject to _____ or salting-out.
11. Two selected pharmaceuticals belonging to the Protein class are:
 - a. Fibrinogen U.S.P.
 - (1) Fibrinogen + Fibrin = _____
 - (2) Use: _____
 - b. Protamine Sulfate Injection U.S.P.
 - (1) Obtained from the sperm of salmon.
 - (2) Use: _____

CARBOHYDRATES

The purpose of this lesson is to acquaint you with the properties and uses of carbohydrates and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and:

1. Define carbohydrates.
2. Classify selected properties of carbohydrates:
 - a. Monosaccharides
 - b. Disaccharides
 - c. Polysaccharides
3. Select pharmaceuticals belonging to these classes and their uses.

QUESTIONS

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1. _____ are aldehydes or ketone derivatives of higher polyhydric alcohols.



2. $\text{R} - \text{C} - \text{H}$ is the general formula for _____.



3. $\text{R} - \text{C} - \text{R}$ is the general formula for _____.

4. _____ is one classification of carbohydrates.

5. The two types of sugars are:

a. _____ - The simplest of all sugars; they cannot be broken down into simpler sugars.

b. _____ are sugars which contain two molecules of the same or different monosaccharides.

6. _____ is the other classification of carbohydrates.

7. Non-sugars are called _____.

8. _____ are complex molecules composed of many monosaccharides.

9. The two major classes of carbohydrates are:

a.

b.

10. Monosaccharides are subject to _____.

11. Monosaccharides are _____ solids, water _____, and have a _____ taste.

12. Disaccharides are subject to _____.

13. Disaccharides are _____ solids, water _____, and have a _____ taste.

14. Polysaccharides are subject to _____.
15. Polysaccharides are _____ solids, many are _____ in water and they are _____.
16. Two selected pharmaceuticals belonging to the monosaccharides class are:
- a. Dextrose U.S.P. (Glucose)
 - (1) _____ found circulating in the blood of animals.
 - (2) Use: _____
 - b. Fructose N.F. (Levulose)
 - (1) Metabolized more rapidly than glucose.
 - (2) Use: _____
17. Two selected pharmaceuticals belonging to the disaccharides class are:
- a. Sucrose U.S.P. (Sugar)
 - Use: _____
 - b. Lactose U.S.P. (Milk Sugar)
 - Use: _____
18. Two selected pharmaceuticals belonging to the polysaccharides class are:
- a. Starch U.S.P. (Corn Starch)
 - Use: _____
 - b. Acacia U.S.P. (Gum Arabic)
 - Use: _____

GLYCOSIDES

The purpose of this lesson is to acquaint you with the properties and uses of glycosides and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and:

1. Define glycosides.
2. Select properties of glycosides.
3. Classify selected pharmaceuticals belonging to these classes and their uses:
 - a. Cardiac
 - b. Cathartic

QUESTIONS

1. _____ are complex compounds consisting of a combination of hydroxyl compounds and sugars.
2. Glycosides are colorless or white, _____ soluble extracts.
3. Glycosides may be _____ soluble and they are _____ active.
4. _____ glycosides have a highly specific action on the heart muscle, they increase tone, excitability and contractability.
5. Three selected pharmaceuticals belonging to the cardiac glycoside class are:

 Digitalis U.S.P. (Foxglove, Whole Leaf)

 Use: _____

 b. Digitoxin U.S.P. (Crystodigin)

 Use: _____

 c. Digoxin U.S.P. (Lanoxin)

 Use: _____

6. _____ glycosides are used widely as they produce catharsis.
7. Two selected pharmaceuticals belonging to the cathartic glycoside class are:
 - a. Cascara Sagrada U.S.P. (Dogwood)
 Use: _____
 - b. Senna N.F. (Senokot)
 Use: _____
8. The two types of glycosides are:
 - a.
 - b.

INFORMATION

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ALKALOIDS

The purpose of this lesson is to acquaint you with the properties and uses of alkaloids and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and:

1. Define alkaloids.
2. Select properties of alkaloids.
3. Classify selected pharmaceuticals belonging to each class and their uses:
 - a. Opium derivatives
 - b. Cinchona derivatives
 - c. Solanaceous derivatives
 - d. Xanthine derivatives
 - e. Ergot derivatives

QUESTIONS

1. _____ are complex (plant) compounds containing nitrogen, which gives them their alkali-like properties.
2. Most alkaloids are of plant origin and usually ending in _____.
3. Because alkaloids are of plant origin they are generally _____ in water and _____ in organic solvents.
4. Alkaloids have many _____.
5. Alkaloids react with acids to form _____ soluble salts.
6. Since alkaloids are of plant origin, they have at least the following elements:
 - a.
 - b.
 - c.
 - d.
7. _____ Compounds which are obtained from the poppy plant (King of the Alkaloids).
8. Five selected pharmaceuticals belonging to the opium derivative class are:
 - a. Morphine Sulfate U.S.P.
 - (1) Phenanthrine derivative
 - (2) Use: _____
 - b. Codeine H.F. (Methylmorphine)
 - (1) Phenanthrine derivative
 - (2) Use: _____
 - c. Hydromorphone Hydrochloride N.F. (Dilaudid)
 - (1) Phenanthrine derivative
 - (2) Use: _____

d. Meperidine Hydrochloride U.S.P. (Demerol)

(1) Phenanthrene derivative

(2) Use: _____

9. _____ Compounds which contain the Quinoline structure as their base.

10. Two selected pharmaceuticals belonging to the cinchona derivative class are:

a. Quinine Sulfate U.S.P.

Use: _____

b. Quinidine Sulfate U.S.P.

Use: _____

11. _____ Compounds characterized by the presence of tropine in the structure.

12. Two selected pharmaceuticals belonging to the solonaceous derivative class are:

a. Atropine Sulfate U.S.P.

Use: _____

b. Cocaine U.S.P.

(1) First _____ anesthetic

(2) Derived from the _____ leaf.

(3) Use: _____

13. _____ Compounds which contain the purine molecule as their base.

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14. Two selected pharmaceuticals belonging to the xanthine derivative class are:

a. Theophylline U.S.P. (Elixophyllin)

Use: _____

b. Caffeine U.S.P.

Use: _____

15. _____ Compounds which contain Lysergic Acid molecule as their base.

16. Two selected pharmaceuticals belonging to the ergot derivative class are:

a. Ergonovine Maleate U.S.P. (Ergotrate Maleate)

Use: _____

b. Ergotamine Tartrate U.S.P. (Gynergen)

Use: _____

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STEROIDS

The purpose of this lesson is to acquaint you with the properties and uses of steroids and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and:

1. Define the basic structure of steroids.
2. Classify selected pharmaceuticals belonging to each class and their uses:
 - a. Adrenal Cortex Hormones
 - b. Bile Salts
 - c. Sterols
 - d. Sex Hormones
 - (1) Female
 - (2) Male

QUESTIONS

1. _____ Compounds which have the perhydrocyclopentanophenanthrene structure as their base.
2. Draw the perhydrocyclopentanophenanthrene structure:
3. Two selected pharmaceuticals belonging to the adrenal cortex hormone class are:
 - a. Cortisone Acetate U.S.P.
Use: _____
 - b. Dexamethasone U.S.P. (Decadron)
 - (1) Synthetic
 - (2) Use: _____
4. Two selected pharmaceuticals belonging to the bile salt class are:
 - a. Ox Bile Extract N.F.
Use: _____
 - b. Dehydrocholic Acid N.F. (Decholin)
Use: _____
5. Two selected pharmaceuticals belonging to the sterol class are:
 - a. Cholesterol U.S.P. (Cholesterin)
Use: _____
 - b. Sitosterols N.F. (Cytellin)
Use: _____
6. Sex hormones are divided into two categories, female and male. Female hormones are subdivided into two categories which are called _____ and _____.

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7. Male hormones are called _____.

8. Five selected pharmaceuticals belonging to the sex hormone class are:

a. Female:

(1) Estradiol Valerate U.S.P. (Delestrogen)

Use: _____

(2) Conjugated Estrogens U.S.P. (Premarin)

Use: _____

(3) Progesterone N.F. (Proluton)

Use: _____

b. Male:

(1) Testosterone N.F. (Androlin)

(a) Injection ONLY (ineffective orally)

(b) Use: _____

(2) Methyltestosterone N.F. (Metandren)

(a) Effective orally

(b) Use: _____

9. The four classes of steroids are:

a.

b.

c.

d.

10. Androgens are _____ sex hormones.

11. The two types of female sex hormones are:

a.

b.

MISCELLANEOUS ORGANIC COMPOUNDS

The purpose of this lesson is to acquaint you with the properties and uses of miscellaneous organic compounds and will further acquaint you with basic fundamentals of pharmaceutical organic chemistry. Specifically, you will identify and:

1. Define
 - a. Phenothiazine derivatives
 - b. Oral hypoglycemic agents
 - c. Antibiotics
2. Classify selected pharmaceuticals belonging to each class and their uses.

QUESTIONS

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1. _____ Compounds which have the phenothiazine structure as their base.
2. Three selected pharmaceuticals belonging to the phenothiazine derivative class are:
 - a. Chlorpromazine U.S.P. (Thorazine)
Use: _____
 - b. Prochlorperazine N.F. (Compazine)
Use: _____
 - c. Thioridazine Hydrochloride U.S.P. (Mellaril)
Use: _____
3. Phenothiazine derivatives are all used as major _____.
4. _____ Compounds which are similar in structure to sulfanilamide, and stimulate the release of insulin from the pancreas.
5. Two selected pharmaceuticals belonging to the oral hypoglycemic agent class are:
 - a. Chlorpropamide U.S.P. (Diabinese)
Use: _____
 - b. Tolbutamide U.S.P. (Orinase)
Use: _____
6. _____ Compounds derived from or produced by a living organism and which inhibit the growth of an organism.
7. Three different types of antibiotics and selected pharmaceuticals belonging to each class are:
 - a. Penicillins:
 - (1) Ampicillin U.S.P. (Polycillin)
 - (2) Procaine Penicillin G U.S.P. (Injection form)

(3) Potassium phenoxymethyl Penicillin U.S.P. (Pen VK and others)

(a) Oral form (Tablets and suspensions)

(b) Resistant to _____ in the stomach.

b. Tetracyclines: (Broad spectrum antibiotics)

(1) Tetracycline U.S.P. (Achromycin)

(2) Oxytetracycline Hydrochloride U.S.P. (Terramycin HCl)

(3) Chlortetracycline Hydrochloride U.S.P. (Aureomycin HCl)

(4) Demethylchlortetracycline HCl N.F. (Declomycin)

c. Miscellaneous Antibiotics:

(1) Erythromycin U.S.P. (Erythrocin, Ilotycin)

(a) A back-up drug for _____.

(b) A drug used for patients who have an allergic reaction
or are known to react to _____.

(2) Chloramphenicol U.S.P. (Chloromycetin)

(a) A _____ antibiotic

(b) Chloramphenicol has many dangerous side effects:

1 Leukopenia -

2 Agranulocytosis -

3 Agranulocytopenia -

4 Aplastic Anemia

GLOSSARY OF TERMS

ACYL RADICAL - Represented by CONH_2 , double bonded oxygen.

ALCOHOLS - R-OH , Organic compounds which contain the hydrocarbon chain and one or more hydroxyl groups.

ALDEHYDES - RCOH , double bonded oxygen, and are the oxidation products of primary alcohols.

ALIPHATIC ACIDS - Organic compounds which contain one or more carboxyl groups in the molecule.

ALIPHATIC HYDROCARBONS - Compounds which contain only Carbon and Hydrogen and are formed in straight or branched open chains.

ALKALOIDS - Complex organic compounds containing Nitrogen which gives them their alkali-like properties.

ALKANES - $\text{C}_n\text{H}_{2n+2}$, also known as the Methane series or the Paraffins.

ALKENES - C_nH_{2n} , also known as the Olefins, name is derived from the Alkanes by changing the ending to ENE and have a double bond.

ALKYNES - $\text{C}_n\text{H}_{2n-2}$, name is derived from the Alkanes by changing the ending to YNE and has a triple bond.

ALKYL HALIDES - Replacement of a hydrogen atom by a Halogen atom on the hydrocarbon chain.

ALKYL RADICAL - Any one of the hydrocarbon radicals of the general formula $\text{C}_n\text{H}_{2n+1}$ and named by the alkane it resembles by changing the ANE ending to YL.

AMIDES - Derivatives of ammonia by replacement of a hydrogen atom by an ACYL group.

AMINES - Derivatives of ammonia by replacement of a hydrogen atom by an ALKYL group.

AMINO ACIDS - Difunctional organic compounds containing an amino group and an acid group.

AMINO RADICAL - Nitrogen (valence state 3) bonded by two hydrogens and an available electron for bonding with another atom or group of atoms. General formula NH_2 .

AMMONIA - Nitrogen (valence state 3) with all of its available electrons bonded by hydrogen. General formula NH_3 .

AMMONIUM RADICAL - Nitrogen (valence state 5) bonded by four hydrogen atoms and donating an electron for ionic bonding. General formula NH_4^+ .

ANALGESICS - Drugs which lessen pain through systemic action.

ANTI-HISTAMINES - Synthetic derivatives of ethanolamine which prevent the effects of histamine.

AROMATIC ACIDS - Compounds which contain an aromatic ring and a CARBOXYL group.

AROMATIC HYDROCARBONS - Compounds which have six carbon atoms, six hydrogen atoms, three double bonds and three single bonds. General formula C_nH_{2n-6} .

ARYL HALIDES - Replacement of a hydrogen atom by a Halogen atom on the aromatic ring.

BARBITURATES - A group of related amines and amides that are used as sedatives and hypnotics.

CARBOHYDRATES - Aldehyde or Ketone derivatives of high polyhydric alcohols, classified as sugars or non-sugars.

CARBONYL GROUP - Represented by $\begin{array}{c} O \\ || \\ -C- \end{array}$

CARBOXYL GROUP - Represented by $\begin{array}{c} O \\ || \\ -C-OH \end{array}$

CARDIAC GLYCOSIDES - Those glycosides which affect the cardiac muscle.

COVALENT BONDING - Chemical bonding in which each atom denotes one or more valence electrons to be shared by the two atoms.

DEHYDRATION - Loss of a water molecule.

DISACCHARIDES - Sugars which contain two molecules of the same or different monosaccharides.

DISUBSTITUTION - Replacement of two atoms.

ESTER - Products formed from the reaction between an alcohol and an acid or an acid chloride or an acid anhydride. General formula $PCOOR$.

ETHERS - $R-O-R$, Dehydration products of two alcohols.

GLYCOSIDES - Complex organic plant principles resulting from the combination of hydroxyl compounds and sugars.

HALOGENS - Fluorine, Chlorine, Bromine, and Iodine.

HOMOLOGOUS SERIES - Each member of the series differs from the next by a set amount (CH_2).

HYDROCARBON - Compounds which contain ONLY Carbon and Hydrogen

HYDROLYZE - Addition of a water molecule.

HYDROXYL GROUP - Represented by $-\text{OH}$.

ISOMERS - Two or more compounds with the same empirical formula but different graphic structures and physical properties.

IUC SYSTEM - International Union of Chemist System for naming organic compounds.

LOCAL ANESTHETICS - Compounds which render nerve fibers temporarily incapable of conducting impulses.

KETONES - RCOR , Oxidation products of secondary alcohols.

META - (m) Means beyond.

MONOSACCHARIDES - The simplest of all sugars which cannot be broken down into simple sugars.

MONOSUBSTITUTION - Single replacement of an atom.

NON-SUGARS - Polysaccharides.

ORGANIC CHEMISTRY - The study of Carbon.

ORGANIC PHARMACEUTICAL CHEMISTRY - The study of compounds containing Carbon and Hydrogen and their derivatives.

ORTHO - (o) Means straight line.

PARA - (p) Means opposite.

PEPTIDES - Combination of two or more amino acids with the removal of a water molecule.

POLYMERS - The product resulting when two or more molecules of the same substance combine.

POLYSACCHARIDES - Non-sugars, complex molecules composed of many monosaccharides.

PROTEINS - Polypeptides forming high molecular polymers of amino acids by the peptide linkage.

RADICAL - A group that preserves its identity throughout a reaction.

SALTS - Products formed from the reaction between organic acids and metals or bases.

STERIODS - Organic compounds which have the perhydrocyclopentanophenanthrene structure as their base.

SULFONAMIDES - Drugs which interfere with the metabolic process of bacteria and are synthetic derivatives of para-aminobenzene sulfonamide.

SURFACE ACTIVE AGENT - See Surfactants.

SURFACE TENSION - The attraction of molecules in a liquid.

SURFACTANTS - Compounds which are intended to modify the surface tension of a liquid in contact with other liquids or solids.

SUGARS - Monosaccharides and disaccharides.

TRISUBSTITUTION - The replacement of three atoms.

VALENCE - The combining capacity of an atom.

SOLUBILITY CHART

DEGREE OF SOLUBILITY

PARTS OF SOLVENT FOR ONE PART OF SOLUTE

Very soluble	Less than 1
Freely soluble	From 1 to 10
Soluble	From 10 to 30
Sparingly soluble	From 30 to 100
Slightly soluble	From 100 to 1,000
Very slightly soluble	From 1,000 to 10,000
Practically insoluble or insoluble	More than 10,000



10-8.

Department of Medicine
School of Health Care Sciences

Medical Service Fundamentals

PREFIXES, ROOTS AND SUFFIXES
OF MEDICAL TERMINOLOGY

April 1974



Sheppard Air Force Base, Texas

Designed For ATC Course Use

74-5-63

DO NOT USE ON THE JOB

PURPOSE OF STUDY GUIDES, WORKBOOKS, PROGRAMMED TEXTS AND HANDOUTS

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PREFIXES, ROOTS AND SUFFIXES OF MEDICAL TERMINOLOGY

This program is an introduction to medical terminology. Although it is not a complete dictionary of medical terms, it does contain a selection of the most common prefixes, roots and suffixes.

The program is not a magical device that will automatically teach you anything. You can only learn medical terminology by applying yourself to the program.

As a medical technician, you will be expected to use medical terminology in dealing with doctors, nurses and other technicians. This is necessary because medicine, like other professions, has its own working language.

Medical terminology was not designed to confuse laymen; instead it was designed to provide uniformity in the meaning of terms. In early medicine, there was little uniformity; consequently, confusion resulted when different words were used to describe the same structure or condition.

Eventually, Greek and Latin words were adopted and certain principles of medical terminology evolved. Those principles are:

1. Each part should have one name only.
2. The names should be as short and simple as possible.
3. Related structures should have similar names.
4. Adjectives, with few exceptions, should be in opposing pairs.

This program will teach you the basic terminology and show you how these principles are applied.

If you have prior knowledge of the terms taught in the program, you may be able to by-pass many, if not all, of the frames. If you feel that you already have a good understanding of medical terminology, turn to page 34, frame 211 and complete the frame according to instructions. Additional instructions will be found in Appendix I located in the back of the program.

If you do not have prior knowledge or if you feel you need the review, begin the program now at frame 1.

This supersedes SW 3AQR90010-1-2

i

NOTE TO THE STUDENT

This program is an introduction to medical terminology. Although it is not a complete dictionary of medical terms, it does contain a selection of the most common prefixes, roots and suffixes.

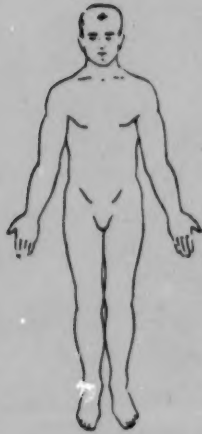
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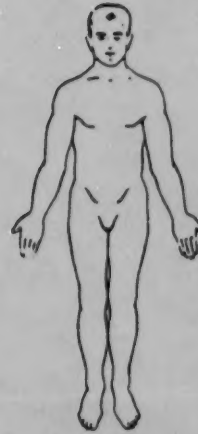
If you do not have prior knowledge or if you feel you need the review, begin the program now at frame 1.

1. Anatomy is a study of the structures of the human body. The arm is a structure of the human body. A study of the structure of the arm then would be a study of _____.
2. (anatomy) The leg, like the arm, is also a _____ of the _____ body.
3. (structure, human) Anatomy is the study of the _____ of the _____ body.
4. (structures, human) Physiology is a study of the functions of the human body. Digestion, respiration and reproduction are all _____ of the human body.
5. (functions) Anatomy is a study of the _____ of the human body. _____ is a study of the _____ of the human body.
6. (structures, Physiology, functions) Movement is a _____ of the human body. A study of movement would be a study of _____.
7. (function, physiology) A study of the structure of the ribs would be a study of _____.
8. (anatomy) Physiology is a study of the (_____ structures, _____ functions) of the human body.
9. (functions) Define anatomy. _____

10. (A study of the structures of the human body) Define physiology.



A.

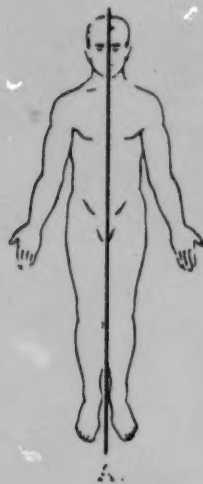


B.

11. (a study of the functions of the human body) Whenever dealing with the human body, references are made with the body in the normal anatomical position. The normal anatomical position has the body at the position of attention with the palms facing forward. Which figure is the normal anatomical position? (☐ both, ☐ A, ☐ B, ☐ neither).
12. (both) The normal anatomical position has the body at the position of (☐ attention, ☐ rest) with the palms facing (☐ rearward, ☐ forward).
13. (attention, forward) The normal anatomical position is
 - a. the position of rest with the palms facing forward.
 - b. the position of attention with the palms facing backward.
 - c. the position of rest with the palms facing forward.
 - d. the position of attention with the palms facing forward.

14. (d, the position of attention with the palms facing forward) For positive identification and location of specific parts of the body, the human form is divided by three anatomical planes. The mid-sagittal or mid-line, transverse, and coronal or frontal are all _____ planes.

300



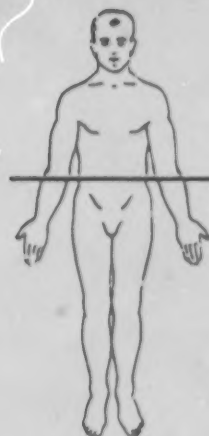
15. (anatomical) The mid-sagittal plane is an imaginary plane which extends the length of the body and divides the body into equal right and left portions. Which drawing above depicts the mid-sagittal plane? (_____ both, _____ A, _____ B, _____ neither)



16. (A) Draw in the mid-sagittal plane on the figure above.
17. The mid-sagittal plane is an imaginary plane which extends the
- a. width of the body and divides the body into a top and bottom section.
 - b. length of the body and divides the body into equal right and left portions.
 - c. length of the body, separating the front of the body from the rear of the body.

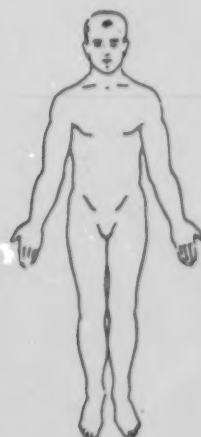


A.



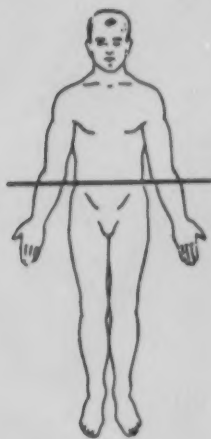
B.

18. (b) A transverse plane is an imaginary plane which extends the width of any section of the body and divides the body into a top and bottom section. Which illustration above indicates a transverse plane? (☐ both, ☐ A, ☐ B, ☐ neither)

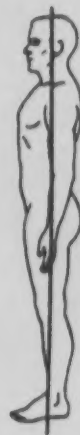


19. (B) Draw in a transverse plane on the figure above.

20. A transverse plane is an imaginary plane which extends the
- a. width of the body and divides the body into a top and bottom section.
 - b. length of the body and divides the body into equal right and left portions.
 - c. length of the body, separating the front of the body from the rear of the body.



A.

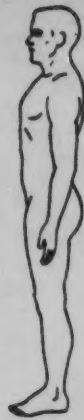


B.

21. (a) A coronal or frontal plane is an imaginary plane extending the length of the body, separating the front of the body from the rear of the body. Which drawing above depicts a coronal plane?

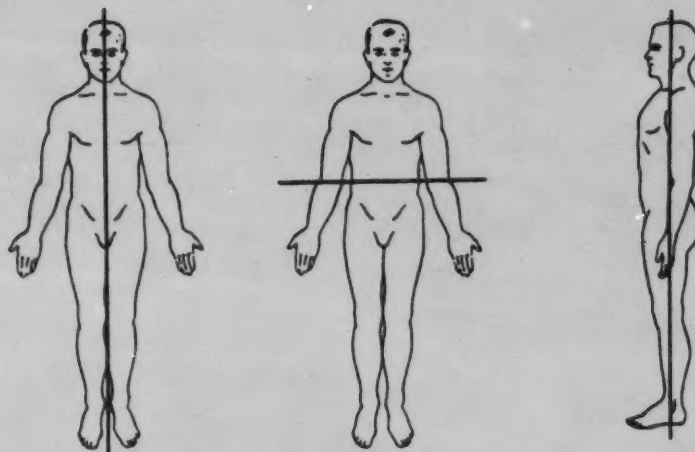
(____ both, ____ A, ____ B, ____ neither)

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22. (B) Draw in a coronal plane on the figure above.
23. A coronal plane is an imaginary plane extending the
- a. width of the body and divides the body into a top and bottom section.
 - b. length of the body and divides the body into equal right and left portions.
 - c. length of the body, separating the front of the body from the rear of the body.

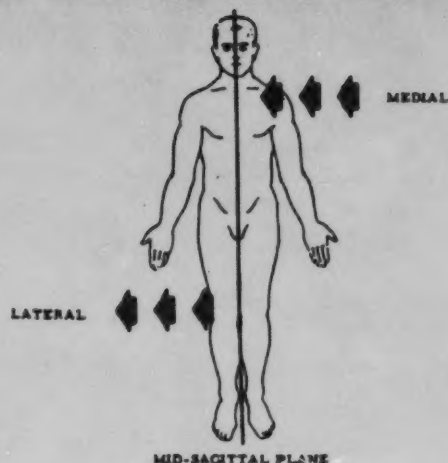
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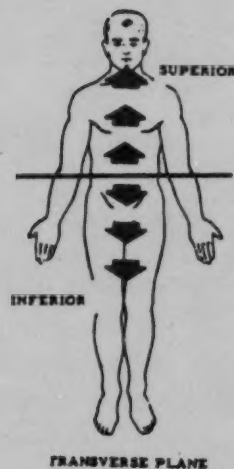
A _____ B _____ C _____

24. (c) Label the planes indicated on the drawings above as mid-sagittal, plane, transverse plane, coronal plane.

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25. (A-mid-sagittal, B-transverse, C-coronal) The mid-sagittal plane has two terms of location connected with it. A part is said to be medial if it is located closer to the mid-sagittal plane than another part. The nose is medial (closer to the mid-sagittal plane) to the ear. The eye would be _____ to the ear.
26. (medial) The tip of the nose is _____ to the eye.
27. (medial) A part is said to be lateral if it is farther from the mid-sagittal plane than another part. The eye is medial to the ear, but the ear is _____ to the eye.
28. (lateral) The tip of the nose is medial to the eye, but the eye is _____ to the tip of the nose.
29. (lateral) The eyes are _____ compared to the ears, but compared to the nose, the eyes are _____.
30. (medial, lateral) Two terms of location connected with the mid-sagittal plane are _____ and _____.

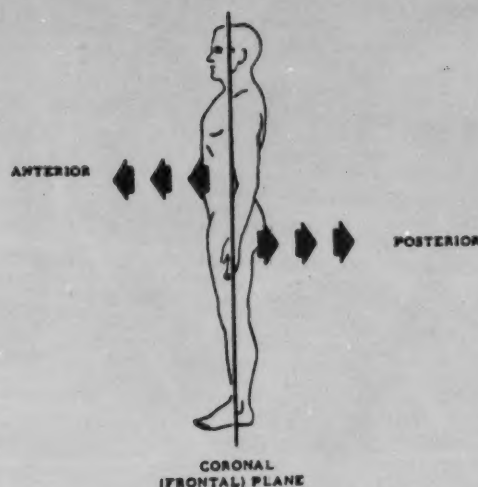


31. (lateral, medial) A transverse plane gives us two terms of location: superior, meaning above, and inferior, meaning below. A thing that is better or above something else is said to be _____.
32. (superior) Something that is inferior is (_____ above, _____ below) standard.

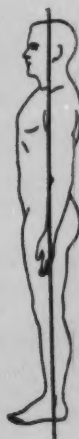


33. (below) With a transverse plane located as in the drawing, the forehead is _____ to the chin.
34. (superior) The two terms of location used in reference to a transverse plane are _____ and _____.
35. (superior, inferior) The two terms of location used in reference to the mid-sagittal plane are _____ and _____.

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36. (medial, lateral) A coronal plane has two terms of location, anterior, meaning in front, and posterior, meaning in back. The nose would be located (____ anterior, ____ posterior) to the back of the head.



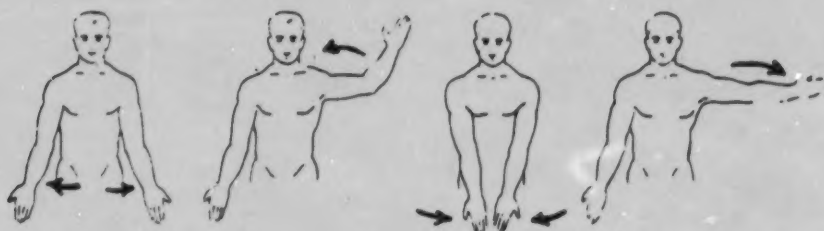
- 37 (anterior) With a coronal plane located as in the drawing, the buttocks are located _____ to the nose.



38. (posterior) The two terms of location connected with a coronal plane are _____ and _____.
39. (anterior, posterior) The terms of location connected with the mid-sagittal plane are _____ and _____. The terms of location connected with a transverse plane are _____ and _____. The terms of location connected with a coronal plane are _____ and _____.
40. (lateral-medial, superior-inferior, anterior-posterior) A point of origin is the beginning of an extremity or system: for example, the mouth is the point of origin for the digestive system, the shoulders would be the point of origin for the _____; while the thigh would be the _____ of _____ for the legs.
41. (arms, point, origin) Two terms of location connected with the points of origin are proximal and distal. If proximal means closest to, then distal must mean _____ from.
42. (farthest) When discussing a part and making reference to the point of origin, the terms _____, meaning closest to, and _____, meaning farthest from, are used.
43. (proximal, distal) The shoulder is the point of origin for the upper extremities. The elbow is closer to the shoulder than the hand. Therefore, the elbow is _____ to the hand.
44. (proximal) In the same light, the hand would be _____ to the elbow.

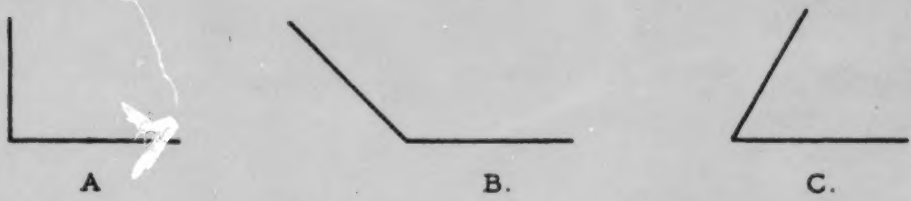
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45. (distal) The point of origin is the beginning of a system or extremity. Two terms of location are _____, meaning closest to the point of origin and _____, meaning farthest from the point of origin.
46. (proximal, distal) An article that is close by is in close proximity or _____.
47. (proximal) An article that is not close by is distant, or _____.
48. (distal) Proximal and distal are used as terms of relationship, i. e., the elbow is distal to the shoulder, but the shoulder is proximal to the arm. Enter the word distal or proximal below.
- The elbow is _____ to the wrist.
 - The fingers are _____ to the elbow.
 - The wrist is _____ to the elbow.
 - The knee is _____ to the heel.
49. (proximal, distal, distal, proximal) Unilateral means pertaining to or affecting only one side. A pain that affects only one arm would be a _____ pain.
50. (unilateral) The removal of one leg could be considered to be a _____ amputation.
51. (unilateral) Bilateral, however, means pertaining to or affecting both sides of the body. Dislocating both shoulders would be a _____ dislocation.
52. (bilateral) The amputation of one arm would be a _____ amputation, while the amputation of both legs would be a _____ amputation.
53. (unilateral, bilateral) Unilateral means pertaining to or affecting _____ side, while bilateral means pertaining to or affecting _____ sides.
54. (one, both) A part that is closest to the point of origin is _____.
55. (proximal) A part that is farthest away from the point of origin is _____.
56. (distal) An item that pertains to or affects only one side of the body is _____, while an item that pertains to or affects both sides of the body is _____.

57. (unilateral, bilateral) Terms of location may sometimes be used together. You have two eyes, so the eyes are (____ unilateral, ____ bilateral). The eyes are located above the tip of the nose, so the eyes are (____ superior, ____ inferior) to the tip of the nose. The eyes are also located behind the tip of the nose, or the eyes are (____ anterior, ____ posterior) to the tip of the nose. In addition, the eyes are farther from the mid-line than the tip of the nose, so the eyes are (____ lateral, ____ medial) to the tip of the nose.
58. (bilateral, superior, posterior, lateral) Using this terminology, you can say the eyes are _____ and the eyes are to the tip of the nose _____, _____ and _____.
59. (bilateral, superior, posterior, lateral) The eye to the ear is (circle the letter that is completely correct)
- anterior - lateral
 - posterior - medial
 - anterior - medial
 - posterior - lateral
60. (c) The foot to the knee is
- superior
 - inferior
 - medial
 - lateral



61. (b) There are four more terms you must learn; these are the terms of movement. Abduction is a movement away from the mid-line. Adduction is a movement toward the mid-line. Flexion is the shortening or closing of an angle. Extension is the lengthening or opening of an angle. Label the drawings: Adduction, Abduction, Flexion, Extension. Note to student: The arrows on all drawings indicate the direction the arm has moved.

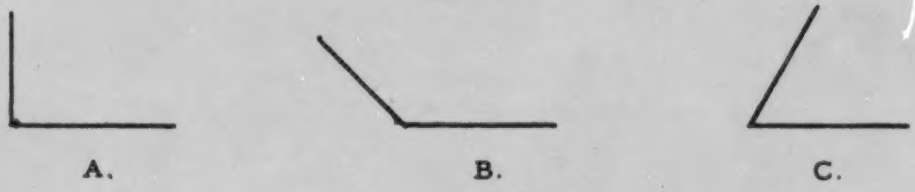
62. (1-Abduction, 2-Flexion, 3-Adduction, 4-Extension) In placing your left hand over your heart, you moved the tips of your fingers toward the mid-sagittal plane. You could say you (___ adducted, ___ abducted) your finger tips.
63. (adducted) A dope addict is drawn towards the dope. A movement towards the mid-line is _____.
64. (adduction) When a person is kidnapped, he is said to have been abducted. Movement away from the mid-line is _____.
65. (abduction) In bringing the hand to the forehead as in the military hand salute, the tips of the fingers are _____, but the elbow is _____. (Consider all movement from the normal anatomical position.)
66. (adducted, abducted) When you contract your arm muscles, the angle formed by the arm and forearm is (___ increased, ___ decreased).



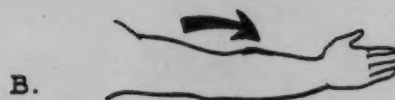
67. (decreased) Flexion is the closing or decreasing of an angle. Starting with the 90° angle of figure A, which angle, B or C, is an example of flexion?



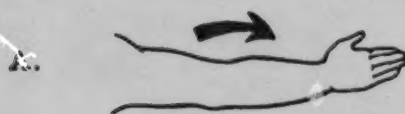
68. (C) Drawing A depicts _____



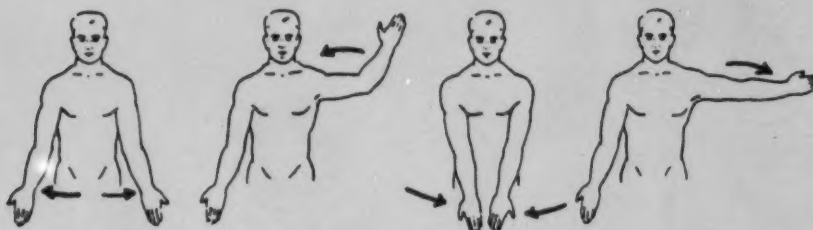
69. (flexion) Extension is the opening or lengthening of an angle. Beginning with the angle in A, which angle, B or C, depicts extension?



70. (B) Which drawing is an example of flexion? (both, A, B, neither).



71. (neither) Which drawing is an example of extension? (both, A, B, neither)



1 , 2 , 3 , 4

72. (both) Label the drawings above as flexion, extension, adduction or abduction.
73. (1-abduction, 2-flexion, 3-adduction, 4-extension) Medical terminology is made up from Latin and Greek roots, suffixes, and prefixes. Learning these roots, suffixes, and prefixes will enable you to understand and form many medical words. Many medical words are formed using and , and .

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74. (Latin, Greek roots, suffixes, prefixes) A prefix is a word used to modify the meaning of the word to which it is attached. It is always placed before the word it modifies. For example, prepaid means paid before. The prefix added to paid in this example is _____.
75. (pre) The prefix meaning without is a or an. Which word or words below mean without something.
- a. abrachia c. abacterial
b. anemia d. diplogen
76. (a, b, c) Abrachia, anemia, abacterial all have a prefix which means _____.
77. (without) The prefix a is used when the root or suffix begins with a consonant. The prefix an is used when the root or suffix begins with a vowel. By adding the proper prefix, change each of the following words into a new form that means without.
1. _____ typical 4. _____ oxia
2. _____ symmetrical 5. _____ otia
3. _____ omaly 6. _____ febrile
78. (1-a, 2-a, 3-an, 4-an, 5-an, 6-a) Which word below could mean without arms?
- a. gelatinase b. abrachia
79. (b) Which word below could mean without blood?
- a. anemia b. napex
80. (a) The prefix ad means to or at. Drawing toward the mid-line is
- a. abduction b. adduction
81. (b) When one substance sticks to another substance, it is said to
- a. adhere b. abort
82. (a) At the mouth would be
- a. aboral b. adoral

83. (b) The prefix meaning without is

- a. ad b. a c. an

84. (b, c) The prefix meaning to or at is _____

- a. ad b. a c. an

85. (a) The prefix meaning before is ante. For each of the following, fill in the word that gives the meaning of the prefix.

- | | |
|-----------------------|------------------------------|
| a. <u>antenatal</u> | _____ birth |
| b. <u>anesthesia</u> | _____ sense of touch or pain |
| c. <u>antecubital</u> | _____ forearm |
| d. <u>aphagia</u> | _____ power of swallowing |
| e. <u>adrenal</u> | _____ kidney |
| f. <u>apnea</u> | _____ breathing |
| g. <u>anoxia</u> | _____ oxygen |
| h. <u>adneural</u> | _____ nerve |

86. (a-before, b-without, c-before, d-without, e-to the (at the), f-without, g-without, h-to the (at the)) Before the arm (forearm) is

- a. abrachial b. anbrachial
c. antebrachial d. adbrachial

87. (c) Antefebriile would mean

- a. after the onset of fever b. before the onset of fever

88. (b) Epi, inter, and intra are three words with similar meanings. They are sometimes difficult to separate. Epi means on or upon, inter means between, and intra means within. Upon the spine is

- a. interspinal b. intraspinal c. epispinal

89. (c) Epicostal means

- a. within a rib b. upon a rib c. between the ribs

90. (b) Inter and intra are the ones most easily mixed up. Inter and enter sound alike; when you enter a gate, you walk between the posts. Between the ribs is

- a. intercostal b. intracostal c. epicostal

91. (a) Upon the skin is _____ while between two or more cartilages would be _____.
- a. interchondral b. epidermal
92. (b, a) Intra means within. Within the skull is _____ and upon the skull is _____.
- a. epicranium b. intracranial
93. (b, a) Fill in the blanks with the correct prefix to match each meaning.
- a. _____ cardium - upon the heart
 b. _____ venous - within the vein
 c. _____ costal - upon the rib
 d. _____ cellular - between the cells
 e. _____ cellular - within the cells
 f. _____ muscular - between muscles
94. (a-epi, b-intra, c-epi, d-inter, e-intra, f-inter) Erythr/o means red. A common word is erythrocyte, meaning _____ blood cell.
95. (red) The abbreviation RBC is frequently used for red blood cell. A red blood cell, then, may be either abbreviated RBC or written _____ cyte.
96. (erythro) Erythrocyte may be abbreviated _____ or written out as _____ blood cell.
97. (RBC, red) Leuk/o and leuc/o mean white. A leukocyte is a _____ blood cell.
98. (white) An abbreviation, WBC, may also be used instead of the prefix-root combination. WBC or _____ cyte means _____ blood cell.
99. (leuko, white) White blood cell may be abbreviated as _____ or written as _____ cyte.

100. (WBC, leuko) You have seen that some prefixes end with the letter o as in leuk/o or erythr/o. Here is the rule for using such prefixes: Add the o when the root or suffix begins with a consonant; drop the o when the root or suffix begins with a vowel. Complete the words below by adding a or an where needed and by retaining or dropping the o in erythr/o and leuk/o.

- a. leuk____cyte b. erythr____emia c. erythro____cyte
d. (a)(an) emic e. (a)(an) symmetrical f. (a)(an) brachi

101. (a-leukocyte, b-erythremia, c-erythrocyte, d-anemic, e-asymmetrical, f-abrachi) Complete the words below using the prefixes you have learned.

- a. _____ costal - between the ribs
b. _____ sexual - without sex
c. _____ blast - a red forming cell
d. _____ neural - upon the nerve
e. _____ emia - white blood disease
f. _____ cranial - within the skull
g. _____ renal - to the kidney
h. _____ emia - without blood
i. _____ brachium - before the arm

102. (a-inter, b-a, c-erythro, d-epi, e-leuk, f-intra, g-ad, h-an, i-ante) Intracranial and endocranial mean the same (inside or within) Inside the skull is _____ or _____.

- a. intracranial b. endocranial
c. epicranial

103. (a, b) Endoscopy means a visual examination

- a. within b. inside
c. upon d. between

104. (a, b) Match the items in column A with the correct meaning in column B.

- A
1. intercostal _____
2. intracranial _____
3. epidermis _____
4. endocardial _____

- B
a. upon the skin
b. within the skull
c. inside the heart
d. between the ribs

Note to student: Although Intra is Latin and Endo is Greek, both meaning within, Intra is usually used to mean within or among while Endo denotes inside or inner.

105. (1-d, 2-b, 3-a, 4-c) Peri means around, retro means behind.
Around the heart is

- a. retrocardial b. pericardial

106. (b) Retronasal means

- a. behind the nose b. upon the nose
c. around the nose

107. (a) Perirectal means _____ the rectum, while retro-ocular
means _____ the eye.

108. (around, behind) A country that is retrogressive may soon be
(____ ahead ____ behind) a similar country that is progressive.

109. (behind) An inflammation around the brain is

- a. retropharyngitis b. periencephalitis
c. endocarditis

110. (b) Label the items "around" or "behind" as applicable.

- a. retrorectal _____ b. periapical _____
c. peribronchial _____ d. retrodural _____

111. (a-behind, b-around, c-around, d-behind) Bi means two. Bicuspid
means _____ cusps.

112. (two) Section means to cut. Cutting into two parts would be
_____.

113. (bisection) The biceps brachii muscle has _____ heads.

114. (two) Bio is a prefix meaning life. A study of life is

- a. hematology b. neurology
c. biology

115. (c) Biogenous means

- a. producing disease b. producing life
c. producing death

116. (b) An examination to determine if life is still present would be
a. endoscopy b. bioscopy
117. (b) Hem/o or hemat means blood. Hematology is a study of _____.
118. (blood) Stasis means standing still. Blood that is standing still would be in a condition of _____ stasis.
119. (hemo) A hemocytoblast is a _____ forming cell.
120. (blood) Hyper and hypo are two prefixes. Hyper means above or an excess; hypo, then, means the opposite, or
a. above or excess b. normal
c. below or deficient
121. (c) Hypertension describes a person who has
a. more tension than normal
b. less tension than normal
c. normal tension
122. (a) A hypodermic needle is a needle that goes
a. above the skin b. below the skin
123. (b) Indicate the meaning of the following words by placing the letter "A" for above and "B" for below, after each word.
a. hyperacute b. hyperacid c. hypochondriac
d. hyperalgia e. hypomorph f. hypotension
124. (a-A, b-A, c-B, d-A, e-B, f-B) Sub is a prefix meaning under. Sub costal would mean _____ the ribs.
125. (under) Glossal refers to the tongue. Under the tongue would be _____.
126. (subglossal) Under the shoulder blade would be _____ scapular.
127. (sub) Pneum/o means air, breath or lung. Pneumonitis is an inflammation of the _____.

128. (lung) Pneumothorax means there is _____ in the chest.
129. (air) Ectomy refers to the surgical removal of a part. Removal of a lung would be a _____ ectomy.
130. (pneum) Listed below are medical words without their prefixes. Add the prefix to make each word mean what the lay term indicates.

- a. _____ brachia - without arms
- b. _____ renal - to the kidney
- c. _____ gastric - upon the stomach
- d. _____ costal - between the ribs
- e. _____ cyte - white blood cell
- f. _____ uria - without urine
- g. _____ cranial - within the skull
- h. _____ cardial - around the heart
- i. _____ cervical - behind the cervix uteri
- j. _____ cardial - inside the heart
- k. _____ logy - a study of life
- l. _____ section - cut in two
- m. _____ logy - a study of blood
- n. _____ tension - over tensed
- o. _____ dermic - beneath the skin
- p. _____ hepatic - under the liver
- q. _____ onitis - inflammation of the lung
- r. _____ cyte - red blood cell
- s. _____ partum - before childbirth

Check the confirmation. Any words you missed or had difficulty with, review before going on.

131. (a-a, b-ad, c-epi, d-inter, e-leuko, f-an, g-intra, h-peri, i-retro, j-endo, k-bio, l-bi, m-hemato, n-hyper, o-hypo, p-sub, q-pneum, r-erythro, s-ante) Periosteum means

- a. behind the bone
- b. upon the bone
- c. around the bone
- d. within the bone

132. (c) Within the cartilage is

- a. perichondrial
- b. interchondral
- c. intrachondrial

133. (c) Epicardial means

- a. within the heart
- b. upon the heart
- c. around the heart

134. (b) Pericardial means
- a. around the heart
 - b. upon the heart
 - c. inside the heart
135. (a) Inside the heart is _____ or _____.
- a. intracardial
 - b. pericardial
 - c. epicardial
 - d. endocardial
136. (a, d) Retrocardiac means
- a. below the heart
 - b. above the heart
 - c. within the heart
 - d. behind the heart
137. (d) Cost is the medical term for rib. Upon the ribs is
- a. epicostal
 - b. intracostal
138. (a) Between the ribs is
- a. intercostal
 - b. intracostal
139. (a) A form referring to the inner surface of the ribs would be
- a. epicostal
 - b. intercostal
 - c. intracostal
140. (c) Cyte is the suffix meaning cell. The most commonly used words employing "cyte" are "leukocyte" and "erythrocyte". These words mean (select two)
- a. bone cells
 - b. red cells
 - c. white cells
 - d. muscle cells
141. (b, c) A red blood cell having no hemoglobin is called a/an
- a. alymphocyte
 - b. anerythrocyte
142. (b) Intraleukocytic means
- a. upon a leukocyte
 - b. inside a leukocyte
 - c. within a leukocyte
 - d. below a leukocyte

143. (b, c) Myo is the medical term for muscle. Myocardium is a/an

- a. arm muscle
- b. neck muscle
- c. heart muscle
- d. head muscle

144. (c) A cell of the muscular tissue is called

- a. myocardium
- b. myocyte

145. (b) Ren and neph both mean kidney. Neph is used most often. Which of the words below pertain to the kidney or heart?

both _____, a _____, b _____, neither _____

- a. nephrocardiac
- b. renicardiac

146. (both) The most common form for kidney is

- a. ren
- b. neph

147. (b) Intrarenal means

- a. within the kidney
- b. inside the kidney
- c. upon the kidney

148.(a-b) Around the kidney is

- a. peribrachial
- b. pericostal
- c. perirenal
- d. pericardial

149. (c) Endonephritis, renal, intrarenal, perirenal, nephrectomy
Looking at the words above, select the correct statement or statements that tell how and when ren and nephro are used.

- a. Ren is always used as a word ending.
- b. Neph is always used as a word beginning.
- c. Ren is never used as a word beginning
- d. Neph is never used as a word ending.
- e. Ren is usually used as a word ending with the suffix al.

150. (d, e) Oste is the medical term for bone. Removing a bone is accomplished through a/an

- a. cardiectomy
- b. pneumonectomy
- c. nephrectomy
- d. ostectomy

151. (d) Ostealgia would be a
- a. headache
 - b. pain in the arm
 - c. pain in the bone
 - d. pain in the neck
152. (c) Osteopathy is a
- a. disease of the skin
 - b. disease of the arm
 - c. disease of the bone
153. (c) Neuro is the medical term for nerve. A neurocyte would be a
- a. muscle cell
 - b. blood cell
 - c. clotting cell
 - d. nerve cell
154. (d) Within a nerve is
- a. subneural
 - b. adneural
 - c. endoneural
 - d. epineural
155. (c) Subneural means
- a. around an arm
 - b. upon a nerve
 - c. upon an arm
 - d. under a nerve
156. (d) Thrombo is the medical term for clot. A thrombocyte is a blood platelet or
- a. red cell
 - b. white cell
 - c. clotting cell
157. (c) A blood clot within the heart is
- a. thrombo-endarteritis
 - b. thrombocytopenia
 - c. thrombo-embolism
 - d. thrombo-endocarditis
158. (d) Producing a clot is
- a. thrombogenic
 - b. thrombocyte

159. (a) You have learned many words. To help you retain this knowledge, the next four frames consist of a review. Column A contains medical terminology and Column B contains lay terminology. Match the medical term with the correct lay term.

A

1. extension _____
2. anatomy _____
3. physiology _____
4. normal anatomical position _____
5. mid-sagittal plane _____
6. transverse plane _____
7. coronal plane _____
8. abduction _____
9. adduction _____
10. flexion _____

B

- a. a study of the functions of the body
- b. an imaginary plane which divides the body into equal right and left halves
- c. an imaginary plane which divides the body into a front and back section
- d. a study of the structures of the body
- e. movement toward the mid-line
- f. lengthening of an angle
- g. movement away from the mid-line
- h. the position of attention with the palms facing forward
- i. the shortening of an angle
- j. an imaginary plane which divides the body into a top and bottom section

160. (1-f, 2-d, 3-a, 4-h, 5-b, 6-j, 7-c, 8-g, 9-e, 10-i) Continue as in the preceding frame.

- | | |
|---------------------|--|
| 1. medial _____ | a. above the transverse plane |
| 2. lateral _____ | b. in front of the coronal plane |
| 3. superior _____ | c. closest to the point of origin |
| 4. inferior _____ | d. pertaining to or affecting but one side |
| 5. anterior _____ | e. pertaining to or affecting both sides |
| 6. posterior _____ | f. nearest to the mid-line |
| 7. proximal _____ | g. farthest from the point of origin |
| 8. distal _____ | h. below the transverse plane |
| 9. unilateral _____ | i. in back of the coronal plane |
| 10. bilateral _____ | j. farthest from the mid line |

161. (1-f, 2-j, 3-a, 4-h, 5-b, 6-i, 7-c, 8-g, 9-d, 10-e) Continue as in the preceding frame.

- | | |
|------------------------|------------------------|
| 1. abrachial _____ | a. situated upon a rib |
| 2. anerythrocyte _____ | b. red blood cell |
| 3. adneural _____ | c. within the heart |
| 4. epicostal _____ | d. white blood cell |
| 5. erythrocyte _____ | e. around the kidney |
| 6. interrenal _____ | f. without red cells |
| 7. leukocyte _____ | g. to a nerve |
| 8. intracardiac _____ | h. behind the heart |
| 9. perinephric _____ | i. without arms |
| 10. retrocardiac _____ | j. between the kidneys |

162. (1-i, 2-f, 3-g, 4-a, 5-b, 6-j, 7-d, 8-c, 9-e, 10-h) Continue as in the preceding frame.

- | | |
|----------------------------|--|
| 1. endocardial _____ | a. destructive to living organisms |
| 2. bilateral _____ | b. excess in the number of white blood cells |
| 3. biocidal _____ | c. under the rib |
| 4. hematology _____ | d. pertaining to the heart and lungs |
| 5. hyperleukocytosis _____ | e. inside the heart |
| 6. hypoleukocytosis _____ | f. a nerve cell |
| 7. subcostal _____ | g. deficiency of white blood cells |
| 8. pneumocardial _____ | h. a blood platlet (clotting cell) |
| 9. neurocyte _____ | i. a study of blood |
| 10. thrombocyte _____ | j. pertaining to both sides |

163. (1-e, 2-j, 3-a, 4-i, 5-b, 6-g, 7-c, 8-d, 9-f, 10-h) Hepat is a root meaning liver. A patient with an inflamed liver would have

- | | |
|-------------|--------------|
| a. neuritis | b. hepatitis |
| c. carditis | d. nephritis |

164. (b) Any disease of the liver would be

- | | |
|----------------|---------------|
| a. hepatopathy | b. osteopathy |
|----------------|---------------|

165. (a) A removal of a portion of the liver would require a

- | | |
|----------------|----------------|
| a. hepatopathy | b. cardiectomy |
| c. hepatectomy | d. ostectomy |

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166. (c) Cephal means head. Medically speaking, if you had a headache you would have
- a. neuralgia
 - b. cardialgia
 - c. cephalalgia
 - d. myalgia
167. (c) Any disease of the head would be classified as
- a. cephalopathy
 - b. hepatitis
 - c. neurology
 - d. osteopathy
168. (a) A headless body would be
- a. bicephalus
 - b. acephalia
169. (b) Chondri is a root meaning cartilage. Under the cartilage is
- a. hypochondrium
 - b. hypochondroplasia
 - c. intrachondrial
 - d. subchondral
170. (d) A cartilage cell is a
- a. chondralgia
 - b. chondrocyte
 - c. chondroblast
 - d. chondritis
171. (b) Cartilage is formed through a process called
- a. myogenesis
 - b. osteogenesis
 - c. neurogenesis
 - d. chondrogenesis
172. (d) The root form for stomach is gastr. An inflammation of the stomach is
- a. nephritis
 - b. gastritis
 - c. cephalitis
 - d. neuritis
173. (b) A word which means pertaining to the heart and stomach is
- a. gastroacephalus
 - b. gastrocardiac
 - c. gastrohepatic
 - d. gastronephritis
174. (b) Which word means an inflammation of the stomach and kidney?
- a. gastrointestinal
 - b. gastrologist
 - c. gastronephritis
 - d. gastrohepatitis

175. (c) Arter means artery. Arteriorenal would be an _____ of the _____.
176. (artery, kidney) Arteritis would be an inflammation of the _____.
177. (arteries) Stenosis means narrowing. Arteriosclerosis means a narrowing of the _____.
178. (arteries) Crani means skull. A craniectomy would be a surgical _____ of the _____.
179. (removal, skull) Crani or cranium means _____.
180. (skull) Pathy means disease. A craniopathy would be a _____ of the _____.
181. (disease, skull) Derma or dermat/o means skin; neur/o means nerve; logy means a study of. Dermatoneurology refers to a _____ of the _____ and the _____.
182. (study, nerves, skin) Itis means an inflamed condition or inflammation. Dermatitis refers to an _____ of the _____.
183. (inflammation, skin) Under the skin is
- a. intradermic
 - b. hypodermic
 - c. subdermic
 - d. epidermic
184. (b-c) Aden is a root meaning gland. A gland can be removed by a/an
- a. neurectomy
 - b. adenectomy
 - c. arterectomy
 - d. nephrectomy
185. (b) An inflammation of the tissues around a gland would be
- a. pericarditis
 - b. periarteritis
 - c. periadenitis
 - d. periangitis
186. (c) A condition of enlarged glands would be
- a. hyperadenosis
 - b. hypoadenia

187. (a) The medical root meaning fat is adip. The condition of a person who is extremely fat is described as

- | | |
|------------------|------------------|
| a. hyperalgesia | b. hyperalgia |
| c. hyperadiposis | d. hyperadenosis |

188. (c) Fat is usually painful to the fat man. A neurotic state in which there are painful areas of fat is

- | | |
|--------------|----------------|
| a. neuralgia | b. adipositis |
| c. neuritis | d. adiposalgia |

189. (d) Inflammation of the fatty tissue is called

- | | |
|--------------|---------------|
| a. neuritis | b. adipositis |
| c. nephritis | d. carditis |

190. (b) Producing fat or fatness is

- | | |
|---------------|-------------|
| a. cytogenic | b. myogenic |
| c. adipogenic | d. ovigenic |

191. (c) A duct is a tube or passage. An egg tube would be an

- | | |
|---------------|--------------|
| a. ovicapsule | b. oviferous |
| c. oviform | d. oviduct |

192. (d) A gland having no excretory passage would be a _____ less gland.

193. (ductless) Either tube or passage is meant by the stem

- | | |
|-----------|----------|
| a. duct | b. ovi |
| c. nephro | d. osteo |

194. (a) Time for a review. The words in column A are lay terms. The words in column B are medical terms. Match the lay terms to the medical terms.

- | <u>A</u> | <u>B</u> |
|---------------------------------------|--|
| 1. an inflammation of the liver _____ | a. ☒ hypochondrial |
| 2. without a head _____ | b. intra-arterial |
| 3. below the cartilage _____ | c. perigastrum |
| 4. around the stomach _____ | d. hepatitis |
| 5. within the artery _____ | e. acephalic |

195. (1-d, 2-e, 3-a, 4-c, 5-b) Complete the next five words as in the last frame.

A

1. egg passage _____
2. fat _____
3. inflamed skin _____
4. skull _____
5. gland _____

B

- a. adipose
- b. cranium
- c. adenal
- d. dermatitis
- e. oviduct

196. (1-e, 2-a, 3-d, 4-b, 5-c) Ophthalm is a root meaning eye. An eye is surgically removed by a/an

- | | |
|----------------|-------------------|
| a. otectomy | b. myectomy |
| c. nephrectomy | d. ophthalmectomy |

197. (d) An inflammation of the eye is

- | | |
|-----------------|-------------|
| a. ophthalmitis | b. neuritis |
| c. nephritis | d. carditis |

198. (a) The study of the eye is called

- | | |
|---------------|------------------|
| a. neurology | b. hematology |
| c. gastrology | d. ophthalmology |

199. (d) Ot is a root word meaning ear. A visual examination of the ear is a/an

- | | |
|--------------|-------------|
| a. endoscopy | b. otoscopy |
|--------------|-------------|

200. (b) A pain in the ear is

- | | |
|--------------|---------------|
| a. neuralgia | b. myalgia |
| c. otalgia | d. nephralgia |

201. (c) Any disease of the ear is called

- | | |
|-------------------|-------------|
| a. ophthalmopathy | b. myopathy |
| c. neuropathy | d. otopathy |

202. (d) Ov is the root word for egg. An egg tube is a/an

- | | |
|---------------|--------------|
| a. oviduct | b. oviferous |
| c. ovigenesis | d. ovination |

203. (a) To bear eggs would be

a. otogenous

b. oviferous

204. (b) An egg cell is an

a. ovigerm

b. ovocyte

c. oviform

205. (b) Vas is the root word meaning vessel. Any nerve and vessel disease would be

a. vasalgia

b. neuropathy

c. vasorrhaphy

d. vasoneuropathy

206. (d) A pain in a vessel is

a. neuralgia

b. vasalgia

c. myalgia

d. nephralgia

207. (b) Any nerve and blood vessel disease is called

a. vasoneuropathy

b. nephropathy

c. myoneurosis

d. pneumonopathy

208. (a) Cyst is the root word meaning bladder. Within the bladder is

a. acystic

b. intracystic

c. pericystic

d. epicystic

209. (b) An inflammation inside the bladder is

a. epicystitis

b. pericystitis

c. endocystitis

d. hypocytosis

210. (c) Inflammation of tissues around the bladder is called

a. cystitis

b. pericystitis

c. epicystitis

d. pericarditis

211. (b) The next group of frames is a review of all the words you have learned. Match the lay term of column A with the medical term of column B.

A

B

- | | |
|---|-------------------------------|
| 1. the study of the structures of the human body _____ | a. normal anatomical position |
| 2. the study of the functions of the human body _____ | b. coronal plane |
| 3. the position of attention with the palms facing forward _____ | c. medial |
| 4. the plane that divides the body into equal right and left halves _____ | d. transverse plane |
| 5. the plane which divides the body into top and bottom sections _____ | e. lateral |
| 6. the plane which divides the body into front and back sections _____ | f. mid-sagittal plane |
| 7. closest to the mid-sagittal plane _____ | g. anatomy |
| 8. farthest from the mid-sagittal plane _____ | h. physiology |

212. (1-g, 2-h, 3-a, 4-f, 5-d, 6-b, 7-c, 8-e) Correct any errors and continue with the next series.

- | | |
|--|---------------|
| 1. above the transverse plane _____ | a. inferior |
| 2. below the transverse plane _____ | b. proximal |
| 3. in front of coronal plane _____ | c. posterior |
| 4. in back of coronal plane _____ | d. unilateral |
| 5. closest to the point of origin _____ | e. superior |
| 6. farthest from the point of origin _____ | f. bilateral |
| 7. affecting one side of the body | g. distal |
| 8. affecting both sides of the body | h. anterior |

- | | |
|---------------------------------------|------------------|
| 3. extreme fatness _____ | c. antebrachium |
| 4. administered below the skin _____ | d. biology |
| 5. under the liver _____ | e. myocardium |
| 6. a surgical removal of a lung _____ | f. hematology |
| 7. forearm _____ | g. sub-hepatic |
| 8. a heart muscle _____ | h. hyperadiposis |

216. (1-d, 2-f, 3-h, 4-a, 5-g, 6-b, 7-c, 8-e) Correct any errors and continue with the next series.

- | | |
|--|----------------|
| 1. under the ribs _____ | a. myocyte |
| 2. muscle cell _____ | b. osteopathy |
| 3. surgical removal of a kidney _____ | c. neuritis |
| 4. bone disease _____ | d. hepatitis |
| 5. pertaining to the kidneys and heart _____ | e. thrombocyte |
| 6. inflammation of the nerves _____ | f. subcostal |
| 7. blood clotting cell _____ | g. nephrectomy |
| 8. inflammation of the liver _____ | h. renocardiac |

217. (1-f, 2-a, 3-g, 4-b, 5-h, 6-c, 7-e, 8-d) Correct any errors and continue with the next series.

- | | |
|----------------------------------|------------------|
| 1. without a head _____ | a. hypochondrial |
| 2. below the cartilage _____ | b. arteriology |
| 3. behind the stomach _____ | c. intracranial |
| 4. a study of the arteries _____ | d. adenitis |
| 5. within the skull _____ | e. adiposis |

213. (1-e, 2-a, 3-h, 4-c, 5-b, 6-g, 7-d, 8-f) Correct any errors and continue with the next series.

- | | |
|--|--------------|
| 1. movement from or away from mid line _____ | a. acephalic |
| 2. movement to or toward mid line _____ | b. epicostal |
| 3. shortening of an angle _____ | c. adduction |
| 4. lengthening of an angle _____ | d. adoral |
| 5. without a head _____ | e. anemia |
| 6. without blood _____ | f. flexion |
| 7. to the mouth _____ | g. abduction |
| 8. upon a rib _____ | h. extension |

214. (1-g, 2-c, 3-f, 4-h, 5-a, 6-e, 7-d, 8-b) Correct any errors and continue with the next series.

- | | |
|--|-----------------|
| 1. red blood cell _____ | a. leukocyte |
| 2. between the ribs _____ | b. pericardium |
| 3. white blood cell _____ | c. retranasal |
| 4. within the skull _____ | d. intercostal |
| 5. a membranous sac around the heart _____ | e. endoneurium |
| 6. behind the nose _____ | f. bisection |
| 7. inside the nerve _____ | g. erythrocyte |
| 8. cut in two _____ | h. intracranium |

215. (1-g, 2-d, 3-a, 4-h, 5-b, 6-c, 7-e, 8-f) Correct any errors and continue with the next series.

- | | |
|---------------------------|----------------|
| 1. a study of life _____ | a. hypodermic |
| 2. a study of blood _____ | b. pneumectomy |

6. inflammation under the tongue _____ f. acephalic
 7. inflammation of a gland _____ g. subglossitis
 8. condition of being fat _____ h. retrogastric
218. (1-f, 2-a, 3-h, 4-b, 5-c, 6-g, 7-d, 8-e) Correct any errors and continue with the next series.
1. egg tube _____ a. ophthalmectomy
 2. the surgical removal of an eye _____ b. oviform
 3. pain in the ear _____ c. vasalgia
 4. egg-shaped _____ d. oviduct
 5. pain in vessels _____ e. cystectomy
 6. removal of a bladder _____ f. otalgia
219. (1-d, 2-a, 3-f, 4-b, 5-c, 6-e) Blast is a word suffix (ending) meaning forming cell. A red forming cell is a/an
- a. leukocyte
 - b. erythroblast
 - c. leukoblast
 - d. erythrocyte
220. (b) Muscle tissue is formed from a
- a. neuroblast
 - b. thromboplast
 - c. myoblast
 - d. osteoblast
221. (c) A white forming cell is a/an
- a. leukocyte
 - b. erythroblast
 - c. leukoblast
 - d. nephrectomy
222. (c) Ectomy attached to a word means a surgical removal of that part. A surgical removal of the liver would be a
- a. hepatectomy
 - b. hepatotomy
223. (a) Removing a part of the heart is accomplished through a
- a. carditis
 - b. cardiectomy
 - c. cardiectomy
 - d. cardiataxia

224. (c) A kidney is removed through a
a. hepatectomy b. nephrectomy
c. arterectomy d. vasectomy
225. (b) Hemat was the prefix meaning blood; emia is the suffix meaning condition of the _____.
226. (blood) The word which means a person is deficient or without blood is
a. erythrocyte b. hypoleukocytosis
c. hyperleukocytosis d. anemia
227. (d) A condition of the white blood cells is called
a. erythremia b. leukemia
228. (b) Itis is a suffix meaning inflammation. Inflammation of the nerves is
a. hepatitis b. nephritis
c. neuritis d. carditis
229. (c) Inflammation of the liver is
a. hepatitis b. nephritis
c. neuritis d. carditis
230. (a) An inflamed kidney would be called nephro_____.
231. (itis) Logy is a suffix meaning a study of. A study of the blood is
a. cardiology b. hematology
c. nephrology d. myology
232. (b) A study of the nerves would be called neuro_____.
233. (logy) Myology is a _____ of the _____.
234. (study, muscle) Pathy is the suffix meaning disease. A disease of the eye would be
a. otopathy b. myopathy
c. cardiopathy d. ophthalmopathy

235. (d) Any disease of the bone is called

- a. osteitis
- b. ostalgia
- c. osteopathy
- d. ostectomy

236. (c) A study of disease is called

- a. pathology
- b. myology

237. (a) A condition is indicated by the suffix osis. A nerve condition is

- a. neuritis
- b. neurosis
- c. neurectomy

238. (b) A condition of a heart muscle would be

- a. myocardium
- b. myocardosis
- c. myocardial
- d. myocele

239. (b) Scopy means a visual examination of. A visual examination inside a part is performed by a/an

- a. episcopy
- b. endoscopy
- c. periscopy
- d. dermatoscopy

240. (b) A visual examination of the eye is performed by a/an

- a. otoscopy
- b. cardioscopy
- c. ophthalmoscopy

241. (c) The suffix _____ is used to indicate a visual examination.

242. (scopy) Stasis is a suffix meaning stoppage or standing still. A stoppage of blood is called

- a. hemostasis
- b. intestinal stasis
- c. ileal stasis

243. (a) Stoppage of the flow of fluid from the kidneys is

- a. hemostasis
- b. urine stasis

244. (b) When the eyes are fixed in one place it is called

- a. hemostasis
- b. venous stasis
- c. ophthalmostasis

245. (c) To the words below add the appropriate ending.

a. erythro _____	red forming cell
b. aden _____	surgical removal of a gland
c. leuk _____	condition of white blood cells
d. neur _____	inflammation of a nerve
e. hemato _____	study of blood
f. osteo _____	bone disease
g. nephro _____	condition of the kidney
h. endo _____	visual examination inside
i. hemo _____	standing or stopped blood

Review any endings you may have missed.

246. (a-blast, b-ectomy, c-emia, d-itis, e-logy, f-path, g-osis, h-scopy, i-stasis) Tomy means surgical incision. A surgical incision into a bladder is a

a. cystotomy

b. myotomy

247. (a) There are three word endings which, when used, place the word in a class, i. e., noun, adjective, past participle. There is a fourth word ending which indicates pertaining to. These word endings are as follows:

Noun - um or ium. For a word not ending with a vowel, use "ium". If the word ends with a vowel, use "um".

Adjective - al

Past participle - ion

Pertaining to - ac

Remember the word noun ends in un, so you add _____ or _____.

248. (um, ium) Make nouns from the following words by adding the correct suffix to each.

cardi _____

gastri _____

chondri _____

249. (cardium, gastrum, chondrium) To form a noun, you add _____ or _____.

330

250. (um, ium) An adjective describes something and is formed by adding al. Make the following words adjectives.

cardi _____
ren _____
brachi _____

251. (cardial, renal, brachial) An adjective is formed by adding _____.

252. (al) The past participle is formed by adding ion; it indicates something that has already happened. Make the following words past participles.

adduct _____
flex _____
abduct _____

253. (adduction, flexion, abduction) Past participles are formed by adding _____.

254. (ion) Pertaining to is denoted by ac. Make each of the following words mean "pertaining to."

cardi _____
chondri _____
brachi _____

255. (cardiac, chondriac, brachiac) The suffix meaning pertaining to is _____.

256. (ac) Complete the following, adding the ending called for

a. cardi _____	noun
b. brachi _____	adjective
c. gastr _____	noun
d. ren _____	adjective
e. chondri _____	pertaining to
f. flex _____	past participle
g. cardi _____	pertaining to
h. adduct _____	past participle

257. (a-um, b-al, c-ium, d-al, e-ac, f-ion, g-ac, h-ion) Ante is a prefix meaning before. The forearm is also called the
- a. antecardium
 - b. antefebrile
 - c. antenatal
 - d. antebrachium
258. (d) Before the heart would be
- a. antebrachium
 - b. antecardium
 - c. anticibum
259. (b) If mortem means death, before death is _____ mortem.
260. (ante) Gen means to produce. Producing eggs is
- a. oviduct
 - b. oviform
 - c. ovigenetic
261. (c) The word below that means producing life is
- a. biochemistry
 - b. biology
 - c. biocidal
 - d. biogenous
262. (d) Producing cells is called
- a. cytogenous
 - b. erythrocyte
 - c. myocyte
 - d. cytoglobin
263. (a) Algia is a suffix meaning pain. A headache would be
- a. neuralgia
 - b. cephalgia
 - c. adenalgia
 - d. gastralgia
264. (b) A pain in the nerves is
- a. neuritis
 - b. neurosis
 - c. neuralgia
 - d. neurology
265. (c) Painful glands might be described as
- a. adenalgia
 - b. adenitis
 - c. adenosis
 - d. adenology

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266. (a) The following is a review of all the words you have learned. You should be able to translate all the medical terminology into lay terminology with little difficulty. Review any words with which you have difficulty. When you complete this program, your instructor will give you additional instructions. Match the terminology in column A with the lay term in column B.

- | | |
|----------------------------|---|
| 1. erythroblastosis _____ | a. excessive blood |
| 2. gastrotomy _____ | b. inflammation of the skin with redness |
| 3. hyperemia _____ | c. a condition of red forming cells |
| 4. pericardectomy _____ | d. a condition of the cartilage and bone |
| 5. osteochondrosis _____ | e. forming blood cells |
| 6. erythrodermatitis _____ | f. surgical removal of the sac around the heart |
| 7. hematocytoblast _____ | g. cutting the stomach |

267. (1-c, 2-g, 3-a, 4-f, 5-d, 6-b, 7-e) Continue with the following words.

- | | |
|---------------------------|--|
| 1. hypochondrium _____ | a. pertaining to the lungs and the heart |
| 2. subdermal _____ | b. pertaining to the arms and head |
| 3. pneumocardial _____ | c. under the cartilage (ribs) |
| 4. periophthalmitis _____ | d. not originating in the liver |
| 5. brachiocephalic _____ | e. an inflammation of the tissues around the eye |
| 6. anhepatogenic _____ | f. a condition in which the blood does not clot |
| 7. athrombosis _____ | g. under the skin |

268. (1-c, 2-g, 3-a, 4-e, 5-b, 6-d, 7-f) Continue with the following words.

- | | |
|---------------------------------|--|
| 1. endoneural _____ | a. inflammation of both eyes |
| 2. perinephritis _____ | b. study of life |
| 3. retro-ocular _____ | c. inflammation of the sac around the kidney |
| 4. bilateral ophthalmitis _____ | d. beneath the liver |
| 5. biology _____ | e. situated within a nerve |
| 6. subhepatic _____ | f. standing blood |
| 7. hepatoscopy _____ | g. examination of the liver |
| 8. hemostasis _____ | h. behind the eye |

269. (1-e, 2-c, 3-h, 4-a, 5-b, 6-d, 7-g, 8-f) Continue with the following words.

- | | |
|-----------------------|---|
| 1. thrombocyte _____ | a. a condition of fat |
| 2. adduction _____ | b. inflammation of the structures above the bladder |
| 3. adiposis _____ | c. within the skull |
| 4. arteriology _____ | d. moving toward the mid-line |
| 5. epicystitis _____ | e. a condition of the upper layer of skin |
| 6. intracranial _____ | f. a cell that helps form a clot |
| 7. epidermosis _____ | g. a study of the arteries |

270. (1-f, 2-d, 3-a, 4-g, 5-b, 6-c, 7-e) Continue with the following words.

- | | |
|-------------------------------------|---|
| 1. anatomy _____ | a. position of attention with palms facing outward |
| 2. physiology _____ | b. imaginary plane that divides the body into front and back sections |
| 3. normal anatomical position _____ | c. imaginary plane that divides body into upper and lower sections |
| 4. mid-sagittal plane _____ | d. study of functions of body |
| 5. transverse plane _____ | e. imaginary plane that divides the body into equal right and left sections |
| 6. coronal plane _____ | f. study of structures of the body |

271. (1-f, 2-d, 3-a, 4-e, 5-c, 6-b) Continue with the following words.

- | | |
|------------------------|--------------------------------------|
| 1. leukocytology _____ | a. an egg-conveying tube |
| 2. intercostal _____ | b. a study of white blood cells |
| 3. oviduct _____ | c. the surgical removal of a bladder |
| 4. otitis _____ | d. inflammation of the ear |
| 5. vasotomy _____ | e. before the arm (forearm) |
| 6. cystectomy _____ | f. cutting of a vessel |
| 7. antebrachium _____ | g. between the ribs |

272. (1-b, 2-g, 3-a, 4-d, 5-f, 6-c, 7-e) Continue with the following words.

- | | |
|--------------------|----------------------------------|
| 1. medial _____ | a. above the transverse plane |
| 2. lateral _____ | b. closest to point of origin |
| 3. superior _____ | c. farthest from the mid-line |
| 4. inferior _____ | d. in back of the coronal plane |
| 5. anterior _____ | e. below the transverse plane |
| 6. posterior _____ | f. closest to the mid-line |
| 7. proximal _____ | g. farthest from point of origin |
| 8. distal _____ | h. in front of the coronal plane |

273. (1-f, 2-c, 3-a, 4-e, 5-h, 6-d, 7-b, 8-g) Continue with the following words.

- | | |
|---------------------|--|
| 1. unilateral _____ | a. movement toward mid-line |
| 2. bilateral _____ | b. closing an angle |
| 3. abduction _____ | c. opening an angle |
| 4. adduction _____ | d. pertaining to or affecting both sides |
| 5. flexion _____ | e. movement from mid-line |
| 6. extension _____ | f. pertaining to or affecting one side |

274. (1-f, 2-d, 3-e, 4-a, 5-b, 6-c) You will be tested on every word taught in this program. Review any words with which you have experienced difficulty. When you are ready, raise your hand and your instructor will give you the examination.

APPENDIX I

Frame 211. If you were able to complete this frame without error, continue with frame 212. Return to this page when you complete frame 212. If you made any errors, return to frame 1 and take the program.

Frame 212. If you were able to complete this frame without error, continue with frame 213. Return to this page when you complete frame 213. If you made any errors, return to page 10, frame 31, and begin the program.

Frame 213. If you were able to complete this frame without error, continue with frame 214. Return to this page when you complete frame 214. If you made any errors, return to page 14, frame 61, and begin the program.

Frame 214. If you were able to complete this frame without error, continue with frame 215. Return to this page when you complete frame 215. If you made any errors, return to page 19, frame 94, and begin the program.

Frame 215. If you were able to complete this frame without error, continue with frame 216. Return to this page when you complete frame 216. If you made any errors, return to page 21, frame 114, and begin the program.

Frame 216. If you were able to complete this frame without error, continue with frame 217. Return to this page when you complete frame 217. If you made any errors, return to page 24, frame 137, and begin the program.

Frame 217. If you were able to complete this frame without error, continue with frame 218. Return to this page when you complete frame 218. If you made any errors, return to page 29, frame 166, and begin the program.

Frame 218. If you were able to complete this frame without error, continue with frame 245, page 40. Return to this page when you complete frame 245. If you made any error, return to page 31, frame 191, and begin the program.

Frame 245. If you were able to complete this frame without error, turn to page 43 and complete frames 266 through 274. Return to this page when you complete this series of frames. If you made any errors, return to page 37, frame 219, and begin the program.

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Frames 266 through 274. If you were able to complete this series of frames and you made no more than 10 errors, you are ready to take the final test. If you made more than 10 errors, return to page 37, Frame 219, and begin the program.



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Technical Training

10-8

Pharmacy Specialist Course

PHARMACEUTICAL INORGANIC CHEMISTRY

January 1976



SCHOOL OF HEALTH CARE SCIENCES, USAF
Department of Biomedical Sciences
Sheppard Air Force Base, Texas 76311

Designed For ATC Course Use

DO NOT USE ON THE JOB

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Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas 76311

WB 3ABR90530-I-5
January 1976

PHARMACEUTICAL INORGANIC CHEMISTRY

OBJECTIVE

Identify the basic concepts, principles, and definitions of pharmaceutical inorganic chemistry. Select the properties of pharmaceutical inorganic chemical elements and compounds. Given the names of specific inorganic elements, correctly write and balance simple chemical equations. Given the necessary data, correctly calculate the milliequivalent concentration of electrolyte solutions.

INTRODUCTION

Chemistry plays many roles in pharmaceutical work. Many of the prescriptions you will be filling will contain two or more ingredients. Sometimes these ingredients will undergo a chemical or physical change. This block of hours will make you aware of these possibilities and provide you with the knowledge to be able to take preventive action. As a result, you will save time and embarrassment by not having to remake the preparation and avoid the possibility of harming yourself or the patient.

PROCEDURES

BASIC CONCEPTS OF INORGANIC CHEMISTRY

QUESTIONS

1. _____ - The study of composition and change in composition of matter.
2. _____ - The science dealing with the elements and mineral matter.
3. _____ - Anything visible or invisible which occupies space and has mass.
4. _____ - Matter which cannot be broken down into simpler matter by ordinary chemical means.
5. _____ - Elements are composed of minute indivisible particles called atoms.
6. _____ - The smallest particle of an element that has the properties of the element and can enter into combination with other elements.
7. _____ - Matter composed of two or more elements combined chemically in definite proportions.
8. _____ - The smallest particle of a compound that can exist and retain the properties of the compound.

COMPOSITION AND PROPERTIES OF MATTER

QUESTIONS

1. Physical States of Matter
 - a. _____ - Has a definite shape and volume.

This supersedes WB 3ABR90530-I-1, October 1974.

b. _____ - Has a definite volume but takes the shape of any container into which it is placed.

c. _____ - Has neither a definite volume or shape. Assumes both the shape and volume of the container into which it is placed.

2. _____ - The capacity to do work.
3. _____ - Energy in motion.
4. _____ - Stored and latent energy.

CLASSIFICATION OF THE ELEMENTS

QUESTIONS

1. _____ - Centrally dense part of an atom that contains the proton and neutron.
2. _____ - Positively charged particle inside the nucleus.
3. _____ - Neutrally charged particle inside the nucleus.
4. _____ - Negatively charged particle orbiting the nucleus.
5. _____ - The number of protons in the nucleus.
6. _____ - A relative system of weight for elements based on carbon 12; equal to the number of protons plus the number of neutrons in the nucleus.
7. _____ - Atoms that have the same atomic number and properties but different atomic weights.
8. _____ - The combining capacity of an element.
9. _____ - The number of electrons of an element involved in the formation of a compound.
10. _____ - Process involving the loss of electrons by an atom accompanied by a gain in the positive valence number.
11. _____ - Process involving the gain of electrons by an atom resulting in a decrease in the positive valence number.
12. _____ - The chemical properties of the elements are periodic functions of their atomic numbers.
13. Lustrous elements which conduct electricity and heat, can be drawn into a fine wire (ductility) or hammered into thin sheets (malleability) are called _____.
14. An element that is a nonconductor of heat and electricity, brittle, and has no characteristic luster is a _____.
15. An element that exhibits properties of both metals and nonmetals is a _____.

16. Using a periodic chart of the elements, list the symbol, atomic weight, and atomic number (rounded off to the nearest whole number) of the following elements.

	<u>SYMBOL</u>	<u>ATOMIC WEIGHT</u>	<u>ATOMIC NUMBER</u>
a. Oxygen	_____	_____	_____
b. Hydrogen	_____	_____	_____
c. Iron	_____	_____	_____
d. Mercury	_____	_____	_____
e. Potassium	_____	_____	_____
f. Silver	_____	_____	_____
g. Sodium	_____	_____	_____
h. Calcium	_____	_____	_____
i. Chlorine	_____	_____	_____
j. Carbon	_____	_____	_____
k. Copper	_____	_____	_____
l. Gold	_____	_____	_____
m. Phosphorous	_____	_____	_____
n. Nitrogen	_____	_____	_____
o. Bromine	_____	_____	_____
p. Iodine	_____	_____	_____
q. Fluorine	_____	_____	_____

17. List the names of the elements corresponding to the following atomic weights.

	<u>NAME OF ELEMENT</u>
a. 40.08	_____
b. 74.922	_____
c. 15.9994	_____
d. 39.948	_____
e. 65.37	_____
f. 39.102	_____
g. 1.00797	_____
h. 4.0026	_____

NAME OF ELEMENT

i.	10.811	_____
j.	35.435	_____
k.	24.312	_____
l.	6.939	_____
m.	22.9898	_____
n.	12.01115	_____
o.	14.0067	_____
p.	18.9984	_____
q.	26.9815	_____

18. If the atomic weight of an element is 35 and its atomic number is 17, how many neutrons are contained within the nucleus of its atom? Show work below.

NAMING INORGANIC COMPOUNDS AND FORMULAS

1. _____ - A combination of symbols representing compounds.
2. _____ - Indicates the number of times the atom (or radical) of the element whose symbol immediately precedes it occurs in the molecule.
3. _____ - Groups of atoms which act or react as a single atom.
4. _____ - An atom or group of atoms (radical) that carry an electrical charge.
5. _____ - The process of dissociation of a substance into its ions.
6. _____ - A condition of net positive charge.
7. _____ - A condition of net negative charge.
8. _____ - Any substance which dissociates into two or more ions when dissolved in water.
9. _____ - The linkage which holds atoms together in a chemical compound.
 - a. _____ or _____ - A type of bonding in compounds in which oppositely charged atoms are held together by attraction.
 - b. _____ - A type of bond in compounds in which pairs of electrons are shared between atoms.

10. The sum of the atomic weights in a chemical formula is the _____.

11. A chemical compound that contains only two elements is termed a _____.

12. _____ - A compound that contains three elements.

13. Rules of cross valence.

a. List the symbols of the _____ or _____ involved and their numbers.

b. Write simple formula inserting _____ above symbols in formula.

c. Take the valence of the element or radical on the left and make it the _____ of the element or radical on the _____.

d. Take the _____ of the element or radical on the _____ and make it the subscript of the _____ or radical on the left.

14. An acid is a compound which contains no _____ ion other than _____.

15. A _____ contains no _____ ion other than hydroxide (OH⁻).

16. A _____ contains positive ions other than _____ and negative ions other than _____.

17. The method of determining acid-base concentration is by measuring _____.

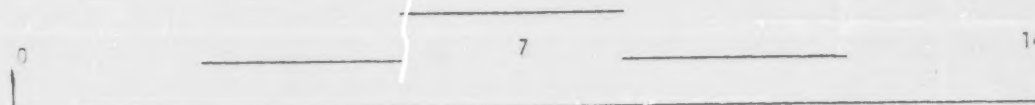
18. The definition of pH is a scale from _____ to _____ indicating the _____ (measure of H⁺ concentration) of a solution.

a. An acid solution has a pH _____ than 7.

b. A basic solution has a pH _____ than 7.

c. A neutral solution has a pH _____ to 7.

19. pH scale:



20. Methods of measuring pH.

a. _____ - A pH meter is used to measure electrical conductivity of a solution. The electrical conductivity of a solution is dependent on the concentration of ions.

b. _____ - Different dyes which change color at different pH values are used as the measuring device.

21. Rules for naming binary compounds.

ACIDS

- a. _____ portion - no name is applied to this portion of the molecule (H^+ ion).
- b. _____ portion.
 - (1) Prefix the name stem of the electronegative element with _____.
 - (2) Suffix the name stem of the electronegative element with _____.
 - (3) Add the word _____ after the above name.

SALTS

- a. Electropositive portion - write out in full the _____ of the element.
- b. Electronegative portion - _____ the name stem of the element with _____.

22. Rules for naming ternary compounds.

ACIDS

- a. _____ portion - no name is applied to this portion of the molecule (H^+ ion).
- b. Electronegative portion
 - (1) For ternary acids made from radicals containing the most common number of oxygen atoms (most common radical), _____ the name stem of the _____ element in the molecule with _____.
 - (2) The acid containing one more oxygen atom than the most common radical retains the suffix _____ and has the prefix _____ added to the name stem of the _____ element.
 - (3) For the acid containing one less oxygen atom than the most common radical, the suffix _____ is changed to _____. No prefix.
 - (4) The acid containing two less oxygen atoms than the most common radical retains the suffix _____ and has the prefix _____ added to the name stem of the _____ element.

SALTS

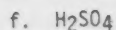
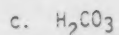
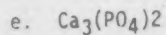
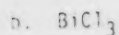
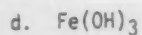
- a. Electropositive portion - write out in full the _____ of the element.
- b. _____ portion.
 - (1) Use the same rules pertaining to the number of oxygen atoms in radicals as used in naming acids except:
 - (a) Change suffix _____ to _____.
 - (b) Change suffix _____ to _____.

- (2) Retain in every case the _____ that was attached to the acid if any.
- c. When a salt (binary or ternary) contains a metal (electropositive element) with more than one valence, suffixes must be applied to the name of the metal to indicate which valence is used. The higher valence is indicated by the suffix _____ on the name of the metal. The lower valence is indicated by the suffix _____ on the name of the metal.

BASES

- a. Write out in full the name of the _____ element of radical.
- b. Always name the OH radical (electronegative portion) _____ for all bases.

23. Compute the molecular weight of the following compounds.



24. A solution with a pH of 7.2 would be slightly _____.

25. A solution with a pH of 5.6 would be slightly _____ while a pH of 2 would indicate a strong _____ solution.

TABLE OF COMMONLY USED VALENCES

NAME	SYMBOL	VALENCES	NAME	SYMBOL	VALENCES
Acetate Radical	$C_2H_3O_2$	-1	Lithium	Li	+1
Aluminum	Al	+3	Magnesium	Mg	+2
Ammonium	NH_4	+1	Manganese	Mn	+2
Barium	Ba	+2	Mercury	Hg	+1, +2
Bicarbonate Radical	HCO_3	-1	Nitrate Radical	NO_3	-1
Bismuth	Bi	+3	Nitrogen	N	+3, +5
Bromine	Br	-1	Oxygen	O	-2
Calcium	Ca	+2	Manganate Radical	MnO_3	-1
Carbon	C	+2, +4	Phosphorous	P	+3, +5
Carbonate Radical	CO_3	-2	Phosphate Radical	PO_4	-3
Chlorine	Cl	-1	Potassium	K	+1
Chlorate Radical	ClO_3	-1	Silver	Ag	+1
Copper	Cu	+1, +2	Sodium	Na	+1
Flourine	F	-1	Sulfate Radical	SO_4	-2
Hydrogen	H	+1	Sulfur	S	-2
Hydroxide Radical	OH	-1	Tin	Sn	+2, +4
Iodine	I	-1	Zinc	Zn	+2
Iron	Fe	+2, +3			
Lead	Pb	+2			

WRITING AND BALANCING EQUATIONS

QUESTIONS

1. Write the chemical symbol for the following:

a. Heat _____

b. Gas _____

- c. Precipitate _____
 - d. Direct Current _____
 - e. Reaction gone to completion _____
2. List the names and symbols of the diatomic elements.
- a. _____
 - b. _____
 - c. _____
 - d. _____
 - e. _____
 - f. _____
 - g. _____
3. When two or more substances combine to form a more complex substance we call this a _____ reaction.
- General formula: _____
4. When a substance is broken down into two or more simple substances, this is called a _____ reaction.
- General formula: _____
5. Rules for decomposition
- a. Metallic carbonates when heated form _____ and _____.
 - b. Metallic chlorates when heated decompose into _____ and _____.
 - c. Some oxides when heated _____.
 - d. The process of decomposing water into hydrogen and oxygen by passing an electrical current through it is called _____.
6. During a reaction when a single element is replaced by a more active element, this type of reaction is termed _____.
- General formula: _____
7. When two compounds exchange positive ions to form two new compounds this is termed a _____ reaction.
- General formula: _____
8. Procedure for writing equations.
- a. Change _____ statement (equation) to _____ statement (symbols and formulas: $XX \rightarrow$).

- b. Check for _____. If there are any, add a subscript of two to them.
- c. Predict resulting _____ ($\rightarrow$ XX).
- d. Check for _____.
- e. Write correct formulas for _____ using _____ (cross valence).
9. Procedure for balancing equations.
- a. Do Not change the _____ (subscripts) of any compound in the equation.
- b. Alter the number of _____ to balance the equation by adding _____ (2X, 3X, etc.) to the molecules.
10. Complete and balance the following equations using the correct chemical symbols, valences and formulas. Do work on separate paper and record answers below.
- Sodium + Chlorine
 - Magnesium - Oxygen
 - Iron (ferrous) + Sulfur
 - Hydrogen + Chlorine
 - Sodium + Hydrochloric acid
 - Potassium Iodide + Chlorine
 - Barium Chloride + Magnesium Sulfate
 - Calcium Chloride + Magnesium Sulfate
 - Hydrochloric acid + Sodium Hydroxide
 - Zinc Chloride + Sodium Sulfate

MOLAR, NORMAL, AND MILLIEQUIVALENT SOLUTIONS

QUESTIONS

- _____ - a liquid consisting of a mixture of two or more substances which are molecularly dispersed through one another in a _____ manner. A solution consists of:
- _____ - the substance dissolved in solution.
- _____ - the medium (_____) in which the substance (solute) is dissolved.
- _____ (GMW) - the molecular weight of a substance expressed in grams. One GMW of a substance equals one _____ of a substance. GMW and mole are used interchangeably.
- _____ - contains one GMW of solute in enough solvent to make one _____ of solution.

6. _____ - the number of moles (GMWs) of solute per liter of solution.

7. Formula for solving molarity problems:

$$M \left(\frac{\text{number of } \underline{\hspace{2cm}} \text{ (GMWs)}}{\text{number } \underline{\hspace{2cm}} \text{ of solution}} \right) = \underline{\hspace{2cm}}$$

8. _____ (GEW) - gram molecular weight divided by the valence of the electropositive element in the molecule.

9. _____ - contains one GEW of solute in enough solvent to make one liter of _____.

10. _____ - the number of GEWs of a solute per liter of solution.

11. Formula for solving normality problems:

$$N \left(\frac{\text{number of } \underline{\hspace{2cm}}}{\text{number } \underline{\hspace{2cm}} \text{ of solution}} \right) = \underline{\hspace{2cm}}$$

12. _____ (mEqwt) - the amount in grams or milligrams of a solute equal to 1/1000 of its gram equivalent weight (GEW).

13. Formulas for solving milliequivalent problems:

a. Solve for number of milliequivalents in a solution.

$$\frac{\# \text{ mg of solute in solution}}{\# \text{ mg per mEqwt of solute}} = \underline{\hspace{2cm}}$$

b. Solve for number of milliequivalents in a specified volume of solution.

(1) Solve for _____ of milliequivalents in solution.

(2) Divide number of milliequivalents in solution by number of _____ of solution - answer in milliequivalents per milliliter.

OXYGEN, HYDROGEN, WATER, AND PEROXIDES

QUESTIONS

- Oxygen is a _____ and _____ gas which is _____ in water.
- Oxygen is a _____ nonmetal and during a chemical reaction it acts as a _____ agent.
- Oxygen will _____ combustion but it is not _____.

4. Oxygen is an allotropic element which means that it _____ in more than one _____.
5. List the names and symbols of the three allotropic forms of oxygen.
 - a. _____
 - b. _____
 - c. _____
6. List the two properties of hydrogen that are similar to oxygen.
 - a. _____
 - b. _____
7. Hydrogen unlike oxygen is _____ flammable and a good _____ agent.
8. Water is a _____, _____ and _____ liquid.
9. Water freezes at _____ degrees C. and _____ F. and boils at _____ degrees C. and _____ degrees F. It is most dense at _____ degrees C.
10. Water combines directly with _____ oxides to form _____ and with _____ oxides to form _____. It is a very _____ compound.
11. Peroxides are _____ compounds whose decomposition yields _____.
12. Peroxides are _____ agent..
13. The uses of the three allotropic forms of oxygen are:
 - a. Oxygen (O_2) - _____
 - b. Ozone (O_3) - _____
 - c. Nascent Oxygen (O) - _____
14. Dilute Hydrochloric Acid is a _____ used for the treatment of achlorhydria and hypochlorhydria.
15. The percentage strength of dilute HCl is 10 percent - it is used as a _____.
16. Only _____ is used as a solvent in pharmaceutical work.
17. _____ - 3 percent is used as an oxidizing agent and _____.
18. Zinc Peroxide is used as a bacteriocide in the control of surgical _____ and in a 40 percent suspension as a _____ agent.

ALKALI METALS (GROUP IA) AND ALKALINE EARTH METALS (GROUP IIA)

QUESTIONS

1. The members of the alkali metals that have compounds related to pharmacy are:
 - a. _____

D.

C.

2. The selected properties of Lithium (Li) are that it is the _____ of all metals, has a _____ luster and has a valence of +1 which makes it _____.

3. Sodium like lithium is a _____ metal with _____ luster. It is also _____, with a valence of +1.

4. Sodium will react with water to decompose it into sodium hydroxide and hydrogen gas. Write and balance the equation for this reaction in the space below.

5. Potassium is a light metal with _____. It is also _____ (valence of +1) and like sodium it will _____ water.

6. _____ - used to treat hyperactive states of manic depression and schizophrenia.

7. The sodium compound that is used as a systemic and nonsystemic antacid is _____.

8. _____ is used as a bacteriostatic and antiseptic.

9. Sodium _____ is used primarily as a urine _____.

10. Sodium Chloride is used as an _____ and to produce _____.

11. Sodium Hypochlorite, a popular bleaching agent is used for its _____ properties.

12. _____ is used by lab technicians to prevent coagulation of _____ samples and as a systemic _____ to relieve mild acidosis.

13. Sodium _____ is a _____ with iodine and a source of the _____ ion.

14. Sodium Nitrate is used as an antidote for _____.

15. Sodium Phosphate is a saline _____ and is also used treating _____ poisoning.

16. Sodium _____ is used with sodium nitrite to treat _____ poisoning and also used as an _____ wash.

17. Potassium _____ is an _____ replenisher and _____.

18. An oxidizing anti-infective is _____.

19. The members of the alkaline earth metals that have compounds related to pharmacy are:

a. _____

b. _____

c. _____

20. Magnesium is a silvery-white metal whose ion is an _____ of many enzyme systems in the body.
21. Magnesium is vital for the function of the _____ system.
22. Calcium ion is indispensable in the function of the _____ and _____ nervous systems. It is also a factor in blood _____ and in the formation of the _____ and like tissue.
23. Soluble barium salts in the body are _____.
24. Magnesium carbonate is used for:
- _____
 - _____
 - _____
25. Magnesium _____ and magnesium _____ are both used as _____ antacids and saline _____.
26. Magnesium _____ (Epsom Salt) is a _____ and anti-_____.
27. Talc is the common name for _____. It is used for:
- _____
 - _____
 - _____
28. _____ has mild _____ qualities and is a nonsystemic _____.
29. Magnesium carbonate, magnesium hydroxide, magnesium oxide and magnesium sulfate can all be used as saline _____.
30. _____ a dentifrice and nonsystemic antacid.
31. Calcium _____ and calcium _____ are both electrolyte _____.
32. Calcium Hydroxide, commonly called _____ is used as an _____, a _____, and an _____.
33. Calcium Phosphate is a _____ of _____.
34. _____ is a radiopaque media in X-ray.

HALOGENS

QUESTIONS

1. The members of the halogens are:
- _____

- b.
- c.
2. Fluorine is a pale _____ which is irritating to _____, _____, and _____.
3. Fluorine is a univalent element that is the _____ active halogen and a powerful _____ agent.
4. List the general properties of chlorine.
- a.
- b.
- c.
- d.
5. Iodine is a bluish black _____ whose most common valence is _____. It serves as an _____ agent and is the _____ active halogen.
6. Fluorine, chlorine, bromine and iodine all exist in nature as _____ elements.
7. List the two selected fluorine compounds and their uses.
- a.
- b.
8. What halogen combines with hydrogen to form a diluted compound used as a stomachic in the treatment of achlorhydria and hypochlorhydria? _____
9. What two chlorine compounds are used as electrolyte replenishers?
- a.
- b.
10. Elemental iodine is used as an _____ externally and internally in the treatment of _____.
11. A compound used as an expectorant in bronchitis and asthma and in the prevention of goiter is _____.
12. _____ increases the solubility of iodine in tinctures and solutions.

SULFUR, NITROGEN AND BORON

QUESTIONS

1. The selected properties of sulfur are:
- a.
- b.

- c.
d.
2. Nitrogen is a _____, _____, and _____ gas which is chemically _____ at room temperature. It is also classed as a _____.
3. Boron is classed as a _____. It has a valence of _____ and is a brown to brownish black _____ or _____.
4. Compounds of sulfur and their uses.
- _____ - saline cathartic, anti-inflammatory.
 - _____ - parasiticide, fungicide, germicide, and keratolytic.
 - _____ - antidote for cyanide, antiseptic wash
 - _____ - parasiticide, fungicide, and germicide because of ability to form hydrogen sulfide on contact with skin. Keratolytic.
5. List the uses of ammonium chloride.
- -
6. Nitrogen _____ oxidation of parenteral solutions.
7. Nitrous Oxide (laughing gas) is used as a _____.
8. Sodium _____ is an antidote for cyanide.
9. _____ is a nonirritating mild antiseptic.
10. Sodium Borate (Borax) is used as an _____ and _____ in cosmetics.

MISCELLANEOUS INORGANIC ELEMENTS

QUESTIONS

- High concentrations of aluminum salts in solution will _____.
- Dilute solutions of soluble aluminum salts when applied topically, cause _____ of blood vessels.
- The bismuth ion is a _____ poison.
- _____ iron (Fe^{+2}) is essential to the hemoglobin of the blood.
- _____ iron (Fe^{+3}) is mainly used externally, as it is a _____ poison.
- Both the silver and zinc ions are _____ poisons.
- The two aluminum compounds used as astringents are _____ and _____.

8. _____ is a nonsystemic antacid and a protective for _____
9. Two aluminum and magnesium preps used as nonsystemic antacids are:
a. _____
b. _____
10. _____ - Bentonite (suspending agent) Kaolin (adsorbant)
11. _____ is used internally for the treatment of dysentery, enteritis, and ulcerative colitis.
12. Two iron compound used as hematinics are:
a. _____
b. _____
13. _____ is an anti-infective used ophthalmically in a one percent solution in newborn babies to combat gonorrhea.
14. _____ is an astringent for _____ and _____. It has powerful escharotic (scab-forming) action.
15. A mild antiseptic and astringent is _____.
16. Zinc _____ is an _____ and emetic.

Technical Training

Fundamentals of Pharmacy

PHARMACEUTICAL INORGANIC CHEMISTRY

January 1976



SCHOOL OF HEALTH CARE SCIENCES, USAF
Department of Biomedical Sciences
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Designed For ATC Course Use

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Department of Biomedical Sciences
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PHARMACEUTICAL INORGANIC CHEMISTRY

Inorganic Chemistry Pharmaceuticals by Most Common Class

1. ADSORBENT - Kaolin
2. ANESTHETIC (general) - Nitrous Oxide
3. ANTACID - Aluminum Hydroxide (nonsystemic)
Aluminum Hydroxide with Magnesium Hydroxide (nonsystemic)
Aluminum Hydroxide with Magnesium Trisilicate (nonsystemic)
Magnesium Hydroxide (nonsystemic)
Magnesium Oxide (nonsystemic) - saline cathartic
Sodium Bicarbonate
4. ANTI-INFLAMMATORY AGENT - Magnesium Sulfate (in hypertonic solution) - saline cathartic
5. ANTISEPTIC - Hydrogen Peroxide (3%) - oxidizing agent
Sodium Thiosulfate - saline cathartic, antidote for cyanide
6. ASTRINGENT - Aluminum Chloride
Ferric Chloride
Zinc Oxide - antiseptic
Zinc Sulfate (ophthalmic)
7. BUFFERING AGENT - Boric Acid - mild antiseptic
8. CATHARTIC - Magnesium Trisilicate (mild) - nonsystemic antacid
9. DENTAL PROPHYLAXIS - Sodium Fluoride - rat and roach poison
Stannous Fluoride
10. DIARRHEA - Bismuth Subcarbonate - dysentery, enteritis, ulcerative colitis
11. ELECTROLYTE REPLENISHER - Calcium Chloride
Potassium Chloride - diuretic
Sodium Chloride
12. EMULSIFYING AGENT - Calcium Hydroxide - astringent, protective
13. EXPECTORANT - Ammonium Chloride - diuretic
14. FILTERING AGENT - Magnesium Silicate (talc) - dusting powder, dispersing agent
15. FUNGICIDE - Elemental Iodine - treatment and prevention of goiter
Sulfur - germicide because of ability to form hydrogen sulfide
Sulfurated Potash - germicide
16. GERMICIDE - Nascent Oxygen (O)
Ozone (O₃)
17. GONORRHEAL INFECTIONS - Silver Nitrate (1% solution used for newborn babies)

Supersedes HO 3ABR90530-I-13, December 1974.

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18. HEMATINIC - Ferrous iron aids formation of hemoglobin of the blood
Ferrous Sulfate
19. HYPERACTIVE STATES OF MANNIC DEPRESSION - Lithium Carbonate
20. OXIDATION RETARDATION OF PARENTERAL SOLUTIONS - Nitrogen
21. OXIDIZING ANTI-INFECTION AGENT - Potassium Permanganate
22. RESPIRATORY FAILURE TREATMENT - Oxygen (O_2)
23. SOLUBILIZING AGENT WITH IODINE - Sodium Iodide
24. SOLVENT, UNIVERSAL - Water (very stable)
Purified Water USP (only one used as a solvent in the pharmacy)
25. STOMACHIC - Dilute Hydrochloric Acid (for achlorhydria and hypochlorhydria)
26. URINE ACIDIFIER - Sodium Biphosphate - mild saline cathartic
27. VASODILATOR - Sodium Nitrate - anti-rust agent
28. X-RAY MEDIA - Barium Sulfate (used in large doses)

368

Molar, Normal and Milliequivalent Solutions

31. If 6 moles of sodium chloride is dissolved in enough water to make 3 liters of solution, what would be the molarity of the solution? (2)

32. What is the molarity of a 5 liter solution containing 3.2 moles of potassium permanganate? (.64)

33. If a 250ml solution contains 4.9 moles of $MnSO_4$, what is the molarity of the solution? (19.6)

34. What is the molarity of a solution containing 2.5 moles of $AgNO_3$ in 3.5 liters of total solution? (.714)

35. If a 1580ml solution contains 5.5 moles of $NaCl$, what is the molarity of the solution? (3.481)

36. If 2 GEW of potassium chloride is dissolved in enough water to make 500ml of solution, what is the normality? (4)

37. What is the normality of a solution containing 4 GEWs of HCl in 250 ml of total solution? (16)

38. What is the normality of a 4 liter solution that contains 4 GMW of MgSO_4 ? (.5)

39. If you put 13.5 GMWs of AlCl_3 in a container and q.s. to 4500ml, what would be the normality of the solution? (1)

40. If you have a 3 liter solution that has a normality of 2, how many GEW of solute were added to it? (6)

41. A 1ml ampul contains .444 Gm of KCl. How many milliequivalents of KCl are there in the ampul? (6)

42. 7.4 Gm of KCl is used to make a 5ml solution. Find the number of milliequivalents per milliliter. (20)

43. 2.4 Gm of MgSO_4 is used to make a 5ml solution. Find the number of milliequivalents per milliliter. (8)

44. A 5 ml ampul contains 2.96 Gm of KCl. How many milliequivalents of KCl are there per milliliter? (8)

45. A 100ml ampul contains 9.88 Gm of NaCl. How many milliequivalents of NaCl are there per milliliter? (1.531)

DEPARTMENT OF BIOMEDICAL SCIENCES

FUNDAMENTALS OF PHARMACY

PHARMACEUTICAL INORGANIC CHEMISTRY

December 1974



SCHOOL OF HEALTH CARE SCIENCES, USAF
SHEPPARD AIR FORCE BASE, TEXAS

Designed For ATC Course Use

DO NOT USE ON THE JOB

PURPOSE OF STUDY GUIDES, WORKBOOKS, PROGRAMMED TEXTS AND HANDOUTS

Study Guides, Workbooks, Programmed Texts and Handouts are training publications authorized by Air Training Command (ATC) for student use in ATC courses.

The STUDY GUIDE (SG) presents the information you need to complete the unit of instruction, or makes assignments for you to read in other publications which contain the required information.

The WORKBOOK (WB) contains work procedures designed to help you achieve the learning objectives of the unit of instruction. Knowledge acquired from using the study guide will help you perform the missions or exercises, solve the problems, or answer questions presented in the workbook.

The STUDY GUIDE AND WORKBOOK (SW) contains both SG and WB material under one cover. The two training publications are combined when the WB is not designed for you to write in, or when both SG and WB are issued for you to keep.

The PROGRAMMED TEXT (PT) presents information in planned steps with provisions for you to actively respond to each step. You are given immediate knowledge of the correctness of each response. PTs may either replace or augment SGs and WBs.

The HANDOUT (HO) contains supplementary training materials in the form of flow charts, block diagrams, printouts, case problems, tables, forms, charts, and similar materials.

Training publications are designed for ATC course use only. They are updated as necessary for training purposes, but are NOT to be used on the job as authoritative references in preference to Technical Orders or other official publications.

Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas

HO 3ABR90530-I-13
December 1974

PHARMACEUTICAL INORGANIC CHEMISTRY

Inorganic Chemistry Pharmaceuticals by Most Common Class

1. ADSORBENT - Kaolin
2. ANESTHETIC (general) - Nitrous Oxide
3. ANTACID - Aluminum Hydroxide (nonsystemic)
Aluminum Hydroxide with Magnesium Hydroxide (nonsystemic)
Aluminum Hydroxide with Magnesium Trisilicate (nonsystemic)
Magnesium Hydroxide (nonsystemic)
Magnesium Oxide (nonsystemic) - saline cathartic
Sodium Bicarbonate
4. ANTI-INFLAMMATORY AGENT - Magnesium Sulfate (in hypertonic solution) - saline cathartic
5. ANTISEPTIC - Hydrogen Peroxide (3%) - oxidizing agent
Sodium Thiosulfate - saline cathartic, antidote for cyanide
6. ASTRINGENT - Aluminum Chloride
Ferric Chloride
Zinc Oxide - antiseptic
Zinc Sulfate (ophthalmic)
7. BUFFERING AGENT - Boric Acid - mild antiseptic
8. CATHARTIC - Magnesium Trisilicate (mild) - nonsystemic antacid
9. DENTAL PROPHYLAXIS - Sodium Fluoride - rat and roach poison
Stannous Fluoride
10. DIARRHEA - Bismuth Subcarbonate - dysentery, enteritis, ulcerative colitis
11. ELECTROLYTE REPLENISHER - Calcium Chloride
Potassium Chloride - diuretic
Sodium Chloride
12. EMULSIFYING AGENT - Calcium Hydroxide - astringent, protective
13. EXPECTORANT - Ammonium Chloride - diuretic
14. FILTERING AGENT - Magnesium Silicate (talc) - dusting powder, dispersing agent
15. FUNGICIDE - Elemental Iodine - treatment and prevention of goiter
Sulfur - germicide because of ability to form hydrogen sulfide
Sulfurated Potash - germicide
16. GERMICIDE - Nascent Oxygen (O)
Ozone (O₃)
17. GONORRHEAL INFECTIONS - Silver Nitrate (1% solution used for newborn babies)

18. HEMATINIC - Ferrous Iron aids formation of hemoglobin of the blood
Ferrous Sulfate
19. HYPERACTIVE STATES OF MANNIC DEPRESSION - Lithium Carbonate
20. OXIDATION RETARDATION OF PARENTERAL SOLUTIONS - Nitrogen
21. OXIDIZING ANTI-INFECTIVE AGENT - Potassium Permanganate
22. RESPIRATORY FAILURE TREATMENT - Oxygen (O_2)
23. SOLUBILIZING AGENT WITH IODINE - Sodium Iodide
24. SOLVENT, UNIVERSAL - Water (very stable)
Purified Water USP (only one used as a solvent in the pharmacy)
25. STOMACHIC - Dilute Hydrochloric Acid (for achlorhydria and hypochlorhydria)
26. URINE ACIDIFIER - Sodium Biphosphate - mild saline cathartic
27. VASODILATOR - Sodium Nitrate - anti-rust agent
28. X-RAY MEDIA - Barium Sulfate (used in large doses)

29. Name the following compounds.

- | | | |
|----------------------------|---|----------------------------------|
| a. HCl | r. KMnO_4 | ii. NH_4NO_3 |
| b. HBr | s. Na_2CO_3 | jj. H_3PO_4 |
| c. H_2S | t. NaHCO_3 | kk. K_3PO_4 |
| d. HClO | u. H_2CO_3 | ll. KNO_3 |
| e. HClO_2 | v. NH_4OH | mm. $\text{Cu}(\text{NO}_3)_2$ |
| f. HClO_3 | w. $\text{Fe}(\text{OH})_3$ | nn. $\text{Al}(\text{NO}_3)_3$ |
| g. HClO_4 | x. FeCl_3 | oo. AgSO_4 |
| h. H_2SO_3 | y. CaCO_3 | pp. FeCl_2 |
| i. H_2SO_4 | z. $\text{Bi}_2(\text{CO}_3)_3$ | qq. CuSO_4 |
| j. HNO_3 | aa. $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ | rr. FeS |
| k. HNO_2 | bb. KI | ss. MnSO_4 |
| l. NaOH | cc. Al_2O_3 | tt. HI |
| m. NaCl | dd. $\text{Ca}_3(\text{PO}_4)_2$ | uu. KMnO_4 |
| n. Na_2S | ee. ZnSO_4 | vv. $\text{Al}_2(\text{SO}_4)_3$ |
| o. NaClO_3 | ff. AgCl | ww. $\text{Al}(\text{OH})_3$ |
| p. NaClO_2 | gg. NaNO_3 | xx. $\text{Ca}(\text{OH})_2$ |
| q. NaClO_4 | hh. $\text{HC}_2\text{H}_3\text{O}_2$ | |

30. Write the formula for the following compounds:

- | | |
|-----------------------|---------------------------|
| a. Sodium Chloride | p. Sulfurous Acid |
| b. Sulfuric Acid | q. Sodium Hydroxide |
| c. Sodium Bicarbonate | r. Nitric Acid |
| d. Calcium Carbonate | s. Carbonic Acid |
| e. Aluminum Oxide | t. Potassium Permanganate |
| f. Calcium Phosphate | u. Hydrobromic Acid |
| g. Zinc Sulfate | v. Ferrous Chloride |
| h. Aluminum Hydroxide | w. Cupric Sulfate |
| i. Potassium Iodide | x. Aluminum Nitrate |
| j. Magnesium Sulfate | y. Cupric Nitrate |
| k. Phosphoric Acid | z. Sodium Carbonate |
| l. Sodium Bromide | aa. Ferric Hydroxide |
| m. Ferric Chloride | bb. Cuprous Sulfate |
| n. Hypochlorous Acid | cc. Manganese Hydroxide |
| o. Sodium Perchlorate | dd. Silver Chloride |

Molar, Normal and Milliequivalent Solutions

31. If 6 moles of sodium chloride is dissolved in enough water to make 3 liters of solution, what would be the molarity of the solution? (2)

32. What is the molarity of a 5 liter solution containing 3.2 moles of potassium permanganate? (.64)

33. If a 250ml solution contains 4.9 moles of MgSO_4 , what is the molarity of the solution? (19.6)

34. What is the molarity of a solution containing 2.5 moles of AgNO_3 in 3.5 liters of total solution? (.714)

35. If a 1580ml solution contains 5.5 moles of NaCl , what is the molarity of the solution? (3.481)

36. If 2 GEW of potassium chloride is dissolved in enough water to make 500ml of solution, what is the normality? (4)

37. What is the normality of a solution containing 4 GEWs of HCl in 250 ml of total solution? (16)

377

38. What is the normality of a 4 liter solution that contains 4 GMW of $MgSO_4$? (.5)
39. If you put 13.5 GMW of $AlCl_3$ in a container and q.s. to 450ml, what would be the normality of the solution? (1)
40. If you have a 3 liter solution that has a normality of 2, how many GEW of solute were added to it? (6)
41. A 1ml ampul contains .444 Gm of KCl. How many milliequivalents of KCl are there in the ampul? (6)
42. 7.4 Gm of KCl is used to make a 5ml solution. Find the number of milliequivalents per milliliter. (20)
43. 2.4 Gm of $MgSO_4$ is used to make a 5ml solution. Find the number of milliequivalents per milliliter. (8)
44. A 5 ml ampul contains 2.96 Gm of KCl. How many milliequivalents of KCl are there per milliliter? (8)
45. A 100ml ampul contains 8.88 Gm of NaCl. How many milliequivalents of NaCl are there per milliliter? (1.531)

318

APPROVAL OFFICE AND DATE HSDM/2 Aug 75 <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 11		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy and Physiology (Cells, Tissues and Glands)			
LESSON DURATION			
CLASSROOM/Laboratory 2 Hours		~~XXXXXXXX~~ Complementary 0	TOTAL 2 Hours
POI REFERENCE			
PAGE NUMBER 8	PAGE DATE 18 July 75		PARAGRAPH 1b
STS/CTS REFERENCE			
NUMBER STS 905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE		DATE	SIGNATURE
<i>[Signature]</i>		32 AUG 1975	
<i>[Signature]</i>		6 OCT 1975	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
None	NA	NA	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HO 3ABR90530-II-1 PT 3ABR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1b. Identify selected cells, tissues and glands pertaining to the human body. (Teaching steps listed in Part II)			

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LESSON PLAN (Part I, Continued)

APPROVAL OFFICE AND DATE MSDB/22 Aug 75 <i>Wilm</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 11		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy and Physiology (Muscular System)			
LESSON DURATION			
CLASSROOM/Laboratory 2 Hours	~~COMPLEMENTARY~~ 0	TOTAL 2 Hours	
POI REFERENCE			
PAGE NUMBER 8	PAGE DATE 18 July 75	PARAGRAPH 1c.	
STS/CTS REFERENCE			
NUMBER STS 905X0	DATE 28 Feb 75		
SUPERVISOR APPROVAL			
SIGNATURE	DATE 1975	SIGNATURE	DATE
<i>[Signature]</i>	8 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	NA	NA	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HO 3ABR90530-II-1 PT 3ABR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1c. Identify selected basic facts and terms about the muscular system. (Teaching Steps listed in Part II)			

ATC FORM 709

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE MSDB/22 Aug 75 <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy and Physiology(Skeletal System)			
LESSON DURATION			
CLASSROOM/Laboratory 1 Hour	~~10000000~~ Complementary 0	TOTAL 1 Hour	
POI REFERENCE			
PAGE NUMBER 8	PAGE DATE 18 July 75	PARAGRAPH 1d.	
STS/CTS REFERENCE			
NUMBER STS 905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	2 AUG 1975		
<i>[Signature]</i>	OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
None	NA	NA	Transparency set #1 Anatomy & Physiology FA Anatomical Chart HO 3ABR90530-II-1 PT 3ABR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1d. Identify selected basic facts and terms about the skeletal system. (Teaching steps listed in Part II)			

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LESSON PLAN (Part I, General)

APPROVAL OFFICE AND DATE MDDR/22 AUG 75 <i>Wilson</i>	INSTRUCTOR
COURSE NUMBER 3ABR90530	COURSE TITLE Pharmacy Specialist
BLOCK NUMBER II	BLOCK TITLE Pharmacology

LESSON TITLE
Anatomy & Physiology (Nervous system)

CLASSROOM/Laboratory 3 Hours	LESSON DURATION ~~3 Hours~~ Complementary 0	TOTAL 3 hours
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PAGE NUMBER 3	PAGE DATE 18 July 75	PARAGRAPH 1e
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NUMBER STS 90570	STS/CTS REFERENCE DATE 28 Feb 75
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SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	22 AUG 1975		
<i>[Signature]</i>	066T 1975		

PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
None	NA	NA	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HO 3ABR90530-II-1 PT 3ABR90530-II-1b Mental Hygiene (Audio Tape)

CRITERION OBJECTIVES AND TEACHING STEPS

1e. Identify selected basic facts and terms about the nervous system.
(Teaching Steps Listed in Part II)

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE MSDB/22 Aug 75 <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 3AER90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy and Physiology (Circulatory System)			
LESSON DURATION			
CLASSROOM / Laboratory 2 hours	~~LABORATORY~~ Complementary 0	TOTAL 2 Hours	
POI REFERENCE			
PAGE NUMBER 8	PAGE DATE 18 July 75	PARAGRAPH 1f	
STS/CTS REFERENCE			
NUMBER STS 905X0	DATE 28 Feb 75		
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	22 AUG 1975		
<i>[Signature]</i>	2 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
None	NA	NA	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HO 3AER90530-II-1 PT 3AER90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1f. Identify selected basic facts and terms about the circulatory system. (Teaching steps listed in Part II)			

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LESSON PLAN (Part I, General)

APPROVAL OFFICE AND DATE MBDB/22 AUG 75 <i>Wilder</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy & Physiology (Respiratory System)			
LESSON DURATION			
CLASSROOM/Laboratory 1 Hour	~~Complementary~~ 0	TOTAL 1 Hour	
POI REFERENCE			
PAGE NUMBER 8	PAGE DATE 18 July 75	PARAGRAPH 1g	
STS/CTS REFERENCE			
NUMBER STS 905X0	DATE 23 Feb 75		
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	22 AUG 1975		
<i>[Signature]</i>	1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	NA	NA	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart H7 3ABR90530-II-1 PT 3ABR90530-II-1b

CRITERION OBJECTIVES AND TEACHING STEPS

1g. Identify selected basic facts and terms about the respiratory system.
(Teaching Steps listed in Part II)

LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE MSDB/22 <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 75		COURSE TITLE	
3ABR90530		Pharmacy Specialist	
BLOCK NUMBER		BLOCK TITLE	
11		Pharmacology	
LESSON TITLE			
Anatomy and Physiology(digestive System)			
LESSON DURATION			
CLASSROOM	/Laboratory	LABORATORY	TOTAL
	2 Hours	~~XXXXXX~~ Complementary	2 hours
POI REFERENCE			
PAGE NUMBER	PAGE DATE	PARAGRAPH	
18	18 July 75	1b	
STS/CTS REFERENCE			
NUMBER		DATE	
STS 90530		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	2 AUG 1975		
<i>[Signature]</i>	10 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	NA	NA	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HO 3ABR90530-II-1 PT 3ABR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1b. Identify the selected basic facts and terms about the digestive system. (Teaching Steps listed in Part II)			

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APPROVAL OFFICE AND DATE <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 7		COURSE TITLE	
		Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Specialist	
LESSON TITLE Pharmacology			
Anatomy & Physiology (Endocrine Systems)			
LESSON DURATION			
CLASSROOM	LABORATORY	TOTAL	
Laboratory	Complementary	1 Hour	
1 HOUR	POI REFERENCE	1 HOUR	
PAGE NUMBER	PAGE DATE	PARAGRAPH	
3	18 JUL 75	11	
SUPERVISOR APPROVAL			
SIGNATURE		DATE	
<i>[Signature]</i>		JUL 1975	
<i>[Signature]</i>		JUL 1975	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	NA	NA	Transparency Set #1 Anatomy & Physiology PA Anatomical Chart HD 3AIR90530-II-1 PT 3AIR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
11. Identify selected basic facts and terms about the endocrine system. (Teaching steps listed in Part II)			

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE HJDB/22 AUG 75 <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 3AER90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy and Physiology(Urinary System)			
LESSON DURATION			
CLASSROOM/Laboratory 1 hour	~~XAR90530~~ Complementary 0		TOTAL 1 Hour
POI REFERENCE			
PAGE NUMBER 8	PAGE DATE 13 July 75		PARAGRAPH 1j
STS/CTS REFERENCE			
NUMBER JTS 905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	22 AUG 75		
<i>[Signature]</i>	6 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	NA	NA	Transparency Set #1 Anatomy & Physiology PA Anatomical Chart HO 3AER90530-II-1 PT 3AER90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1j. Identify selected basic facts and terms about the urinary system. (Teaching steps listed in Part II)			

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE HS 17/22 Aug 75 <i>Wilson</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Anatomy & Physiology (Reproductive System)			
LESSON DURATION			
CLASSROOM/Laboratory 1 Hour	SEPARATE Complementary 0		TOTAL 1 Hour
POI REFERENCE			
PAGE NUMBER # 9	PAGE DATE 18 Jul 75		PARAGRAPH 1k
STS/CTS REFERENCE			
NUMBER STS 905X0		DATE 23 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE		DATE	SIGNATURE
<i>[Signature]</i>		10 AUG 1975	
<i>[Signature]</i>		0 OCT 1975	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	na	na	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HO 3ABR90530-II-1 PT 3ABR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
1k. Identify selected basic facts and terms about the reproductive system. (Teaching Steps Listed in Part II)			

ATC 5554

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE <i>Kilmer</i>		INSTRUCTOR	
COURSE NUMBER		COURSE TITLE Specialist	
BOOK NUMBER		174961516/gy	
REPRESENTATIVE Physiology (Eye & Ear)			
LABORATORY		~~XXXXXXXXXXXXXXXXXXXX~~	
214961516	LABORATORY	0	214961516
POI REFERENCE			
PAGE NUMBER	15 FEB 75		PARAGRAPH
STS/CTS REFERENCE			
NUMBER		DATE	
STS 915X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE		DATE	
<i>Don E. [Signature]</i>		22 AUG 1975	
<i>Don E. [Signature]</i>			
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NONE	na	na	Transparency Set #1 Anatomy & Physiology FA Anatomical Chart HD 3AHR90530-II-1 PT 3AHR90530-II-1b
CRITERION OBJECTIVES AND TEACHING STEPS			
11. Identify selected basic facts and terms about the eye and ear. (Teaching steps listed in Part II)			

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APPROVAL OFFICE AND DATE <i>16 Aug 74</i>		INSTRUCTOR	
MSDB <i>7M84 74</i>		COURSE TITLE Pharmacy Specialist	
COURSE NUMBER 3ABR90530		BLOCK TITLE Pharmacology	
BLOCK NUMBER II		LESSON TITLE Drug Abuse	
CLASSROOM/Laboratory <i>42 hrs/0 hrs</i>		LESSON DURATION ~~0000000000~~ Complementary 0 hrs	TOTAL 42 hrs
POI REFERENCE			
PAGE NUMBER 10	PAGE DATE 18 July 75	PARAGRAPH 14a	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	<i>16 Aug 74</i>	<i>[Signature]</i>	<i>16 OCT 75</i>
<i>[Signature]</i>	<i>16 Aug 74</i>		
<i>[Signature]</i>	<i>8 Aug 75</i>		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	Drug Abuse Trans- parency Set I Films: Weed, Acid, Last Minute to Choose Hooks, The Perfect Drug, Speedscene, Drug Abuse, Problems of Amphetamine Abuse WB3ABR90530-II+Pharma- cology
CRITERION OBJECTIVES AND TEACHING STEPS			
2a. Describe the drugs subject to abuse and the symptoms of drug abuse. (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE MSDB <i>16 Aug 74</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Pharmaceutical and Medicinal Agents			
LESSON DURATION			
CLASSROOM/Laboratory 20 hrs/0 hrs	~~XXXXXXXXXX~~ Complementary 8 hrs	TOTAL 28 hrs	
POI REFERENCE			
PAGE NUMBER <i>8</i>	P 18 July 75	PARAGRAPH <i>A 5a</i>	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	16 Aug 74	<i>[Signature]</i>	8 OCT 1975
<i>[Signature]</i>	16 Feb '5		
<i>[Signature]</i>	8 Aug '5		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	WB3ABR90530-II-1 Pharmacology Pharmacology Trans- parency Set #2 Films: The Digestive System & Ascariasis
CRITERION OBJECTIVES AND TEACHING STEPS			
5a. Classify and describe the properties of locally acting drugs, gastro-intestinal drugs, local anesthetics and anti-infective drugs. (Teaching steps listed in Part II)			

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APPROVAL OFFICE AND DATE MSDB Wilson 27 Sept 74		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 11		BLOCK TITLE Pharmacology	
LESSON TITLE Pharmaceutical and Medicinal Agents			
LESSON DURATION			
CLASSROOM/Laboratory 10 Hrs/0 Hrs	~~XXXXXXX~~ Complementary 4 Hrs	TOTAL 14 Hrs	
POI REFERENCE			
PAGE NUMBER 8 / 10	PAGE 18 July 75	PARAGRAPH K 58.6656	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	27 SEPT 74	<i>[Signature]</i>	6 OCT 1975
<i>[Signature]</i>	27 MAR 75		
<i>[Signature]</i>	12 SEPT 75		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	NB 3ABR90530-II-1 Pharmacology Pharmacology Transpar- ency Set #3 Films: Fundamentals of the Nervous System; Halothane, Abnormal Behavior.
CRITERION OBJECTIVES AND TEACHING STEPS			
9b. Classify and describe the properties of drugs acting on the central nervous system. (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE 1211N <i>Hilson 29 Oct 79</i>		INSTRUCTOR	
COURSE NUMBER 3AHR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER II		BLOCK TITLE Pharmacology	
LESSON TITLE Pharmaceutical and Medicinal Agents			
LESSON DURATION			
CLASSROOM/Laboratory 16 hrs/0 hrs	~~5000000000~~ Complementary 6 hrs	TOTAL 22 hrs	
POI REFERENCE			
PAGE NUMBER 11	PAGE DATE 18 Jul 75	PARAGRAPH 5c	
STS/CTS REFERENCE			
NUMBER STS905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	24 Oct 7		
<i>[Signature]</i>	24 April 75		
<i>[Signature]</i>	6 Oct 3		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	WB3AHR90530-II-1 Pharmacology; Pharmacology Transpare- ncy Set #4; Films: Work of the Heart; The Blood; Common Heart Disorders; Congestive Heart Failure
CRITERION OBJECTIVES AND TEACHING STEPS			
5c. Classify and describe the properties of drugs acting on the autonomic nervous system and circulatory system. (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)

APPROVAL OFFICE AND DATE Wilson 25 Nov 75		INSTRUCTOR	
COURSE NUMBER SABER		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 11		BLOCK TITLE Pharmacology	
LESSON TITLE Pharmaceutical and Medicinal Agents			
LESSON DURATION			
CLASSROOM / Laboratory 16 hrs / 0 hrs	COMPLEMENTARY 6 hrs		TOTAL 22 hrs
POI REFERENCE			
PAGE NUMBER 11	PAGE DATE 18 July 75		PARAGRAPH 5d
STS/CTS REFERENCE			
NUMBER STS 905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE <i>Ante Delgado</i>		DATE 2 Dec 75	
SIGNATURE		DATE	
SIGNATURE		DATE	
SIGNATURE		DATE	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	WB3AER90530-II-1 Pharmacology Pharmacology Trans- parency Set #5 Films; Endocrine System, Menstrual Cycle, Vitamins and Some Deficiency Diseases, Immunization
CRITERION OBJECTIVES AND TEACHING STEPS			
5d. Classify and describe the properties of drugs acting on the endocrine system and miscellaneous therapeutic drugs. (Teaching steps listed in Part II)			

ATC

FORM 7-66

7-66

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE ICM/12 Nov 75 <i>Kilson</i>		INSTRUCTOR	
COURSE NUMBER 3A1820.30		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER 11		BLOCK TITLE Pharmacology	
LESSON TITLE Dispensing Laboratory			
LESSON DURATION			
CLASSROOM/Laboratory 0/42	~~XXXXXXX~~ Complementary 14	TOTAL 56	
POI REFERENCE			
PAGE NUMBER 11	PAGE DATE 18 July 1975	PARAGRAPH 6a, b, c, d.	
STS/CTS REFERENCE			
NUMBER STS 905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE <i>Ron E. Wegman</i>	DATE 20 Nov 75	SIGNATURE	DATE
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
Typewriter Numbering Machines Prescription Files Prescriptions Drugs References Telephones			
CRITERION OBJECTIVES AND TEACHING STEPS			
6a. Given instructor assistance and placed in the dispensing pharmacy (model pharmacy and pharmacology research area), correctly interpret, fill and label prescriptions in accordance with AFM 168-4 and complete handouts for locally acting drugs, gastro-intestinal drugs, local anesthetics and anti-infective drugs. 6b. Given instructor assistance and placed in the dispensing pharmacy (model pharmacy and pharmacology research area), correctly interpret, fill and label prescriptions in accordance with AFM 168-4 and complete handouts for drugs acting on the central nervous system. 6c. Given instructor assistance and placed in the dispensing pharmacy (model pharmacy and pharmacology research area), correctly interpret, fill and label prescriptions in accordance with AFM 168-4 and complete handouts for drugs acting on the autonomic nervous system and circulatory system. 6d. Given instructor assistance and placed in the dispensing pharmacy (model pharmacy and pharmacology research area), correctly interpret, fill and label prescriptions in accordance with AFM 168-4 and complete handouts for drugs acting on the endocrine system and miscellaneous. (Teaching steps 1-4)			

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HANDOUTS, II - 3 through 7

Course 10-8

10-8 396

Department of Biomedical Science
School of Health Care Sciences, USAF
Sheppard AFB, Texas

HANDOUT 3ABR90530-11-3
November 1974

PHARMACOLOGY

1. Match Column A with Column B

<u>COLUMN A</u>	<u>COLUMN B</u>
1. Alphabetical index of brand names (PDR)	a. Nerve damage may occur
2. Drug Classification Index (PDR)	b. Combination drug
3. Surfact	c. Green pages
4. UBI-TD	d. Inhibitory effect on the nervous system
5. Robitussin-DM	e. Time released
6. Alphabetical listing according to generic name only (Index)	f. Remington's Pharmaceutical Sciences
7. Indications	g. Dryness of the mouth
8. Precautions	h. Uses
9. Action	i. Not to be used for patients with cardiac disorders
10. Adverse reaction	j. 240mg red gelatin capsule
11. Contraindication	k. Pink pages

2. Using the Product Identification Section of the PDR, describe the Mergesic tablet.

3. a. In Remington's Pharmaceutical Science, the alphabetical listing of generic names is located in the (front/back) of the book.
b. In the PDR, drugs are indexed in the (front/back) of the book.
4. In Section Five of the PDR (Product Information), drugs are listed alphabetically according to brand name. They are listed alphabetically within the heading of a specific _____ also appearing in alphabetical order.
5. List two references which specify whether or not a drug is a controlled item.
a. _____
b. _____

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6. Determine the schedule for each of the following drugs:
- Decadron _____
 - Phelantix _____
 - Solfo-Serpine _____
 - Biphetamine _____
 - Ambar _____
7. Using the Federal Supply Catalog determine which of the following drugs is a controlled item.
- Methylphenidate HCl Tablets USP _____
 - Sodium Amobarbital capsules USP _____
 - Prednisolone NaPO₄ Inj. USP _____
 - Thiopental anesthesia kit _____
 - Reserpine tablets USP _____
8. Trade names (are/are not) given in the Remington's Pharmaceutical Sciences.
9. What is the name for Diethylpropion which indicates, in detail, its chemical structure? _____ (use Remington's)
10. What topic heading in Remington's Pharmaceutical Science contains adverse reactions, uses, precautions and warnings for a particular drug?
- Uses
 - How supplied
 - Dose
 - Descriptions
11. Using the PDR and Remington's Pharmaceutical Sciences, give a trade name for each of the following generic names.
- Phenylephrine _____
 - Phenobarbital _____
 - Pheniramine _____
 - Phenacitin _____
12. Draw the symbol, located at either the upper or lower right of a drug's name, which indicates that this name is a trade name?

Department of Biomedical Science
School of Health Care Sciences, USAF
Sheppard AFB, Texas

HANDOUT 3ABR90530-II-4
November 1974

PHARMACOLOGY

- A. Aniline Derivative
- B. Aniline Derivative, causes Methemoglobinemia
- C. Anti-depressant, Monomine Oxidase Inhibitor
- D. Anti-Depressant, Nonmonomine Oxidase Inhibitor
- E. Cerebral Stimulant
- F. Found in coffee
- G. Intermediate Acting Barbiturate
- H. Long Acting Barbiturate
- I. Major Tranquillizer, Phenothiazine Derivative
- J. Major Tranquillizer, Non-Phenothiazine Derivative
- K. Minor Tranquillizer
- L. Non-Barbiturate Sedative Hypnotic
- M. Non-Opiate Analgesic, Narcotic
- N. Non-Opiate Analgesic, Treat Heroin Addiction
- O. Opiate Analgesic
- P. Pyrazolon Derivative, Causes Agranulocytosis
- Q. Salicylate
- R. Semi-Synthetic Opiate
- S. Semi-Synthetic Opiate, Narcotic Antagonist
- T. Short Acting Barbiturate
- U. Treat Grand Mal Epilepsy
- V. Treat Petit Mal Epilepsy
- W. Ultra Short Acting Barbiturate

- | | | |
|-----------------------|---------------------|--------------|
| 1. Acetophenetidin | 26. Hypan | 51. Tridione |
| 2. Amytal | 27. Indocin | 52. Trilafon |
| 3. Ascriptin | 28. Librium | 53. Tylenol |
| 4. Atarax | 29. Lorfan | 54. Valium |
| 5. Aventyl | 30. Luminal | 55. Vistaril |
| 6. Azolid | 31. Mebaral | 56. Vivactil |
| 7. Beta Chlor | 32. Mellaril | 57. Zactane |
| 8. Bufferin | 33. Methylnorphine | 58. Zarontin |
| 9. Butazolidin | 34. Miltown | |
| 10. Caffeine Citrated | 35. Mysoline | |
| 11. Cendex | 36. Nalline | |
| 12. Compazine | 37. Noludar | |
| 13. Daro Tab | 38. Nembutal | |
| 14. Darvon | 39. Oxalid | |
| 15. Demerol | 40. Paracetaldehyde | |
| 16. Dexedrine | 41. Parnate | |
| 17. Dihycon | 42. Pentothal | |
| 18. Dilantin | 43. Ritalin | |
| 19. Dilaudid | 44. Seconal | |
| 20. Dimindol | 45. Somnes | |
| 21. Dolophine | 46. Sparine | |
| 22. Doriden | 47. Stelazine | |
| 23. Elavil | 48. Tandearil | |
| 24. Equanil | 49. Thorazine | |
| 25. Haldol | 50. Tofranil | |

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CNS STIMULANTS

_____ Caffeine
 _____ Dextroamphetamine

CNS DEPRESSANTS

_____ Amobarbital
 _____ Chloral Betaine
 _____ Glutethimide
 _____ Methypylon
 _____ Mephobarbital

_____ Paraldehyde
 _____ Phenobarbital
 _____ Pentobarbital
 _____ Secobarbital
 _____ Thiopental

NON-NARCOTIC ANALGESIC

_____ Acetaminophen
 _____ Aspirin
 _____ Etchoheptazine
 _____ Indomethacin

_____ Oxyphenbutazone
 _____ Phenacetin
 _____ Phenylbutazone
 _____ Propoxyphene

NARCOTIC ANALGESIC

_____ Codeine
 _____ Hydromorphone
 _____ Levallorphan

_____ Meperidine
 _____ Methadone
 _____ Nalorphine

ANTIEPILEPTICS

_____ Diphenylhydantoin
 _____ Ethosuximide

_____ Primidone
 _____ Trimethadione

PSYCHOTHERAPEUTICS

_____ Amitriptyline
 _____ Chlordiazepoxide
 _____ Chlorpromazine
 _____ Diazepam
 _____ Haloperidol
 _____ Hydroxyzine
 _____ Methylphenidate
 _____ Meproamate
 _____ Imipramine

_____ Nortriptyline
 _____ Perphenazine
 _____ Prochlorperazine
 _____ Promazine
 _____ Protriptyline
 _____ Thioridazine
 _____ Tranlycypromine
 _____ Trifluoperazine

Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas

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HANDOUT 3ABR90530-II-5
November 1974

PHARMACOLOGY

CARDIAC GLYCOSIDES

_____ Digitoxin
_____ Digoxin

Antiarrhythmia

_____ Quinidine
_____ Procainamide

CORONARY VASODILATORS

_____ Nitroglycerin
_____ Pentaerythritol Tetranitrate

HYPERTENSION

_____ Reserpine
_____ Isoxsuprine
_____ Hydralazine
_____ Methyldopa
_____ Guanethadine

HEMATINICS

_____ Ferrous Sulfate
_____ Ferrous Fumarate and
_____ Dioctyl Sodium Sulfosuccinate

COAGULANTS

_____ Menadione Sodium
_____ Menadione

ANTICOAGULANTS

_____ Warfarin Sodium
_____ Heparin Sodium
_____ Bishydroxycoumarin

1. Aldomet
2. Apresoline
3. Coumadin
4. Crystodigin
5. Davoxin
6. Dicumarol
7. Ferro-Sequels
8. Glyceryl Trinitrate
9. Ismelin
10. Lanoxin
11. Lipo-Heparin
12. Myodigin
13. Panheprin
14. Peritrate
15. Pentritol
16. Pronestyl
17. Purodigin
18. Sandril
19. Saroxin
20. Serfin
21. Serpasil
22. Vasodilan

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SYMPATHOMIMETIC DRUGS (Adrenergic)

_____ Epinephrine
 _____ Levarterenol
 _____ Metaraminol
 _____ Phenylephrine
 _____ Phenylpropanalamine
 _____ Dextroamphetamine
 _____ Diethylpropion
 _____ Phenmetrazine
 _____ Nyldrin
 _____ Isoproterenol

SYMPATHOLYTIC DRUGS (Adrenergic Blocking)

_____ Tolazoline
 _____ Phentolamine
 _____ Ergotamine
 _____ Ergotamine with Caffeine
 _____ Methysergide

PARASYMPATHOMIMETIC DRUGS (Cholinergic)

_____ Bethanechol
 _____ Neostigmine

PARASYMPATHOLYTIC DRUGS (Cholinergic Blocking)

_____ Propantheline
 _____ Benztropine
 _____ Dicyclomine
 ✓ _____ Trihexyphenidyl

MUSCLE RELAXANTS

_____ Meprobamate
 _____ Mephensin
 ✓ _____ Methocarbamol
 _____ Carisprodol
 _____ Chlorzoxazone
 _____ Succinylcholine

1. Adrenalin
2. Anectine
3. Aramine
4. Arlidin
5. Artane
6. Bently
7. Caffergot
8. Cogentin
9. Dexedrine
10. Equanil
11. Gynergen
12. Isuprel
13. Levophed
14. Miltown
15. Norisodrine
16. Paraflex
17. Pipanol
18. Preludin
19. Pressonex
20. Priscoline
21. Probanthine
22. Propadrine
23. Prostigmin
24. Quelicin
25. Rela
26. Regitine
27. Robaxin
28. Sansert
29. Soma
30. Sucostrin
31. Tenuate
32. Tolseram
33. Tolserol
34. Tremin
35. Urecholine

- a. Antiarrhythmia
- b. Appetite depressant
- c. Asthma
- d. Causes constipation
- e. Central acting muscle relaxant
- f. Central acting muscle relaxant, chief use as tranquilizer
- g. Diagnosis of pheochromocytoma
- h. Elevate blood pressure during shock
- i. Hematinic
- j. Narcolepsy
- k. Nasal decongestant
- l. Neuromuscular blocking agent, used in surgery
- m. Ocular decongestant
- n. Parkinson's disease
- o. Possibility of hemorrhage
- p. Prophylaxis for migraine headache
- q. Toxicity called CHINCHONISM
- r. Treat angina pectoris
- s. Treat congestive heart failure
- t. Treat migraine headaches
- u. Used during blood transfusions
- v. Treat moderately severe hypertension
- w. Treat myasthenia gravis
- x. Treat severe hypertension, ganglionic blocker
- y. Treat urinary retention
- z. Treatment of peptic ulcer
- aa. Treat mild hypertension
- bb. Used with local anesthetics
- cc. Vasodilator in peripheral vascular disease
- dd. Vitamin K injectable
- ee. Vitamin K oral

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Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas

HANDOUT 3ABR90530-II-6
November 1974

PHARMACOLOGY

ENDOCRINE AND MISCELLANEOUS WORKSHEET

- a. Most common glucocorticoid
- b. Treats prostate cancer
- c. Passive immunity
- d. Active immunity
- e. Suppresses uterine contraction
- f. Used in emergency treatment of diabetic coma
- g. Treatment of mild diabetes mellitus
- h. Oral form of testosterone
- i. Treatment of hyperthyroidism
- j. Most commonly used insulin preparation
- k. Synthetic thyroid
- l. Used to increase the absorption of drugs
- m. Antitussive with local anesthetic effect
- n. Most potent diuretics
- o. Aldosterone antagonist
- p. Toxin
- q. Sedative expectorant
- r. Not to be used in combination with oxytocin
- s. Used to treat motion sickness
- t. Treats beri-beri
- u. Needed for the prevention of night blindness
- v. Treats pellegra
- w. Anabolic agent
- x. Sequential oral contraceptive
- y. Treats menopausal symptoms and dysmenorrhea
- z. Antiemetic, augment and potentiate the CNS depressants

- | | | | |
|---------------------|-------------------|--------------------|-------------------|
| 1. Aldactone | 22. Dymelor | 43. Nilevar | 65. Riboflavin |
| 2. Alpha-tocopheryl | 23. Elase | 44. Norlestrin | 66. Romilar |
| 3. Aristocort | 24. Enovid E | 45. NPH Iletin | 67. Sabin (Oral) |
| 4. Ascorbic acid | 25. Estinyl | 46. Ojeovitamin | 68. Shick Test |
| 5. Benadryl | 26. Ergotrate | 47. Oncovin | 69. Spartocin |
| 6. Bonine | 27. Guaiacol | 48. Oracon | 70. Stillbesterol |
| 7. Chlortrimeton | 28. Lasix | 49. Oreton | 71. Synalar |
| 8. Cortone | 29. Lente Iletin | 50. Ortho-Novum | 72. Tapazole |
| 9. Cortril | 30. Leukeran | 51. Ovulen | 73. TAT |
| 10. Cyanocobalamin | 31. Medrol | 52. Parathormone | 74. Temaril |
| 11. Cytomel | 32. Menadione | 53. Periactin | 75. Terpinol |
| 12. Cytosan | 33. Mercuhydrin | 54. Phenergan | 76. Theelin |
| 13. DBI | 34. Methergine | 55. Pitocin | 77. Thiamine |
| 14. Decadron | 35. Methotrexate | 56. PPD | 78. Thimerin |
| 15. Diabinese | 36. Metandren | 57. Premarin | 79. Thytropar |
| 16. Diamox | 37. Meticortelone | 58. Provera | 80. Tolinase |
| 17. Dick Test | 38. Meticorten | 59. Purinethol | 81. TNTC |
| 18. Dimetane | 39. Mustargen | 60. Pyribenzamine | 82. Varidase |
| 19. Diuril | 40. Myleran | 61. Pyridoxine | 83. Wydase |
| 20. DPT | 41. Neutrapen | 62. PZI Iletin | |
| 21. Dramamine | 42. Niacin | 63. Regular Iletin | |

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ENDOCRINE AND MISCELLANEOUS WORKSHEET

Adrenal Hormones

- ___ Cortisone
- ___ Hydrocortisone
- ___ Prednisone
- ___ Prednisolone
- ___ Methylprednisolone
- ___ Methylprednisolone
- ___ Dexamethasone
- ___ Fluocinalone
- ___ Triamcinalone

Insulin Hormones

- ___ Regular Insulin
- ___ Protamine Zinc Insulin
- ___ Isophane Insulin
- ___ Insulin Zinc Suspension

Oral Hypoglycemics

- ___ Acetohexamide
- ___ Chlorpropamide
- ___ Tolazamide
- ___ Phenformin

Thyroid Hormones

- ___ Thyroid USP
- ___ Liothyronine
- ___ Thyrotropin

Antithyroid Hormones

- ___ Propylthiouracil
- ___ Methimazole

Parathyroid Hormone

- ___ Parathyroid Inj.

Androgen Drugs

- ___ Testosterone
- ___ Methyltestosterone
- ___ Norethandrolone

Estrogen and Progesterone Drugs

- ___ Ethinyl Estradiol
- ___ Estrone
- ___ Conjugated Estrogens
- ___ Diethylstilbesterol
- ___ Medroxyprogesterone
- ___ Ethinyl Estradiol and Dimethisterone with Ethinyl Estradiol
- ___ Ethynodiol Diacetate and Mestranol
- ___ Norethynodrel and Mestranol
- ___ Norethindrone and Mestranol
- ___ Norethindrone Acetate and Ethinyl Estradiol

ENDOCRINE AND MISCELLANEOUS WORKSHEET

Dietary Supplements

- ☐ Vitamine A
- ☐ Vitamin E
- ☐ Vitamin K
- ☐ Vitamin B₁
- ☐ Vitamin C
- ☐ Nicotinic Acid
- ☐ Vitamin B₆
- ☐ Vitamin B₂
- ☐ Vitamin B₁₂

Immunological Agents

- ☐ Polio Vaccine
- ☐ Diptheria Toxin USP
- ☐ Scarlet Fever Streptococcus Toxin USP
- ☐ Purified Protein Derivative of Tuberculin USP
- ☐ Diptheria and Tetanus Toxoids and Pertussis Vaccine
- ☐ Tetanus Antitoxin USP

Diuretics

- ☐ Mercaptopurin
- ☐ Meralluride
- ☐ Chlorothiazide
- ☐ Furosemide
- ☐ Spironalactone
- ☐ Acetazolamide

Expectorants and Antitussives

- ☐ Potassium Iodide
- ☐ Ammonium Chloride
- ☐ Glyceryl Guaiacolate
- ☐ Terpin Hydrate
- ☐ Dextromethorphan
- ☐ Benzonatate

Oxytocics

- ☐ Ergonovine
- ☐ Methylergonovine
- ☐ Sparteine Sulfate
- ☐ Oxytocin

Enzymes

- ☐ Hyaluronidase
- ☐ Penicillinase
- ☐ Fibrinolysin and Desoxyribonuclease
- ☐ Streptokinase-Streptodornase

Antineoplastic Agents

- ☐ Aminopterin
- ☐ Busulfan
- ☐ Mercaptopurine
- ☐ Methchloroethamine
- ☐ Chlorambucil
- ☐ Cyclophosphamide
- ☐ Vincristine
- ☐ Estrogens

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ENDOCRINE AND MISCELLANEOUS WORKSHEET

Antihistamines

Diphenhydramine

Triphenennamine

Chlorpheniramine

Brompheniramine

Dimenhydrinate

Meclizine

Promethazine

Trimeprazine

Cyproheptadine

Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas

HANDOUT 3ABR90530-II-7
November 1974

PHARMACOLOGY

(ENDOCRINE AND MISCELLANEOUS WORKSHEET)

Match Generic names in COLUMN A with drug categories in COLUMN B

COLUMN A	COLUMN B
1. Prednisolone.....	a. Adrenal Hormone
2. Ammonium Chloride.....	b. Antithyroid Hormone
3. Dexamethasone.....	c. Androgen Drug
4. Propylthiouracil.....	d. Estrogen
5. Fluocinolone.....	e. Progesterone
6. Methylprednisolone.....	f. Dietary Supplement
7. Busulfan.....	g. Antihistamine
8. Dextromethorphan.....	h. Antitussive
9. Furosemide.....	i. Sedative Expectorant
10. Brompheniramine.....	j. Oxytocic
11. Estrone.....	k. Enzyme
12. Protamine Zinc Insulin.....	l. Antineoplastic agent
13. Folic Acid.....	m. Acidifying diuretic
14. Acetazolamide.....	n. Osmotic diuretic
15. Mercaptopurine.....	o. Carbonic Anhydrase
16. Ergonovine.....	p. Onset 4-6 hrs., duration 24-43 hrs (Insulin)
17. Cyclophosphamide.....	q. Onset 2 hrs., duration 16-24 hrs (Insulin)
18. Chlorpheniramine.....	r. Oral contraceptive
19. Vincristine.....	s. Thiazide diuretic
20. Nicotinic Acid.....	t. Miscellaneous diuretic
21. Fibrinolysin and Desoxyribonuclease.....	
22. Ethinyl Estradiol.....	
23. Insulin Zinc Suspension.....	
24. Methchlorethamine.....	
25. Methimazole.....	
26. Prednisone.....	
27. Streptokinase-Streptornase.....	
28. Aminopterin.....	
29. Diphenhydramine.....	
30. Triamcinolone.....	
31. Testosterone.....	
32. Mannitol Inj.	
33. Codene.....	
34. Penicillinase.....	
35. Methylegonovine.....	
36. Androgens.....	
37. Conjugated Estrogens.....	
38. Oxytocin.....	
39. Ethynodiol Diacetate and Mestranol.....	
40. Potassium Iodide.....	
41. Chlorambucil.....	
42. Tripeleminamine.....	
43. Chlorothiazide.....	
44. Medroxyprogesterone.....	

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10-8

Technical Training

Pharmacy Specialist

PHARMACOLOGY

October 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
Department of Biomedical Sciences
Sheppard Air Force Base, Texas 76311

10
Designed For ATC Course Use

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Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas 76311

WB 3ABR90530-II-1
October 1975

PHARMACOLOGY

OBJECTIVE

Given information pertaining to pharmacological principles of selected drug groups, complete questions in WB 3ABR90530-II pertinent to each day's instruction.

EQUIPMENT

Selected Transparencies
Overhead Projector
Selected Motion Pictures
Motion Picture Projectors

PROCEDURE

Define and identify selected drugs in relation to their groups.

The object of this lesson is to acquaint you with the primary drug groups and their actions.

INSTRUCTIONS

For the most part, answers to the following questions must be obtained from the class lecture. Remington's Pharmaceutical Sciences and Cutting's Handbook of Pharmacology contain useful, supplementary information. By completing the workbook you will more easily understand the lesson. You will also have an excellent source of review material for the test.

STUDY REFERENCES

1. Remington's Pharmaceutical Sciences.
2. Cutting's Handbook of Pharmacology.

This supersedes WB 3ABR90530-II-1, October 1974

QUESTIONS

1. Define Pharmacology.

2. Define a drug

3. List six general drug uses.

4. Name four sources of drugs.

SOURCES

- a. _____
- b. _____
- c. _____
- d. _____

5. Broadly speaking, medications may be administered externally or _____.

6. Ointments, creams, and lotions are examples of _____ drug administration.

7. The simplest, most painless way to give a drug internally is _____.

8. List two advantages and two disadvantages of administering drugs orally.

ADVANTAGES

- a. _____
- b. _____

DISADVANTAGES

- a. _____
- b. _____

9. Drugs are inserted into the rectum in the form of _____
or _____.
10. Parenteral refers to the administration of drugs by _____.
11. List two advantages and two disadvantages of administering drugs by injection.
ADVANTAGES
 - a. _____
 - b. _____DISADVANTAGES
 - a. _____
 - b. _____
12. Sublingual medications are placed under the _____, whereas
buccal medications are placed between the _____ and
the _____.
13. Three factors which determine how often a drug must be administered are
absorption, _____, and _____.
14. List five overall effects of drug action.
 - a. _____
 - b. _____
 - c. _____
 - d. _____
 - e. _____
15. A drug which stimulates a cell function does so by _____
the activity of cells.
16. Drugs which produce sleep are examples of drugs acting by _____.
The activity of cells is _____ in this method of drug action.
17. A third method of drug action is irritation which causes _____
of cells.
18. Insulin and Thyroid administration are examples of _____
action.
19. Penicillin is an example of a drug which acts by _____.

20. List four types of drug action:
- a. _____
 - b. _____
 - c. _____
 - d. _____
21. In reflex action, the effects are produced as a result of a local _____.
22. Before general action is produced, the drug must be absorbed into the _____.
23. The type of action produced only at the point of contact is _____.
24. An undesirable or secondary effect of a drug is called a _____ action.
25. List six factors which modify the action or dose of a drug.
- a. _____
 - b. _____
 - c. _____
 - d. _____
 - e. _____
 - f. _____
26. These factors alter the _____ of a drug necessary to produce the desired results.
27. A drug interaction is when the effect of one drug _____ or _____ the effects of another drug.
28. _____ is when two drugs exert a greater combined effect than either one by itself.
29. _____ is when the effect of one drug cancels or reduces the effects of the other drug.
30. Drug interactions may be either _____ or _____.

TOXICOLOGY

QUESTIONS

1. Toxicology is the science of _____.
2. Explain cumulative effect of poison.
3. Define acute poisoning.
4. Chronic poison is the condition brought on by taking _____ doses of a poison or as a result of the _____ of a poison over a _____.
5. Hydrochloric acid, Phenol, and Sodium Hydroxide are examples of corrosive poisons. This type of poison _____ or _____ body tissues at the point of contact.
6. Zinc Sulfate, Silver Nitrate, and Iodine are examples of irritant poisons. This type of poison sets up an _____ process at the point of contact. They do not directly destroy body tissue.
7. The two types of neurotic poisons are _____ and _____ . They act on the central nervous system.
8. Gaseous poisons may destroy the capability of the blood to carry _____ and may _____ or _____ the tissues of the air passages, lungs and skin.

9. NEVER induce vomiting or use a stomach tube when _____ poisons have been ingested. In this type of poisoning there is a danger of rupturing the weakened wall of the stomach or esophagus.

10. When poisoning is due to irritants or neurotics, remove the poison from the stomach as soon as possible by inducing _____ or by gastric tube.

11. In the case of gaseous poisoning, get the victim into fresh air and start _____ promptly if patient is not breathing.

12. A clearing house for poison information is called a _____. It provides valuable information on the treatment of poisonings. Every medical facility should try to utilize the service nearest them.

DRUG ABUSE

QUESTIONS

1. This lesson is concerned with drugs subject to abuse which act on the central nervous system to produce changes in mood and _____.
2. When it takes a larger dose of a drug to produce the same effects as the original dose, we can say the person has developed a _____.
3. In habituation the harmful effects of a drug are primarily on the _____
_____. There is little or no physical dependence; however, psychic dependence does occur.
4. Physical dependence is a component of the condition known as addiction.
In this condition the harmful effects of the drug are on _____
as well as the individual.
5. List three basic causes of drug abuse.
 - a. _____
 - b. _____
 - c. _____
6. List the six groups of drugs subject to abuse.
 - a. _____
 - b. _____
 - c. _____
 - d. _____
 - e. _____
 - f. _____
7. Narcotics are used in medicine to relieve or modify _____
_____, suppress _____, and control _____.
8. Narcotics are used illegally for their ability to produce a feeling of _____
_____ and an escape from _____.
9. The primary class of drugs in the depressant category are the _____.
10. Barbiturates are used in medicine for their _____
and _____ effects.

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11. Barbiturates are abused for their _____ effects.
12. Amphetamines are used in medicine to _____ mood, suppress _____, and to treat _____.
13. Stimulants are abused for their mood _____ effects and their ability to overcome _____.
14. Hallucinogenic drugs cause distortions of _____, dream _____, and radically alter _____.
15. Hallucinogenic drugs have a legitimate medical use.
True _____ False _____
16. Marijuana is abused for its _____ and _____ effects.
17. LSD is abused for its so-called " _____ " effects.
18. There is a danger of _____ in inhaling solvent fumes.

LOCALLY ACTING DRUGS

QUESTIONS

1. The outer integument or covering of the body, consisting of two basic layers of tissue is the definition of _____.
2. The external or outer surface of the skin is the _____.
3. The bottom layer of skin that contains blood and lymph vessels, hair follicles, sebaceous glands, and sweat glands is the _____.
4. The functions of skin are _____, heat regulation, prevention of bacterial invasion, metabolic processes, and excretion.
5. Melanin prevents tissue damage caused by _____ light.
6. The body excretes waxes and oils through the _____ glands located in the skin.
7. The tissue lining the cavities and canals of the body is called _____.
8. Mucous membranes are more _____ than skin and have mucous glands rather than cutaneous glands.
9. The functions of mucous membranes are protection, secretion, and _____.
10. Bland, fatty, or oleaginous substances that soften the skin and protect the skin are called _____.
11. List three emollients.
 - a.
 - b.
 - c.
12. Demulcents are protective agents which are employed primarily to alleviate irritation and protect the _____.
13. List three demulcents.
 - a.
 - b.
 - c.
14. Flexible Collodion USP and Absorbable Gelatin Film (Gelfilm) are classified as _____.
15. Protectives are administered to the _____ to form an adherent, continuous coat which may be either flexible or semi-rigid, depending on the substances and the manner in which they are applied.

16. Locally applied protein precipitants that contract and wrinkle the skin or tissues are classified as _____.

17. List two astringents

a.

b.

18. _____ are used to remove unwanted substances and microorganisms from living tissue and inanimate objects.

19. Detergents, Solvents, and Abrasives are all examples of _____ agents.

20. Which of the above remove unwanted substances by mechanical action?

21. Keratolytics _____ the horny layer of skin where most fungi and bacteria reside.

22. Most keratolytics are _____.

23. List three keratolytics.

a.

b.

c.

24. Irritants are drugs which act locally on the skin and mucous membranes to promote healing by increasing _____ flow to the injured area.

25. List five irritants.

a.

b.

c.

d.

e.

26. Certain relatively indifferent (inert and insoluble) substances that are used to protect epithelial surfaces, ulcers, and wounds by the absorption of skin moisture and decreasing friction are called _____.

27. List three absorbents

a.

b.

c.

GASTROINTESTINAL DRUGS

QUESTIONS

1. The Gastrointestinal system acts on food both _____ and chemically.
2. Food and drug absorption takes place in the _____ intestine.
3. Food and drug enter the bloodstream through millions of tiny fernlike projections in the small intestine called _____.
4. _____ are drugs promoting or aiding the digestive processes in the gastrointestinal tract.
5. Hydrochloric Acid (HCl), Bile, Bile Salts, and Bile Acids are classified as _____.
6. Which digestant can cause damage to the teeth if not taken properly?
7. Bile has a variety of functions but the most important is the digestion of _____ and fat soluble vitamins. Also, it reduces surface tension of fats and activates pancreatic _____.
8. Pepsin, Pancreatin, and Papain are digestive enzymes. (True) (False)
9. Gastric Antacids are agents that neutralize or remove _____ from gastric contents.
10. Sodium Bicarbonate is highly soluble, acts immediately, causes rebound hyperacidity, and is absorbed _____.
11. Aluminum, Calcium, and Magnesium salts are used as nonsystemic _____.
12. Activated Charcoal USP, Kaolin NF, and Pectin NF are all examples of _____.
13. Kaopectate is used to treat diarrhea and is a combination of Kaolin and _____.
14. Drugs that facilitate the passage and elimination of feces from the colon and rectum are called _____.

15. List the five classes of cathartics.

- a.
- b.
- c.
- d.
- e.

16. Cascara Sagrada USP, Senna NF, Castor Oil USP, and Bisacodyl NF are all _____ cathartics.

17. Magnesium Sulfate USP, Milk of Magnesia USP, and Fleet Enema are all _____ cathartics.

18. Psyllium Hydrophilic Mucilloid (Metamucil) is a _____ bulk cathartic.

19. Mineral Oil USP (Heavy Liquid Petrolatum) and Cottonseed Oil USP are _____ cathartics.

20. Dioctyl Calcium Sulfosuccinate NF (Surfak) and Dioctyl Sodium Sulfosuccinate USP (Colace) are _____ softeners or "surface-acting" agents.

21. An emetic is a drug which induces _____.

22. _____ HCl is a systemic emetic that acts directly by stimulating the medulla oblongata and is given by injection.

23. Sodium Chloride USP (Table Salt), and Cupric Sulfate NF (Copper Sulfate) are nonsystemic emetics that act directly on the stomach lining to cause _____ to the gastric mucosa.

LOCAL ANESTHETICS

QUESTIONS

1. Local Anesthetics are drugs which produce a condition of anesthesia in a _____ area around the site of application or injection of the drug. They interfere with the initiation and transmission of the nerve impulse.
2. Pain is a specific _____ experience which is separate from those which mediate other sensations such as touch, pressure, heat and cold.
3. Sensory nerve fibers terminate as _____ nerve endings.
4. Define the three classes of local anesthetics.
 - a. Refrigerants:
 - b. Protoplasmic Poisons:
 - c. Specific Anesthetics:
5. Local anesthetics have a mechanism of action that prevents passage of impulses through sensory nerve endings (Bare nerve endings) or prevents passage of impulses through the nerve _____ (Bundle of nerve fibers).
6. Match the following Methods of Administration:

_____ Applied to the skin and mucous membrane surface	a. Saddle Block
_____ Injection directly into the area that is painful or to be subjected to surgical trauma	b. Topical
_____ Injection into a nerve trunk	c. Spinal
_____ Injection into the spine between the 3rd and 4th or 4th and 5th lumbar vertebrae. Mixes with spinal fluid.	d. Infiltration
_____ A form of low spinal that affects the perineal area.	e. Epidural and Caudal Block
_____ Spinal fluid is not affected.	f. Block

7. Cocaine HCl is a potent anesthetic that is highly effective on mucous membranes and is never to be _____ because of its high toxicity.
8. In addition to its anesthetic properties, Cocaine is also a powerful _____.
9. Procaine (Novacaine) is ineffective when applied _____.
10. Ethyl Aminobenzoate (Benzocaine) is insoluble. Therefore, it is not to be _____ but should be used topically only.
11. _____ (Ophthaine) (Ophthalmic) is an ophthalmic anesthetic that should be kept in the refrigerator upon opening.
12. Dibucaine (Nupercaine) is _____ more toxic and potent than procaine if injected and is used mainly on the mucous membranes in a 1 percent ointment.
13. Ethyl Chloride USP is skin _____.
14. Phenol USP (Carbolic Acid) is a local anesthetic that is applied topically and is classed as a _____ poison because it kills the cell by the precipitation of protein.

ANTI-INFECTIVE DRUGS

QUESTIONS

1. Those parasites or pathogenic organisms that invade or infest the body causing a reaction of the tissues by the toxins generated by them is the definition of _____
2. The smallest of all infectious organisms are the _____ which are characterized by a lack of metabolism and proliferate only in the presence of living tissue.
3. Diseases caused by viruses are
 - a.
 - b.
 - c.
 - d.
4. Minute rod-shaped parasites that cause disease in man by arthropod vectors are called _____.
5. Rocky Mountain Spotted Fever, Typhus, and Mite Fever are diseases caused by _____.
6. Bacteria are one-celled microorganisms that belong to the _____ kingdom.
7. Bacteria are classified by a procedure called "_____ Stain."
8. Bacteria that stains blue with Gram Stain is classified as Gram _____.
9. List six diseases caused by Gram Positive organisms:
 - a.
 - b.
 - c.
 - d.
 - e.
 - f.
10. Bacteria that stains red or pink is classified as Gram _____.

11. List five diseases caused by Gram Negative organisms:

- a.
- b.
- c.
- d.
- e.

12. The fungi are small _____ that have no roots, stems, or chlorophyll.

13. Diseases caused by fungi are:

- a.
- b.
- c.
- d.
- e.

14. A protozoa is a true member of the _____ kingdom and does not have a cell wall.

15. Four classes of protozoa are the Amoeba, Flagellates, Ciliates, and _____.

16. Diseases caused by protozoa are:

- a.
- b.
- c.
- d.

17. An intestinal worm or worm-like parasite is the definition of _____.

18. There are three classifications of Helminths: the Nematodes, Cestodes and _____.

19. List the following Nematodes:

- a. Roundworm
- b.
- c. Whipworm
- d.

20. List the following Cestodes:

- a. Beef tapeworm
- b. Pork tapeworm
- c.

21. List the following Trematodes:

- a.
- b.

22. Lice (crabs) are also referred to as _____.

23. Diseases caused by lice are Typhus, Lapsing Fever, and _____.

24. Minute animals related to the spiders and are parasitic on man and domestic animals producing various irritations on the skin are called _____.

25. Scabies, inflammation, and secondary infections are caused by _____.

26. The treatment of disease by administering chemicals which affect the causative organism unfavorably but do not injure the patient is the definition of _____.

27. Antibiotics include a large class of drugs chemically produced by _____ able to inhibit growths of or destroy bacteria and other disease-causing pathogens.

28. Penicillin is the most widely used of all the antibiotics and is effective against most Gram _____ organisms and against some Gram Negative organisms such as gonococci and spirochetes.

29. Penicillin G is a _____ penicillin.

30. Penicillin V has the same properties as Penicillin G, but unlike Penicillin G, it is more stable in an acid medium causing it to be better _____ in the gastrointestinal tract.

31. Methicillin (Staphcillin) is a semi-synthetic penicillin intended to combat _____-producing staphylococci resistant to other penicillins.

32. Methicillin cannot be taken _____.

33. Oxacillin (Prostaphlin) is intended to combat penicillinase producing staphylococci but, unlike Methicillin, it is resistant to gastric acids so it can be taken _____.

34. Nafcillin (Unipen) is a semi-synthetic penicillin primarily used in treatment of Penicillin _____-resistant staphylococcal infections.

35. Ampicillin has a wider range of activity than Penicillin G, but is destroyed by _____.

36. The use of the drugs in the tetracycline class during tooth development (last half of pregnancy, infancy, and up to 8 years of age) may cause permanent discoloration of the _____.

37. Monilia _____ sometimes occurs with oral use, therefore, the tetracyclines are often combined with an antimycotic (antifungal) agent.

38. Demeclocycline HCl has the slowest rate of excretion of all the tetracyclines and also has been known to cause photodynamic and _____ reactions.

39. Streptomycin Sulfate is given by injection only and is used primarily to treat tuberculosis. However, this drug may produce toxic effects in the liver or kidneys and cause damage to the 8th cranial nerve which results in tinnitus, vertigo, and an eventual loss of _____.

40. Chloramphenicol (Chloromycetin) is highly effective against certain rickettsial infections such as Typhus and Rocky Mountain Spotted Fever but serious and sometimes fatal blood _____ may occur with its use.

41. Erythromycin (Ilotycin, Erythrocin, Ilosone) is mainly used in the treatment of _____ sensitive organisms where the patient has a penicillin sensitivity.

42. Neomycin is not _____ when taken orally.

43. Neomycin is taken orally to treat infections of the _____ tract.

44. If injected, Neomycin may cause severe kidney damage and _____ loss.

45. The trade name for Lincomycin HCl Monohydrate is _____.

46. Cephalothin (Keflin) is a broad spectrum antibiotic somewhat similar to the _____ group.

47. Keflin is only given by _____.

48. Kanamycin Sulfate (Kantrex) is chemically related to Neomycin and Streptomycin and can also cause _____ loss.

49. Polymixin B Sulfate is used to treat Gram _____ bacteria.

50. Nitrofurantion USP (Furadantin) and Nalidixic Acid NF (Neg Gram) are used to treat infections of the _____ tract.

51. Nalidixic Acid NF (Neg Gram) has an outstanding cure rate against Proteus infections but bacterial _____ develops rapidly.

52. Sulfonamides are classified as either _____ or non _____.
53. The systemic Sulfonamides are either rapidly or slowly _____.
54. Match the following:

_____ Sulfadiazine	a. Systemic
_____ Sulfamerazine	b. Nonsystemic
_____ Sulfamethazine	c. Topical
_____ Sulfisoxazole (Gantrisin)	
_____ Succinylsulfathiazole (Sulfasuxidine)	
_____ Phthalylsulfathiazole (Sulfathalidine)	
_____ Sulfacetamide Sodium (Sulamyd)	
_____ Mafenide (Sulfamylon)	
_____ Salicylazosulfapyridine (Azulfadine)	
_____ Sulfamer (Sulla)	
_____ Sulfadimethoxine (Madrison)	
55. Trisulfapyrimidines Suspension (Sulfose, Triple Sulfa) is a combination of Sulfamerazine, Sulfamethazine, and _____.
56. Because Gantrisin is very soluble, the incidence of _____ is very slight.
57. Sulfacetamide Sodium (Sulamyd) Ophthalmic Solution must be stored in the _____.
58. A topical sulfonamide used in the treatment of burns is _____.
59. Azulfadine is nonsystemic and used in the treatment of _____.
60. Those agents which kill or inhibit fungi are called _____.
61. Amphotericin B USP is used both systemically and topically and has the widest _____ of antifungal activity of all the systemic antifungal drugs.
62. Griseofulvin (Fulvicin, Grifulvin, Grisactin) is an _____ antifungal agent that is used in the treatment of superficial fungus infections.
63. Griseofulvin is very effective in the treatment of a skin disease called _____ worm.
64. Nystatin USP (Mycostatin) is used both systemically and topically. It is used orally to treat intestinal _____.
65. Tolnaftate USP (Tinactin) is applied locally and affects those diseases of the skin that are susceptible to _____ therapy.

66. Undecylenic Acid NF has the trade name of _____.

67. The three categories of Antiprotozoals are the Antimalarials, _____, and miscellaneous Antiprotozoals.

68. Match the following:

_____ Chloroquine	a. No trade name given
_____ Carbarsone	b. Diodoquin, Floraquin
_____ Diiodohydroxyquin	c. Humatin
_____ Emetine HCl	d. Vioform
_____ Idochlorhydroxyquin	e. Aralen
_____ Paromomycin Sulfate NF	

69. Because Emetine HCl is concentrated and stored in the liver, it is highly effective in the treatment of amebic _____.

70. Idochlorhydroxyquin USP (Vioform) and Diiodohydroxyquin USP (Diodoquin, Floraquin) are nonsystemic and are effective in the treatment of _____ amebiasis.

71. A patient should be cautioned about drinking _____ beverages when taking Metronidazole USP (Flagyl).

72. Tryparsamide USP XVII is an antiprotozoal agent that can cause _____.

73. Those drugs used to combat any type of helminthiasis are called _____.

74. Antimony Potassium Tartrate USP is used in the treatment of _____ flukes and the patient may exhibit the effects of heavy metal poisoning during treatment.

75. Bephenium Hydroxynaphthoate (Alcopara) is the drug of choice in the treatment of _____.

76. Diethyl Carbamazine Citrate USP has the trade name of _____ and is used to treat roundworms.

77. Hexylresorcinol NF (Crystoids) can cause a painful ulceration of the oral mucous membrane and caustic burns if the pills are not swallowed _____.

78. Lucanthone HCl USP is used to treat blood and liver _____.

79. Piperazine Citrate USP (Antepar Citrate) and Pyriwinium Pamoate USP (Povan) are the drugs of choice in the treatment of _____.

80. In addition to treating _____, Antepar is also effective in the treatment of roundworms.

81. A _____ is highly acceptable to the patient when taking Povan, but the drug colors the stools and teeth bright red.

82. Tetrachloroethylene USP is used in the treatment of hookworms, _____, pinworms, and flukes.

83. An agent that destroys the itch mite and eggs on the skin of man is called a _____.
84. Gamma Benzene Hexachloride (Kwell, Lindane) is a pediculocide and a _____.
85. Benzyl Benzoate USP (Zylate) Crothamiton BP (Eurax) and Precipitated Sulfur USP are classed as _____.
86. Pediculocides are agents that destroy body _____ and their eggs.
87. Gamma Benzene Hexachloride (Kwell, Lindane) and Chlorophenothane (DDT) are effective _____.
88. Agents that kill microbes on contact are called _____ and are classified into two categories called Antiseptics and Disinfectants.
89. Antiseptics are applied to _____ to kill or prevent the growth of pathogens.
90. Disinfectants are used on _____ to kill or prevent the growth of pathogens.
91. Which germicide has a residual effect that is destroyed only by organic solvents?

92. Methenamine Mandelate USP (Mandelamine) is a urinary tract antiseptic that depends upon the liberation of _____ for its action.

DRUGS ACTING ON THE CENTRAL NERVOUS SYSTEM

QUESTIONS

Divisions of the Central Nervous System (CNS) and their Functions

1. The CNS is composed of the _____ and the _____.
2. Sites of consciousness, memory, and sensation are located in the _____.
3. The part of the CNS which relays sensory impulses from the body to sensory areas of the brain is known as the _____.
4. The part of the CNS which regulates body temperature is known as the _____.
5. The _____, _____, and _____ control centers are located in the medulla oblongata.
6. The _____ controls posture and coordinates motor responses concerned with maintenance of equilibrium.
7. What are the two functions of the spinal cord?
 - a. _____
 - b. _____

Central Nervous System Stimulants

8. List the two categories of CNS stimulants.
 - a. _____
 - b. _____
9. List the three types of general stimulants according to the site of action.
 - a. _____
 - b. _____
 - c. _____
10. In therapeutic doses, medullary stimulants are used to stimulate the _____ center. Overdose may produce _____.
11. Aromatic Spirit of Ammonia is a _____ stimulant. It indirectly stimulates the respiratory center by _____ of the nose and throat.

12. Strychnine is a _____ stimulant. In large doses it causes _____.
13. Cerebral stimulants such as Amphetamine are used primarily to _____ the mood and _____ mental alertness.
14. In large doses, cerebral stimulants also stimulate the _____.
15. Match the generic names in column A with the classifications in column B.

Column A

_____ Nikethamide
 _____ Aromatic Spirit of Ammonia
 _____ Pentylenetetrazol
 _____ Strychnine
 _____ Doxapram
 _____ Caffeine
 _____ Amphetamine

Column B

a. Reflex stimulant
 b. Medullary stimulant
 c. Cerebral stimulant
 d. Spinal cord stimulant

16. Therapeutic doses of the medullary stimulants act specifically on the _____ center within the medulla.
17. Hypnotics are agents which induce _____.
18. Agents which produce a calming or quieting effect without sleep are called _____.
19. Barbiturates are classified according to their _____ and _____ of action.
20. List the four classes of barbiturates.
- _____
 - _____
 - _____
 - _____
21. In barbiturate poisoning, death is due to _____.
22. List two barbiturates which are of particular value in controlling epileptic seizures.
- _____
 - _____
23. Long action barbiturates are used mainly as _____ while short acting barbiturates are used mainly for their _____ effects.
24. Ultra short acting barbiturates are used primarily for _____.

25. Short acting barbiturates are used to produce sedation when given prior to _____.

26. Match the barbiturates in column A with the durations of action in column B.

Column A

_____ Pentobarbital

_____ Phenobarbital

_____ Amobarbital

_____ Thiopental

_____ Mephobarbital

_____ Secobarbital

Column B

a. Long acting

b. Ultra short acting

c. Intermediate acting

d. Short acting

27. List the generic name of three nonbarbiturate sedative-hypnotic drugs.

a. _____

b. _____

c. _____

28. _____ are the most widely used sedative-hypnotic drugs.

Antiepileptic Drugs

29. Epilepsy is characterized by muscular _____, or by transient episodes of sensory or psychic _____, and loss of _____.

30. List the two types of epilepsy which are described in this lesson.

a. _____

b. _____

31. List the characteristics of Grand Mal and Petit Mal epilepsy.

Grand Mal

a. _____

b. _____

c. _____

d. _____

e. _____

f. _____

Petit Mal

a. _____

b. _____

c. _____

32. Match the drugs in column A with the type of epilepsy they control in column B.

Column A

Column B

_____ Diphenylhydantoin

a. Grand Mal

_____ Phenobarbital

b. Petit Mal

_____ Trimethadione

_____ Mephobarbital

_____ Ethosuximide

33. State the mode of action for drugs used in the treatment of Grand Mal epilepsy.

a. _____

b. _____

General Anesthetics

34. List the three primary actions produced by general anesthetics.

a. _____

b. _____

c. _____

35. List the four stages of general anesthesia.

a. _____

c. _____

b. _____

d. _____

36. List the two major classes of general anesthetics and their routes of Administration.

a. _____

b. _____

37. _____ anesthetics are not metabolized and are readily removed from the body.

38. _____ anesthetics are metabolized and are not readily removed from the body.

39. Intravenous anesthetics have a _____ analgesic effect. They are used only for _____ procedures.

40. Basal anesthetics are used to produce _____ prior to administration of general anesthetics.

41. Preanesthetic medication prepares the patient for surgery in the following ways.

- a. _____
- b. _____
- c. _____
- d. _____

42. A major drawback of Ether and several other inhalation anesthetics is the danger of _____.

43. Cyclopropane and Chloroform sensitize the heart to _____.

44. Match the general anesthetics in column A with the classes in column B.

Column A

_____ Thiopental

_____ Chloroform

_____ Halothane

_____ Cyclopropane

_____ Ether

Column B

a. Inhalation

b. Intravenous

45. When using nitrous oxide there is a danger of _____ due to its weak anesthetic potency.

Analgesics

46. Analgesics are drugs used to relieve _____ without causing the patient to lose _____.

47. Nonnarcotic analgesics _____ elevate body temperature. This is called an _____ effect.

48. Nonnarcotic analgesics are effective in pain arising from skeletal _____. They are ineffective in relieving _____ pain.

49. Analgesics such as Aspirin are also used to relieve symptoms of _____ fever.

50. Salicylates such as Aspirin and Sodium Salicylate have _____, and _____ effects.

51. Aniline derivatives such as Phenacetin produce a condition of the blood known as _____ after long periods of high dosage.

52. Pyrazolon derivatives such as Phenylbutazone and Oxyphenbutazone produce a condition of the blood known as _____ after long periods of high dosage.

53. Colchicine is used to relieve pain in acute _____.

54. Match the generic names in column A with the drug groupings in column B.

Column A

- _____ Acetaminophen
- _____ Phenylbutazone
- _____ Acetylsalicylic acid
- _____ Phenacetin
- _____ Indomethacin
- _____ Ethoheptazine
- _____ Oxyphenbutazone
- _____ Propoxyphene

Column B

- a. Salicylate
- b. Aniline derivation
- c. Pyrazole derivative
- d. Miscellaneous

Analgesics - Narcotic

55. Narcotic analgesics differ from nonnarcotic analgesics in that they produce _____, _____, and _____.

56. Morphine is effective in almost all types of pain, but is used primarily for _____ pain.

57. Match the sites of action in column A with the effects of Morphine in column B.

Column A

- _____ Respiratory center
- _____ Vomiting center
- _____ Cough center
- _____ Pupils
- _____ Gastrointestinal tract
- _____ Spinal cord
- _____ Pain center

Column B

- a. Stimulation
- b. Depression
- c. Constriction
- d. Constipation

58. Death in narcotic overdose is usually due to _____.

59. Codeine is most widely used in _____ preparations.

60. Naloxone and Levallorphan are narcotic _____.

61. Apomorphine is used for its _____ effects.

62. Meperidine and Methadone are _____ opiates.

63. _____ is used in the treatment of Heroin addiction.
64. Match the generic names in column A with the drug groupings in column B.

Column A

- _____ Camphorated opium tincture
- _____ Hydromorphone
- _____ Naloxone
- _____ Meperidine
- _____ Codeine
- _____ Levallorphan
- _____ Methadone
- _____ Apomorphine

Column B

- a. Opiate alkaloid
- b. Semi-synthetic opiate
- c. Synthetic or nonopiate narcotic

Psychotherapeutic Drugs

65. List the two types of mental illness.
- a. _____
- b. _____
66. _____ constitutes the lesser degree of mental illness, and are those for which commitment to a mental institution is usually not necessary.
67. List the two types of neuroses.
- a. _____
- b. _____
68. _____ constitutes a more severe degree of mental illness, and are those for which commitment to a mental institution is usually necessary.
69. An example of organic psychosis would be _____.
70. Two examples of functional psychosis are _____ and _____.
71. The two major types of drugs used to treat mental illness are _____ and _____.
72. Tranquilizers have a _____ effect without dulling the senses.
73. Major tranquilizers are used to treat _____; whereas, minor tranquilizers are used to treat _____.
74. Phenothiazine derivatives are used in the treatment of _____ and for relief of _____ and _____.

75. _____ is an example of a phenothiazine.
76. Chlordiazepoxide and Diazepam are _____ tranquilizers and are used to treat _____.
77. Antidepressants are used in mental illness when the predominant emotional symptom is _____.
78. List the two major classes of antidepressant drugs.

- a. _____
- b. _____

79. List the generic name of drugs used in the treatment of depression.

- a. _____
- b. _____
- c. _____
- d. _____
- e. _____

80. List the major and minor tranquilizers.

Minor Tranquilizers

- a. _____
- b. _____
- c. _____
- d. _____

Major Tranquilizers

Nonphenothiazine derivatives

- a. _____
- b. _____

Phenothiazine derivatives

- a. _____
- b. _____
- c. _____
- d. _____
- e. _____
- f. _____

DRUGS ACTING ON THE AUTONOMIC NERVOUS SYSTEM

QUESTIONS

1. The autonomic nervous system is a _____ system innervating visceral organs.
2. The sympathetic and parasympathetic divisions of the ANS innervate the following structures.
 - a. _____
 - b. _____
 - c. _____
3. Preparing the body for fight, flight and fright is the general function of the _____ division.
4. Acting to protect, conserve and store body energy is the general function of the _____ division.
5. The sympathetic division of the ANS has (long/short) preganglionic fibers.
6. The parasympathetic division of the ANS has (long/short) postganglionic fibers.
7. The neurohormone between the postganglionic fiber and effector organ in the sympathetic division of the ANS is _____.
8. The neurohormone between the postganglionic fiber and effector organ in the parasympathetic division of the ANS is _____.
9. State the effects of the sympathetic and parasympathetic system on the following body structures.

	<u>SYMPATHETIC</u>	<u>PARASYMPATHETIC</u>
a. Cardiac Muscle		
(1) Rate	_____	_____
(2) Force of Contraction	_____	_____
b. Smooth Muscle		
(1) Gastrointestinal System		
(a) Peristalsis	_____	_____
(b) Sphincters	_____	_____
(2) Blood Vessels		
(a) Gastrointestinal	_____	_____
(b) Skeletal Muscle	_____	_____
(c) Skin	_____	_____
(d) Coronary Vessel	_____	_____

	<u>SYMPATHETIC</u>	<u>PARASYMPATHETIC</u>
c. Gland Secretion		
(1) Sweat	_____	_____
(2) Gastrointestinal, Nasal and Lacrimal	_____	_____
d. Blood Pressure	_____	_____
e. Mental Activity	_____	_____
f. Metabolism	_____	_____
g. Pupils	_____	_____

10. Match the following.

- | | |
|---------------------------|-------------------------------|
| _____ Sympathomimetic | a. Cholinergic blocking agent |
| _____ Sympatholytic | b. Adrenergic agent |
| _____ Parasympathomimetic | c. Adrenergic blocking agent |
| _____ Parasympatholytic | d. Cholinergic agent |

11. List in the blanks below whether the following class of ANS drugs have a sympathetic effect or parasympathetic effect.

- a. Sympathomimetic _____
- b. Sympatholytic _____
- c. Parasympathomimetic _____
- d. Parasympatholytic _____

12. Sympathomimetic drugs produce _____ of sympathetic receptor sites.

13. Supply the uses for the following sympathomimetic drugs.

- a. Epinephrine (1) _____
- (2) _____
- (3) _____
- (4) _____
- b. Levarterenol _____
- c. Metaraminol _____
- d. Isoproterenol (1) _____
- (2) _____

- e. Ephedrine (1) _____
(2) _____
(3) _____
(4) _____
- f. Phenylephrine (1) _____
(2) _____
(3) _____
(4) _____
- g. Phenylpropanolamine (1) _____
(2) _____
- h. Dextroamphetamine (1) _____
(2) _____
(3) _____
- i. Diethylpropion _____
- j. Tetrahydrozoline (1) _____
(2) _____
- k. Pseudoephedrine _____
- l. Oxymetazoline _____
- m. Xylometazoline _____

14. The mechanism of action of sympatholytic drugs is _____ of sympathetic receptor sites.

15. Depression of one division of the ANS produces effects similar to _____ of the other.

16. Supply the uses of the following sympatholytic drugs.

- a. Tolazoline _____
- b. Phentolamine _____
- c. Propranolol (1) _____
(2) _____

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17. List the two mechanisms of action for parasympathomimetic drugs.

a. _____

b. _____

18. Supply the uses for the following parasympathomimetic drugs.

a. Bethanochol _____

b. Pilocarpine _____

c. Neostigmine _____

d. Echothiophate _____

e. Edrophonium (1) _____

(2) _____

19. The mechanism of action for the parasympatholytic drugs is _____ of parasympathetic receptor sites.

20. Supply the uses for the following parasympatholytic drugs.

a. Atropine (1) _____

(2) _____

(3) _____

b. Propantheline _____

c. Benztropine _____

d. Dicyclomine (1) _____

(2) _____

(3) _____

e. Trihexyphenidyl _____

f. Scopolamine (1) _____

(2) _____

g. Clindinium (1) _____

(2) _____

(3) _____

(4) _____

h. Glycopyrrolate _____

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21. Three toxic symptoms of atropine poisoning are
 - a. _____
 - b. _____
 - c. _____
22. Muscle relaxants are defined as drugs which _____
23. Centrally acting muscle relaxants block nerve impulses in the _____ and _____.
24. _____ blocking agent exert their action between the nerve and the muscle where acetylcholine is released.
25. The sequence of muscles to be affected by the neuromuscular blocking agents are as follows:
 - a. _____
 - b. _____
 - c. _____
26. The drug used as an antidote for an overdose of tubocurarine is _____.
27. The two chief uses of muscle relaxants are
 - a. _____
 - b. _____
28. Match drugs in column A with modes of action in column B.

Column A

- _____ Meprobamate
- _____ Mephensin
- _____ Tubocurarine
- _____ Carisoprodol
- _____ Succinylcholine
- _____ Chlorzoxazone
- _____ Orphenadrine
- _____ Methocarbamol

Column B

- a. Centrally acting muscle relaxants
- b. Neuromuscular blocking agents

DRUGS ACTING ON THE CIRCULATORY SYSTEM

QUESTIONS

1. The section of the heart wall which serves as a protective covering is the _____.
2. The _____ is the heart muscle itself.
3. The _____ is the interior of the myocardial walls, and forms the chambers of the heart.
4. Deoxygenated blood is received into the heart in the _____ and passes to the _____ which pumps deoxygenated blood to the lungs for oxygenation. The oxygenated blood is received in the _____ and passes to the _____ where the blood is pumped throughout the body.
5. Define heart failure.

6. An _____ is any disturbance in the normal rhythm of the heart.
7. A heart rate of approximately 50 beats per minute would be described by the term _____.
8. A heart rate of approximately 90 beats per minute would be described by the term _____.
9. Match the following.

_____ Auricular flutter	a. 180-400 uniform beats per minute
_____ Congestive heart failure	b. Prophylactic coronary dilator
_____ Embolism	c. Transfer of an intravascular mass from point of origin
_____ Coronary Thrombosis	d. Cinchonism
_____ Symptom of Digitalis overdose	e. Formation of a clot in coronary artery, resulting in obstruction of that vessel.
_____ Symptom of Quinidine overdose	f. Administered sublingually for angina pectoris
_____ Arteriosclerosis	g. Greenish-yellow vision
_____ Procainamide (Pronestyl)	h. Loss of elasticity of arteries
_____ Nitroglycerin	i. Heart can't pump all blood out that is supplied to it.
_____ Pentaerythritol Tetranitrate (Peritrate)	j. Treatment of ventricular arrhythmias

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10. Hypertension can be divided into two classes called _____
and _____.

11. Match the following to their class of hypertension.

- | | |
|-------------------------------------|--------------|
| _____ About 90 percent of the cases | a. Primary |
| _____ Cause unknown | b. Secondary |
| _____ About 10 percent of all cases | |
| _____ Cause well defined | |

12. Guanethidine Sulfate is a _____ agent.

13. Name four functions of the blood.

- a. _____
- b. _____
- c. _____
- d. _____

14. Anemia is a lack of _____ or red blood cells in the blood.

15. Match the following anemias with causes.

- | | |
|-------------------------|--|
| _____ Aplastic anemia | a. Inadequate quantities of iron for the formation of hemoglobin |
| _____ Pernicious anemia | b. Bone marrow ceases to function as producer of blood cells |
| _____ Hemolytic anemia | c. Results from excessive loss of blood |
| _____ Iron deficiency | d. Young red blood cells do not mature due to vitamin B ₁₂ deficiency |

16. _____ increase the hemoglobin content of the blood.

17. List two coagulant drugs and their route of administration.

- | <u>Drug</u> | <u>Route of Administration</u> |
|-------------|--------------------------------|
| a. _____ | _____ |
| b. _____ | _____ |

ENDOCRINE SYSTEM

QUESTIONS

1. Adrenocorticoids are secreted by the _____ of the adrenal glands.
2. The two classes of adrenal hormones are the _____ and _____.
3. Glucocorticoids have a pronounced effect on metabolism of _____, _____, and _____.
4. The output of glucocorticoids is greatly _____ during periods of physical stress.
5. Mineralocorticoids primarily affect the _____ and _____ balance of the body.
6. Mineralocorticoids cause sodium retention and _____ loss.
7. When the adrenal cortex has an abnormally low secretion, characteristic symptoms known as _____ disease develop.
8. General steroid therapy (a term to indicate glucocorticoids) is used to treat
 - a. _____
 - b. _____
 - c. _____
 - d. _____
 - e. _____
9. Insulin is secreted by the beta cells of the _____.
10. Insulin regulates _____ metabolism.
11. Diabetes mellitus is a condition marked by the _____ of insulin.
12. Unlike NPH and PZI Insulin, Lente Insulin contains no foreign _____.
13. Match the following insulins with their onset and durations of actions.

_____ Onset 4 to 6 hours, duration 24-48 hours	a. Regular Insulin
_____ Onset 2 hours, duration 16-24 hours	b. Lente Insulin
_____ Onset 1 hour, duration 6-8 hours	c. NPH Insulin
_____ Onset and duration same as NPH Insulin	d. PZI Insulin
14. Thyroid deficiency from birth produces dwarfism, called _____.
15. Myxedema results from severe _____ in adults.
16. _____ parathyroidism results in hypocalcemia, while _____ parathyroidism cause decalcification, a condition marked by soft and fragile bones.

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17. The antithyroid agents inhibit the synthesis of the _____ hormone.
18. The parathyroid hormone regulates the exchange of _____ between the bones and body fluids.
19. The primary female sex hormone is _____.
20. Progesterone suppresses _____.
21. The male sex hormones are called _____.
22. (Testosterone/Methyltestosterone) is the drug of choice for oral administration.
23. Match the following.

- | | |
|---|-------------------------------------|
| _____ Testosterone | a. Androgen |
| _____ Dimethisterone with Ethinyl Estradiol | b. Estrogen |
| _____ Norethandrolone | c. Progesterone |
| _____ Norethindrone | d. Sequential oral contraceptive |
| _____ Estrogenic substances conjugated | e. Nonsequential oral contraceptive |
| _____ Norethindrone with Mestranol | |
| _____ Ethynodiol Diacetate with Mestranol | |
| _____ Estradiol | |
| _____ Norethynodrel with Mestranol | |

MISCELLANEOUS DRUGS

QUESTIONS

1. _____ are organic dietary substances necessary for the maintenance of normal body function.

2. Since most vitamins are not synthesized in the body, they must be supplied from _____.

3. Match the following vitamins with the appropriate classification.

- | | |
|-------------------------------|------------------|
| _____ Vitamin B ₁₂ | a. Water soluble |
| _____ Vitamin B ₆ | b. Fat soluble |
| _____ Vitamin B ₁ | |
| _____ Vitamin E | |
| _____ Vitamin K | |
| _____ Vitamin C | |
| _____ Niacin | |
| _____ Vitamin B ₂ | |
| _____ Vitamin A | |
| _____ Vitamin D | |

4. Nonavitamin tablets, Heptavitamin tablets, and Octavitamin drops are all used to treat _____.

5. Match the following vitamins with the appropriate deficiency disease.

- | | |
|-------------------------------|------------------------------|
| _____ Vitamin B ₁ | a. Megaloblastic anemia |
| _____ Vitamin B ₂ | b. Blood coagulation |
| _____ Vitamin B ₆ | c. Pellagra and vasodilation |
| _____ Vitamin B ₁₂ | d. Beri Beri |
| _____ Vitamin K | e. Pernicious anemia |
| _____ Niacin | f. Rickets |
| _____ Folic acid | g. Nightblindness |
| _____ Vitamin A | h. Nausea |
| _____ Vitamin D | i. Pellagra |

6. A condition that exempts the body from contracting a contagious disease or which enables it to resist infection is known as _____.

7. That immunity which normally exists in a human (species-race) is _____ immunity.
8. A specific immunity that does not occur naturally, but is induced actively or passively into the body is known as _____ immunity.
9. Two types of acquired immunity are _____ and _____.
10. _____ acquired immunity is when the body itself produces the antibodies.
11. _____ acquired immunity is a temporary immunity provided by antibodies not produced by the organism's own body cells.
12. An antibody that neutralizes a toxin is called an _____.
13. A substance that causes the formation and appearance of specific antibodies in the bloodstream is an _____.
14. Antigens produce _____ immunity.
15. Toxoids are usually modified with _____ to reduce its toxicity but not its antigenicity.
16. Match the type of drug with its definition
- | | |
|------------------------|---|
| _____ Vaccine | a. A specific antibody capable of neutralizing a specific toxin |
| _____ Toxin | b. Poisonous substances liberated by microorganisms |
| _____ Toxoid | c. Sterile solutions or suspensions of killed or attenuated live microorganisms |
| _____ Antitoxin | d. A detoxified toxin, chemically modified to be low in toxicity but high in antigenicity |
| _____ Antisera (serum) | e. Blood serum of an animal or human that contains antibodies against an infectious disease |
17. Match the class of drug with the type of immunity it provides.
- | | |
|-----------------|---------------------|
| _____ Vaccine | a. Diagnostic agent |
| _____ Toxin | b. Passive immunity |
| _____ Toxoid | c. Active immunity |
| _____ Antitoxin | |
| _____ Antisera | |

18. Match the tests with the appropriate disease.

- | | |
|---|------------------|
| _____ Schick test | a. Scarlet fever |
| _____ Dick test | b. Diphtheria |
| _____ Tuberculin, Purified Protein Derivative | c. Tuberculosis |
| _____ Tine test | |

19. A sensitizing protein or antigen when introduced into the body gives rise to the formation of _____.

20. A _____ in blood pressure is an effect of histamine poisoning.

21. Histamine poisoning causes _____ of the stomach, intestine, and uterus, therefore _____ may occur.

22. Because of its marked stimulation of gastric juice secretion, histamine is used to differentiate between true or false _____.

23. Direct antagonism of histamine is the mode of action of the _____ drugs.

24. _____ is the most common side effect of the antihistamines.

25. Two antihistamines widely used in the prevention and treatment of motion sickness are _____ and _____.

26. Drugs that increase the rate of flow of urine are called _____.

27. Diuretics are used to _____, _____, and _____.

28. The four classes of diuretic drugs are

- _____
- _____
- _____
- _____

29. Hydrochlorothiazide is approximately _____ times more potent than Chlorothiazide.

30. Match the following drugs with their appropriate classification.

- | | |
|-------------------------|---------------------------------|
| _____ Meralluride | a. Osmotic |
| _____ Mannitol | b. Acid forming salt |
| _____ Spironolactone | c. Mercurial |
| _____ Ammonium chloride | d. Carbonic anhydrase inhibitor |
| _____ Acetazolamide | e. Aldosterone antagonist |

31. A drug given to increase bronchial secretions and to facilitate the expulsion of sputum is classified as an _____.
32. Drugs that act to soothe acute inflammation by aiding the secretion of protective mucous are classified as _____ expectorants.
33. Drugs that act to stimulate repair in chronic inflammatory processes of the mucous membranes of the respiratory tract are called _____ expectorants.
34. Drugs that act through central depression of the cough center are classified as _____.
35. Agents used to contract the uterus and speed delivery are referred to as _____.
36. Prolonged therapy with Ergonovine Maleate may lead to _____.
37. Methylergonovine Maleate has the same action as Ergonovine Maleate but its _____ and _____ are greater.
38. Oxytocin acts directly on the smooth muscle to produce rhythmic contractions. Its action is more _____ and of _____ duration than that of Ergonovine.
39. _____ can be employed in the control of postpartum bleeding.
40. Organic catalysts produced by living organisms are called _____.
41. Enzymes are named using three methods.
 - a. By adding _____ to the root of the substance being acted upon
 - b. By adding _____ to the root of the reaction taking place.
 - c. _____; that is calling it a name without regard to anything else.
42. The enzyme that acts against Hyaluronic acid to increase the area of distribution of liquids in tissue spaces and facilitates absorption is _____.
43. An enzyme preparation used as an emergency drug to inactivate penicillin is _____.
44. Two nitrogen mustard preparations used as antineoplastics are _____ and _____.
45. The antineoplastic agents are sometimes referred to as _____ agents.
46. Prostate cancer is sometimes treated with _____ and breast cancer with _____.

GLOSSARY

PREFIXES AND SUFFIXES USED IN MEDICAL TERMINOLOGY

The following prefixes and suffixes should be studied carefully. They will be used throughout your career in the pharmacy in addition to being used in this workbook.

<u>Prefix or Suffix</u>	<u>Definition</u>	<u>Example</u>
a -	negative prefix (before consonants)	asepsis, asexual
an-	negative prefix (before vowels)	anaerobic
ad-	near or to	adrenal
alb-	white	albino
andro-	man	androgen
ante-	before	ante mortem, ante cibos
anti-	against (antacid - i dropped before vowels)	anticoagulant
brady-	slow	bradycardia
carcinoma	malignant tumor	
cardio	heart	electrocardiogram
cephal	head	encephalitis
cer, cera	wax	
-cid, cide	kill	bacteriacide
-cise	cut	incise, excise
contra-	against, counter	contraindication
corti-, cortex	bark, rind	cerebral cortex
cost	rib	intercostal
cuti	skin	subcutaneous
cyst	bladder, fluid filled sac	cystitis
cyt	cell	erythrocyte
dermo-, dermat	skin	dermis, dermatology
dextr	on the right	oculus dexter
dys-	bad, pain, improper	dysentery
ect-, ecto-	outside	ectoplasm, ectoparasite
ede-	swelling	edema

<u>Prefix or Suffix</u>	<u>Definition</u>	<u>Example</u>
enter-	intestine	dysentery
erythro-	red	erythrocyte
extra-	outside	extracellular
-facient	make	rubifacient
febr-	fever	febrile
flav-	yellow	riboflavin
gastr-	stomach	gastritis
gloss	tongue	
glyc-, gluco-	sweet	hyperglycemia
gyn-	woman	gynecology
hemo-	blood	hemoglobin, hematology
hepat	liver	hepatitis
hist-	tissue	histology
hyper-	over, too much	hyperirritable, hypertension
hypo-	under, too little	hypotension, hypodermic
hystero-	uterus, womb	hysterectomy
-iasis	condition of	helminthiasis
inter-	between	intercostal
intra-	within	intravenous
-itis	inflammation	appendicitis
kerat	horn	keratolytic
lep-	seizing	epilepsy, narcolepsy
leuco-, leuko-	white	leukocyte
lingu	tongue	sublingual
lipo-	fat	lipid
-lysis	loosen, dissolve	hemolysis
mega-, macro	large	
micro-	small	microbe
mast	breast	mastitis

<u>Prefix or Suffix</u>	<u>Definition</u>	<u>Example</u>
men-	month	menstrual
myco-	funqus	mycosis
-myelo	marrow, spinal cord	osteomyelitis
myo-	muscle	myocardium
narco-	sleep	narcolepsy, narcotic
necro-	death	
neo-	new	neophyte, neonatal
nephro-	kidney	nephritis
neuro-	nerve	neuralgia, neurology
oculo-	eye	oculist, oculus dexter
-oma	tumor or swelling	hematoma
-osis	condition of	mycosis
-osteo	bone	osteology
oto	ear	otologist, otoscope, otic
path	sickness, disease	psychopathic
ped (Greek)	child	pediatrics
-penia	too few	leukopenia
-phagia	eat, swallow, speak	dysphagia
phleb-	vein	phlebitis
plasty	operative revision	rhinoplasty
post	after	post mortem postoperative
pyr-	heat	pyrogen
pseud-	false	pseudoreaction
renal	kidney	adrenal
-rrhea	flow	diarrhea
rhino	nose	rhinitis
-sclera	hard	arteriosclerosis
sep	rot, decay	sepsis
-sinister, sinistra	on the left side	oculus sinister
-stasis	stoppage	bacteriostasis
stoma	mouth	stomatitis
tach	fast	tachycardia
thromb-	lump, clot	thrombosis

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HANDOUT

3ABR90530-11-1

Technical Training

10-6

Pharmacy Specialist

PHARMACOLOGY
(ANATOMICAL DRAWINGS)

April 1976



SCHOOL OF HEALTH CARE SCIENCES, USAF
Department of Biomedical Sciences
Sheppard Air Force Base, Texas 76311

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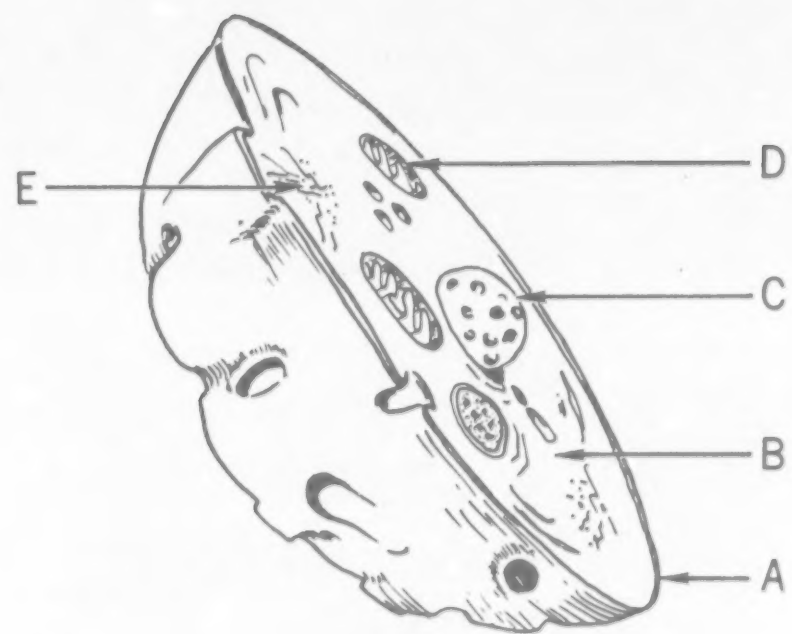
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School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas 76311

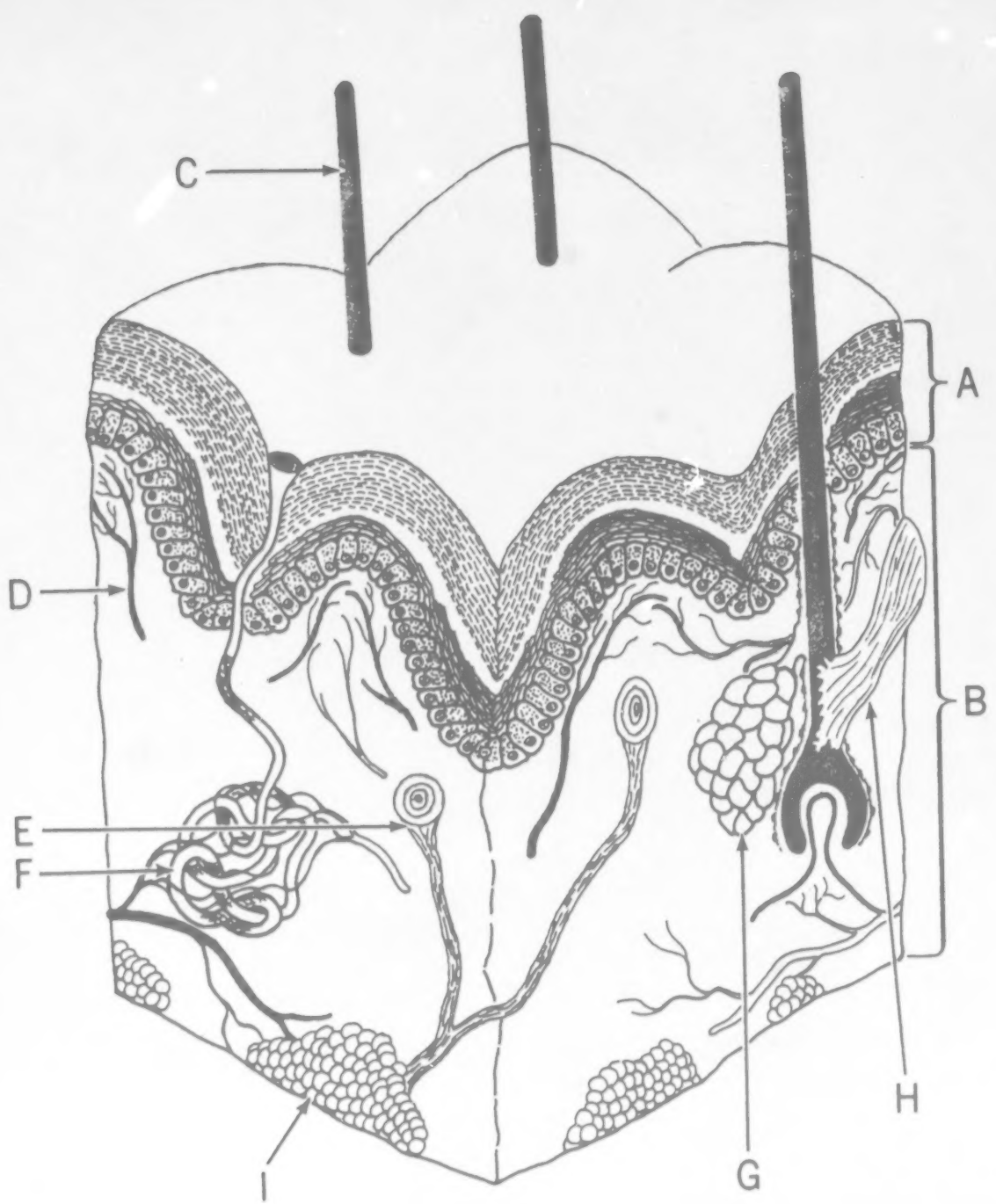
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April 1976

PHARMACOLOGY
(ANATOMICAL DRAWINGS)

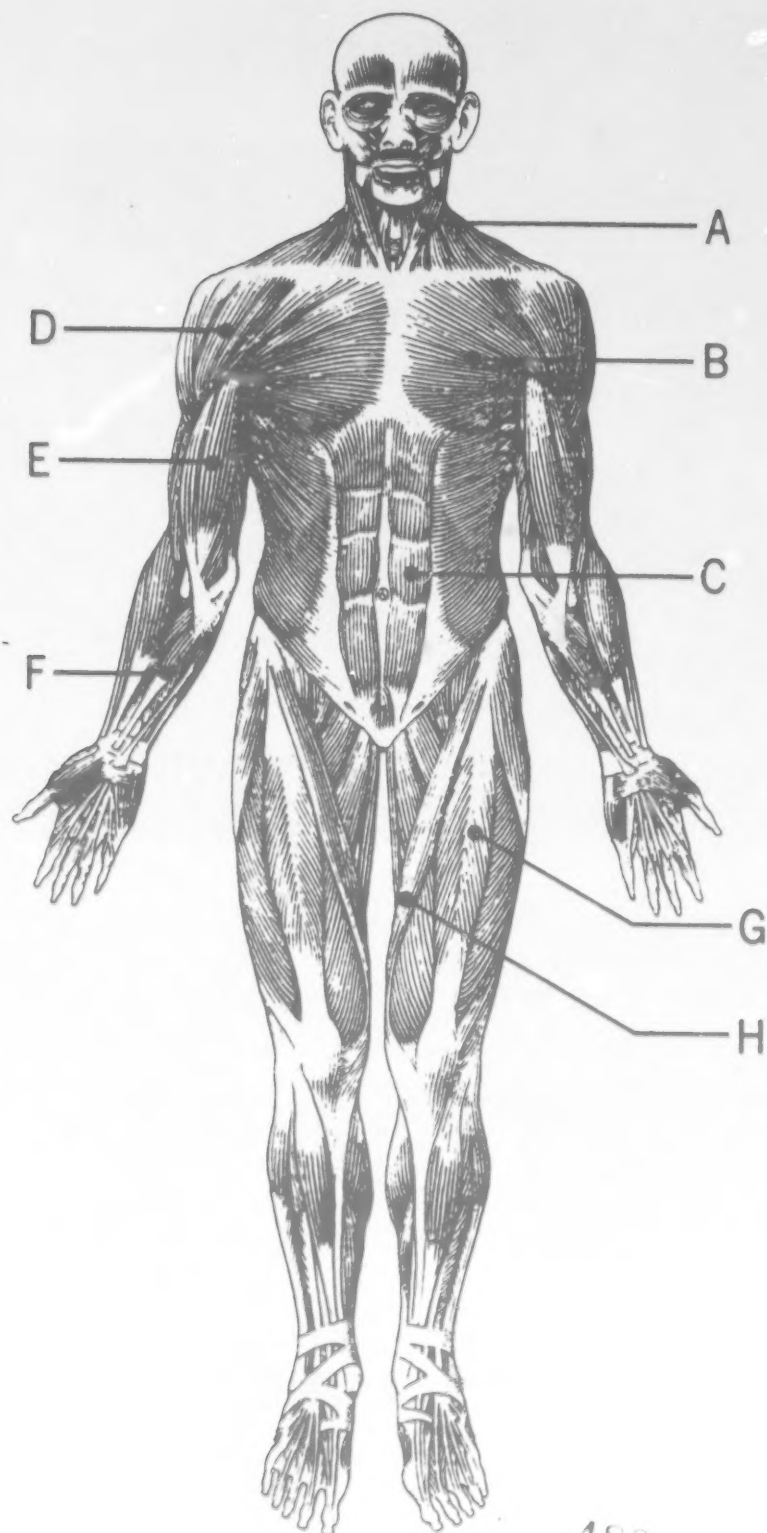


This supersedes HO 3ABR90530-II-1, July 1975

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26. Air is pushed out of the lungs by the compression of the
- a. sac.
 - b. alveoli.
 - c. lungs.
 - d. bronchi.

////////////////////////////////////

Correct response c

27. Match the function in Column B with its structure in Column A

<u>Column A</u>	<u>Column B</u>
a. lungs	<u> </u> 1. Relaxes
b. pleural sac	<u> </u> 2. Air pressure equal to outside of body
c. diaphragm	<u> </u> 3. Is pulled on to thin out confined air
	<u> </u> 4. Contracts
	<u> </u> 5. Compresses the lungs
	<u> </u> 6. Carbon dioxide starts out of the respiratory system

////////////////////////////////////

Correct response: 1. c, 2. a, 3. b, 4. c, 5. b, 6. a

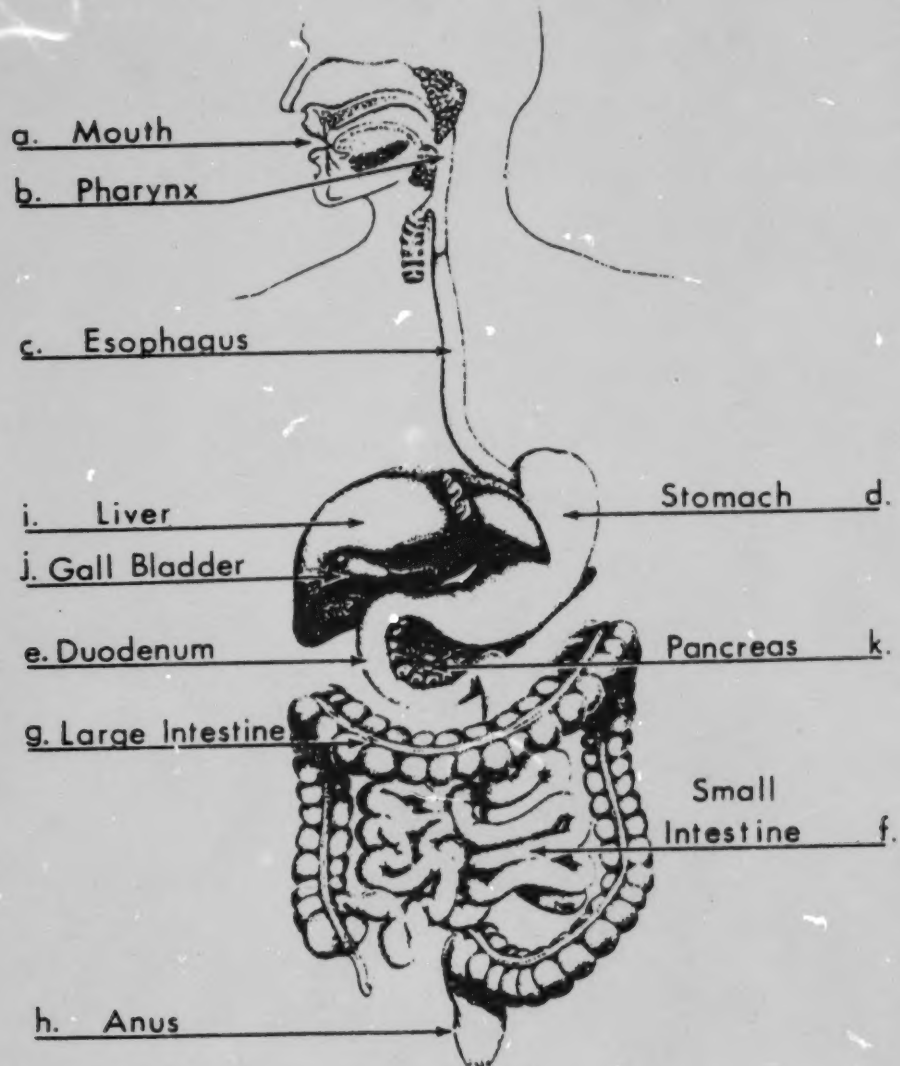
This concludes the chapter on the respiratory system. Now we know how we breathe, how oxygen is taken on by the body and carbon dioxide is given off. If you have any questions, feel free to ask your instructor.

Chapter Seven

DIGESTIVE SYSTEM

The digestive system is one of the most important systems of the human body. It is this system that supplies the body with the food and energy to maintain life. Just like the other systems of the body, the digestive system is made up of many parts. It is these parts and their functions with which we are concerned.

In the following diagram the main organs of digestion are listed. Remember them and their order in the system.



1. We will now look at the main organs of the digestive system. The mouth is the first structure of the alimentary canal (digestive tract). Mechanical digestion (or chewing) occurs here. A chemical action takes place in the mouth also, but that will be discussed later.

Complete the following statement. Food is received in the mouth and chewed. This is called _____.

////////////////////////////////////

Correct response: mechanical digestion

2. The mechanical digestion that occurs in the mouth is sometimes referred to as _____.

////////////////////////////////////

Correct response: chewing

3. The next two organs are passageways, the pharynx and the esophagus. Their function is to allow food to pass from the mouth to the stomach. When food enters the esophagus it produces a dilatation that stimulates contraction. This contraction of smooth muscle is referred to as peristalsis, a ripple like motion. Peristalsis starts at the esophagus and continues down through the entire alimentary canal.

The organs of the digestive system that are passageways are the

_____ and _____.

////////////////////////////////////

Correct response: pharynx and esophagus

4. From the diagram on page 7-1, select the letters that indicate the pharynx and the esophagus.

////////////////////////////////////

Correct response: b, c

5. The next organ in the digestive system is the stomach. This is the widest portion of the alimentary canal and is also the most muscular organ. A churning action in the stomach mixes the food with gastric juice. Only partial chemical digestion of proteins occurs in the stomach.

From the diagram on page 7-1, locate and write the letter that indicates the stomach. _____

////////////////////////////////////

Correct response: d

6. Write in your own words the function and actions that occur in the stomach.

////////////////////////////////////

Correct response: The function of the stomach is to receive food from the esophagus and start a churning action. Gastric juices are mixed with the food and begin chemical digestion of proteins.

7. Let us go back and review the organs that we have discussed up to now. The first structure is the mouth in which food is chewed. This is called mechanical digestion. The next two organs are the pharynx and the esophagus. They serve as passageways through which food may travel on its way to the stomach. When the food arrives at the stomach, a churning action occurs and gastric juices are added to the food mixture.

After food leaves the stomach it enters the small intestine. The small intestine is similar to a long tube, approximately 20 feet in length. The main function of the small intestine is the absorption of digested food. As we have said before, the digestion process starts in the mouth and continues to break down the food so that the body can use it.

Some digestion occurs in the first part of the small intestine. This part is referred to as the duodenum and represents the first ten inches of the small intestine. Two organs empty their digestive juices into the duodenum of the small intestine.

Refer to the diagram on page 7-1 and identify the small intestine. Write the letter in the space provided. _____

////////////////////////////////////

Correct response: f

8. Write in your own words why food is digested.

////////////////////////////////////

Correct response: Food is broken down because in its solid state it cannot be used by the body.

9. After the food leaves the small intestine, it enters the large intestine. This organ is five feet in length and is larger in diameter than the small intestine. When food enters the large intestine, it is in a semi-liquid form; as it travels through the large intestine the liquid is absorbed through the intestinal walls and is utilized to maintain proper body fluid balance. By the time the food reaches the distal end of the large intestine, all of the necessary materials have been removed. That which remains is of no use to the body. This portion of the large intestine is called the anus, and the material that leaves the body is called fecal matter.

In the diagram on page 7-1, select the letter that indicates the large intestine. _____

////////////////////////////////////

Correct response: g

10. The main function of the large intestine is

- a. absorption of waste material.
- b. absorption of liquid.
- c. absorption of digested food.
- d. reabsorption of digestive juices.

////////////////////////////////////

Correct response: b

11. The name of the part at the distal end of the large intestine is the _____.

////////////////////////////////////

Correct response: anus

12. We now have finished with the main organs of digestion. Up until now, we have only discussed the alimentary canal. There are other organs which aid digestion but they lie outside of the alimentary canal. The function of these organs is to break down the food either by chemical or mechanical digestion. These organs are called accessory organs of digestion.

The first of these accessory organs is the teeth, tongue, and salivary glands, all located in the mouth. When food is taken in, the teeth and tongue chew and grind up the food. This is mechanical digestion. Mechanical digestion continues on through the alimentary canal due to the squeezing action of peristalsis.

The salivary glands secrete an enzyme, which is a substance necessary for chemical digestion of food. The name of this enzyme is saliva and its function is to start the digestion of carbohydrates. Carbohydrates, such as bread and potatoes, are used by the body to supply heat and energy.

The accessory organs to digestion are mainly located _____

////////////////////////////////////

Correct response: outside of the alimentary canal.

13. The two types of digestion are _____ and _____.

////////////////////////////////////

Correct response: mechanical and chemical

14. Does mechanical digestion stop in the mouth? _____

////////////////////////////////////

Correct response: no

15. The teeth and tongue are used in _____ digestion.

////////////////////////////////////

Correct response: mechanical

16. The salivary glands produce an enzyme called _____.

////////////////////////////////////

Correct response: saliva

17. Saliva works on which of the following types of food?

- a. fats
- b. proteins
- c. minerals
- d. carbohydrates

////////////////////////////////////

Correct response: d

18. Carbohydrates are used by the body for

- a. supplying heat and energy.
- b. growth and repair of body tissue.
- c. regulating fluid balance.
- d. regulating solid balance.

////////////////////////////////////

Correct response: a

19. The next accessory organ is the liver. The liver is one of the largest and most important organs of the human body. It has many functions, but its most important function is the production of an emulsifier called bile. Up to now, we have talked about enzymes. What is an emulsifier? An emulsifier is a chemical substance which breaks large particles into many smaller particles but does not chemically change them. Bile breaks up fat to enable further digestion and storage. Bile is being

produced by the liver constantly, and must be stored until needed by the body. Bile is stored in the gall bladder, a small gland located just below the liver. Fats are used by the body the same way as carbohydrates as a supply of heat and energy.

The liver secretes a substance called _____.

////////////////////////////////////

Correct response: bile

20. Bile is not an enzyme, it is an _____.

////////////////////////////////////

Correct response: emulsifier

21. Bile is secreted by the _____ and is stored in the _____.

////////////////////////////////////

Correct response: liver, gall bladder

22. Fats are used by the body for

- a. growth and repair of body tissue.
- b. regulating fluid balance.
- c. regulating solid balance.
- d. supply of heat and energy.

////////////////////////////////////

Correct response: d

23. The pancreas is the last accessory organ and is located just below the stomach. It secretes an enzyme called pancreatic juice. The function of this enzyme is to digest fats, proteins and carbohydrates. Proteins are needed by the body for the repair and growth of body tissues. Both pancreatic juice and bile are emptied into the duodenum. The pancreas is known as a dual function gland and will be discussed more fully later when we talk about the endocrine system.

The pancreas empties its enzymes into what portion of the alimentary canal? _____

////////////////////////////////////

Correct response: duodenum

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24. Pancreatic juice works on fats, carbohydrates and _____.

////////////////////////////////////

Correct response: proteins

25. Proteins are used by the body for

- a. growth and repair of body tissues.
- b. regulating fluid balance.
- c. regulating solid balance.
- d. supply of heat and energy.

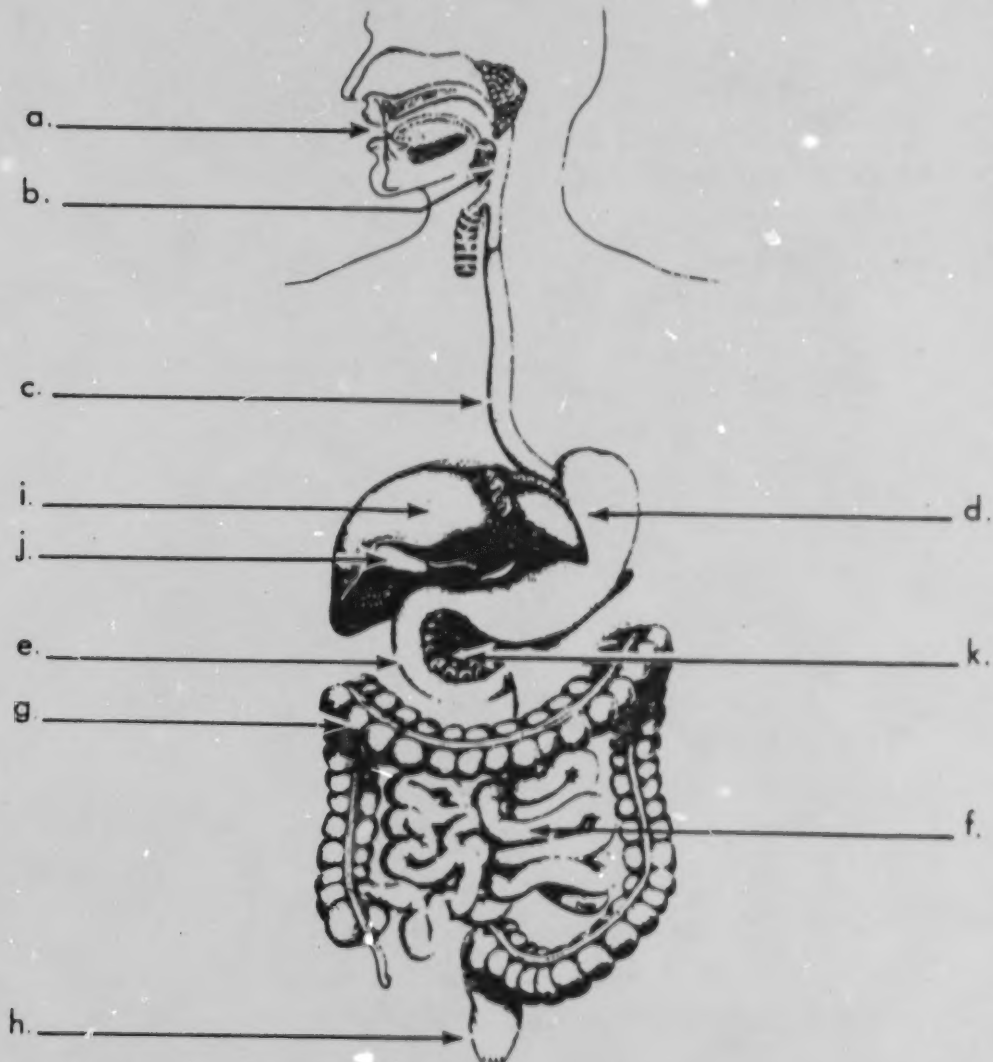
////////////////////////////////////

Correct response: a

26. You now have been given all of the information on the digestive system, both main organs and accessory organs. As a means of review, we now will trace a bit of food from start to finish.

Mouth	{	chews food, mechanical digestion.
	{	saliva, works on carbohydrates, chemical digestion.
Pharynx & Esophagus	{	passageway for food to the stomach.
	{	churning action.
Stomach	{	partial digestion of proteins.
	{	production of gastric juices.
Duodenum	{	the first ten inches of the small intestine.
	{	receives pancreatic juice and bile.
Liver	{	secretes an emulsifier to break down fats.
Gall bladder	{	stores bile for the liver.
Pancreas	{	secretes an enzyme to digest fats, carbohydrates and proteins.
Small intestine	{	absorption of digested food.
Large intestine	{	absorption of liquids.
	{	elimination of waste through the anus.

The diagram on this page is the same as the one on page 7-1. As a test to see if you know the main units of the digestive system, you should label all of the listed parts without referring to your notes. Also, on page 7-2 you will find a list of terms we have used in this chapter. Beside each organ, write its function. Try to do this without referring to your notes.



LIST OF TERMS

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Mouth

Saliva

Pharynx

Esophagus

Stomach

Duodenum

Liver

Gall bladder

Pancreas

Small Intestine

Large Intestine

Anus

DO NOT WRITE ON THIS PAGE

Chapter Eight

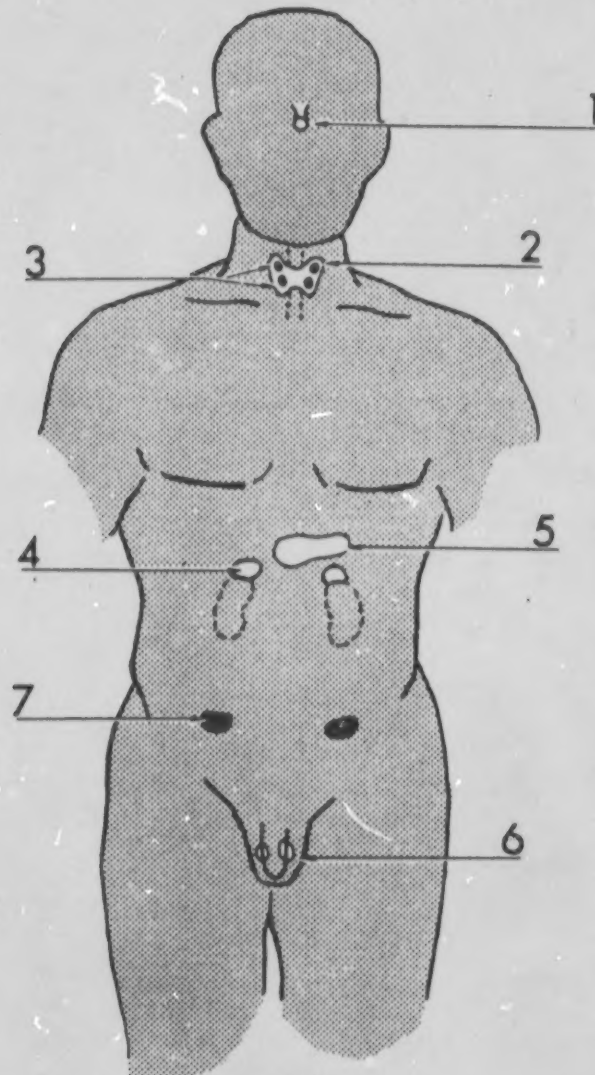
ENDOCRINE SYSTEM

The endocrine system is important because of the secretions produced by its glands. These secretions are responsible for the proper functioning of the body and many of its organs. An excess or deficiency of any one of these secretions can have serious effects on the body.

1. Let's take a look at the glands of the endocrine system, their location and function.

<u>GLAND</u>	<u>LOCATION</u>	<u>FUNCTION</u>
Pituitary	Base of the brain just behind the eyes.	Produces hormones which regulate function of thyroid, adrenals, and gonads; influences growth.
Thyroid	Anterolateral neck on both sides located just below the larynx.	Regulates rate of body metabolism (rate at which the body uses oxygen to burn food taken in).
Parathyroids	Imbedded in posterior surface of thyroid.	Regulates calcium level of blood.
Adrenals Cortex Medulla	On top of each kidney. Outermost part. Innermost part.	Cortical hormone regulates salt and water balance for the body. Adrenalin stimulates cardiac rate and influences blood pressure. Prepares the body for flight or fight.
Pancreas	Posterior to stomach.	Produces insulin which regulates sugar metabolism.
Gonads:		
Testes (Male)	Suspended from the body in the scrotum.	Hormone influences secondary sexual characteristics of the male; i.e., beard, deep voice, coarse skin, hair, etc.
Ovaries (Female)	In the pelvis.	One hormone influences secondary sexual characteristics of the female; i.e., lack of beard, high pitched voice, development of breasts, fat distribution. One hormone prepares and maintains the uterus for pregnancy.

2. Write in the name of each gland described on page 8-1.



////////////////////////////////////
Correct response: 1. Pituitary, 2. Thyroid, 3. Parathyroid, 4. Adrenal,
5. Pancreas, 6. Testes, 7. Ovaries

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Chapter Nine

URINARY SYSTEM

1. The body is constantly metabolizing food, when this occurs there is always some waste produced. The body is unable to use this material so it has to rid itself of it. We have already mentioned that solid waste is removed through the anus, but there is more waste than just the solid waste. The body also filters the blood and removes liquid waste products. This waste is filtered from the blood by the urinary system for elimination from the body.

The first organs to be discussed are the kidneys; they are the main organs of the urinary system. The kidneys are two bean shaped organs located superior to the waistline and posterior to the digestive organs. Blood is filtered by these organs and the liquid removed is called urine. Urine is considered a waste product.

Urine (contains/does not contain) waste products.

////////////////////////////////////

Correct response: contains

2. The kidneys are located (superior/inferior) to the waistline.

////////////////////////////////////

Correct response: superior

3. The human body contains how many kidneys? Answer: _____

////////////////////////////////////

Correct response: 2

4. Using the diagram on page 9-2, identify the kidneys by letter. _____

////////////////////////////////////

Correct response: a

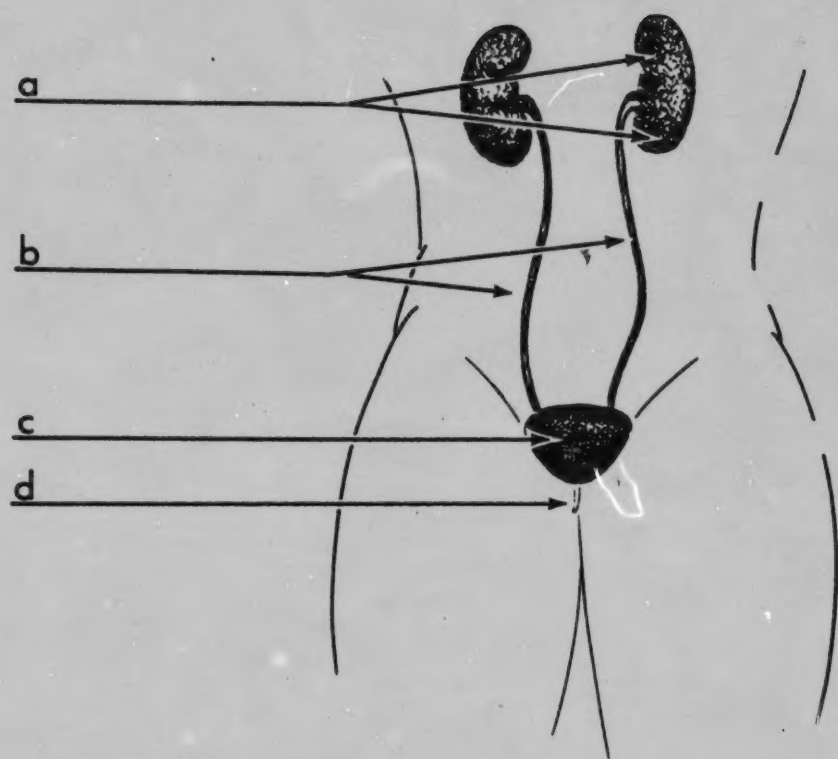
5. After urine is secreted by the kidneys, it is transported to the bladder by two small tubes called ureters, one ureter leading from each kidney. Urine moves through these tubes by way of peristalsis, just as food moves through the digestive system.

Urine is moved through the ureters by a wave like motion called _____

////////////////////////////////////

Correct response: peristalsis

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6. The human body has how many kidneys?

////////////////////////////////////

Correct response: 2

7. The next two organs to be considered are the bladder and the urethra. The bladder is a hollow muscular organ. Its function is to store the urine which has been filtered by the kidneys until it is time to be released from the body. This organ is capable of expanding to hold the urine. When the bladder is full a signal is sent to the brain telling the person that the bladder needs to be emptied. The urine then flows through a small tube called the urethra. The male urethra is approximately 6-8 inches long, whereas the female urethra is approximately 1-1/2 to two inches in length. It is through this tube that urine leaves the body.

The tubes that lead from the kidneys to the bladder are called _____.

////////////////////////////////////

Correct response: ureters

7. The function of the ureters is to

////////////////////////////////////

Correct response: Transport the urine from the kidneys to the bladder.

8. Using the diagram on page 9-2, locate by letter the ureters.

////////////////////////////////////

Correct response: b

9. What is the shape of the bladder?

////////////////////////////////////

Correct response: hollow and muscular

10. Using the diagram on page 9-2, locate by letter the bladder.

////////////////////////////////////

Correct response: c

11. The tube leading outwardly from the bladder is called _____.

////////////////////////////////////

Correct response: urethra

12. The function of the urethra is to

////////////////////////////////////

Correct response: transport urine from the bladder to the outside of body.

13. The size of the urethra in the male is _____ inches long,
and in the female it is _____ inches long.

////////////////////////////////////

Correct response: 6-8 inches in male. 1-1/2 to 2 inches in female.

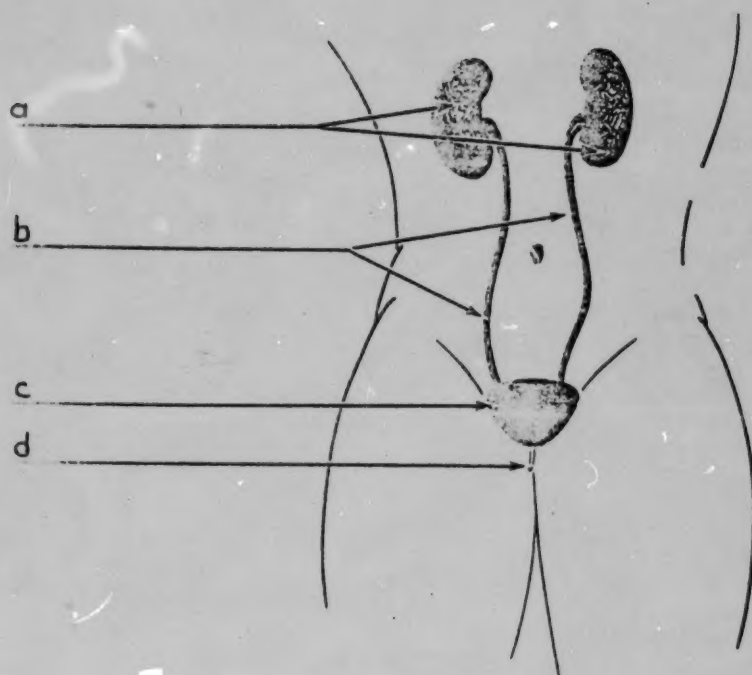
14. Using the diagram on page 9-2, locate by letter the urethra.

////////////////////////////////////

Correct response: d

15. On page 9-5, you will find a diagram of the urinary system identical to the one on page 9-2. Identify the parts in order and write their names below.

- a.
- b.
- c.
- d.



////////////////////////////////////

response: a. kidneys, b. ureters, c. bladder, d. urethra

Chapter Ten

REPRODUCTIVE SYSTEM

This chapter is divided into two sections. The first is on the male system, and the second is on the female system.

REPRODUCTIVE SYSTEM

The testes are the sperm producing organs of the male reproductive system. These organs are located in a pouch called the scrotum. The sperm they produce fertilize the egg in the female.

Label the testes and scrotum on the diagram on page 10-2.

Test response: testes a, scrotum b

What do the testes produce?

Test response: sperm

Where are the testes located?

Test response: scrotum

The egg is fertilized by the

- a. scrotum.
- b. testes.
- c. sperm.
- d. offspring.

Test response: c

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5. Sperm leaves the testes and enters the epididymis, a multicoiled tube that lies on top of the testes. They are stored in this tube until they mature.

Label the epididymis on the diagram on page 10-2.

////////////////////////////////////

Correct response: c

6. Name the structure on top of the testes.

////////////////////////////////////

Correct response: epididymis

7. The function of the epididymis is to

- a. produce sperm.
- b. fertilize the egg.
- c. store sperm.
- d. support the testes.

////////////////////////////////////

Correct response: c

8. The mature sperm leaves the epididymis and enters a passageway known as the vas deferens. The sperm is moved through this passageway by muscular action until it reaches the urethra and is ejected.

On the diagram on page 10-2, label the vas deferens.

////////////////////////////////////

Correct response: d

9. What is the name of the structure that moves the sperm from the epididymis?

////////////////////////////////////

Correct response: vas deferens

10. The vas deferens is a passageway to the

- a. bladder.
- b. urethra.
- c. scrotum.
- d. testes.

////////////////////////////////////

correct response: b

11. How is the sperm moved through the vas deferens?

////////////////////////////////////

Correct response: by muscular action

12. Enroute to the urethra, the sperm is mixed with seminal fluid, a fluid secreted by the seminal vesicles. This fluid provides nutrients and protection for the sperm. At this point the sperm and fluid enters the ejaculatory duct which connects with the urethra. The prostate gland also aids in the protection of the sperm. It is located immediately below the bladder and surrounds the urethra. It secretes an alkaline fluid that neutralizes the acid content in the female vagina, thus protecting the sperm. It also aids in neutralizing the acid in the male urethra.

Label the seminal vesicles and ejaculatory duct on the diagram on page 10-2.

////////////////////////////////////

Correct response: e, f

13. The seminal vesicles add what type of fluid to the sperm?

////////////////////////////////////

Correct response: seminal fluid

14. Check the correct function(s) below that pertain(s) to the seminal fluid.

- a. Protects the sperm.
- b. Provides nutrients for the testes.
- c. Protects the vas deferens.
- d. Provides nutrients for the sperm.

////////////////////////////////////

Correct response: a, d

15. Label the prostate gland on the diagram on page 10-2.

////////////////////////////////////

Correct response: g

16. The main function of the prostate gland is to

////////////////////////////////////

Correct response: neutralize the acid in the female vagina.

17. The alkaline fluid also neutralizes

////////////////////////////////////

Correct response: the urine in the male urethra.

18. The urethra, as studied in the urinary system, is the duct through which urine passes to reach the external surface of the body. In the male, the urethra serves an additional purpose, the passage of sperm. The urethra is located within the male organ of copulation, the penis.

Label the urethra and penis on the diagram on page 10-2.

////////////////////////////////////

Correct response: urethra h, penis i

19. In the male reproductive system the urethra is used to carry

- a. acid.
- b. sperm.
- c. urine.
- d. neutralizing fluids.

////////////////////////////////////

Correct response: b, c, d.

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FEMALE REPRODUCTIVE SYSTEM

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20. The ovaries in the female reproductive system are responsible for producing ova or ovum. Ova are eggs which must be fertilized by the male sperm in order to produce offspring. The ovaries are located in the lower pelvic cavity of the female. Near, but not attached to the ovaries, are the fallopian tubes which convey the ova to the uterus of the female. All fertilization takes place in the fallopian tubes.

Label the ovaries and fallopian tubes on the diagram on page 10-6.

////////////////////////////////////

Correct response: ovaries a, fallopian tubes b

21. The ovary produces the

////////////////////////////////////

Correct response: ovum or egg.

22. Where will the ovum be fertilized?

////////////////////////////////////

Correct response: The ovum will be fertilized in the fallopian tube.

23. The ovum is transported away from the ovary by the

////////////////////////////////////

Correct response: fallopian tube.

24. When a fertilized ovum reaches the uterus, a pear shaped organ, it attaches itself to the wall of the uterus to receive nourishment. The uterus also protects the fertilized ovum during its development. The vagina is the female organ of copulation and also serves as the birth canal.

Label the uterus and vagina on the diagram on page 10-6.

////////////////////////////////////

Correct response: uterus c, vagina d

25. From the list below, check the function(s) of the uterus.

- a. protects the ovum.
- b. neutralizes the ovum.
- c. nourishment of the ovum.
- d. passageway of the ovum.

////////////////////////////////////

Correct response: a, c

26. The vagina also serves as the

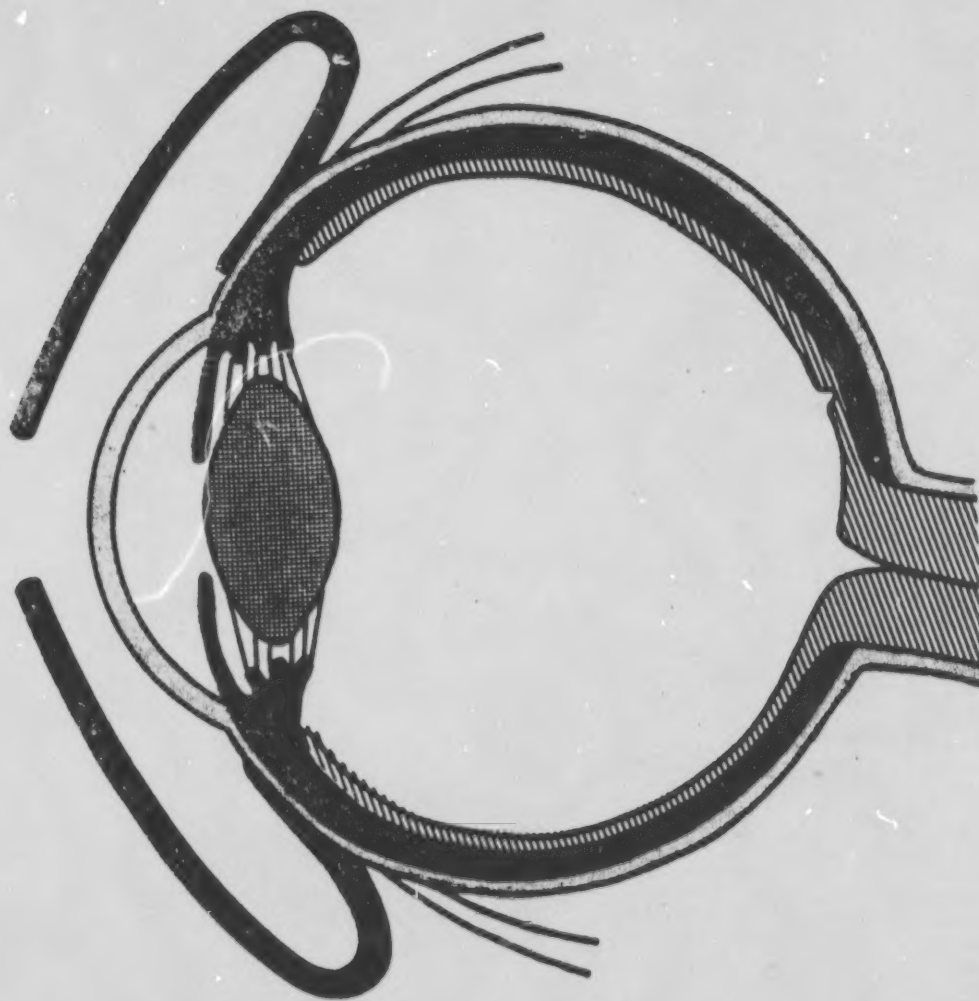
////////////////////////////////////

Correct response: birth canal.

This concludes the chapter on the reproductive system. If you should have any questions about the structures or functions of the male or female system, feel free to ask your instructor.

Chapter Eleven

THE EYE AND EAR



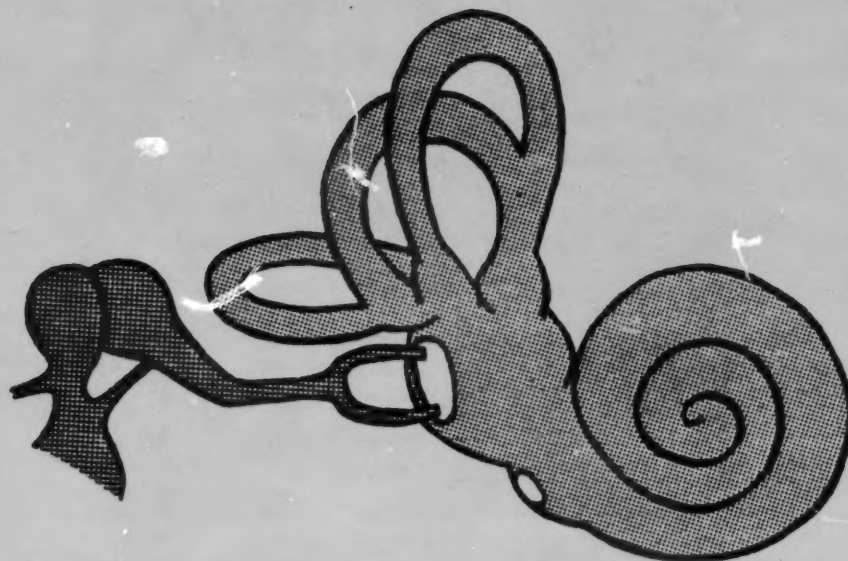
11-1

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APPROVAL OFFICE AND DATE MSDB <i>Wilson</i> 1 OCT 1974		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE Pharmaceutical Calculations II			
CLASSROOM/Laboratory 16 hrs/0 hrs		LESSON DURATION ~~Laboratory~~ Complementary 6 hrs	TOTAL 22 hrs
POI REFERENCE			
PAGE NUMBER 12 13	PAGE DATE 18 July 75	PARAGRAPH 1a	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>R. E. Vegeau</i>	1 OCT 7	<i>R. E. Vegeau</i>	6 OCT 1975
<i>R. E. Vegeau</i>	31 Mar 5		
<i>R. E. Vegeau</i>	12 SEPT 5		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
NA	NA	NA	NA 3ABR90530-III-1 Pharmaceutical Prep- arations Transparency Set #1 Pharmaceutical Cal- culations II
CRITERION OBJECTIVES AND TEACHING STEPS			
1a. Solve problems in reducing and enlarging formulas, specific gravity, percentage preparations, concentration and dilution, alligation, and temperature conversion. (Teaching steps listed in Part II)			

APPROVAL OFFICE AND DATE MSDB <i>Walton 17 Oct 74</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE Pharmaceutical Calculations II Laboratory			
LESSON DURATION			
CLASSROOM/Laboratory 0/3	~~XXXXXX~~ Complementary 2 hrs	TOTAL 5 hrs	
POI REFERENCE			
PAGE NUMBER 13 14	PAGE DATE 18 July 75	PARAGRAPH 2a	
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>R. E. Heggen</i>	22 OCT 74		
<i>R. E. Heggen</i>	1 Apr 75		
<i>R. E. Heggen</i>	6 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
N/A	N/A	N/A	SW WB 3ABR90530-III-1 Pharmaceutical Preparations Transparency Set #1, Pharmaceutical Calcu- lations II
CRITERION OBJECTIVES AND TEACHING STEPS			
2a. Given information pertaining to reducing and enlarging formulas, specific gravity, percentage preparations, concentration and dilution, alligation, and temperature conversion, solve problems in each area in WB 3ABR90530-III-1 with 60% accuracy. SW (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE <i>W. Nelson 31 Oct 74</i>		INSTRUCTOR	
COURSE NUMBER 3 ABR 90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE Techniques of Pharmaceutical Compounding			
LESSON DURATION			
CLASSROOM / Laboratory 8/0	LABORATORY Complimentary 2	TOTAL 10	
POI REFERENCE			
PAGE NUMBER 14	18 July 75		PARAGRAPH 4a
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	8 1 OCT 1974		
<i>[Signature]</i>	25 APR 75		
<i>[Signature]</i>	6 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
Laboratory Equipment	N/A	N/A	Transparency Set #1 Pharmaceutical Preparations WB 3ABR90530-III-2 Pharmaceutical Preparations
CRITERION OBJECTIVES AND TEACHING STEPS			
4a. Identify laboratory equipment, equipment user maintenance procedures, metrology procedures, incompatibilities, and methods of comminution. (Teaching steps listed in Part II)			

LESSON PLAN (Part I, General)

APPROVAL OFFICE AND DATE MSDB Wilson 31 Oct 74		INSTRUCTOR	
COURSE NUMBER 3 ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE PHARMACEUTICAL PREPARATIONS AND THEIR MANUFACTURE	
LESSON TITLE Pharmaceutical Dosage Forms			
LESSON DURATION			
CLASSROOM/Laboratory 4/0	~~***~~ Complimentary 2	TOTAL 6	
POI REFERENCE			
PAGE NUMBER 15	P/ 18 July 75	PARAGRAPH 5a	
STS/CTS REFERENCE			
NUMBER STS905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	8 1 OCT 1974		
<i>[Signature]</i>	25 APR 75		
<i>[Signature]</i>	8 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
N/A	N/A	N/A	WB3ABR90530-III-2 Pharmaceutical Preparations

CRITERION OBJECTIVES AND TEACHING STEPS

5a. Identify the properties, preparation techniques and incompatibilities of waters, spirits, solutions and syrups.

(Teaching steps listed in Part II)

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LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE HSTN <i>Wilson 31 Oct 74</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE Pharmaceutical Dosage Forms			
CLASSROOM/Laboratory 8/C		LESSON DURATION ~~XXXXXX~~ Complimentary 2	TOTAL 10
POI REFERENCE			
PAGE NUMBER 15	PAGE DATE 18 July 75		PARAGRAPH 56
STS/CTS REFERENCE			
NUMBER STS 905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	8 1 Oct 1974		
<i>[Signature]</i>	25 APRIL 75		
<i>[Signature]</i>	6 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
N/A	N/A	N/A	WB3ABR90530-III-2 Pharmaceutical Preparations
CRITERION OBJECTIVES AND TEACHING STEPS			
5b. Identify the properties, preparation techniques and incompatibilities of eye, ear, and nose preparations, elixirs, tinctures, mixtures, magmas, suspensions, gels, lotions and liniments. (Teaching steps listed in Part II)			

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APPROVAL OFFICE AND DATE MSDB <i>Wilson 31 Oct 74</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE Pharmaceutical Dosage Forms			
LESSON DURATION			
CLASSROOM/Laboratory 10/0		~~LABORATORY~~ Complimentary 3	TOTAL 13
POI REFERENCE			
PAGE NUMBER 15		PAGE DATE 18 July 75	PARAGRAPH 5c
STS/CTS REFERENCE			
NUMBER STS905X0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE		DATE	SIGNATURE
<i>[Signature]</i>		1 OCT 1974	
<i>[Signature]</i>		25 APR 75	
<i>[Signature]</i>		6 OCT 1975	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
N/A	N/A	N/A	WB3ABR90530 III-2 Pharmaceutical Preparations

CRITERION OBJECTIVES AND TEACHING STEPS

5c. Identify the properties, preparation techniques and incompatibilities of powders, capsules, emulsions, ointments, pastes, creams and suppositories.

(Teaching steps listed in Part II)

586

LESSON PLAN (Part I, General)

APPROVAL OFFICE AND DATE MSDB <i>Wilson</i> 31 Oct 74		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE Pharmacy Specialist ✓	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE Pharmaceutical Dosage Forms			
LESSON DURATION			
CLASSROOM/Laboratory 12/0	~~4~~ Complimentary 4	TOTAL 16	
POI REFERENCE			
PAGE NUMBER 15	PAGE DA 18 July 75	PARAGRAPH 5d	
STS/CTS REFERENCE			
NUMBER STS905X0		28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	1 OCT 1974		
<i>[Signature]</i>	25 APRIL 75		
<i>[Signature]</i>	6 OCT 1975		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
N/A	N/A	N/A	WB3ABR90530-III-2 Pharmaceutical Preparations Slides, Ohio State Admixture Slides and cassette Tape
CRITERION OBJECTIVES AND TEACHING STEPS			
5d. Identify the properties, preparation techniques and incompatibilities of parenterals, bulk compounding, prepackaged items and intravenous admixtures. (Teaching steps listed in Part II)			

611

587

APPROVAL OFFICE AND DATE MSDH <i>Wilson 6 Oct 75</i>		INSTRUCTOR	
COURSE NUMBER 3ABR90530		COURSE TITLE PHARMACY SPECIALIST	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations and Their Manufacture	
LESSON TITLE PHARMACEUTICAL DOSAGE FORMS			
LESSON DURATION			
CLASSROOM 0	laboratory 0	~~LABORATORY~~ CTT .5	TOTAL .5
POI REFERENCE			
PAGE NUMBER 15	PAGE DATE 18 July 75	PARAGRAPH 5e	
STS/CTS REFERENCE			
NUMBER STS905x0		DATE 28 Feb 75	
SUPERVISOR APPROVAL			
SIGNATURE <i>[Signature]</i>		DATE 6 OCT 1975	
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
N/A	N/A	N/A	WB/SG 3ABR9XXXX/ MOBM0104-1
CRITERION OBJECTIVES AND TEACHING STEPS			
5e. Identify information as classified, unclassified, or possible intelligence value, Top Secret, Secret, Confidential, or for Official Use Only. (Teaching Steps Listed In Part II)			

LESSON PLAN (Part I, General)			
APPROVAL OFFICE AND DATE <i>Wilson 21 Oct 74</i>		INSTRUCTOR	
COURSE NUMBER 24100000		COURSE TITLE Pharmacy Specialist	
BLOCK NUMBER III		BLOCK TITLE Pharmaceutical Preparations	
LESSON TITLE Compounding Laboratory			
CLASSROOM/Laboratory <i>0/42</i>		LESSON DURATION ~~LABORATORY~~ Complementary <i>16</i>	TOTAL <i>58</i>
POI REFERENCE			
PAGE NUMBER <i>16/17</i>	PAGE <i>18 July 75</i>	PARAGRAPH <i>6a, 6b, 6c</i>	
STS/CTS REFERENCE			
NUMBER SPS 905X0		28 Feb 75 1974	
SUPERVISOR APPROVAL			
SIGNATURE	DATE	SIGNATURE	DATE
<i>[Signature]</i>	<i>22 OCT 74</i>		
<i>[Signature]</i>	<i>18 APR 75</i>		
<i>[Signature]</i>	<i>6 OCT 75</i>		
PRECLASS PREPARATION			
EQUIPMENT LOCATED IN LABORATORY	EQUIPMENT FROM SUPPLY	CLASSIFIED MATERIAL	GRAPHIC AIDS AND UNCLASSIFIED MATERIAL
Laboratory equipment Laminar flow hood IV admixture materials Chemicals Class A balances Typewriters Prescriptions Also p filter-tank unit (over)	N/A	N/A	N/A
CRITERION OBJECTIVES AND TEACHING STEPS			
6a. Given instructor assistance, necessary references and selected formulas; compound waters, spirits, solutions, and syrups IAW AF Form 2380 and AF Form 2381. Then package and label IAW AFM 168-4. 6b. Given instructor assistance, necessary references and selected formulas; compound ear and nose preparations, elixirs, tinctures, suspensions, lotions and liniments IAW AF Form 2380 and AF Form 2381. Then package and label preparations IAW AFM 168-4. 6c. Given instructor assistance, necessary references and selected formulas; compound capsules, emulsions, ointments, creams IAW AF Form 2380 and AF Form 2381. Then package and label preparations IAW AFM 168-4. (Teaching steps listed in Part II)			

Equipment Located
in Laboratory
(Cont'd)

Tablet counting machine
Bottle filling machine
Label imprinter

CRITERION OBJECTIVES AND TEACHING STEPS (Continued)

6d. Given instructor assistance, necessary references and selected prescriptions; compound intravenous admixtures, correcting any incompatibilities, using accepted methods and techniques as outlined in checklist 3ABR90530-III-6d.

6e. Given instructor assistance, rotate through the outpatient, inpatient, supply and administrative work areas of the USAF Regional Hospital Sheppard Pharmacy IAW local directives and policies.

HANDOUTS III - I through IV

Course 10 - 8

1. Pine Tar Ointment (unofficial)

Pine Tar.....	150.90 Gms
Zinc Oxide.....	54.18 Gms
Petrolatum.....	248.62 Gms
Makes about	<u>454.00 Gms</u>

Calculate the quantity of each ingredient needed to prepare
1 Kilogram of the above ointment.

2. How many mgsof each ingredient are needed to make one Kgm.
of the following ointment?

Benzoic Acid.....	12 Gms
Sulfur.....	20Gms
Coal Tar.....	15Gms
Zinc Oxide.....	45Gms
Petrolatum.....	135Gms
To make	227Gms

3. Calculate the number of Grams of Magnesium Oxide and
Calcium Carbonate needed to make 4 ounces of the
following mixture.

Calcium Carbonate.....	370Gms
Magnesium Oxide.....	70Gms
Sodium Bicarbonate.....	280Gms
Bismuth Subcarbonate.....	188Gms
M. ft. powder	

4. From the following formula, calculate the number of grains
of Sodium Chloride needed to make one gallon of Normal
Saline Solution.

Sodium Chloride.....	gr. xxxiiss
Aqua qs.....	flz viii

5. Hydrocortisone..... 1 part
 Sulfur.....20 parts
 Zinc Oxide Paste.....79 parts

Mix and make 30 Gms

In the above formula calculate the number of Gms of each ingredient needed to make the formula.

6. Calculate the number of Grams each of Camphor and of Starch needed to make 60 Gms of the following:

Camphor.....	.8 parts
Calamine powder.....	8.0 parts
Starch.....	9.2 parts
Talc.....	30.0 parts

7. Calculate the number of mgs of Methylparaben contained in 2.5 ~~g~~ of the following:

Methylparaben.....	0.26Gms
Propylparaben.....	0.14Gms
Purified Water..... QSAD.....	1000.00mls

8. Calculate the number of grains of Thyriod contained in one dose of the following:

Thyriod.....	199.4mg
Phenobarbital.....	100mg
Ascorbic Acid.....	50mg
Makes 20 capsules	
Sig: Take 2 capsules TID	

9. If 53 milliliters of a liquid weighs 61.48Gms, what is it's specific gravity?

10. Calculate the specific gravity of a liquid if one pint weighs one pound.

11. Sixteen fluid ounces of a liquid weighs 11/3 lbs.(apoth). Calculate the specific gravity of the liquid.

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12. How many grams will 40 mls of Chloroform weigh, if the specific gravity is 1.46?

13. How many pounds(AV) does one gallon of glycerin, sp.gr. 1.25 weigh?

14. 45 Gms of Glycerin (sp.gr. 1.25) will have a volume of ____.

15. What volume will 1 Kgm of Lactic Acid measure if the sp.gr. is 1.264.

1. Calculate the number of Grams of Sodium Citrate needed to prepare 1 Kg. of 40% (w/v) solution of Sodium Citrate in water.
2. How many grains of Phenol are there in 2 fl $\frac{3}{4}$ of a 4% (w/v) solution?
3. How many mgs of Sodium Chloride are required to prepare $\frac{3}{4}$ of a 1:10% (w/w) preparation?
4. With 14,560 grs of Potassium Iodide, how many fl $\frac{3}{4}$ of a 2% solution can you prepare?
5. How many liters of a 8% solution can be prepared from 22 Grams of Gentian Violet?
6. If 8 liters of a solution of Iodine in water contains 14 Grams of Iodine, what is the percentage strength (w/v) of the solution?
7. If 3 $\frac{3}{4}$ of Boric Acid is dissolved in enough water to make .5 liters, what is the percentage of this solution?
8. How many ml of Peppermint Oil should be added to 1 quart of water to make a 4% solution?
9. How many grains of Sodium Chloride should be mixed with 2 $\frac{3}{4}$ of Potassium Iodide to make a .3% preparation?
10. If you need to prepare a 14% powder, how many grams of active ingredient should be mixed with 22 $\frac{3}{4}$ of your base?
11. A saturated solution of Sodium Chloride boils at 227.1°F. Express this temperature on the centigrade scale.
12. Theobroma Oil melts between 30° and 35°C. What is the range of it's melting point on the Fahrenheit scale?
13. Convert the following:
 - a. -20°C.
 - b. -14°F.
 - c. 32°F.
 - d. -40°C.

1. Convert the following to %.
a. 1:500 b. 1:250 c. .001 d. 1:800
2. If 600 ml of a 25% solution is diluted to 3 Liters, what will be the percentage strength?
3. How many ml of a 1:25 stock solution of a chemical should be used to prepare 500 ml of a 1:4000 solution?
4. How many mls of a .5% Benzalkonium Chloride solution are needed to fill a prescription requiring 30 ml of 1:10,000 Benzalkonium Chloride?
5. How many mls of 1:1000 solution of a drug are needed to prepare 50 ml of a 1:20,000 solution?
6. How many mls of water should be added to a quart of 1:500 solution to make a 1:4000 solution?
7. How much water should be added to a Liter of 1:250 solution to make a .1% solution?
8. How many milliliters of 25%(w/v) mild silver protein solution and how many milliliters of 5%(w/v) mild silver protein solution are required to make 500 ml of a 10%(w/v) solution?
9. How many mls of water must be mixed with 500 ml of 70%(w/v) alcohol in order to reduce the strength to 50%(w/v) alcohol?
10. Calculate the number of Grams of 2% Boric Acid ointment needed to be added to 40Grams of 10% Boric Acid ointment in order to prepare some 5% Boric Acid ointment.
11. How many Grams of Petrolatum should be added to 180Grams of 45% Sulfathiazole ointment to make a 8% Sulfathiazole ointment?
12. What percentage of Boric Acid is contained in a mixture of 20Grams of 30% Boric Acid ointment, 35Grams of 40% Boric Acid ointment and 50Grams of 75% Boric Acid ointment?
13. What is the percentage of Alcohol in a mixture of 800ml of 95% Alcohol, 300 ml of 65% Alcohol, 75ml of 30% Alcohol and 250ml of pure Alcohol?

1 Camphor and Soap Liniment, NF

Hard Soap, powdered.....	60 Gm
Camphor, small pieces.....	45 Gm
Rosemary oil.....	10ml
Alcohol USP.....	700 ml
Purified water.... QSAD.....	1,000 ml

Calculate the quantity of each ingredient needed to make 200 ml of the above formula.

2. Calculate the number of ml of each ingredient listed below that will be required to make one liter.

Aromatic Cascara.....	1 part
Liquid Petrolatum.....	3 parts
Milk of Magnesia.....	4 parts

3. 20 Grams of Acetic Acid solution is needed for a preparation and has a specific gravity of 1.05. Calculate the amount of Acetic Acid in ml that is needed for this preparation.

4. Calculate the specific gravity of a liquid if one-half liter weighs 730 Grams.

5. Calculate the number of Grams that should be used to make two ounces (Apoth) of an ointment if it contains 10% (w/w) of active ingredient.

6. Calculate the percent of Mercurochrome (w/v) in one gallon of solution containing 227 Grams of Mercurochrome.

7. Calculate the number of Grams of Potassium Iodide that should be added to 250ml of water, so that the finished solution will be 12%.

8. Convert the following degreesCentigrade to Fahrenheit:

a. 82 C

b. -15 C

1.

9. Convert the following degrees Fahrenheit to Centigrade:

a. 16 F

b. -31 F

10. Calculate the number of Grams of petrolatum ointment base that should be added to 180 Gms. of 10% Ammoniated Mercury Ointment to dilute its strength to 3%.

11. Calculate the number of Grams of Lactose that is needed to dilute 25Gms of 1:10 trituration of Atropine Sulfate to a 1:50 strength.

12. Calculate the number of mls of 95% Alcohol and of 50% Alcohol needed to prepare 180mls of 70% Alcohol.

13. Calculate the percentage strength of the following mixture:

24oz. containing 86% opium

8oz. containing 73% opium

3oz. containing 70% opium

10-8

Technical Training

Pharmacy Specialist

PHARMACEUTICAL PREPARATIONS

November 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
Department of Biomedical Sciences
Sheppard Air Force Base, Texas 76311

Designed For ATC Course Use

DO NOT USE ON THE JOB

600

Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas 76311

SW 3ABR90530-III-1
November 1975

PHARMACEUTICAL PREPARATIONS

OBJECTIVE

Solve problems in reducing and enlarging formulas, specific gravity, percentage preparations, concentration and dilution, and temperature conversion.

INTRODUCTION

Each type of problem you may encounter will be explained by the instructor. Fill in each blank in the example sections as the information is given to you. This will assist you in working the practice problems. These problems will be evaluated by the instructor to insure you are working them correctly. Complete all problems assigned. SHOW ALL WORK!

INFORMATION

REDUCE AND ENLARGE FORMULAS

EXAMPLE: Reduce this formula to make 120 ml.

Peppermint Oil.....	2 ml
Talc.....	15 Gm
Purified Water.....OSAD.....	1000 ml

Ratio and Proportion Method

- Step 1. How much will the original (old) formula make? _____ . How much Peppermint Oil does the "old" formula call for?

Now write these values side by side.

2 ml 1000 ml

- Step 2. How much do you wish to make of the original formula? _____ . Place this value over the 1000 ml. (Total Amounts)

$\frac{120 \text{ ml (New amount)}}{1000 \text{ ml (Old amount)}}$

Use an "x" for the number of ml Peppermint Oil wanted. Place the "x" over the 2 ml. (Active Ingredients)

A.I. T

$\frac{x \text{ ml} = 120 \text{ ml}}{2 \text{ ml} \quad 1000 \text{ ml}}$

This supersedes SW 3ABR90530-III-1, September 1974.

Previous editions may be used until the existing supply is exhausted.

Step 3. Since the problem is set up, now use the Ratio Proportion Method and solve for the unknown.

$$\frac{x \text{ ml}}{2 \text{ ml}} = \frac{120 \text{ ml}}{1000 \text{ ml}}$$

Cross multiply:

$$1000 x = 240$$

Divide both sides of the equation by 1000, maintaining an equation of equality.

$$\frac{1000 x}{1000} = \frac{240}{1000}$$

Step 4. Your answer for the value of "x" is:

$$x = 0.24 \text{ ml}$$

Therefore, 0.24 ml of Peppermint Oil is needed in the formula to make 120 ml.

Step 5. Repeat this procedure for each ingredient, making sure that each value is placed over the proper value to be found.

A.I. T

$$\frac{x \text{ Gm}}{15 \text{ Gm}} = \frac{120 \text{ ml}}{1000 \text{ ml}}$$

$$1000 x = 1800$$

$$x = 1.8 \text{ Gm of Talc}$$

Step 6. Your new formula is: Peppermint Oil..... 0.24 ml
Talc..... 1.8 Gm
Purified Water....QSAD..... 120.0 ml

NOTE: To enlarge formulas, use exactly the same procedures. Remembering to place the new value over the old value.

Factor Method of Solving the Same Example That is On Page One

Step 1. How much will the original (old) formula make? _____.

Step 2. How much of the original formula do you wish to make (new)? _____.

Step 3. Place these values over each other as illustrated.

$$\frac{120 \text{ ml (New Amount)}}{1000 \text{ ml (Old Amount)}}$$

Step 4. Solve for your FACTOR.

FORMULA FOR FACTORING

NEW
OLD

Step 5. Divide the Old into the New.

$$120 \div 1000$$

Your FACTOR is:

$$0.12$$

Step 6. Once this factor is found, to find the amount of active ingredient to be used in the new formula, multiply each ingredient by this FACTOR.

$$\begin{array}{r} \text{Peppermint Oil} \\ 2 \text{ ml} \\ \times .12 \text{ (factor)} \\ \hline 24 \end{array}$$

Answer: 0.24 ml Peppermint Oil

$$\begin{array}{r} \text{Talc} \\ 15 \text{ gm} \\ \times .12 \text{ (factor)} \\ \hline 180 \end{array}$$

Answer: 1.80 gm Talc

Step 7. Your new Formula is: Peppermint Oil..... 0.24 ml
Talc..... 1.8 gm
Purified Water.....QSAD.....120.0 ml

NOTE: To enlarge formulas, use exactly the same procedure. Remember to place the New over the Old to find your factor and multiply this number by your active ingredient to find your new amount.

In reducing formulas the value of your factor will be less than one, in enlarging formulas the value of your factor will be greater than one.

PROCEDURE

The object of this lesson is to solve problems in pharmaceutical calculations. Specifically, you will calculate the amount of medicinals in a preparation by:

1. Reducing and enlarging formulas.
2. Preparing formulas when given proportionate parts.
3. Calculating the amount of ingredient in a dose.

QUESTIONS

1. Reduce this formula to make 100 ml:

Liquid Coal Tar.....	4 ml	Answer _____
Sulfur.....	10 ml	Answer _____
Lime Water.....	50 ml	Answer _____
Bentonite Magma.....QSAD.....	120 ml	Answer _____

2. Reduce this formula to make 30 ml:

Ephedrine Sulfate.....	30 Gm	Answer _____
Chlorobutanol.....	5 Gm	Answer _____
Sodium Chloride.....	3.6 Gm	Answer _____
Purified Water.....QSAO.....	1000 ml	Answer _____

3. Enlarge this formula to make 1 gallon:

Talc.....	12 Gm	Answer _____
Bentonite.....	3.5 Gm	Answer _____
Zinc Oxide.....	25 Gm	Answer _____
Distilled Water.....QSAO.....	100 ml	Answer _____

4. Enlarge this formula to make 1 liter:

Orange Oil.....	12 ml	Answer _____
Lemon Oil.....	3 ml	Answer _____
Coriander Oil.....	1.2 ml	Answer _____
Anise Oil.....	0.3 ml	Answer _____
Alcohol USP.....Q.S.A.D.....	60 ml	Answer _____

INFORMATION

SOLVING PROBLEMS USING PROPORTIONATE PARTS

EXAMPLE: Prepare 130 Gm. of this ointment.

Starch.....	5 parts
Zinc Oxide.....	10 parts
Hydrophilic Ointment Base.....	50 parts

Step 1. Add up all the parts to find the _____ number of parts in the formula.

5
10
50
65 Total Parts

Step 2. What is the _____ number of Gms. you wish to make? 130 Gms.

Step 3. Since 130 Gms. represents the total weight and 65 parts represents the _____ parts, therefore, 130 Gms. equals 65 parts.

To find the weight of one part, divide the number of parts into the total weight. 130 Gms. ÷ 65

Each part will weigh 2 Gms.

Step 4. Multiply the weight of one part times the number of parts required for each ingredient.

$$\begin{array}{r} \text{Starch:} \quad 2 \text{ gm (Weight of 1 part)} \\ \quad \times 5 \\ \hline 10 \text{ gms} \end{array}$$

$$\begin{array}{r} \text{Zinc Oxide:} \quad 2 \text{ gm} \\ \quad \times 10 \\ \hline 20 \text{ gms} \end{array}$$

$$\begin{array}{r} \text{Hydrophilic Oint:} \quad 2 \text{ gm} \\ \quad \times 50 \\ \hline 100 \text{ gms} \end{array}$$

Step 5. Your new formula: Starch..... 10 gms
Zinc Oxide..... 20 gms
Hydrophilic Oint.... 100 gms
Total Weight 130 gms

QUESTIONS

1. From the following formula, calculate the quantity of each ingredient required to make 2 ounces (Apoth) of this ointment.

Zinc Oxide.....	2 parts	Answer _____
Coal Tar.....	2 parts	Answer _____
Starch.....	15 parts	Answer _____
Petrolatum.....	25 parts	Answer _____

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2. From the following formula, calculate the quantity of each ingredient required to prepare 120 ml of this solution.

Witch Hazel.....	4 parts	Answer _____
Glycerin.....	1 part	Answer _____
Boric Acid Solution.....	15 parts	Answer _____

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INFORMATION

CALCULATING THE AMOUNT OF MEDICATION IN A DOSE

EXAMPLE: How many mg. of Codeine Phosphate will each dose contain?

Codeine Phosphate..... 240 mg
Elixir of Terpin Hydrate..... OSAD..... 120 ml

SIG: 1 teaspoonful as needed.

Step 1. How much active ingredient is in the "old" formula? _____ mg. How many ml will the "old" formula make? _____ ml. Write these values side by side.

240 mg 120 ml

Step 2. How much active ingredient will each dose contain? _____ This is unknown so use "x" and place it over the 240 mg. How many ml will be in each dose (new)? _____ ml. Write this value over 120 ml.

A.I.	T
x mg	5 ml
240 mg	120 ml

Step 3: Solve by using the Ratio and Proportion method.

Cross multiply:

120 x = 1200

Divide both sides of the equation by 120.

x = 10

Step 4. The value of "x" is your answer which is 10 mg. Always write the prescription doses using the amount of active ingredient per dose.

10 mg/5 ml

NOTE: When calculating for one active ingredient it is not necessary to use the Factor Method. The Factor Method is a short cut in finding many active ingredients.

QUESTIONS

1. In the following prescription calculate the number of mg of Phenobarbital in each 2 teaspoonfuls dose.

Phenobarbital.....	4 Gm
Alcohol.....	150 ml
Glycerin.....	450 ml
Distilled Water..QSAD	1000 ml

Answer _____

2. In the following prescription calculate the number of Grams of Terpin Hydrate and the number of ml of Alcohol the patient will receive in each dose.

Terpin Hydrate.....	17.00 Gm
Alcohol.....	430.00 ml
Syrup USP.....	100.00 ml
Glycerin.....	400.00 ml
Benzaldehyde.....	0.05 ml
Distilled Water.....QSAD.....	1000.00 ml

SIG: $\mathcal{Z}$ ss QID prn.

Answer _____

Answer _____

INFORMATION

SPECIFIC GRAVITY

Specific Gravity is the ratio of the weight of a liquid to the weight of an equal volume of water. Water has the Specific Gravity of 1.000.

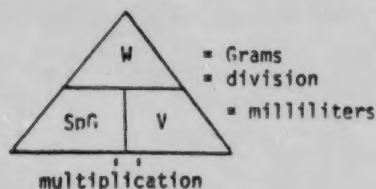
The general formulas used in calculating Specific Gravity problems are:

$$\text{Specific Gravity} = \frac{\text{Weight (in Grams)}}{\text{Volume (in milliliters)}}$$

$$\text{Weight} = \text{Specific Gravity} \times \text{Volume (in mls)}$$

$$\text{Volume} = \frac{\text{Weight (in Grams)}}{\text{Specific Gravity}}$$

Instead of learning three separate formulas, use the triangle below with the proper labeling:



This triangle can be used to solve any specific gravity problem.

CALCULATING THE SPECIFIC GRAVITY OF A LIQUID

EXAMPLE: What is the Specific Gravity of Glycerin if 100 ml weighs 125 Gm?

Step 1. Draw the triangle and label.

Step 2. Assign values to the appropriate terms:

$$W = 125 \text{ Gms}$$

$$V = 100 \text{ mls}$$

Step 3. Substitute the assigned values for the terms.

$$\text{SpG} = W (\text{Gm}) \text{ divided by } V (\text{ml})$$

Step 4. Solve by the process indicated:

$$\text{SpG} = 125 \text{ Gm divided by } 100 \text{ ml}$$

Step 5. Your answer:

$$\text{SpG} = 1.250$$

QUESTIONS

1. If 125 ml of a liquid weighs 160 Gm, what is its Specific Gravity?

Answer _____

2. If 134 Grams of a liquid measures 142.6 ml, what is its Specific Gravity?

Answer _____

3. What is the Specific Gravity of a liquid if 2 liters weighs 1.75 kilograms?

Answer _____

4. Six pounds (Apoth) of a liquid measures 128 fluid ounces. What is its Specific Gravity?

Answer _____

INFORMATION

CALCULATING THE WEIGHT OF LIQUIDS

EXAMPLE: What is the weight of 200 ml of Castor Oil (SpG 0.96)?

Step 1. Draw the triangle and label.

Step 2. Assign values to the appropriate terms: $SpG = 0.96$
 $V = 200 \text{ ml}$

Step 3. Substitute the assigned values for the terms, and place them in the triangle

Step 4. Always make sure that you have the volume in MILLILITERS before you solve your problem.

$$W = SpG \times V$$

Step 5. Solve by the process indicated.

$$W = 0.96 \times 200 \text{ ml}$$

Step 6. Your answer:

$$W = 192$$

Step 7. Since you are solving for weight and you have the Specific Gravity given, and the Volume is in ml, therefore, your answer will be in Grams.

$$W = 192 \text{ Gm}$$

NOTE: Always label your answer to what it is. This will tell you where you are at all times. If the problem desires a different unit of measure, by labeling it will assist you in what to do.

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QUESTIONS

1. How many grams does 225 ml of an acid weigh if the Specific Gravity of the acid is 1.83?

Answer _____

2. The Specific Gravity of a liquid is 0.75 and the liquid measures 3 quarts. What is its weight in grams?

Answer _____

3. A liquid has a Specific Gravity of 1.50. What is the weight of 1.5 Liters?

Answer _____

4. If an oil has a Specific Gravity of 1.55, what is the weight of 240 ml?

Answer _____

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INFORMATION

CALCULATING THE VOLUME OF LIQUIDS

EXAMPLE: What is the volume of 156 Gm of Isopropyl Alcohol (SpG 0.78)?

Step 1. Draw the triangle and label.

Step 2. Assign values to the appropriate terms: $W = 156 \text{ Gm}$

$\text{SpG} = 0.78$

Step 3. Substitute the assigned values for the terms, and place them in the triangle.

Step 4. Always make sure that you have the weight in GRAMS before you solve your problem.

$V = W \text{ divided by SpG}$

Step 5. Solve the process indicated.

$V = 156 \text{ Gm divided by } 0.78$

Step 6. Your answer:

$V = 200$

Step 7. Since you are solving for volume and you have the Specific Gravity given, and the Weight is in Gm, therefore, your answer will be in milliliters.

$V = 200 \text{ ml}$

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QUESTIONS

1. What is the volume in ml of 227 grams of a liquid having the Specific Gravity of 1.230?

Answer _____

2. A formula for 1000 ml of a preparation calls for 800 grams of Cottonseed Oil with a Specific Gravity of 0.920. How many ml of Cottonseed Oil should be used in preparing 5 liters of this formula?

Answer _____

3. What is the volume, in pints, of 40 lb. of a liquid with the Specific Gravity of 1.32?

Answer _____

4. How many ml will 3 kilograms of oil be if its Specific Gravity is 1.11?

Answer _____

633

615

INFORMATION

PERCENTAGE PREPARATIONS

Three types of percentage preparations:

Weight in Volume.....	W/V
Volume in Volume.....	V/V
Weight in Weight.....	W/W

The General Formulas used in calculating percentage preparations are:

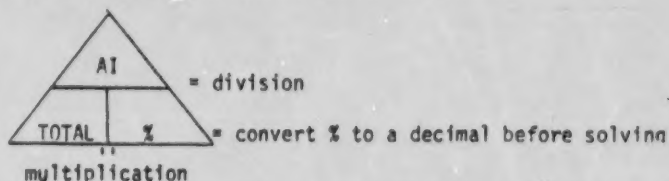
$$\text{Active Ingredient (AI)} = \text{Total Amount} \times \text{Percent (\%)}%$$

$$\text{Percent (\%)} = \frac{\text{Active Ingredient (AI)}}{\text{Total Amount}}$$

$$\text{Total Amount} = \frac{\text{Active Ingredient (AI)}}{\text{Percent (\%)}}$$

In working these problems, the percent is converted to a decimal before solving. The general rule for changing percent to a decimal is to divide by 100. The general rule for changing a decimal to a percentage is to multiply by 100. Watch your decimal point!

Instead of learning the three general formulas above, use the triangle below with the proper labeling:



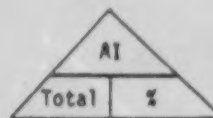
NOTE: Change your percentage (%) to a decimal before solving your problem in the triangle.

616

CALCULATING THE AMOUNT OF ACTIVE INGREDIENT IN A PERCENTAGE PREPARATION
WHEN GIVEN THE PERCENTAGE STRENGTH AND THE TOTAL AMOUNT

EXAMPLE: How many grams of Sodium Chloride will be required to prepare 100 ml of a 15% (W/V) solution?

Step 1. Write the complete formula or draw your triangle and label.



Step 2. Assign the values to the appropriate terms.

- | | |
|--|------------------------------|
| a. What are you looking for?
Therefore, the Active ingredient becomes: | Active Ingredient
X Grams |
| b. What is your total amount? | 100 milliliters |
| c. What is your percentage? 15%
Change the percentage to a decimal by dividing it by 100. | 0.15 |

Step 3. Rewrite the values, substituting them for the terms in the triangle.

$$X = 100 \times 0.15$$

Step 4. Solve by the process indicated:

$$X = 15$$

Step 5. Since you have assigned the value of X as Grams, your answer is:

15 Grams

NOTE: Notice in your problem after 15% you see the symbol (W/V). Where have you seen this term before? Specific Gravity. Therefore, if the volume is given to you as milliliters, your weight will be in Grams. In solving these problems, make sure you are working in the proper system - Metric, Apothecary, or Avoirdupois. This will depend on what is given to you and what you are to solve.

QUESTIONS

1. How many grams of Mercuric Chloride are required to prepare 250 ml. of a 5% (W/V) solution?

Answer _____

2. How many grams of Boric Acid are there in 30 ml. of a 2% (W/V) solution?

Answer _____

3. How many grams of Phenol are required to prepare 480 ml. of a 1/10% (W/V) solution?

Answer _____

4. How many grains of Silver Nitrate will be required to prepare 6 fluid ounces of a 0.25% (W/V) solution?

Answer _____

5. How many grains of Gentian Violet should be used in preparing 2 fluid ounces of a 1/2% solution?

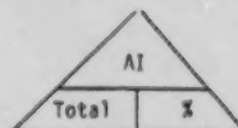
Answer _____

INFORMATION

CALCULATING THE PERCENTAGE STRENGTH OF A PREPARATION WHEN GIVEN
THE TOTAL AMOUNT AND THE AMOUNT OF ACTIVE INGREDIENT

EXAMPLE: What is the percentage strength of 240 ml (V/V) of a solution containing 12 ml of Orange Oil?

Step 1. Write the complete formula or draw your triangle and label.



Step 2. Assign the values to appropriate terms.

- | | |
|--|---------------------|
| a. What is your total amount? | 240 ml |
| b. What is your active ingredient? | 12 ml of Orange Oil |
| c. What are you looking for?
Therefore, let X represent the
percentage strength. | percentage strength |

Step 3. Rewrite the values, substituting them for the terms in the triangle.

$$X = \frac{12 \text{ ml}}{240 \text{ ml}}$$

Step 4. Solve by the process indicated:

$$X = 0.05$$

Step 5. $X = 0.05$, is this your final answer?

NO, it is a decimal answer.

Step 6. Changing a decimal to a percentage (%), multiply by 100. Therefore, your answer is:

5%

NOTE: Always make sure you know what you are looking for. As in this case you were looking for percentage (%). Therefore an additional step is indicated. Changing a decimal to a percent (%) by multiplying your answer by 100.

QUESTIONS

1. If 425 Grams of Sucrose is dissolved in enough water to make 500 ml, what is the percentage strength of this solution?

Answer _____

2. If 2 liters of a solution of Iodine in Alcohol contains 7 Grams of Iodine, what is the percentage strength (W/V) of the solution?

Answer _____

3. If 1 gallon of a solution contains 474 Grams of solute, what is the percentage strength of the solution?

Answer _____

4. What are the percentages (W/V) of the ingredients in the following prescription?

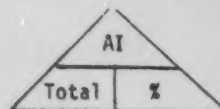
Zinc Sulfate	2 grains	Answer _____
Boric Acid	20 grains	
Distilled Water ...QSAD.....	1 fluid ounce	Answer _____

INFORMATION

CALCULATING THE TOTAL AMOUNT OF A PREPARATION WHEN GIVEN THE PERCENTAGE STRENGTH AND THE AMOUNT OF ACTIVE INGREDIENT

EXAMPLE: How many Grams of a 15% (W/W) Sulfur Ointment can be made from 30 Grams of Sulfur powder?

Step 1. Write the complete formula or draw your triangle and label.



Step 2. Assign the values to the appropriate terms.

a. What are you looking for?
Therefore, the total amount becomes:

Total Amount

X Grams

b. What is your active ingredient?

30 Grams

c. What is your percentage? 15%
Change the percentage to a decimal by dividing by 100.

0.15 (decimal form of 15%)

Step 3. Rewrite the values, substituting them for the terms in the triangle.

X = Active Ingredient

X

X = 30 Grams
0.15

Step 4. Solve by the process indicated:

X = 200

Step 5. Since you have assigned the value of X as Grams, your answer is:

200 Grams

NOTE: Notice in your problem after 15% you see the symbol (W/W). This means that there are solids involved in this problem. If one is given to you as Grams, then the other will be in Grams. Make sure that you stay in the proper system when solving and convert to any other system after solving for the unknown if necessary.

QUESTIONS

1. How many ml of a 0.1% solution can be made from one gram of Atropine Sulfate?

Answer _____

2. How many Liters of a 2% (W/V) Iodine Tincture can be made from 123 grams of Iodine?

Answer _____

3. How many fluid ounces of a 0.55% solution can be prepared from 75 grains of Scopolamine Hydrobromide?

Answer _____

4. How many milliliters of a 6% solution can be prepared from 14 grams of Neomycin Sulfate?

Answer _____

5. With 43 grams of Hydrocortisone Powder, how many grams of 1.5% ointment could you make?

Answer _____

622

INFORMATION

CALCULATING THE AMOUNT OF ACTIVE INGREDIENT WHEN GIVEN THE PERCENTAGE STRENGTH AND THE AMOUNT OF THE SOLVENT

EXAMPLE: How many grams of Potassium Iodide should be added to 180 ml of water, so that the finished solution will be 10%?

Step 1. What percent of the finished product will the number of Grams represent 10%

Step 2. What percent of the finished product does the 180 ml of water represent? 90%

Step 3. Since you are adding to the 180 ml, your total amount of the preparation will be increasing. Therefore, the triangle CANNOT BE USED. Set up a ratio and proportion, using 10% over 90% and X over 180 ml. (Water can be changed to Grams).

$$\begin{array}{c|c} 10\% & X \text{ gm} \\ \hline 90\% & 180 \text{ gm} \end{array}$$

Step 4. Solve by the process indicated:

$$\begin{array}{c|c} 10 & X \\ \hline 90 & 180 \end{array}$$

Cross multiply:

$$90x = 1800$$

Divide both sides by 90:

$$x = 20$$

Step 5. Since you are solving for X, and X is in grams, your answer is:

20 Grams

647

QUESTIONS

1. How many mg of Boric Acid should be added to 240 ml of water to make a 2% solution?

Answer _____

2. How many grams of Vioform should be added to 1 pound (Avoirdupois) ointment base to yield a 3% Vioform ointment?

Answer _____

3. How many ml of Orange Oil should be added to 1 pint of water to make a 0.02% Solution?

Answer _____

4. How many grains of Phenobarbital should be added to 16 fl $\frac{1}{2}$ of Cherry Syrup to make a 0.3% Phenobarbital Syrup?

Answer _____

5. If you need to prepare a 15% powder, how many grams of active ingredient should be added to 8 $\frac{3}{4}$ of your powder base?

Answer _____

624

INFORMATION

There is one general formula used in conversion of temperatures.

$$\text{General Formula: } 9C = 5F - 160$$

This formula can be used for converting either Centigrade or Fahrenheit. All that is necessary is to place the value of your given temperature to the proper symbol (C or F) and solve the mathematical equation.

CONVERTING FAHRENHEIT DEGREES TO CENTIGRADE DEGREES

EXAMPLE: Convert -40°F to Centigrade.

- | | |
|---|---------------------------------|
| Step 1. Write the complete formula: | $9C = 5F - 160$ |
| Step 2. Substitute the -40°F for the "F" in the formula: | $9C = 5(-40) - 160$ |
| Step 3. Multiply 5 times -40 and place this product in the equation: | $9C = -200 - 160$ |
| Step 4. Since you have like signs, bring down your sign and proceed as in addition: $-200 - 160 = -360$ | $9C = -360$ |
| Step 5. Divide both sides of the equation by 9 to find the value of C: | $C = -40$ |
| Step 6. Since C stands for Centigrade, your answer is: | $-40^{\circ}\text{ Centigrade}$ |

NOTE: The following are some general mathematical rules:

1. When multiplying like signs, your product will be positive (+).
2. When multiplying unlike signs, your product will be negative (-).
3. When adding like signs, bring down the sign and proceed as in addition.
4. When adding unlike signs, take the sign of the larger number and proceed as in subtraction.

640

QUESTIONS

1. Convert 98.6°F to Centigrade.

Answer _____

2. Convert 32°F to Centigrade.

Answer _____

3. Convert 210°F to Centigrade.

Answer _____

4. Convert -60°F to Centigrade.

Answer _____

5. Convert 212°F to Centigrade.

Answer _____

6. Convert 55°F to Centigrade.

Answer _____

626

INFORMATION

CONVERTING CENTRIGRADE TO FAHRENHEIT

Use the same General Formula: $9C = 5F - 160$

EXAMPLE: Convert 30°C to Fahrenheit.

- | | |
|--|--------------------------------|
| Step 1. Write the complete formula: | $9C = 5F - 160$ |
| Step 2. Substitute the 30°C for the "C" in the formula: | $9(30) = 5F - 160$ |
| Step 3. Multiply 9 times 30 and place this product in the equation: | $270 = 5F - 160$ |
| Step 4. Solving for "F" you must add 160 to both sides of the equation to maintain its equality. What is done to one side of an equation MUST be done to the other side. | $160 + 270 = 5F - 160 + 160$ |
| Step 5. Do the mathematical processes indicated: | $430 = 5F$ |
| Step 6. Divide both sides of the equation by 5 to find the value of "F": | $86 = F$ |
| Step 7. Since F stands for Fahrenheit, your answer is: | $86^{\circ}\text{ Fahrenheit}$ |

NOTE: Remember in calculating for Fahrenheit, one must add 160 to both sides of the equation before solving.

651

QUESTIONS

1. Convert 60°C to Fahrenheit.

Answer _____

2. Convert -47°C to Fahrenheit.

Answer _____

3. Convert 92°C to Fahrenheit.

Answer _____

4. Convert 13°C to Fahrenheit.

Answer _____

5. Convert 44°C to Fahrenheit.

Answer _____

6. Convert -15°C to Fahrenheit.

Answer _____

INFORMATION

CONCENTRATION AND DILUTION OF
STOCK SOLUTIONS AND STOCK TRITURATIONS

Stock solutions are solutions of known concentration that are frequently prepared by the pharmacist for convenience in dispensing.

Stock Triturations are dilutions of potent substances prepared by mixing finely powdered medicaments with finely powdered Lactose in a definite proportion by weight.

General Formula:

$$\text{Amt}_1 \times \%_1 = \text{Amt}_2 \times \%_2$$

EXAMPLE: If 500 ml of a 10% solution was diluted to make a 2% solution, how many ml. will the new solution measure?

Step 1. What information is given concerning the first solution?

500ml of 10%

Step 2. What information is given concerning the second solution?
(use "X" for the unknown)

X ml of 2%

Step 3. Write the complete Formula:

$$\text{Amt}_1 \times \%_1 = \text{Amt}_2 \times \%_2$$

Step 4. Assign values to the appropriate terms:

$$\text{Amt}_1 = 500 \text{ ml}$$

$$\%_1 = 0.10 \text{ (as a decimal)}$$

$$\text{Amt}_2 = "x" \text{ ml}$$

$$\%_2 = 0.02 \text{ (as a decimal)}$$

Step 5. Rewrite the formula substituting the assigned values for the terms:

$$500\text{ml} \times 0.10 = x\text{ml} \times 0.02$$

Step 6. Solve by the processes indicated:

$$50.00 = 0.02x$$

Step 7. Divide both sides of the equation by 0.02:

$$2500 = x$$

Step 8. Since x is mls, your answer is:

$$2500 \text{ ml}$$

QUESTIONS

1. How many ml of a 25% (W/V) solution can be made from 750 ml of a 65% (W/V) solution?

Answer _____

2. If 30 Grams of a 45% (W/W) powder was diluted to make a 30% (W/W) powder, how many grams will the new preparation weigh?

Answer _____

3. If you dilute two pints of a 65% (W/V) solution to 30% (W/V), how many fl ~~oz~~ will the new preparation measure?

Answer _____

4. How many grams of 10% (W/W) Phosphoric Acid can be made from one Kilogram of 85% (W/W) Phosphoric Acid?

Answer _____

5. How many gallons of 70% Alcohol can be made from 10 gallons of 95% (V/V) Alcohol.

Answer _____

INFORMATION

CALCULATING THE AMOUNT OF DILUENT

EXAMPLE: How much water should be added to 1 liter of a 70% solution to make a 35% solution?

- Step 1. Write the complete formula: $Amt_1 \times \%_1 = Amt_2 \times \%_2$
- Step 2. Assign values to the appropriate terms:
- $Amt_1 = 1000 \text{ ml (1 liter)}$
 $\%_1 = 0.70 \text{ (as a decimal)}$
 $Amt_2 = "x" \text{ ml}$
 $\%_2 = 0.35 \text{ (as a decimal)}$
- Step 3. Rewrite the formula substituting the assigned values for the terms: $1000 \text{ ml} \times 0.70 = x \text{ ml} \times 0.35$
- Step 4. Solve by the processes indicated: $700.00 = 0.35x$
 $2000 = x$
- Step 5. Since x is in ml your answer for the new amount is: 2000 ml
- Step 6. The total volume of the new solution is 2000 ml.
- How many ml did you start with (the amount of the first solution)? 1000 ml
- Step 7. Subtract the amount of the first solution from the amount of the new solution to find how much water was added.
- 2000 ml
 $- 1000 \text{ ml}$
 1000 ml
- Step 8. The difference between the volume of the first solution and that of the second solution is the amount of water added. 1000 ml of water added

QUESTIONS

1. If 55 ml of an 18% (W/V) solution are diluted to 330 ml, what will the percentage strength be?

Answer _____

2. If 1 pint of a 1:500 (W/V) solution is diluted to 24 fluid ounces, what will the percentage strength be?

Answer _____

3. How many ml of normal saline solution (0.9% W/V) can be prepared from 250 ml of 25% (W/V) Sodium Chloride solution?

Answer _____

4. How much water should be added to a pint of a 1:2000 (W/V) solution to make a 1:2500 (W/V) solution?

Answer _____

5. How much Lactose and how much Strychnine Sulfate should be used to make 4 drams of a 1:10 trituration of Strychnine Sulfate?

Answer _____

Answer _____

INFORMATION

ALLIGATION

Alligation Alternate is a method by which we may calculate the number of parts of two or more components of a given strength when they are to be mixed to prepare a mixture of desired strength.

Alligation Medial is the method by which the weighted average percentage strength of a mixture of two or more substances, whose quantities and concentrations are known, is calculated.

Calculating Problems Using Alligation Alternate

EXAMPLE: If you wished to make 1000 ml of a 40% solution, using a 10% solution and a 50% solution, how many milliliters of each would be required?

Step 1. What is the percentage strength of the solution you are going to make and its amount?
This is called the WANTED %.

40% 1000 ml

Step 2. What are the strengths of the solutions you are going to use? These are called the HAVES %s.

50% and 10%

Step 3. Draw the following configuration to set up your problem. It is similar to the game of Tic-Tac-Toe.

Step 4. Label each section of the Tic-Tac-Toe configuration as follows: Have, Want, Parts, Amounts. Then place the "WANTED" in the proper location. It always goes in the center box.

HAVE WANT PARTS AMOUNTS

	40		1000 ml

Step 5. Place the largest percentage of the HAVES in the upper LEFT corner; and the lower percentage in the lower LEFT corner.

HAVE WANT PARTS AMOUNTS

50		30	
	40		1000 ml
10		10	

Step 6. What is the difference between 50 and 40, _____; place this number in the bottom section of the PARTS. What is the difference between 40 and 10, _____; place this number in the top section of the PARTS.

NOTE: Since the question asks for the number of milliliters of the 50% and 10%, you must use the parts section of this Tic-Tac-Toe structure. However, in solving any ratio proportion problem you must have three knowns and solve for the fourth. Therefore:

Step 7. In this case we must find the total parts the solution will contain. To do this, all that you do is add the parts you already have and place it into the center.

HAVE	WANT	PART	AMOUNTS
50		30	x ml
	40	40	1000 ml
10		10	x ml

Step 8. Reduce the parts section if possible. In this case it can be reduced by 10.

HAVE	WANT	PART	AMOUNTS
50		3	x ml
	40	4	1000 ml
10		1	x ml

Step 9. Since 4 parts equal 1000 ml (the total), how many ml will each part contain? This can be calculated by setting up a ratio and proportion since you have three sections to the problem.

$$\frac{3 \text{ parts}}{4 \text{ parts}} = \frac{x \text{ ml}}{1000 \text{ ml}}$$

Step 10. Solve by the processes indicated:

$$4x = 3000$$

$$x = 750$$

Step 11. Since x equals the number of ml of the 50% solution, your answer is:

750 ml of 50% solution

Step 12. Solve for the 10% solution the same way:

$$\frac{4 \text{ parts}}{1 \text{ part}} = \frac{1000 \text{ ml}}{x \text{ ml}}$$

$$4x = 1000$$

$$x = 250$$

Step 13. Since x equals the number of ml of the 10% solution your answer:

250 ml of 10% solution

634

QUESTIONS

1. How many grams of Sulfathiazole should be added to 3400 grams of a 10% Sulfathiazole Cream to prepare a cream containing 15% Sulfathiazole?

Answer _____

2. How many grams of Coal Tar should be added to 925 grams of Zinc Oxide Paste to prepare a 6% Coal Tar Ointment?

Answer _____

3. In what proportions should 95% Alcohol be mixed with 30% Alcohol to make 70% Alcohol?

Answer _____

4. In what proportions should solutions of 1.2% (W/V) and 0.38% (W/V) be mixed to make a 0.5% (W/V) solution?

Answer _____

5. How many grams of Petrolatum should be added to 250 grams of 20% Sulfathiazole Ointment to make a 5% Sulfathiazole Ointment?

Answer _____

635

6. How many grams of Coal Tar should be added to 2700 grams of an Ointment Base to prepare a 10% Ointment of Coal Tar?

Answer _____

7. How many grams of Iodochlorhydroxyquin Powder should be added to 2000 grams of a Water Soluble Ointment, to prepare an ointment containing 3% Iodochlorhydroxyquin?

Answer _____

8. How many mls of 70% Alcohol and 30% Alcohol should be used to make 5 gallons of 55% Alcohol?

Answer _____

Answer _____

9. How many grams of Precipitated Sulfur Ointment 20% and Precipitated Sulfur Ointment 5% should be used to make 908 Gms of 8% Sulfur Ointment?

Answer _____

Answer _____

10. How much 95% Alcohol should be mixed with Purified Distilled Water to obtain 7568 mls of 70% Alcohol?

Answer _____

INFORMATION

CALCULATING PROBLEMS USING ALLIGATION MEDIAL

EXAMPLE: If you mixed the following solutions together, what would the percentage strength of the total be?

100 ml of a 50% (W/V)
200 ml of a 10% (W/V)
50 ml of water 0%

Step 1. List all the amounts and their percentages (converted to a decimal).

100 ml 0.50
200 ml 0.10
50 ml 0.00

Step 2. Multiply each of the volumes times its respective percentage strength and place the answer to the right, on the same line.

100 ml x 0.50 = 50.00
200 ml x 0.10 = 20.00
50 ml x 0.00 = 00.00

Step 3. Add the first column up and write the answer under it.

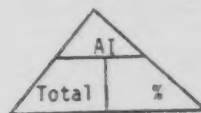
100 ml x 0.50 = 50.00
200 ml x 0.10 = 20.00
50 ml x 0.00 = 00.00

350 ml 70.00 Gm

Now add the third column up and write its answer under it.

Step 4. The 350 ml is the total volume of the mixture and the 70 Grams will be the total amount of Active Ingredient in that total.

Step 5. Remember in percentage preparations you had a triangle, if you had the total amount and the amount of Active Ingredient (AI) you could find the percentage strength.



Step 6. Assign the values to the appropriate terms:

$$X = \frac{70 \text{ Gms}}{350 \text{ Gms (ml to Gm)}}$$

Step 7. Solving for the process indicated

$$X = 0.2$$

Step 8. The answer of 0.2 is a decimal and you are asked for a percentage. To change a decimal to a percent multiply by 100. Therefore, your answer is:

20%

QUESTIONS

1. What is the percentage strength of Alcohol in a mixture of 900 ml of a 40% Alcohol, 500 ml of a 60% Alcohol, 300 ml of a 75% Alcohol and 600 ml of a 35% Alcohol?

Answer _____

2. What would be the percentage strength if you mixed 700 ml of a 55% Alcohol, 330 ml of a 33% Alcohol, 40 ml of a 60% Alcohol and 3000 ml of a 90% Alcohol?

Answer _____

3. What would be the percentage strength if the following Coal Tar Ointments were mixed together?

3% Coal Tar	45 Grams
15% Coal Tar	15 Grams
10% Coal Tar	120 Grams

Answer _____

ADDITIONAL QUESTIONS

REDUCING AND ENLARGING FORMULAS

1. From the following formula calculate the quantities required to make 5 gallons of Ferrous Sulfate Syrup.

Ferrous Sulfate	40 Gm	Answer _____
Citric Acid	2.1 Gm	Answer _____
Peppermint Spirits	2 ml	Answer _____
Sucrose	825 Gm	Answer _____
Purified Water. . QSAD	1000 ml	Answer _____

2. In the above original formula, how many Grams of Ferrous Sulfate will the patient receive in a 2 teaspoonfuls dose?

Answer _____

3. From the below formula for 100 capsules, calculate the quantity of each ingredient required to prepare 35 capsules.

Codeine Phosphate	1.6 Gm	Answer	_____
Phenacetin	4 Gm	Answer	_____
Aspirin	16 Gm	Answer	_____
Atropine Sulfate	0.025 Gm	Answer	_____

4. In the above original formula, how many mg of Codeine and how many mg of Atropine Sulfate will be contained in each capsule?

Answer _____
 Answer _____

5. Reduce the following formula to make 120 ml.

Terpin Hydrate	17	Gm	Answer	_____
Codeine Sulfate	2	Gm	Answer	_____
Orange Peel Tincture	20	ml	Answer	_____
Benzaldehyde	0.35	ml	Answer	_____
Glycerin	400	ml	Answer	_____
Alcohol	430	ml	Answer	_____
Syrup	100	ml	Answer	_____
Distilled water . . QSAD	1000	ml	Answer	_____

SIG: Dram one QID

6. In the above prescription, calculate the prescription dose the patient will receive for the Terpin Hydrate and Codeine.

Answer _____

Answer _____

ADDITIONAL QUESTIONS

SPECIFIC GRAVITY

1. If 250 ml of Alcohol weighs 203 Gm, what is its Specific Gravity?

Answer _____

2. If 125 ml of a liquid weighs 165 Gm, what is its Specific Gravity?

Answer _____

3. If 500 ml of Ferric Chloride solution weighs 650 Gm, what is its Specific Gravity?

Answer _____

4. If a liter of syrup weighs 1313 Gm, what is its Specific Gravity?

Answer _____

5. If Olive Oil has a Specific Gravity of 0.912 and weighs 225 Gm, what is its volume?

Answer _____

6. A product weighs 3 ounces and has a Specific Gravity of 1.25, what is its volume?

Answer _____

7. If a gallon of oil has a Specific Gravity of 0.888, what is its weight?

Answer _____

8. A liquid has a Specific Gravity of 0.91 and a volume of 16 fluid ounces, what is its weight?

Answer _____

9. A liquid has a Specific Gravity of 1.33 and weighs 455 Gm, what is its volume?

Answer _____

10. If 30 ml of a certain liquid weighs 570 grains, what is its Specific Gravity?

Answer _____

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ADDITIONAL QUESTIONS

PERCENTAGE PREPARATIONS

1. How many mg of Cocaine Hydrochloride will be required to prepare 60 ml of a 4% aqueous solution?

Answer _____

2. How many grains of Mercuric Chloride will be required to make 1 gallon of a 1:10,000 solution?

Answer _____

3. How many Gm of Silver Nitrate will be required to make 8 fluid ounces of a 0.5% solution?

Answer _____

4. Calculate the percentage strength (W/V) of a solution, 1 pint of which contains 1 ounce of solute.

Answer _____

5. How would you prepare 8 fluid ounces of a 12% solution of Wintergreen Oil in Alcohol?

Answer _____

5. How many Gm of Potassium Iodide must be added to 8 fluid ounces of water to make a 15% (W/W) solution?

Answer _____

7. How many ml of Wintergreen Oil must be used to prepare 180 ml of a 6% solution in Alcohol?

Answer _____

8. How many mg of Potassium Permanganate are required to make 1 pint of a 1:400 solution?

Answer _____

9. Two pounds of a mixture of Zinc Chloride contain 75 Gm of Zinc Chloride. What is the percentage strength (W/W) of this preparation?

Answer _____

10. How many Gm of a chemical should be added to 200 ml of water to make a 15% (W/W) solution?

Answer _____

ADDITIONAL QUESTIONS

TEMPERATURE CONVERSION

1. Convert -162° F. to Centigrade.

Answer _____

2. Convert 425° F. to Centigrade.

Answer _____

3. Convert 70° F. to Centigrade.

Answer _____

4. Convert 37° C. to Fahrenheit.

Answer _____

5. Convert 200° C. to Fahrenheit.

Answer _____

6. Convert -6° C. to Fahrenheit.

Answer _____

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ADDITIONAL QUESTIONS

CONCENTRATION AND DILUTION

1. If 250 ml of a 1:800 (V/V) solution are diluted to 1000 ml, what will be the ratio strength (V/V)?

Answer _____

2. If 400 ml of a 20% (W/V) solution is diluted to 2 liters, what will be the percentage strength?

Answer _____

3. How many ml of 0.45% (W/V) Sodium Hypochlorite solution can be prepared from 800 ml of an 11.25% (W/V) solution?

Answer _____

4. How many ml of 10% (W/V) solution can be made from 50 ml of an 85% (W/V) solution?

Answer _____

5. How many pounds of 10% (W/W) Sulfuric Acid can be made from 9 pounds of 94% (W/W) Sulfuric Acid?

Answer _____

6. How many ml of Normal Saline solution 0.9% (W/V) can be prepared from 250 ml of 25% (W/V) salt solution?

Answer _____

7. How many ml of a 1:8000 Potassium Permanganate Solution can be prepared from 20 ml of a 1% solution?

Answer _____

8. How many fluid ounces of 6% (W/V) solution can be made from 2 fluid ounces of a 36% (W/V) solution?

Answer _____

9. How many ml of 2.5% (W/V) stock solution of Iodine should be used to prepare 5 liters of a 1:5000 (W/V) solution?

Answer _____

ADDITIONAL QUESTIONS

ALLIGATION

1. How many grams of Sulfathiazole should be added to 4300 grams of a 10% Sulfathiazole Cream to prepare a cream containing 15% Sulfathiazole?

Answer _____

2. How many grams of Coal Tar should be added to 908 Grams of Zinc Oxide Paste to prepare a 9% Coal Tar Ointment?

Answer _____

3. In what proportions should 95% Alcohol be mixed with 50% Alcohol to make 70% Alcohol?

Answer _____

4. In what proportions should solutions of 3.8% (W/V) and 0.12% (W/V) be mixed to make a 0.5% (W/V) solution?

Answer _____

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5. How many Grams of Petrolatum should be added to 340 Grams of 35% Sulfathiazole Ointment to make a 10% Sulfathiazole Ointment?

Answer _____

6. How many Grams of Coal Tar should be added to 2270 Grams of an ointment base to prepare a 4% Coal Tar Ointment?

Answer _____

7. What percentage of Sulfur is contained in a mixture of 10 Grams of 20% Sulfur Ointment, 20 Grams of 10% Sulfur Ointment, and 60 Grams of 5% Sulfur Ointment?

Answer _____

8. What is the percentage of Alcohol in a mixture of 700 ml of 95% Alcohol, 250 ml of 70% Alcohol, 150 ml of 60% Alcohol and 200 ml of 25% Alcohol?

Answer _____

10-8
DEPARTMENT OF BIOMEDICAL SCIENCES

PHARMACY SPECIALIST

PHARMACEUTICAL PREPARATIONS

August 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
SHEPPARD AIR FORCE BASE, TEXAS

Designed For ATC Course Use

DO NOT USE ON THE JOB

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Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas 76311

SW 3ABR90530-III-2
August 1975

PHARMACEUTICAL PREPARATIONS

OBJECTIVES

Upon the completion of the lessons in this workbook, you will have a working knowledge of pharmaceutical compounding and dispensing.

Given information pertaining to pharmaceutical heating, measuring and filtering techniques, complete questions in SW 3ABR90530-III-2.

Given information pertaining to the properties and techniques of preparing pharmaceutical dosage forms, complete questions in SW 3ABR90530-III-2.

EQUIPMENT

SW 3ABR90530-III-2
General Laboratory Equipment
Pen or Pencil
Martin's Dispensing of Medication
Remington's Pharmaceutical Sciences
United States Pharmacopeia
National Formulary

INTRODUCTION

The following list of rules have been compiled to insure that you will be able to perform your assigned projects in the laboratory correctly and safely. It is essential that you understand these rules and follow them completely. Your success and safety is our prime concern. When in doubt as to any procedure, ask an instructor.

INFORMATION

1. Before Coming to Class

a. Each student will complete the section of the workbook that applies to the lecture and laboratory work to be done that day.

b. Each student will observe proper wear of the uniform, proper haircuts, and maintain a high standard of hygiene.

2. During Classroom Lecture

a. Each student will wear low quarter shoes only! It has been observed that any other type of military shoes scuff and mar the floors in the classroom.

b. Each student will take his seat quietly and be prepared for the lecture.

c. Each student will be responsible for all information presented during the lecture, research, and compounding hours.

d. After the lecture, each student will complete the laboratory sheets with the aid of the issued reference books. The preparation sheet will be used for this purpose. The instructor staff will be available to answer questions.

This supersedes WB 3ABR90530-III-2, May 1974

e. Upon completion of the laboratory worksheets, the student must have them pre-checked by an instructor. The worksheets must be rated satisfactory before starting laboratory work. Failure to comply with this responsibility will result in an unsatisfactory on the final preparation.

f. After receiving a precheck, the student will correct any mistakes annotated by the instructor and then start working in the laboratory on the preparation. Any mistakes pointed out on the precheck will be corrected before turning in the prep sheet. Failure to do this will result in an unsatisfactory on the final preparation.

3. During the Pharmacy Laboratory

a. All work must be done individually unless otherwise specified. Make only the quantities requested. When you have completed a preparation, turn it in to your lab instructor. DO NOT, under any circumstances, touch any preparation once it has been placed on the table by an instructor. No student is authorized to touch anything on the preparation table.

b. No talking among students while lab is in session; consult the instructor instead of another student.

c. There will be no horseplay, radios, eating, or drinking in the laboratory at any time. Breaks will be provided every hour and the breaks will be mandatory.

d. All drugs and reagents used in the laboratory are potentially hazardous and extreme caution will be observed in their use. Improper handling and carelessness can result in serious injury and a poor preparation. Make sure that you read the label of all drugs and reagents before using and be familiar with proper techniques for handling them. Use only clean equipment to avoid contamination of the preparation and to prevent any adverse effects. After using the chemical, secure the stopper or lid, clean the bottle, and return it to the proper storage place as soon as possible. Whenever ALCOHOL U.S.P. is used in a preparation, the amount used is to be recorded on an AF Form 582 which will be provided. Accurate records in this area are mandatory.

e. All drugs and chemicals will be labeled at all times. Any material not labeled will be destroyed immediately by the instructor. Under no circumstances will a student take any drug, reagent, chemical, or preparation out of the laboratory unless specifically authorized by the officer in charge of the pharmacy training school or his representative.

f. Each student's work area will be neat at all times. Any material or liquids that are spilled will be cleaned up immediately. All preparation sheets will be in a document protector and will remain in the personnel area at all times. Each student will be assigned to a group, and each group will have specific tasks and general housekeeping duties that must be accomplished on a daily basis. Because the laboratory is subject to both formal and informal inspections, visits, and tours, it is essential that the laboratory classroom be kept clean and neat at all times.

g. Students will be held responsible for excessive breakage of equipment. If any equipment is broken, you will inform an instructor and then will be required to enter your name, date, and type of equipment broken in a special book provided. Any broken, chipped, or cracked equipment will be replaced by the student as soon as possible. Remember, safety depends on reliable equipment, and it is your responsibility to maintain your equipment properly.

h. Read the compounding instructions CAREFULLY!!! Observe all time provisions, temperature indications, order of mixing, and other specifications. Remember that there is a reason behind every step taken in compounding!!! Students will budget their time and strive to develop a technique that promotes efficiency and meets the strict requirements of laboratory safety and pharmaceutical accuracy. Remember, accuracy and neatness will never be sacrificed for speed.

i. When a container of a drug/chemical is completely used, the student will order another container on an AF Form 1517 (supply card) and turn it in to an instructor. The empty container will then be disposed of properly.

j. Each student will turn in the completed preparation sheet along with the completed preparation so that both can be evaluated as satisfactory or unsatisfactory by the laboratory instructor.

k. After a preparation has been evaluated, it will be placed on the display table and you will not handle it again until you are directed by the instructor to dispose of it.

LABELING PROCEDURE

1. AFM 168-4 specifies the labeling requirements of a prescription. Each student will be required to meet these requirements when typing any label in the pharmacy laboratory. The following format will be used:

SCHOOL OF HEALTH CARE SCIENCES SHEPPARD A.F.B., TEXAS	
	
No.	Code # (J-3) Date/Initials
For:	PHARMACY STOCK
GENERIC NAME STRENGTH QUANTITY	
Lot# (Julian date + code #)	
KEEP OUT OF THE REACH OF CHILDREN	

2. Labels will be typed accurately and with a minimum of typing errors. The label will be affixed to the container neatly, and auxiliary labels (if any) will be attached directly under the main label. The auxiliary labels will be attached according to order of importance. An example of this would be a POISON label being placed above a SHAKE WELL label.

3. Remember, the finished appearance of any preparation handed to a patient indicates the attitude, proficiency, and integrity of the pharmacy personnel. Just like an artist signs his name to a picture or work of art, the prescription is initialed by the pharmacy specialist, and the label is initialed by the typist. The last line in the hospital chain, from physician to the pharmacy, is YOU, the pharmacy specialist.

EVALUATION PROCEDURES

Each preparation made in the laboratory will be evaluated by the instructor staff. The evaluation procedure is not difficult, but to preclude any misunderstanding, the following is a copy of the checklist used by the instructor staff to evaluate your preparations.

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GENERAL PURPOSE CHECKLIST		Page 1 of 6 Pages		
TITLE/SUBJECT/ACTIVITY Dispensing Pharmacy Checklist, SICS, MSDB, Pharmacy		DATE		
ITEM (Assign a paragraph number to each item. Draw a horizontal line between each major paragraph.)		Yes	No	
Code No. _____ Date _____ Class No. _____ Evaluator _____ <u>Instructions</u> * Check mark () the complied (yes or no) blocks. * A "No" response requires a brief explanation. 1. AF Form 2380-Pharmacy Manufacture Control Data a. Properly filled out at home and presented at work showing: (1) Ingredients, amounts, stock solutions, all math calculations and initialed by instructor <u>prior</u> to starting compounding. b. Annotated while compounding to show: (1) Product lot number (julian date and code number), manufacturer, manufacturer's lot number, compounder's initials and use of schedule drugs. c. Reviewed for accuracy, signed and dated by compounder.				

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GENERAL PURPOSE CHECKLIST		Page 2 of 6 Pages	
TITLE/SUBJECT/ACTIVITY Dispensing Pharmacy Checklist, SHCS, MSDB, Pharmacy		DATE	
ITEM (Assign a paragraph number to each item. Draw a horizontal line between each major paragraph.)	Yes	No	
2. AF Form 2381-Pharmacy Master Formula a. Factor amounts in directions b. Read directions thoroughly c. Follow directions <u>EXACTLY</u>			
3. AF Form 2382-Pharmacy Bulk Compounding Chronological Control Log a. List preparation lot number b. Include name of the preparation c. Include amount of the preparation NOTE: THIS IS TO BE DONE WHEN PREPARATION IS TURNED IN.			

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GENERAL PURPOSE CHECKLIST		Page 3 of 6 Pages	
TITLE/SUBJECT/ACTIVITY		DATE	
Dispensing Pharmacy Checklist, SHCS, MSDR, Pharmacy			
ITEM	Yes	No	
(Assign a paragraph number to each item. Draw a horizontal line between each major paragraph.)			
4. Label			
a. Follow format as outlined by display poster and workbook. The following will be included:			
(1) Facility and location			
(2) Date/initials			
(3) Code number			
(4) Pharmacy stock			
(5) Generic name, strength and amount			
(6) Include N.F., U.S.P., or local formula			
(7) Preparation lot number			
(8) Keep out of the reach of children			
(9) Auxiliary labels			
(10) Insure that typing area is left neat			
(11) Insure that label does not contain strikovers, is straight and is proportionately placed on container			

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GENERAL PURPOSE CHECKLIST		Page 4 of 6 Pages	
TITLE/SUBJECT/ACTIVITY Dispensing Pharmacy Checklist, SHCS, MSDB, Pharmacy		DATE	
ITEM (Assign a paragraph number to each item. Draw a horizontal line between each major paragraph.)	Yes	No	
5. Preparing prescription for controlled drugs a. The prescription must include the following: (1) Date (2) Prescription number (3) Pharmacy stock (4) Item and amount (5) Name and amount of preparation being compounded (6) Used by (7) Time and initials (8) Instructor's signature NOTE: THIS MUST BE DONE PRIOR TO WITHDRAWING ALCOHOL			

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GENERAL PURPOSE CHECKLIST		Page 5 of 6 Pages	
TITLE/SUBJECT/ACTIVITY		DATE	
Dispensing Pharmacy Checklist, SHCS, MSDS, Pharmacy			
ITEM (Assign a paragraph number to each item. Draw a horizontal line between each major paragraph.)		Yes	No
6. AF Form 582-Pharmacy Stock Record Card a. The Pharmacy Stock Record Card should include: (1) Date (2) Prescription number (3) Physician's name (4) Amount dispensed (5) Balance (6) Student's initials NOTE: PRESCRIPTION MUST BE POSTED AS SOON AS THE ALCOHOL IS DRAWN			

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GENERAL PURPOSE CHECKLIST		Page 6 of 6 Pages	
TITLE/SUBJECT/ACTIVITY		DATE	
Dispensing Pharmacy Checklist, SHCS, MSDB, Pharmacy			
ITEM	Yes	No	
(Assign a paragraph number to each item. Draw a horizontal line between each major paragraph.)			
7. Techniques			
a. Balance			
(1) Put into equilibrium			
(2) Weights on right-hand pan			
(3) Balance locked when adding or subtracting weights			
(4) Rider-arm and weights returned when not in use			
(5) Cleaned with Isopropyl alcohol			
(6) May be "zero-ed in" at beginning of day and left. However, if balance moved it must be recalibrated.			
b. Uses correct pouring technique			
c. Returns clean stock containers to side shelves			
d. Properly maintain glassware and area while compounding			
e. Labels all drugs and chemicals not in stock container			
f. Completed in allotted time			

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INTRODUCTION TO PHARMACEUTICAL DISPENSING

1. EVAPORATING DISH - Used for sand baths, fusion, and incineration.
2. CASSEROLE DISH - Used for sand baths, fusion, and incineration.
3. MEDICINE DROPPER - Used to measure phenols, acids, and other caustics or when measurement of liquids is too difficult to use other glassware.
4. PILL TILE (OINTMENT SLAB) - Used to make ointments, pastes, and creams. Provides a smooth surface for levigating or spatulating and can be cleaned easily.
5. STAINLESS STEEL SPATULA - Used for levigation and spatulation of ointments, pastes, and creams.
6. RUBBER OR PLASTIC SPATULA - Used for levigation and spatulation of ointments, pastes, and creams when one or more of the ingredients are mercury, iodine, tannic acid, or heavy metals and their salts.
7. LABORATORY BEAKER - Used to mix and heat liquids. Graduations on this glassware are only approximate, and it will never be used for measuring.
8. ERLLENMEYER FLASK - Used to mix, heat, and macerate. Graduations on this glassware are only approximate, and it will never be used for measuring.
9. CYLINDRICAL GRADUATE - Used for measuring liquids. The most accurate of all the graduates due to the "flat" meniscus obtained.
10. MORTAR - Used to mix, triturate, levigate, and comminute in many cases.
11. PESTLE - Used with the mortar to mix, triturate, levigate, and comminute in many cases.
12. STIRRING ROD - Used to stir or mix liquids.
13. CONICAL GRADUATE - Used for measuring liquids. This glassware is calibrated in both apothecaries and metric systems.
14. RING STAND AND BASE - Used for heating and filtering procedures.
15. ASBESTOS PAD (WIRE GAUZE) - Used to provide an even distribution of heat when heating various types of glassware.

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16. FISHER BURNER - A gas operated burner used for heating procedures.
17. GLASS FUNNEL - Used for filtering. Do not use for pouring.
18. SUPPOSITORY MOLD - Used to make suppositories by fusion-molding method.
19. PIPETTE - Used for measuring small amounts of liquids. Never to be used when measuring phenol, caustic agents or poisons.

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SUMMARY OF TABLES OF WEIGHTS AND MEASURES

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1. Metric System

Unit of Length: Meter (M)
Unit of Volume: Liter (L)
Unit of Weight: Gram (Gm. or g.)

These units may be preceded by these prefixes

Kilo - 1000	deci - 0.1	micro - 0.000001
Hecto - 100	centi - 0.01	
Deka - 10	milli - 0.001	

2. Avoirdupois Weight

437.5 grains (gr.)	1 ounce (oz.) (oz. av.)
16 ounces (oz.) (av. oz.)	1 pound (lb.) (#)
2000 pounds	1 hundred weight (C.W.T.)
	1 ton (T)

3. Apothecary Weight

20 grains (gr.)	1 scruple ()
3 scruples ()	1 dram ()
60 grains (gr.)	1 dram ()
8 drams ()	1 apothecary ounce () (oz. apoth)
12 apoth. ounces ()	1 apoth. pound ()

4. U. S. Fluid ("Wine") and Apothecary Fluid Measure

60 minims (min.) ()	1 fluidram (f.)
8 fluidrams (fl.)	1 fluidounce (f.)
16 fluidounces (fl.)	1 pint (pt.) (O.)
2 pints (pt.)	1 quart (qt.)
4 quarts (qt.)	1 gallon (gal.) (C)
f. or fl. frequently omitted and "understood"	

Conversion Equivalents

1 ml.	1 cc.
1 inch	2.54 cm.
1 M.	39.37 in.
1 gal.	231 cu. in.
1 gal.	3785 ml.
1 qt.	946 ml.
1 pt.	473 ml.
1 gr. av.	1 gr. apoth.
1 fluidounce	29.57 ml.
1 Kg.	2.2 lbs (av.)
1 Gm.	15.432 grains
1 grain 64.8 mg.	64.8 mg.
1 ml.	16.23 minims.
1 min. of water	0.95 grains

Conversion Equivalents (cont.)

1 apoth. ounce	480 grains
1 apoth. fluidounce of water	454.6 gr.
1 apoth. ounce	31.1 Gm.
1 av. ounce	28.35 Gm.
1 pound (av.)	7000 grains
1 pound (av.)	453.6 Gm.
1 pound (apoth.)	5760 grains
1 gal. water	8.32 lbs.
1 gr. of water	1.05 minims
1 gallon (C., Cong., gal.)	128 fluidounces
1 L.	2.113 pints
1 gamma	1 microgram
1 micron ()	0.001 mm.
1 ml. of water at 4°C	1 Gm.
1 cu. ft. water	62.5 lbs. av.
1 drop (gtt) has no definite weight or size.	

5. Household Measure*

1 teaspoonful (tspf.) (℥)	5 ml.	1/6 f.
1 dessertspoonful (ii)	10 ml.	1/3 f.
1 tablespoonful (tbsp.) (ss)	15 ml.	1/2 F
1 wineglassful	60 ml.	F. ii
1 teacupful	120 ml.	f. iv
1 tumblerful	240 ml.	f. viii

*Highly inaccurate. The (dram) sign is normally read Teaspoonful when it appears in the "sig."

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INTRODUCTION TO PHARMACEUTICAL DISPENSING

QUESTIONS

1. A beaker is used for mixing, heating or stirring. Never used for _____ liquids.
2. An _____ is used for mixing, macerating or heating of liquids.
3. Conical and Cylindrical _____ are used to measure liquids.
4. The mortar and pestle is used for mixing, grinding, and various forms of _____.
5. The _____ is used for measuring small amounts of liquids. It is never to be used when measuring phenol, caustic agents, or poisons.
6. Wire gauze (asbestos pad) is used for providing an even distribution of _____ on the bottom of containers.
7. The process of strongly heating solid or semi-solid substances to a definite and limited degree (the residue of this is the product sought) is called _____.
8. The process of the removal of water of crystallization, or moisture, from a solid crystalline substance by heating strongly is called _____.
9. The process of roasting certain organic substances in order to modify some of their constituents is called _____.

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10. The process of liquifying solid substances by the application of heat without the use of a solvent is called _____.
11. Converting a liquid or solid into a vapor is called _____.
12. Driving off as a vapor, volatile portion of a liquid by the application of heat is called _____.
13. Separation of the constituents of a liquid mixture by vaporization and subsequent condensation of the vapors is called _____.
14. Separation of volatile solids from non-volatile solids is called _____ and the product obtained is the sublimate.
15. A Cool Place is _____ C or _____ F.
16. A Cold Place does not exceed _____ C or _____ F.
17. Refrigerated means the temperature is _____ C or _____ F.
18. Excessive heat is any temperature that exceeds _____ C or _____ F.
19. The primary concern when using pharmaceutical heating devices is _____.
20. _____ are used to check excessive heat and not to injure certain medicinal products.

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21. Match the following:

For temperatures not to exceed 100°C (212°F) _____

a. Oil Bath

Maximum temperature is 250°C _____

b. Water Bath

Used to maintain a CONSTANT temperature of 100°C (212°F) _____

c. Steam Bath

Used for moderate to extremely high temperatures _____

d. Sand Bath

Same as heat source _____

Decomposes under high temperatures _____

METROLOGY

QUESTIONS

1. The Troemner and Torsion Balances are Class _____ Prescription Balances.
2. The balance that works on the knife-edge principle is a _____ Balance.
3. The maximum amount weighable on a Class A Prescription Balance is _____ Gm.
4. The minimum amount weighable on a Class A Prescription Balance is _____ mg.
5. Rider Scale on the Torsion and Troemner Balances is divided into _____ mg increments and the maximum amount weighable using the rider weight is _____ Gm.
6. The Harvard Trip (Laboratory Balance) is a Class _____ Prescription Balance.
7. The maximum amount weighable on a Class B Prescription Balance is _____ Gm. (_____ lb.)
8. The minimum amount weighable on a Class B Prescription Balance is _____ mg. (_____ gr.)
9. The Rider Scale on the Harvard Trip (Laboratory Balance) is divided into _____ mg increments and the maximum amount weighable using the rider (poise) weight is _____ Gm.
10. Because of rapid evaporation, the solvent of choice for cleaning balances is _____.

11. Weights are put on the _____ pan of the balance and the drug to be weighed is put on the _____ pan.

12. Match the following:

Used to measure small amounts of liquids _____

a. 30 ml Conical Graduate

Phenol not measured in this _____

b. 10 ml Cylindrical Graduate

Never used to measure amounts less than 5 ml _____

c. Pipette

Never used to measure amounts less than 2 ml _____

d. Medicine Dropper

Most accurate device to measure large amounts of liquids _____

e. Beaker

Never used for measuring liquids _____

f. Graduate

13. Cylindrical Graduates are more accurate than Conical Graduates.
(TRUE or FALSE)

14. Liquids that cannot be used in a pipette because of their toxic or caustic properties or too small to be measured in a pipette will be measured in a _____.

15. "TD" means "To _____".

16. "TC" means "To _____".

17. A Cylindrical Graduate has a "flatter" meniscus than a Conical Graduate and as a result, the chances of an ERROR of _____ is reduced.

COMMINUTION

QUESTIONS

1. The process of physically reducing solid substances into smaller fragments or particles is called _____.
2. The purpose of Comminution is to _____ the rate of solution of solids, obtain a uniform powder or mixtures of powders, and to increase ease and thoroughness of _____.
3. Water sifting is called _____.
4. Hand picking or sorting is called _____.
5. Sieves or screens are used when _____.
6. The process of placing a substance in a heavy mortar and crushing it by pounding with a heavy pestle is called _____.
7. Rasping, grating, cutting, slicing, and chopping are methods of preparing _____ drugs.
8. The process of reducing substances to a powder by rubbing them in a mortar with a pestle is called _____.
9. Reducing the particle size of a solid by first forming a mass of the solid with a liquid and then grinding in a Mortar and Pestle or by spatulation is called _____.
10. The process of reducing a solid to a powder through the use of a foreign substance from which the powder is freed by some simple method is called _____.
11. Comminution can never increase solubility but can increase the _____ of _____.

INCOMPATIBILITIES

QUESTIONS

1. When drugs or chemicals are not capable of acting in harmony or concert with another, they are said to be _____.
2. 3 types or classifications of incompatibilities are:
 - a. _____
 - b. _____
 - c. _____
3. _____ incompatibilities are evidenced by:
 - a. Clouding
 - b. Precipitation
 - c. Liquefaction
 - d. Change in consistency
 - e. Immiscibility
 - f. Insolubility
4. A change in the physical state of a preparation resulting from mixing two or more drugs together is called a _____ incompatibility.
5. The result of a reaction taking place when two or more ingredients of a prescription are mixed and forming a new compound or compounds is called a _____ incompatibility.
6. A condition in a prescription that results in a dosage different from that intended by the prescriber is called a _____ incompatibility.

7. _____ incompatibilities are evidenced by:
- a. Explosion or Implosion
 - b. Liberation of a gas
 - c. Change of color
 - d. Formation of a precipitate
8. "SA" is translated as "_____ to the _____".
9. _____ incompatibilities are evidenced by:
- a. An overdose or improper dose of a single drug
 - b. An undesirable combination of two or more drugs
 - c. A contraindicated drug
 - d. The wrong drug
10. The 2 methods of correcting incompatibilities are:
- a. According to the Art
 - b. _____ the prescriber
11. The prescriber is always contacted in the case of _____ incompatibilities.
12. When correcting physical incompatibilities, care should be taken to insure that the _____ effect is not altered.

GENERAL METHODS OF CORRECTING INCOMPATIBILITIES

1. On any question of therapeutics, consult the prescriber.
2. Modify order of mixing.
3. Alter solvents.
4. Change form of one or more ingredients.
5. Alter volume.
6. Emulsify or suspend.
7. Add or omit therapeutically inactive substances.
8. Change dosage form.
9. Replace one of reacting ingredients.
10. Control reaction during mixing.
11. Add color inhibiting or masking agent.

Type of incompatibility	Possible methods of correction
-------------------------	--------------------------------

Therapeutic:

Overdose	1
Additive combinations	1
Antagonistic combinations.	1
Wrong drug	1
Contraindicated medication	1

Physical:

Immiscibility.	2,3,4,5,6,7,8
Insolubility	2,3,4,5,6,8
Precipitation.	3,5,6,8,9
Liquefaction of solids	7,8,9
Solidification of solids	5,7,8,9

Chemical:

Precipitation.	3,4,6,7,8,9
Effervescence.	8,9,10
Decomposition.	3,4,7,8,9
Color formation.	7,8,9,11
Explosion.	2,8,9

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- Sig: i dr q.i.d.

Directions for correcting:

- M. Ft. oint.

Directions for correcting:

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674

3. Aminophylline gr iss
Sugar of Milk gr iii

Per capsule Make 24

Incompatibility(ies):

Directions for correcting:

4. Ferric Chloride 40
Water qs ad 120

M. Ft. sol.

Incompatibility(ies):

Directions for compounding:

WATERS AND SPIRITS

QUESTIONS

1. _____ USP is not suitable for pharmaceutical work because of the considerable amount of dissolved solids present.
2. _____ water USP does not contain the dissolved mineral matter that Water USP contains and is the water used for pharmaceutical compounding unless otherwise specified.
3. Water for _____ USP is pyrogen free, used in making parenterals and is not sterile.
4. _____ Water for Injection USP is used for preparing parenterals and is sterile, but does not have an antimicrobial agent in it.
5. _____ Water for Injection USP is Sterile Water for Injection USP with 1 or more antimicrobial agents added to it.
6. Solutions of volatile oils or other aromatic or volatile substances in purified water are called _____ Waters.
7. Unless otherwise specified, Aromatic Waters are saturated and their percentage strength is _____.
8. The 3 methods of preparing Aromatic Waters are:
 - a. _____
 - b. Simple _____
 - c. _____ Solution

9. When preparing Aromatic Waters by the Simple Solution method, the volatile substance and water must set for _____ hours.
10. The Alternate Solution method requires only 10 minutes of agitation because of the addition of _____ that acts as a dispersing agent.
11. The use of Talc in preparing Aromatic Waters speeds up the saturation by dispersing the aromatic substance and also acts as a _____ bed.
12. Alcoholic or hydroalcoholic solutions of volatile substances are called _____.
13. The 4 methods of preparing Spirits are:
- _____ Solution
 - Solution with _____
 - _____ Reaction
 - _____
14. When water is added to a Spirit, _____ occurs.
15. Match the Following:
- | | |
|--|----------------------------|
| Carminative, anesthetic, and antiseptic in eye preparations. Pharmaceutical Solvent. _____ | a. Aromatic Ammonia Spirit |
| Flavored Vehicle _____ | b. Cinnamon Water |
| Carminative and Flavored Vehicle _____ | c. Camphor Water |
| Reflex Stimulant _____ | d. Compound Orange Spirit |
| Local irritant _____ | e. Peppermint Water |
| Flavoring Agent _____ | f. Camphor Spirit |

SOLUTIONS AND SYRUPS

QUESTIONS

1. Aqueous solutions of nonvolatile substances are called _____.
2. Because solutions are used for whatever the therapeutic effect of the substances dissolved, no general use can be stated. (TRUE/FALSE)
3. When preparing solutions, the solute is dissolved into the solvent without the aid of any catalyst and this method is called _____ Solution.
4. Solution by _____ is a method of preparing solutions in which the solutes react to form the solution.
5. The method of sterilizing a simple solution is called Simple Solution with _____.
6. The method of macerating the solute to obtain the active ingredient or constituent in solution is called Solution by _____.
7. 4 factors affecting the _____ are:
 - a. Particle size
 - b. Agitation
 - c. Heat
 - d. Degree of Saturation
8. List the degrees of saturation:
1 Gm of solute in:
Less than 1 ml _____
1-10 ml _____

- 10-30 ml _____
- 30-100 ml _____
- 100-1000 ml _____
- 1000-10,000 ml _____
- 10,000 ml or more _____

- 9. Supersaturation _____ the rate of solution.
- 10. Hyposaturation _____ the rate of solution because the percentage strength is lowered.
- 11. When Supersaturation occurs by evaporation, the percentage strength is _____.
- 12. In addition to evaporation, supersaturation may occur by _____.
- 13. Match the Following:

- Anti-Infective _____
- No Dose _____
- Source of Iodine _____
- Expectorant _____
- 5% _____
- Topical _____
- 1% _____
- 0.3 ml 3 times daily _____
- 100% (w/v) _____
- Local Formula _____
- 0.3 ml _____

- a. Amaranth Solution USP
- b. Saturated Solution of Potassium Iodide NF
- c. Gentian Violet Solution
- d. Potassium Permanganate Solution
- e. Strong Iodine Solution USP

Coloring Agent _____

Antifungal _____

SSKI _____

Lugol's Solution _____

FD&C Red No. 2 _____

14. A nearly saturated aqueous solution of sugar with or without a medicinal or flavoring agent is called a _____.
15. _____ Syrups contain a medicinal ingredient or ingredients designed for a therapeutic effect on the body or system.
16. Non-Medicated Syrups are used just as a sweetner or _____.
17. The 2 methods of preparing syrups are as follows:
- a. Solution _____ Heat
 - b. Solution _____ Heat.
18. Overheating of Syrup USP will result in _____.
19. Syrups should be stored in a cool place or _____ if possible.
20. Syrups USP will not need a preservative if the concentration of Sucrose is _____%.
21. Match the Following:
- | | |
|--------------------------------------|----------------------------|
| Vehicle, Sweetening Agent _____ | a. Simple Syrup USP |
| Sympathomimetic, Antiasthmatic _____ | b. Ephedrine Sulfate Syrup |
| 20 mg/5ml _____ | |
| 85% _____ | |
| No Dose _____ | |
| 5 ml 4 times daily _____ | |

EYE, EAR, AND NOSE PREPARATIONS

QUESTIONS

1. When a concentrated solution is separated from a less concentrated solution by a semipermeable membrane, the solvent passes through the membrane to the more concentrated solution and the force which brings this about is called _____ pressure.
2. A substance is said to be _____ if it has the same osmotic pressure as body fluids (the same number of particles in solution as body fluids).
3. A substance is said to be _____ if it has a lower osmotic pressure than body fluids (a lesser concentration of particles than the body fluids).
4. A substance is said to be _____ if it has a higher osmotic pressure than body fluids (a greater concentration of particles than body fluids).
5. _____ is the negative common logarithm of the hydrogen ion concentration.
6. A pH of 7-14 is _____.
7. _____ are substances that resist a change in pH of a preparation.
8. Group I (Boric Acid) is used to buffer anesthetics and _____.
9. Group II (Modified Sorenson's Solution) is used to buffer _____ and similar drugs.

10. Ophthalmic preparations must be _____, have the correct pH, tonicity, viscosity, free from foreign particles, and sterile.
11. The vehicles for ophthalmic preparations are either _____ aqueous solutions or _____ ointment bases.
12. To prolong a drug's contact with the eye you would increase the _____.
13. No single preservative is sufficiently free from incompatibilities to be used in all cases but they should be _____.
14. Benzalkonium Cl, Chlorobutanol, Methyl and Propyl Paraben, and Thimersol are used as ophthalmic _____.
15. The 5 methods of applying ophthalmic preparations are:
- a. _____
 - b. _____
 - c. _____
 - d. _____
 - e. _____
16. Match the following:
- | | |
|------------------------|-----------------------|
| astringent _____ | a. Zinc Sulfate |
| mydriatic _____ | b. Tetracaine |
| anesthetic _____ | c. Fluorescein Sodium |
| miotic _____ | d. Atropine Sulfate |
| diagnostic agent _____ | e. Pilocarpine |
17. Otic preparations should be _____, non-allergenic and nonsensitizing, have the correct pH, and be sterile.

18. Vehicles used for otic preparations are:
 - a. Glycerin
 - b. Propylene Glycol
 - c. Ethyl Alcohol
 - d. Vegetable Oils
 - e. Acidic _____ Solutions
19. To prevent bacterial growth, aqueous preparations in the ear should be _____.
20. The 2 methods of applying or instilling ear preparations are:
 - a. _____
 - b. _____
21. Benzocaine, Antipyrine, Glycerin, and Menthol ear drops is classified as an Anesthetic, _____.
22. Acetic Acid 2% in Ethyl Alcohol ear drops is classified as a _____.
23. Nose preparations should be stable, non-allergenic and nonsensitizing, have the correct _____, and correct pH.
24. The pH range of nose preparations is _____ to _____.
25. The most commonly used vehicle in nasal preparations is _____.
26. Mineral oil and other petroleum distillates were used as vehicles in the past but it was found to cause oil-aspiration _____.
27. The 2 methods of applying or instilling nose preparations are:
 - a. _____
 - b. _____

8. Ephedrine Sulfate Solution is used as a nasal _____.
9. Phenylephrine HCL Solution (Neosynephrine) is used as a nasal _____.
30. Commonly used preservatives in nasal preparations are:
- a. _____
- b. _____

ELIXIRS AND TINCTURES

QUESTIONS

1. A clear, sweetened, hydro-alcoholic liquid, medicated or non-medicated, intended for oral use is the definition of an _____.
2. Elixirs are colored to match taste, pleasant tasting, sweet, used therapeutically or as a vehicle and have an alcohol range of _____ to _____%.
3. There is no set procedure for preparing Elixirs but the _____ soluble ingredients are dissolved in the water portion and the _____ soluble ingredients are dissolved in the alcohol.
4. Raising the _____ content will cause the water soluble ingredients to precipitate or come out of solution.
5. Lowering the _____ content will cause the alcohol soluble ingredients to precipitate or come out of solution.
6. Match the Following:

Expectorant _____	a. Terpin Hydrate Elixir NF
Antihistamine, Pediatric Sedative _____	b. Aromatic Elixir USP (Simple Elixir)
Anticonvulsant, Sedative, Hypnotic _____	c. Phenobarbital Elixir (Luminal Elixir)
Flavored Vehicle _____	d. Diphenhydramine HCl Elixir USP (Benadryl Elixir)
20 mg/5 ml _____	
85 mg/5 _____	
21-23% Ethyl Alcohol _____	
12.5 mg/5 ml _____	

30 mg 4 times daily _____

No dose _____

10 ml 4 times daily _____

5 ml as needed _____

7. Alcoholic or hydroalcoholic solutions prepared from vegetable drugs or chemical sources are called _____.

8. _____ Tinctures do not exceed 10% of the active drug,

9. _____ Tinctures do not exceed 20% of the active drug.

10. _____ Tinctures do not exceed 50% of the active drug.

11. Tinctures are prepared by _____ (Process M).

12. Match the Following:

2% _____

a. Iodine Tincture USP

Pharmaceutical Necessity _____

b. Sweet Orange Peel Tincture

Flavoring Agent _____

No Usual Dose _____

Topical Use

50% _____

Germicide _____

SUSPENSIONS

QUESTIONS

1. Two-phase liquid preparations containing solid insoluble material for oral, injection, or ophthalmic use are called _____.
2. The dispersed phase in suspensions should settle slowly and should be easily _____.
3. Suspensions should not cake upon _____.
4. Suspensions should pour easily and be _____.
5. Suspensions should have good patient _____ in taste and color qualities.
6. List the 3 types of suspensions:
 - a. _____
 - b. _____
 - c. _____
7. Nonviscous aqueous preparation containing insoluble material intended for internal use are called _____.
8. Viscous aqueous preparations, containing insoluble material intended for internal use are called _____.
9. Aqueous liquid preparations containing insoluble material of nearly colloidal size intended for internal use are called _____.
10. Magmas are prepared by _____ or chemical reaction.
11. There is no set method for preparing _____.

12. The presence of a suspending agent is required to overcome agglomeration of the dispersed particles and to increase the _____ of the medium so that particles settle slowly.
13. Many suspending agents are also _____ agents.
14. Acacia USP is a suspending agent used for insoluble substances in water and is susceptible to microbial attack and, therefore, needs a _____.
15. Bentonite USP is a protective _____ used for the stabilization of suspensions.
16. _____ is used as a dispersing, thickening, emulsifying, and coating agent. It is not susceptible to microbial growth like the natural gums such as Acacia and Tragacanth but should still have a preservative added.
17. Tragacanth USP hydrates very slowly and is only partially soluble in water. It is susceptible to microbial attack and must have a _____ added.
18. List the preservatives used in suspensions:
- a. _____ 7%
 - b. _____ 0.2%
 - c. _____ 0.2%
 - d. _____
19. The auxiliary label required for Mixtures, Magmas, Suspensions, and Gels is "_____".
20. Magmas and Gels must be kept from _____.

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21. Match the Following:

Antibacterial _____

Antacid _____

Suspending Agent _____

Expectorant _____

3-4 Gm Stat then

1-2 Gm q4-6 hours _____

No Dose _____

15 ml _____

5 ml _____

a. Brown Mixture NF

b. Bentonite Magma

c. Aluminum Hydroxide Gel

d. Acetylsulisoazole Suspension

e. Chalk Mixture NF

LOTIONS AND LINIMENTS

QUESTIONS

1. Aqueous or hydroalcoholic preparations that contain insoluble solids held in suspension and that are intended for external use by application to the skin without friction are called _____.
2. Lotions are filtered before dispensing. (TRUE/FALSE)
3. The 2 methods of preparing lotions are Trituration, and _____.
4. No suspending agent is needed when preparing a lotion by the precipitation method since the precipitate is nearly _____ size.
5. Commonly used preservatives used in Lotions are:
 - a. _____ 7%
 - b. _____ 0.2%
 - c. _____ 0.2%
 - d. Methyl and _____ parabens
6. Oily or alcoholic preparations intended for external use to be applied with friction and are liquid or semi-solid in form are called _____.
7. An external preparation having the poorest patient acceptance is a _____ because:
 - a. Application must be-repeated over a period of time.
 - b. The patient is required to expend effort on application.
 - c. The preparation may leave a film on skin or stain clothing.
 - d. Removal may be difficult.
 - e. The appearance and odor may not be pleasing.

8. There is no set method of preparing a liniment. Therefore, the method of preparation depends on the _____.
9. Because of a slight effervescent action, _____ should be allowed to stand for a few minutes before capping.
10. Match the Following:
- | | |
|--------------------------------|------------------------------------|
| a. Protectant _____ | 1. Medicinal Soft Soap Liniment NF |
| b. Detergent _____ | 2. White Lotion USP |
| c. Astringent- _____ | 3. Calamine Lotion USP |
| d. Counterirritant _____ | 4. Camphor Liniment NF |
| e. Stand before capping _____ | |
| f. Alcoholic base | |
| g. Oily base | |
| h. Made by precipitation _____ | |
| i. Oily base _____ | |
11. The purpose of levigating powders when making a lotion is to reduce _____ size.
12. Liniments with oily or alcoholic bases are usually prepared by the _____ Solution method.

POWDERS AND CAPSULES

QUESTIONS

1. Mixtures of drugs or chemicals in a dry, pulverized form intended for internal or external use are called _____.
2. The 2 types of powders are Bulk powders and _____ powders.
3. Effervescent powders, dusting powders, dentifrices, and insufflations are _____ powders.
4. Divided powders are measured and packaged by the _____.
5. Bulk powders are measured by the _____.
6. Chartula is another name for _____ powders.
7. The _____ activity of a medication is affected by the degree of fineness of the powder.
8. In all prescriptions, the powders should be in a fine state of _____.
9. Trituration is the method of choice for _____ powders and also the most common method used for mixing powders.
10. Spatulation can be used for mixing _____ quantities of powders.
11. _____ is the method of choice for light powders.
12. Tumbling is the mixing method used to advantage where _____ on the powder is undesirable.
13. Powders that are blown into body cavities are called _____.

14. Powder papers containing deliquescent, efflorescent, hygroscopic, or volatile substances should be _____ wrapped to protect them from the atmosphere.
15. _____ powders absorb enough moisture from the atmosphere to become a liquid.
16. Efflorescent powders lose their waters of _____ to the atmosphere.
17. _____ powders absorb enough moisture from the atmosphere to become moist.
18. Salts which contain a large amount of water of crystallization may need to be _____ before mixing to prevent reactions.
19. Phenols, phenolic compounds, aldehydes, and ketonic compounds form _____ mixtures.
20. Eutectic mixtures can be corrected by the addition of Magnesium Carbonate or Light _____ Oxide.
21. _____ are shells of gelatin used for containing individual doses of medication.
22. Hard Gelatin capsules are used primarily for _____ compounding.
23. _____ Gelatin capsules usually have a liquid in them.
24. _____ Coated capsules are intended to dissolve in the intestine rather than the stomach.
25. The lactose or diluent is added to the active ingredient by _____ dilution.

26. Capsules should be stored in a cool place of _____
humidity.
27. A warm dry place will cause capsules to lose water and become
_____.
28. Match the Following:
- | | |
|---------------------|--|
| Antispasmodic _____ | a. Phenobarbital and Belladonna Capsules |
| Analgesic _____ | b. Aspirin and Codeine Capsules |
29. After punching capsules, fingerprints may be removed by polishing them
in _____.

EMULSIONS

QUESTIONS

1. Preparations containing two _____ liquids, one of which is uniformly dispersed as globules within the other are called _____.
2. Emulsions are classified as either _____ in _____ or oil in water.
3. If oil in the internal phase and water in the external phase, the emulsion is a _____ in water emulsion.
4. Acacia (Gum Arabic) is a natural gum and a true emulsifying agent. The emulsions produced by acacia are rather _____.
5. When preparing oil-in-water emulsions, _____ is the most efficient emulsifying agent but it is the least stable and prone to rapid decomposition.
6. Special Gelatin Pharmagel A is used in an _____ pH.
7. Special Gelatin Pharmagel B is used in a _____ pH.
8. Tragacanth _____ is seldom used alone as an emulsifying agent but in combination with other emulsifying agents it increases the viscosity of the preparation to such an extent that it is _____ times more powerful than Acacia.
9. List the Miscellaneous Emulsifying Agents:
 _____: Soap, sulfonates, and sulfates
 _____: Benzalkonium Cl

- _____ : Spans and Tweens
- _____ : Cholesterol and Wool Fat
- _____ : Bentonite, Silica, Magnesium Hydroxide,
and other fine powders
10. To prevent spoilage of carbohydrate emulsifying agents, they should be kept in a TLRC, preserved and _____.
11. The Dry Gum (Continental Method) is used to prepare emulsions of fixed oils, _____ oils, or nonvolatile oils.
12. The ratio of the Dry Gum method is 4 parts _____, 2 parts _____, 1 part emulsifying agent.
13. The Dry Gum method requires you to add oil to the emulsifying agent and add the _____ ALL AT ONCE.
14. When making an emulsion, the primary emulsion is "rested" so that the emulsifying agent can _____.
15. The Wet Gum (English Method) is used to prepare emulsions of fixed oils, _____ oils, or nonvolatile oils.
16. The Wet Gum method is more difficult than the Dry Gum method in that the oil is added after the _____.
17. The Bottle Method (Forbes Method) is used to prepare emulsions of _____ oils, or nonviscous oils.
18. The ratio of the Bottles method is 2 parts _____, 2 parts _____, and 1 part emulsifying agent.
19. The Chemical Reaction Method is a reaction between fatty acids and a weak _____ base. No emulsifying agent is needed.

20. The gums (acacia, pectin, and tragacanth) have a common storage problem which is _____ and as a result, they must be _____ and preserved.
21. The preservatives used in emulsions must be soluble in the _____ phase.
22. List the following preservatives:
- _____ : 7% to 15% of the TOTAL of the water used in the entire preparation.
- _____ : 0.2% of the aqueous phase of the primary emulsion.
- _____ : 0.1% of the aqueous phase of the primary emulsion.
23. Acacia Emulsions are stable over a wide range (2 to _____ pH).
24. Tragacanth Emulsions and Pectin Emulsions are stable only in an acid (1- _____ pH).
25. Match the Following:
- | | |
|--|------------------------------|
| Palatable source of Vitamins A&D _____ | a. Mineral Oil Emulsion N.F. |
| Irritant Cathartic _____ | b. Castor Oil Emulsion |
| Emollient and Protectant _____ | c. Cod Liver Oil Emulsion |
| Cathartic _____ | d. Calamine Emulsion |
| 30 ml dose _____ | |
| 15 ml dose _____ | |
| External Use Only _____ | |

OINTMENTS

QUESTIONS

1. Soft, semi-solid preparations usually containing medicinal agents intended for application to the skin or mucous membrane with or without rubbing is the definition of _____.
2. Ointments are classified by composition or by therapeutic action based on _____.
3. _____ ointment bases includes fixed oils of vegetable origin, fats obtained from animals and semisolid hydrocarbons obtained from petroleum.
4. Oleaginous ointment bases are highly compatible, good emollients but they are difficult to remove from skin and clothes, unstable, and not good _____ absorbers.
5. _____ bases are generally anhydrous substances which have the property of absorbing considerable quantities of water and still retain their ointment-like consistency. They are highly compatible, relatively heat stable but are unpleasant to use and not _____.
6. Emulsion bases are actually _____ emulsions.
7. Emulsion bases are water washable, easily applied and removed but they must be preserved and are subject to _____ loss.
8. Water soluble ointment bases have a wide range of compatibilities, do not support mold growth, are nonirritating, adhere well to skin, and are easily _____ off.

9. Match the Following:

Acts on the surface of the skin to produce local effect _____

- a. Epidermic
- b. Endodermic

Penetrates into the deeper layers of the skin but not through the skin _____

- c. Diadermic

Penetrates through the skin and medication may be absorbed _____

- 10. Ointments are prepared by Mechanical Incorporation or _____.
- 11. Solid substances having high melting points are incorporated into ointments by _____.
- 12. When the materials used to prepare an ointment are all soft at room temperature, the method of preparation would be mechanical _____.
- 13. Ointment bases of fats and oils may become _____.
- 14. Emulsion bases will support _____ growth.
- 15. Sulfur Ointment USP is a _____.
- 16. Whitfield's Ointment is an _____.
- 17. Ointment-like mixtures which generally contain a higher percentage of powdered materials are called _____.
- 18. Semisolid emulsions containing suspensions or solutions of medicinal agents for external use are called _____.
- 19. Cold Cream USP is used as a base, _____, and cleansing agent.
- 20. Vioform HC (Iodochlorhydroxyquin and Hydrocortisone) is an anti-infective _____.

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21. Vioform and Hydrocortisone cream should be prepared by using a
_____ spatula.

SUPPOSITORIES

QUESTIONS

1. Solid bodies of various weights and shapes adapted for introduction into one of the orifices of the body and usually melting, softening or dissolving at body temperature is the definition of _____.
2. Rectal suppositories are bullet shaped, cylindrical and tapered to a point with a weight of _____ Gm.
3. Vaginal suppositories are globular or balloon shaped, weigh _____ Gm and are sometimes called a pessary.
4. _____ suppositories are rod or pencil shaped and have a weight of 2 Gm (Female) and 4 Gm (Male).
5. A disadvantage of suppositories is inconvenience of _____.
6. Suppositories are made by Cold Compression, Hand Shaping, or by _____.
7. Match the Following:

For Migraine headaches _____	a. Glycerin Suppositories
Antiasthmatic _____	b. Aspirin Suppositories
Analgesic, Antipyretic _____	c. Ergotamine Tartrate and Caffeine Suppositories (Cafergot)
Laxative _____	d. Aminophylline Suppositories
Rectal Evacuant _____	e. Tannic Acid Suppositories
Astringent _____	f. Bisacodyl (Dulcolax)

PARENTERALS

QUESTIONS

1. Sterile solutions, suspensions or emulsions for injection under or through one or more layers of skin or mucous membrane are called _____.
2. An injection into the substance of the muscles is called _____.
3. An injection into the vein (most rapid onset of action) is called _____.
4. An injection into the corium or substance of the skin is called _____.
5. An injection administered beneath the skin but not into the muscle is called _____.
6. The therapeutic introduction of a fluid, usually a large volume, into a vein by gravity flow is called _____.
7. Parenterals are classified by their _____ characteristics.
8. Sterile _____ have dissolved medicaments in aqueous, oily, or organic solvents.
9. Sterile _____ have solid medicaments in aqueous or oily vehicles.
10. Sterile Solid Medications yield a _____ upon the addition of a suitable vehicle.

11. Sterile Solid Medications can also yield a _____ upon the addition of a suitable vehicle.
12. Sterile _____ are water-in-oil or oil-in-water and are not suitable bases for antibiotics.
13. To increase the local action, _____ are used with local anesthetics.
14. Soluble, filterable, thermostable, substances resulting from the decomposition of certain strains of bacteria are called _____.
15. Parenteral administration is available for many drugs which are _____ by gastric juices.
16. Ampules are all glass, have a constriction at the neck, and are usually _____ dose.
17. _____ are stoppered glass containers that can be single or multiple dose.
18. Stoppered glass containers which preserve, in a sterile condition, multiple doses of parenteral medications are called _____.
19. Syringe containers (Tubex) are the most convenient in the administration of emergency, life-saving drugs. (TRUE/FALSE)
20. The potency of antibiotic parenterals can usually be extended by keeping them in the _____.
21. _____ are chemical agents which are added to parenteral solutions to enable them to resist a change in pH.

22. A long-acting parenteral would probably have an _____
base.
23. The container of choice for the administration of Narcotics would be
a _____.

BULK COMPOUNDING AND PREPACKAGING

QUESTIONS

1. Bulk compounded items are prepared for inpatient use, outpatient use, and _____.
2. Bulk compounding saves time, saves money, and provides a _____ system.
3. When using pharmaceutical equipment, the most important consideration is _____.
4. A water still can either be steam operated or _____.
5. The water produced by a still should be _____.
6. The food blender (Waring Blender) is an excellent device for making _____, lotions, mixtures, magmas, suspensions, and gels.
7. The Laboratory Magnetic Stirrer-Hot Plate (Thermo Magna Stir) is used for mixing preparations of low viscosity and is used with an _____ flask because of the vortex caused by the magnet.
8. The Alsop Mixer is used for mixing and filtering amounts over 4,000 ml and is extremely efficient and excellent for making Simple Syrup USP by the _____ method.
9. The main advantage of using a Suppository Compression Machine in the preparation of Cocoa Butter Suppositories is that the _____ of the base is not changed.

10. The major factor to be considered when determining the quantity of an item to be compounded in bulk is the amount to be used in the _____.
11. The 3 factors of Quality Control in Bulk Compounding are:
- a. Quality Control Forms
 - b. Lot Numbers
 - c. _____ Letters.
12. Any drug reported in the Air Force Medical Materiel Office Letter as unsuitable for use can only be identified by a _____ number.
13. The expiration date of any dated item compounded in the pharmacy will be determined by the expiration date of the _____ drug to expire in that preparation.
14. All labels made for bulk compounded preparations will be written using the _____ name only.
15. Guidelines for prepackaging medications are established in AFM _____.
16. The primary factor to be considered when determining which medications to prepackage for outpatient dispensing is _____ of use.
17. Prepackaged medication labels must have the prescribing physician's name. (TRUE/FALSE)
18. Prepackaged medications must have the original manufacturer's lot number on the label OR have a lot number assigned by the _____.

19. A disadvantage that may arise from prepackaging bulk compounds is that it can cause a critical _____ problem.
20. The bottle filling machine is used to fill _____ bottles.
21. The Mini-Counter is designed to count _____ or Capsules.
22. The Prescription Label Imprinter is primarily used for printing _____ labels.
23. Which of the following would not normally be prepackaged?
- a. Ointments
 - b. Parenterals
24. When the pharmacy transfers the contents from an original stock container of 500 tablets, to five stock containers of 100 tablets each, the procedure is called _____.
25. Quality Control is lost once the medication is dispensed to the _____.

INTRAVENOUS ADMIXTURES

1. Asepsis is the prevention of the access of _____.
2. Aseptic technique does not assure _____ but prevents further contamination by microorganisms.
3. When preparing I.V. admixtures, the disinfectant of choice is _____.
4. The possibility of upper respiratory microorganisms contaminating the aseptic environment could be reduced by wearing a _____.
5. The primary purpose of the laminar flow hood is to provide an _____ air flow.
6. To slow down or to prevent bacterial growth, I.V. admixtures are stored in the _____.
7. A needle and syringe with a 2 way valve is used as the equipment of choice for _____ reconstitution.
8. In-line filtration and random sampling is used to check for _____ contamination.
9. Light and Dark field observation is used to check for _____ matter.
10. I.V. admixtures must be sealed with a _____ cap.
11. I.V. admixtures will automatically expire within _____ hours from time of preparation.

12. When preparing I.V. admixtures, most contamination is caused by poor _____ technique.
13. After receiving an I.V. admixture prescription, the next step in the workflow pattern would be to research _____.
14. List the 4 transfer techniques:
- a.
 - b.
 - c.
 - d.
15. When the exact quantity of an additive is required from an ampule or vial, the _____ and _____ transfer method is used.
16. Total parenteral feeding of a patient by the use of protein hydrolysate, carbohydrates, vitamins, and electrolytes is called _____.
17. The method used to prepare I.V. Hyperalimentation solutions would involve the use of a _____ set.

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IV CHECKLIST

1. Clean equipment
 - a. Was isopropyl alcohol 70 percent used on everything except plexiglass?
 - b. Was Benzalkonium Cl used on plexiglass?
2. Check for incompatibilities
 - a. Was an incompatibility found using proper reference materials?
 - b. Was an instructor notified if an incompatibility was found?
3. Label
 - a. Start before (typed) - date and time left blank
 - b. Patient's name, ward, and bed number
 - c. Today's date
 - d. Bottle number
 - e. Start time
 - f. Ingredients (same as on the prescription)
 - g. Amounts (same as on the prescription)
 - h. In... (base solution)
 - i. Infusion rate
 - (1) How many hours
 - (2) Gtts/min
 - (3) ml/hr (if over 150 ml)
 - j. Doctor's name
 - k. Facility
4. Assemble tray (checked by the instructor)
 - a. Are there proper ingredients?
 - b. Is there correct equipment for reconstitution of each drug, if necessary?
 - (1) Size of syringe
 - (2) Proper diluent
 - c. Is there correct equipment for transfer?
 - (1) Double needle for the entire contents of vial
 - (2) Needle and syringe for specific quantity
 - (a) Same needle and syringe may be used for reconstitution and transfer of the same drug
 - (b) Same N&S may be used for reconstitution of the same drug
 - d. Was a N&S requisition form turned in?
 - e. Is there a correct base solution and quantity?
 - f. Is there a tamper-proof cap?
 - g. Is there a scorer if necessary?
 - h. Is the prescription included?
 - i. Is the label included?
 - j. Is there a mask?
 - k. Were all ingredients taken out of cardboard containers?
5. Prepare IV under laminar flow hood
 - a. Was hood cleaned with isopropyl alcohol 70 percent and Benzalkonium Cl?
 - b. Was tray cleaned?
 - c. Were glass containers cleaned?
 - d. Was mask put on?
 - e. Were all ingredients taken off tray except cap, prescription, label, and scorer?

- f. Were ingredients placed in a line parallel to back of hood with no item blocking another from the air flow?
 - g. Are all procedures done in front of the 6" line?
 - h. All stoppers cleaned with isopropyl alcohol 70 percent before each needle insertion?
 - i. Are all items (gauze wrapper, syringe packages, and ampuls) opened away from the working area, preferably over the tray?
 - j. Were all items using a double needle inserted first to utilize the vacuum?
 - k. Were lyophilized drugs reconstituted?
 - (1) Proper dilution
 - (2) Proper diluent
 - l. Were needle and syringe assembled and used properly?
 - (1) Sterile paper wrap opened like a banana
 - (2) Was student aware of sterility of needle hub, needle, and syringe end?
 - (3) Was protective cap on syringe kept on until needle was ready to be attached?
 - (4) Was needle sheath kept on when not in use?
 - (5) Was the plunger kept sterile, i.e., not touched?
 - m. Were all needles inserted at a 45 degree angle to prevent coring?
 - n. Were all additives inserted in proper area on IV solution stopper?
 - o. With double needle transfers, were vials raised to keep needle end in solution?
 - p. With a needle and syringe transfer using an ampule, was the needle kept in solution by tilting the ampule at a 45 degree angle with the needle end at the top or deepest part?
 - q. Were ampules opened properly?
 - (1) Scored once
 - (2) Use of gauze
 - (3) Opened away from technician
 - r. Was the IV shaken after each additive?
 - s. Were all paper scraps thrown away, while all other materials (needle, syringes, ampules, vials, etc.) kept on the tray after use?
 - t. Was IV stopper cleaned with isopropyl after the last additive and prior to the tamper-proof cap being added?
6. Check for particulate matter using light and dark field examination
- a. Was the solution turned on end and swirled gently?
 - b. Was the admixture held up to a dark field to detect light particles?
 - c. Was the admixture held up to a light field to detect dark particles?
 - d. Was the instructor notified of any particulate matter?
7. Label affixed
- a. Was the label affixed so it could be read while the admixture was hanging?
 - b. Was the label affixed so it would not cover the manufacturer's bold faced print on label?
 - c. Was expiration time and date put on label?
 - d. Was particulate matter notated on label?
 - e. Was "Refrigerate" label affixed?
8. Tray check
- a. Was tray checked by partner?
 - b. Was tray checked for proper ingredients and proper amounts?
 - c. Was tray checked against label and prescription?
9. Area clean up
- a. Were needles and syringes destroyed?
 - b. Were all other materials thrown away?
10. Delivery
- a. Was IV admixture delivered to the instructor with prescription and checklist?

GLOSSARY OF PHARMACEUTICAL LAB TERMS

1. AROMATIC WATER. Solutions of volatile oils or other aromatic or volatile substances in purified water.
2. ANTAGONISM. The effect of two or more drugs having opposite actions when administered together.
3. BUFFERING AGENTS. Those substances which resist a change in hydrogen ion concentration (pH) by reducing the ionization of acids or alkalies.
4. BULK POWDERS. Powders dispensed in bulk, usually measured out by the patient.
5. CAPSULE. Shells of gelatin used for containing individual doses of medication.
6. CHEMICAL INCOMPATIBILITY. When a new compound of undesirable nature forms from the interaction of two or more drugs, the incompatibility is chemical.
7. CHEMICAL REACTION METHOD. An emulsion formed by the reaction between a weak alkaline solution and a fatty acid.
8. COLD PLACE. Any temperature not exceeding 8 degrees C. or 46 degrees F.
9. COLLOID. A gelatinous substance made up of very small, insoluble, non-diffusible particles, larger than molecules but small enough so they remain suspended in a fluid medium without settling to the bottom; a colloid does not affect the freezing point, boiling point, or vapor tension of the medium in which it is suspended.
10. COMMINATION. The process of physically reducing solid substances into smaller fragments or particles.
11. CONGEALING POINT. The point at which a melted solid becomes a solid again.
12. CONTINENTAL METHOD. "Dry Gum Method" for the use with fixed oils only. Use the ratio to form the primary emulsion only. (4:2:1)
13. CONTUSION. The process of placing in a heavy mortar and pounding with a heavy pestle to break down the cellular structure of fresh drugs.
14. COOL PLACE. Any temperature between 8 and 15 degrees C. (46 to 59 degrees F.)
15. CREAMS. Semi-solid emulsions containing suspension or solutions of medicinal agents for external application.
16. DELIQUESCENT SUBSTANCE. A substance that absorbs moisture from the atmosphere but to a greater degree than hygroscopic substances. Deliquescent substances finally liquify.
17. DISPERSING AGENT. A substance that breaks down or reduces globule size of oils which results in more complete mixing in an aqueous phase.
18. DISTILLATION. Separation of the constituents of a liquid mixture by vaporization and subsequent condensation of the vapors.
19. DIVIDED POWDERS. Powders dispensed with the dosage premeasured by the pharmacist.
20. EFFLORESCENT SUBSTANCE. A substance that is opposite in its reaction in the atmosphere and gives up moisture (water of crystallization). It spontaneously changes from crystalline nature to amorphous powder.

21. ELIXIR. A clear, sweetened, hydro-alcohol liquid, medicated or nonmedicated and intended for oral use.
22. ELUTRIATION. Water softing.
23. ENGLISH METHOD. "Wet Gum Method" for use with a fixed oil only. Use the ratio to form the primary emulsion, but the water is added to the emulsifying agent to make a mucilage and the oil is added gradually. (4:2:1)
24. EUTECTIC MIXTURES. When certain drugs, solid at room temperature, are mixed, a lowering of the fusing or melting point occurs, causing the mixture to spontaneously liquefy without the aid of a solvent.
25. EVAPORATION. Driving off as a vapor the volatile portion of a liquid by application of heat.
26. EXCESSIVE HEAT. Any temperature above 40 degrees C. or 104 degrees F.
27. EXSICCATION. Removal of water of crystallization, or moisture, from a solid crystalline substance by heating strongly.
28. FORBES METHOD. "Bottle Method" for use with volatile and nonviscous oils. Made entirely in an appropriate size bottle in the ratio of 2:2:1.
29. FRESH DRUG TINCTURE. Must not exceed 50 percent active ingredient.
30. FUSION. Liquefying solid substances by the application of heat, without the use of a solvent. (Melting)
31. GARBLING. Hand picking or sorting.
32. GEL. Aqueous liquid preparations, containing suspended insoluble material of nearly colloid size intended for internal use.
33. GEOMETRIC DILUTION. The potent drug is first placed in the mortar with an equal bulk of diluent and triturated until mixed well, then an amount of diluent equal to the combined bulk of the potent drug along with its diluent is added and so forth until all of the necessary diluent has been added.
34. HYPERTONIC. More than isotonic; having a greater concentration of dissolved particles, fluid will be drawn into this solution from the less concentrated (hypotonic or isotonic) area.
35. HYPOTONIC. Less than isotonic; having a lesser concentration of dissolved particles, fluid will be drawn from this solution to the more concentrated (hypertonic or isotonic) area.
36. IGNITION. Process of strongly heating solid or semi-solid substances to a definite and limited degree. The residue on this is the product sought.
37. IMMISCIBLE. When two or more liquids are physically unable to mix homogenously, they are said to be immiscible.
38. ISOTONIC. Having the same osmotic pressure and the same number of particles (molecules or ions) in solution as another solution. Thus, we say that blood is isotonic with tear fluid and both blood and tear fluids have the same tonicity as a 0.9 percent solution of sodium chloride. They are isotonic.

39. LINIMENTS. Oily or alcoholic preparations for external use to be applied WITH friction.
40. LOTIONS. Liquid preparations, usually aqueous, containing insoluble material, intended for external use, and are to be applied WITHOUT friction.
41. MACERATION. Maceration is extraction by soaking. More exactly, maceration is the process of soaking the properly comminuted drugs or substance in the menstrum until the cellular structure is thoroughly penetrated and the soluble portions softened and dissolved.
42. MAGMAS. Viscous, aqueous liquid preparations containing suspended insoluble material intended for internal use.
43. MIXTURES. Non-viscous aqueous preparation containing insoluble material intended for internal use.
44. NON-POTENT TINCTURE. Must not exceed 20 percent active ingredient.
45. OIL-IN-WATER EMULSION. When the oil is uniformly dispersed within the water or the internal phase is oil and the external phase is water.
46. OSMOTIC PRESSURE. That force causing a liquid to pass through a semi-permeable membrane from a lower to a higher concentration. The passing of fluid through the membrane is known as osmosis.
47. PHYSICAL INCOMPATIBILITY. When a change in physical state occurs in one or more substances in a mixture, producing a cloudy, unsightly, or otherwise undesirable product, the incompatibility is physical.
48. PHASE. This is a term which refers to either of the two liquid portions of the emulsion.
49. PASTES. Ointment-like mixtures which generally contain a higher percentage of powdered materials for external application.
50. PRESERVATIVES. A substance added to prevent the growth of microorganisms, to inhibit oxidation and other changes not desirable in the product.
51. PYROGEN. Soluble, filterable, thermostable, substances resulting from the decomposition of certain strains of bacteria.
52. REFRIGERATE. A cold place in which the temperature is held between 2 and 3 degrees C. (36 and 46 degrees F.)
53. RATE OF SOLUTION. The rate or speed of the solute (solid) to completely dissolve in the solvent (liquid) is referred to as the rate of solution.
54. RASPING, GRATING, CUTTING, SLICING AND CHOPPING. These procedures are also known as the "solid process," and they are all methods of preparing fresh drugs.
55. SOLUBILITY. The extent to which the solute (solid) dissolves in the solvent (liquid) is referred to as its solubility.

SOLUBILITY TABLE

<u>DESCRIPTIVE TERMS</u>	<u>PARTS OF SOLVENT FOR 1 PART OF SOLUTE</u>
--------------------------	--

Very soluble	Less than 1
Freely soluble	From 1 to 10
Soluble	From 10 to 30
Sparingly soluble	From 30 to 100
Slightly soluble	From 100 to 1000
Very slightly soluble.	From 1000 to 10,000
Practically insoluble or insoluble . .	More than 10,000

- 56. SOLUTIONS. Aqueous solutions of non-volatile substances.
- 57. SPIRITS. Alcoholic or hydroalcoholic solutions of volatile substances.
- 58. SUBLIMATION. Separation of volatile solids from non-volatile solids; the product obtained is the sublimate.
- 59. SUPPOSITORIES. Solid bodies of various weights and shapes adapted for introduction into one of the orifices of the body and usually melting, softening at body temperature.
- 60. SUSPENSION. Liquid preparation containing suspended material for oral injection or ophthalmic use.
- 61. SYNERGISM. A joint action of two or more drugs combined so that their total effect is greater than would be expected from the sum of their individual effects.
- 62. SYRUP. A nearly saturated aqueous solution of sugar, with or without a medicinal or flavoring agent.
- 63. THERAPEUTIC INCOMPATIBILITY. When medications administered together produce a response different from that occurring upon individual administration, the incompatibility is therapeutic. Overdoses as well as the "wrong drug" also represent a therapeutic incompatibility.
- 64. TINCTURE. Alcoholic or hydro-alcoholic solution prepared from vegetable drugs or chemical sources.
- 65. TONICITY. The tension or concentration of a solution or substance.
- 66. TORREFACTION. Roasting certain organic substances in order to modify some of their constituents.
- 67. TRACER. a coloring agent used to color a potent drug and show its presence when diluted.
- 68. TRITURATION. The process of reducing substances to a powder by rubbing them in a mortar with a pestle.

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69. **TURBIDITY.** When water is added to a spirit and the alcohol content is lowered, the oil is separated from the alcoholic phase, and as a result, the preparation appears murky with a slight pearlescent sheen.
 70. **VAPORIZATION.** Converting a liquid or solid into a vapor.
 71. **VOLATILE SUBSTANCE.** A substance that evaporates at room temperature usually giving off a characteristic odor.
 72. **WATER-IN-OIL EMULSION.** When the water is uniformly dispersed within the oil or the internal phase is water and the external phase is the oil.

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DEPARTMENT OF BIOMEDICAL SCIENCES

10-8

PHARMACY SPECIALIST

PHARMACEUTICAL PREPARATIONS

August 1975



SCHOOL OF HEALTH CARE SCIENCES, USAF
SHEPPARD AIR FORCE BASE, TEXAS

Designed For ATC Course Use

DO NOT USE ON THE JOB

Department of Biomedical Sciences
School of Health Care Sciences, USAF
Sheppard Air Force Base, Texas 76311

WB 3ABR90530-III-3
August 1975

PHARMACEUTICAL PREPARATIONS

OBJECTIVE

Given instructor assistance, necessary references, selected formulas, and laboratory preparation sheets, complete the preparation sheets, compound waters, spirits, elixers, tinctures, solutions, syrups, ear, eye, and nose preparations, mixtures, suspensions, gels, lotions, liniments, capsules, emulsions, ointments, pastes & creams, correcting any incompatibilities, using accepted methods and techniques. Student will then package the preparation in a suitable container and label in accordance with AFM 168-4.

EQUIPMENT

United States Pharmacopia
National Formulary
Martin's Dispensing of Medication
Necessary Laboratory Equipment

PROCEDURES

The object of this lesson is to acquaint you with proper extemporaneous pharmaceutical compounding procedures. Specifically, you will:

1. Calculate and reduce formulas given on AF Form 2381.
2. Record all pertinent information on AF Form 2380.
3. Properly utilize AF Forms 2382, 582, and AF Form 781.
4. Label and package all preparations in accordance with AFM 168-4.

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SAFETY CHECKLIST FOR PHARMACY
SPECIALIST COURSE 3ABR90530

I acknowledge that I have been informed about the safety hazards involved in the use of the following areas and equipment.

1. Pharmacy Laboratory Area

- a. Pipettes
- b. Pill tiles
- c. Fisher burners
- d. Chemicals or caustic agents
- e. All pharmacy equipment:
 - (1) Alsop mixer
 - (2) Bottle filler
 - (3) Mini-counter
 - (4) Magnetic-heat stirrer
 - (5) Labeling machine
 - (6) Balances
 - (7) Glassware
 - (8) Needles and Syringes
 - (9) Laminar flow hood
 - (10) I.V. preparation bottles
 - (11) Ampules and Vials

Signature

Code #

Date

719

PHARMACY MASTER FORMULA		COST		ATTACH LABEL HERE
		12 OZ.	8 OZ.	
		1 OZ.	16 OZ.	
		3 OZ.	32 OZ.	
PRODUCT		PEPPERMINT WATER		USP p816
INGREDIENTS				AMOUNT
1	Peppermint oil			2 ml
2	Talc			15 Gm
3	Distilled water			1000 ml
4				
5				
6				
7				
8				
9				
10				
11				
12				
DIRECTIONS FOR MANUFACTURE				
Place 2 ml () of peppermint oil in a small mortar to which has been added 15 Gm () of talc. Then add a small amount of the 1000 ml () of distilled water to form a paste. Add the remainder of the water to wash the contents of the mortar into a flask. Stopper and shake intermittently for ten minutes. Filter until clear.				
LABELING				
Keep out of the reach of children				
SPECIAL CONTAINER REQUIREMENTS				
TLRC Store in a cool place				
THEORETICAL YIELD				
1000 ml				
DATE	PREPARED BY		DATE	CHECKED BY

720

PHARMACY MANUFACTURING CONTROL DATA						FOR TRAINING PURPOSES ONLY	
PRODUCT		LOT NUMBER			ATTACH LABEL HERE		
INGREDIENTS		MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY	
1							
2							
3							
4							
5							
6							
7							
8							
9							
10							
11							
12							

AF FORM 2380 JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (if any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-438-192

D-24844

FOR TRAINING PURPOSES ONLY

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT SYRUP USP p706			
INGREDIENTS			AMOUNT
1	Sucrose		850 Gm
2	Purified water qsad		1000 ml
3			
4			
5			
6			
7			
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE			
Heat about 450 ml () of purified water to a slow boil. Remove the heat. Add the sucrose. Stir until dissolved. Gentle heat may be applied to aid solution. Add a sufficient quantity of purified water to make the product measure 1000 ml (). Filter through gauze. NOTE: Do not heat for a prolonged period and do not boil as either will cause sucrose to invert and ultimately caramelize.			
LABELING Refrigerate			
SPECIAL CONTAINER REQUIREMENTS TLRC			
THEORETICAL YIELD 1000 ml			
DATE	PREPARED BY	DATE	CHECKED BY

[illegible]

AF 2380

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD		ACTUAL YIELD	REASON FOR DISCREPANCY (If any)
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-438-392

9-24544

747

723

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT LERHAM'S SOLUTION Local			
INGREDIENTS			AMOUNT
1	Ethylaminobenzoate		10.72 Gm
2	Phenol (liquid)		10.72 ml
3	Menthol		21.44 Gm
4	Alcohol		100 ml
5	Glycerin qsad		1000 ml
6			
7			
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE Dissolve the ethylaminobenzoate, menthol and phenol in the 100 ml () of alcohol. Add a sufficient quantity of glycerin to make the product measure 1000 ml ().			
LABELING Keep out of the reach of children (For the ear)			
SPECIAL CONTAINER REQUIREMENTS TLRC			
THEORETICAL YIELD 1000 ml			
DATE	PREPARED BY	DATE	CHECKED BY

724

PHARMACY MANUFACTURING CONTROL DATA FOR TRAINING PURPOSES ONLY					
PRODUCT	LOT NUMBER			ATTACH LABEL HERE	
INGREDIENTS	MEAS.	LOT NUMBER	AMOUNT	RECEIVED BY	CHECKED BY
1					
2					
3					
4					
5					
6					
7					
8					
9					
10					
11					
12					

AF FORM 2380
JUN 73

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (if any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-435-392

B-24844

728

PHARMACY MASTER FORMULA		COST		ATTACH LABEL HERE
		12 OZ.	8 OZ.	
		1 OZ.	16 OZ.	
		2 OZ.	32 OZ.	
PRODUCT PHENYLEPHRINE HCL NASAL SOLUTION 1/4% local				
INGREDIENTS				AMOUNT
1	Phenylephrine HCl 1%			250 ml
2	Normal Saline qsad			1000 ml
3				
4				
5				
6				
7				
8				
9				
10				
11				
12				
DIRECTIONS FOR MANUFACTURE Place 250 ml () of phenylephrine HCl 1% in a graduate. Add a sufficient quantity of normal saline to make the product measure 1000 ml (). Filter three times through long strand cotton that has been previously wetted with normal saline.				
LABELING Keep out of the reach of children				
SPECIAL CONTAINER REQUIREMENTS TLRC				
THEORETICAL YIELD 1000 ml				
DATE	PREPARED BY		DATE	CHECKED BY

726

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT		LOT NUMBER			ATTACH LABEL HERE	
	INGREDIENTS	MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

FORM 2380
AF JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (IF ANY)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-626-102

8-24944

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT DIPHENHYDRAMINE HYDROCHLORIDE ELIXIR USP p209			
INGREDIENTS			AMOUNT
1	Diphenhydramine HCl		2.5 Gm
2	Orange oil		0.24 ml
3	Cinnamon oil		0.11 ml
4	Clove oil		0.08 ml
5	Coriander oil		0.03 ml
6	Anethole (**substitute anise oil**)		0.03 ml
7	Amaranth solution		1.6 ml
8	Alcohol		150 ml
9	Syrup		350 ml
10	Purified water qsad		1000 ml
11			
12	Talc		15 Gm
DIRECTIONS FOR MANUFACTURE Dissolve the diphenhydramine HCl in 250 ml () of purified water, and the flavoring oils and the anethole* in the alcohol. Mix the two solutions, and add the syrup, the amaranth solution, and sufficient purified water to make 1000 ml (). Mix, and filter if necessary. (May use talc) NOTE: Add the alcoholic solution to the aqueous solution.			
LABELING Diphenhydramine HCl 12.5mg/5ml			
SPECIAL CONTAINER REQUIREMENTS TLRC			
THEORETICAL YIELD 1000 ml			
DATE	PREPARED BY	DATE	CHECKED BY

720

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT		LOT NUMBER			ATTACH LABEL HERE	
INGREDIENTS		MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

AF FORM 2380 JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (If any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-456-362

B-24844

72

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT SWEET ORANGE PEEL TINCTURE USP p463			
INGREDIENTS			AMOUNT
1	Sweet orange peel		500 Gm
2	Alcohol USP		900 ml
3	Alcohol USP qsad		1000 ml
4	Talc		15 Gm
5			
6			
7			
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE Process M (see pg 816 USP) Macerate 500 Gm () of the sweet orange peel (NOTE - EXCLUDE THE INNER WHITE PORTION OF THE RIND) in 900 ml () of alcohol and allow to set for three days. Complete the preparation with alcohol to make the product measure 1000 ml (). Use talc as the filtering medium.			
LABELING			
SPECIAL CONTAINER REQUIREMENTS TLRC Store in a cold place			
THEORETICAL YIELD 1000 ml			
DATE	PREPARED BY	DATE	CHECKED BY

730

PHARMACY MANUFACTURING CONTROL DATA FOR TRAINING PURPOSES ONLY					
PRODUCT	LOT NUMBER			ATTACH LABEL HERE	
INGREDIENTS	MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY
1					
2					
3					
4					
5					
6					
7					
8					
9					
10					
11					
12					

AF FORM 2380 JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (If any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-438-392

B-24944

731

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT 3% IODOCHLORHYDROXYQUIN SUSPENSION (Local)			
INGREDIENTS			AMOUNT
1	Iodochlorhydroxyquin		30 Gm
2	Talc		50 Gm
3	Zinc Oxide		50 Gm
4	Bentonite Magma		250 ml
5	Purified water qsad		1000 ml
6			
7			
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE Triturate the powders until a fine state of subdivision is obtained. Dilute the 250 ml () of bentonite magma with an equal volume of purified water. Slowly add the diluted mixture to the powders forming a thick paste. Gradually add 200 ml () of purified water, with constant trituration. Transfer to a graduate and add a sufficient quantity of purified water to make the product measure 1000 ml ().			
LABELING Shake well. For external use only. Keep from freezing			
SPECIAL CONTAINER REQUIREMENTS TLRC			
THEORETICAL YIELD 1000 ml			
DATE	PREPARED BY	DATE	CHECKED BY

732

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT		LOT NUMBER			ATTACH LABEL HERE	
	INGREDIENTS	MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

AF FORM 2380 JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (If any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-430-192

B-24844

733

PHARMACY MASTER FORMULA		COST		ATTACH LABEL HERE
		12 OZ.	8 OZ.	
		1 OZ.	16 OZ.	
		2 OZ.	32 OZ.	
		4 OZ.		
PRODUCT CALAMINE LOTION USP p87				
INGREDIENTS				AMOUNT
1	Calamine			80 Gm
2	Zinc oxide			80 Gm
3	Glycerin			20 ml
4	Bentonite magma			250 ml
5	Calcium hydroxide solution qsad			1000 ml
6				
7				
8				
9				
10				
11				
12				
DIRECTIONS FOR MANUFACTURE				
Dilute the bentonite magma with an equal volume of calcium hydroxide solution. Mix the powders intimately with the glycerin and about 100 ml () of the diluted magma, triturating until a smooth, uniform paste is formed. Gradually add the remainder of the diluted magma. Finally add enough calcium hydroxide solution to make 1000 ml (), and shake well. If a more viscous consistency in the lotion is desired, the quantity of bentonite magma may be increased to not more than 400 ml (). <u>NOTE:</u> Shake Calamine Lotion well before dispensing.				
LABELING Shake well, For external use				
SPECIAL CONTAINER REQUIREMENTS TLRC				
THEORETICAL YIELD 1000 ml				
DATE	PREPARED BY		DATE	CHECKED BY

7139

PHARMACY MANUFACTURING CONTROL DATA						FOR TRAINING PURPOSES ONLY	
PRELIMINARY		LOT NUMBER			ATTACH LABEL HERE		
INGREDIENTS		WGT	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY	
1							
2							
3							
4							
5							
6							
7							
8							
9							
10							
11							
12							

AF FORM 2380

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (if any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-638-392

3-24844

735

PHARMACY MASTER FORMULA	COST 12 OZ. 8 OZ. 1 OZ. 16 OZ. 2 OZ. 32 OZ. 4 OZ.	ATTACH LABEL HERE
PRODUCT ASPIRIN CAPSULES 200 mg (local)		
INGREDIENTS		AMOUNT
1	Aspirin	20 Gm
2	Lactose sufficient quantity	
3		
4		
5		
6		
7		
8		
9		
10		
11		
12		
DIRECTIONS FOR MANUFACTURE Punch a trial capsule using a #2 capsule. Weight of trial capsule x 10 () - Total weight of Act Ing x 10 20 Gm () Amount of lactose to use as filler Gm () Triturate the aspirin and lactose. (Mix by geometric dilution) Add 10 drops of yellow food color as a tracer. Punch and weigh each capsule. (Each capsule should equal the weight of the trial capsule) Clean each capsule with sucrose.		
LABELING Keep out of the reach of children		
SPECIAL CONTAINER REQUIREMENTS TLRC		
THEORETICAL YIELD 100 capsules		
DATE	PREPARED BY	DATE CHECKED BY

739

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT		LOT NUMBER			ATTACH LABEL HERE	
INGREDIENTS		MFG.	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

AF FORM 2380
JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (if any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-638-392

B-24644

737

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT CASTOR OIL EMULSION 50% (local)			
INGREDIENTS		AMOUNT	
1	Castor oil	60 ml	
2	Acacia	15 Gm	
3	Peppermint oil	0.02 ml	
4	Glycerin	11 ml	
5	Cocoa	7 Gm	
6	Purified water qsad	120 ml	
7			
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE			
Place 15 Gm of acacia in a mortar. Slowly add 60 ml of castor oil and triturate well. Add 0.2 ml of a 1:10 stock solution of peppermint oil and mix well. Add 30 cc of purified water all at once, and triturate rapidly until a thick, creamy emulsion is formed. Let rest. In a separate mortar mix the glycerin and the cocoa and combine this with the primary emulsion. Transfer to a graduate and add a sufficient quantity of purified water to make the product measure 120 ml. <u>NOTE:</u> The mortar should be warm, rough and dry.			
LABELING			
Shake well, Refrigerate			
SPECIAL CONTAINER REQUIREMENTS			
TLRC			
THEORETICAL YIELD			
120cc			
DATE	PREPARED BY	DATE	CHECKED BY

135

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT	LOT NUMBER			ATTACH LABEL HERE		
INGREDIENTS	MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY	
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

AF FORM 2380 JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (If any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-438-392

B-24844

739

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	10 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT 3% IODOCHLORHYDROXYQUIN - 1% HYDROCORTISONE OINTMENT (local)			
INGREDIENTS		AMOUNT	
1	Iodochlorhydroxyquin	30 Gm	
2	Hydrocortisone	10 Gm	
3	Propylene Glycol	10 ml	
4	Water soluble ointment base	950 Gm	
5			
6	to make about	1000 Gm	
7			
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE			
Levigate the iodochlorhydroxyquin and the hydrocortisone with the propylene glycol. Add the ointment base in small portions mixing well after each addition. <u>NOTE:</u> Do not use a metal spatula.			
LABELING			
For external use only			
CONTAINER REQUIREMENTS			
Avoid metal caps			
NET WEIGHT			
1000 Gm			
DATE	PREPARED BY	DATE	CHECKED BY

740

PHARMACY MANUFACTURING CONTROL DATA FOR TRAINING PURPOSES ONLY					
PRODUCT	LOT NUMBER			ATTACH LABEL HERE	
INGREDIENTS	MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY
1					
2					
3					
4					
5					
6					
7					
8					
9					
10					
11					
12					

AF FORM 2380 JUN 51

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (IF ANY)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1951 O-438-135

8-24864

741

PHARMACY MASTER FORMULA	COST 12 OZ. 8 OZ. 1 OZ. 16 OZ. 2 OZ. 32 OZ. 4 OZ.	ATTACH LABEL HERE
PRODUCT SULFUR OINTMENT USP p701		
INGREDIENTS		AMOUNT
1	Precipitated sulfur	100 Gm
2	Mineral oil	100 Gm
3	White ointment	800 Gm
4		
5		
6		
7		
8		
9		
10		
11		
12		
DIRECTIONS FOR MANUFACTURE Levigate the sulfur with the mineral oil to a smooth paste, and then incorporate with the white ointment.		
LABELING		
SPECIAL CONTAINER REQUIREMENTS Preserve in well closed containers and avoid prolonged exposure to excessive heat.		
THEORETICAL YIELD 1000 Gm		
DATE	PREPARED BY	DATE CHECKED BY

742

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT	LOT NUMBER			ATTACH LABEL HERE		
INGREDIENTS	MFG	LOT NUMBER	AMOUNT	REISHED BY	CHECKED BY	
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

AF FORM 2380
JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (If any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-420-192

B-24544

743

PHARMACY MASTER FORMULA	COST		ATTACH LABEL HERE
	12 OZ.	8 OZ.	
	1 OZ.	16 OZ.	
	2 OZ.	32 OZ.	
PRODUCT COLD CREAM USP pg143			
INGREDIENTS		AMOUNT	
1	Spermaceti	125 Gm	
2	White wax	120 Gm	
3	Mineral oil	560 Gm	
4	Sodium borate	5 Gm	
5	Purified water	190 ml	
6			
7	to make about	1000 Gm	
8			
9			
10			
11			
12			
DIRECTIONS FOR MANUFACTURE Reduce the spermaceti and white wax to small pieces, melt them on a steam bath with the mineral oil, and continue heating until the temperature of the mixture reaches 70°C. Dissolve the sodium borate in the purified water, warmed to 70°C, and gradually add the warm solution to the melted mixture, stirring rapidly and continuously until it has congealed.			
LABEL AS External use only			
SPECIAL CONTAINER REQUIREMENTS Preserve in tight containers			
THE TOTAL YIELD 1000 Gm			
DATE PREPARED BY		DATE CHECKED BY	

749

PHARMACY MANUFACTURING CONTROL DATA					FOR TRAINING PURPOSES ONLY	
PRODUCT	LOT NUMBER			ATTACH LABEL HERE		
INGREDIENTS	MFG	LOT NUMBER	AMOUNT	WEIGHED BY	CHECKED BY	
1						
2						
3						
4						
5						
6						
7						
8						
9						
10						
11						
12						

AF FORM 2380
JUN 71

LABELING			
CONTAINERS UTILIZED		TYPE	SIZE
SPECIAL SPECIFICATIONS			
THEORETICAL YIELD	ACTUAL YIELD	REASON FOR DISCREPANCY (If any)	
MANUFACTURED BY		TIME	CONTROL ACTION
REMARKS			
DATE	PREPARED BY	DATE	CHECKED BY

FOR TRAINING PURPOSES ONLY

U.S. GOVERNMENT PRINTING OFFICE: 1971 O-436-392

B-24844

745

AF FORM 781

When an AF Form 781 is required for compounding, it will be completed using the example below. All AF Forms 781 will be completed by the student prior to presentation to an instructor for his signature.

AF FORM
AUG 73 781

REPLACES DD FORM 1289, 1 NOV 71,
WHICH IS OBSOLETE IN THE USAF.

MULTIPLE ITEM PRESCRIPTION

1. <i>Cross out unused blanks below</i>		STRENGTH	AMOUNT	DIRECTIONS	R NUMBER	REFILL
ALCOHOL USP		-----	54 ml	Used to make Elixir of Terpin Hydrate 120 ml	Initials Rx # Time	-----
~~2.~~		~~-----~~	~~-----~~	~~Used by: Student's name~~	~~-----~~	~~-----~~
~~3.~~		~~-----~~	~~-----~~	~~-----~~	~~-----~~	~~-----~~
~~4.~~		~~-----~~	~~-----~~	~~-----~~	~~-----~~	~~-----~~

NAME OF PATIENT PHARMACY STOCK		SIGNATURE OF PRESCRIBER INSTRUCTOR'S SIGNATURE	
ADDRESS SHCS / SAFB		PRESCRIBER IDENTIFICATION (Name, SSAN or BYDD, Grade, Degree, Service and Facility)	
TELEPHONE IAB	AGE -----	DATE TODAY	
REMARKS (If to reverse, if necessary)			

FOR TRAINING PURPOSES
ONLY

770

144
NAME _____

CLASS NUMBER _____

1. Please consider each area of the course carefully and rate the elements of instruction according to the following scale:

- a. Excellent
- b. Good
- c. Fair
- d. Poor

2. After rating the elements, make specific comments which you feel will help to improve the course. Feel free to praise a deserving instructor or to offer constructive suggestions to aid him.

a. OVERALL COURSE

	<u>E</u>	<u>G</u>	<u>F</u>	<u>P</u>
(1) Course Length	_____	_____	_____	_____
(2) Course Content	_____	_____	_____	_____
(3) Difficulty Level	_____	_____	_____	_____
(4) Examinations	_____	_____	_____	_____
(5) Instructors	_____	_____	_____	_____
(6) Field Trip	_____	_____	_____	_____

Comments:

b. Pharmaceutical Calculations I

	<u>E</u>	<u>G</u>	<u>F</u>	<u>P</u>
(1) Course Content	_____	_____	_____	_____
(2) Training Aids, etc.	_____	_____	_____	_____
(3) Instructors	_____	_____	_____	_____
(4) Handouts, Study Guides	_____	_____	_____	_____
(5) Examinations	_____	_____	_____	_____
(6) Overall Effectiveness	_____	_____	_____	_____

Comments:

c. Pharmaceutical Inorganic Chemistry

(1) Course Content

(2) Training Aids, etc.

(3) Instructors

(4) Handouts, Study Guides

(5) Examinations

(6) Overall Effectiveness

<u>E</u>	<u>G</u>	<u>F</u>	<u>P</u>
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

Comments:

d. Pharmaceutical Organic Chemistry

(1) Course Content

(2) Training Aids, etc.

(3) Instructors

(4) Handouts, Study Guides

(5) Examinations

(6) Overall Effectiveness

<u>E</u>	<u>G</u>	<u>F</u>	<u>P</u>
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

Comments:

e. Pharmacy Administration

(1) Course Content

(2) Training Aids, Films

(3) Instructors

(4) Handouts, Study Guides

(5) Examinations

(6) Overall Effectiveness

<u>E</u>	<u>G</u>	<u>F</u>	<u>P</u>
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

Comments:

f. Pharmaceutical Dispensing	<u>E</u>	<u>F</u>	<u>G</u>	<u>P</u>
(1) Course Content	_____	_____	_____	_____
(2) Training Aids	_____	_____	_____	_____
(3) Instructors	_____	_____	_____	_____
(4) Handouts, Study Guides	_____	_____	_____	_____
(5) Individual Lab Supervision	_____	_____	_____	_____
(6) Written Examinations	_____	_____	_____	_____
Comments:				

g. Pharmaceutical Calculations II	<u>E</u>	<u>F</u>	<u>G</u>	<u>P</u>
(1) Course Content	_____	_____	_____	_____
(2) Training Aids	_____	_____	_____	_____
(3) Instructors	_____	_____	_____	_____
(4) Handouts, Study Guides	_____	_____	_____	_____
(5) Examinations	_____	_____	_____	_____
(6) Overall Effectiveness	_____	_____	_____	_____
Comments:				

h. Pharmacology	<u>E</u>	<u>F</u>	<u>G</u>	<u>P</u>
(1) Course Content	_____	_____	_____	_____
(2) Training Aids	_____	_____	_____	_____
(3) Instructors	_____	_____	_____	_____
(4) Handouts, Study Guides	_____	_____	_____	_____
(5) Examinations	_____	_____	_____	_____
(6) Overall Effectiveness	_____	_____	_____	_____
Comments:				

749

1. Course Administration

(1) Scheduling

(2) Were you kept informed?

(3) Textbooks (Make specific
comments below)

(4) Counseling

<u>E</u>	<u>F</u>	<u>G</u>	<u>P</u>
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

Comments:

774